

Synthesis and Characterization of a New Microporous Cesium Silicotitanate (SNL-B) Molecular Sieve.

M. Nyman[§], B. X. Gu[†], L. M. Wang[†] and R. C. Ewing[†], T. M. Nenoff^{§*}

[§] Sandia National Laboratories, Catalysis and Chemical Technologies; P.O. Box 5800,
M.S. 0710; Albuquerque, NM 87185-0710.

[†] University of Michigan, Department of Nuclear Engineering and Radiological Sciences;
Ann Arbor, MI 48109-2104.

*Author to whom correspondence should be addressed (email: tmnenof@sandia.gov).

Abstract

Ongoing hydrothermal Cs-Ti-Si-O-H₂O phase investigations has produced several new ternary phases including a novel microporous Cs-silicotitanate molecular sieve, SNL-B with the approximate formula of Cs₃TiSi₃O_{9.5}•3H₂O. SNL-B is only the second molecular sieve Cs-silicotitanate phase reported to have been synthesized by hydrothermal methods. Crystallites are very small (0.1 x 2 microns) with a blade-like morphology. SNL-B is confirmed to be a 3-dimensional molecular sieve by a variety of characterization techniques (N₂ adsorption, ion exchange, water adsorption/desorption, solid state CP-MAS NMR). SNL-B is able to desorb and adsorb water from its pores while retaining its crystal structure and exchanges Cs cations readily. Additional techniques were used to describe fundamental properties (powder X-ray diffraction, FTIR, ²⁹Si and ¹³³Cs MAS NMR, DTA, SEM/EDS, ion selectivity, and radiation stability). The phase relationships of metastable SNL-B to other hydrothermally synthesized Cs-Ti-Si-O-H₂O phases are discussed, particularly its relationship to a Cs-silicotitanate analogue of pharmacosiderite, and a novel condensed phase, a polymorph of Cs₂TiSi₆O₁₅ (SNL-A).

Keywords: molecular sieve, silicotitanates, Cs and Sr radionuclides, electron beam irradiation.

RECEIVED
OSTI
APR 10 2000

DISCLAIMER

This report was prepared as an account of work sponsored by an agency of the United States Government. Neither the United States Government nor any agency thereof, nor any of their employees, make any warranty, express or implied, or assumes any legal liability or responsibility for the accuracy, completeness, or usefulness of any information, apparatus, product, or process disclosed, or represents that its use would not infringe privately owned rights. Reference herein to any specific commercial product, process, or service by trade name, trademark, manufacturer, or otherwise does not necessarily constitute or imply its endorsement, recommendation, or favoring by the United States Government or any agency thereof. The views and opinions of authors expressed herein do not necessarily state or reflect those of the United States Government or any agency thereof.

DISCLAIMER

**Portions of this document may be illegible
in electronic image products. Images are
produced from the best available original
document.**

Introduction

The CST ion exchanger, a crystalline silicotitanate material jointly developed by Sandia National Laboratories (SNL) and Texas A & M University, is highly selective for Cs over a broad pH range and in the presence of competitive ions (i.e. Na).[1-5] Additionally, it is stable in extreme radioactive and chemical environments.[4] These properties make the CST ion exchanger useful for removal of the ^{137}Cs radionuclide from defense wastes such as those stored at the Hanford site. However, the Cs-loaded CST is not easily back-exchanged without destroying the framework structure.[4] Therefore, the CST material must be incorporated into a waste form for immobilization of ^{137}Cs that has been selectively sorbed. Direct thermal conversion of the Cs-CST produces a viable waste form by dehydration and phase transformation to form a mixture of crystalline products.[6, 7] In order to fully characterize the phases which 1) compose the original waste form material, and 2) precipitate as secondary alteration phases, exploratory syntheses are underway using the component oxides of Cs-CST. In particular, synthesis of Cs-Ti-Si-O-H₂O ternary phases directed toward Cs-CST waste form characterization is the current focus of the ongoing studies.

In addition to supplementing Cs-CST waste form characterization [6, 7], ongoing hydrothermal synthesis experiments in the system Cs-Ti-Si-O-H₂O have led to the identification of new materials [8-10]. Herein, we report a novel Cs-silicotitanate ion exchanger phase named SNL-B, with the approximate formula of $\text{Cs}_3\text{TiSi}_3\text{O}_{9.5}\cdot 3\text{H}_2\text{O}$. SNL-B is the second reported ion exchanger Cs-silicotitanate phase, the first being an analogue to the mineral pharmacosiderite ($\text{HCs}_3\text{Ti}_4\text{Si}_3\text{O}_{16}\cdot 4\text{H}_2\text{O}$) [11-13] [*Cs/Si/Ti-pharmacosiderite*]. In fact, the only other previously reported Cs-Si-Ti-O ternary phases are $\text{CsTiSi}_2\text{O}_{6.5}$ (pollucite) and $\text{Cs}_2\text{TiSi}_6\text{O}_{15}$, both of which are condensed tectosilicate phases in which Cs is not exchangeable. These phases were synthesized at PNNL by high temperature, solid-state techniques. [9, 10, 14, 15] More recently, a polymorph of $\text{Cs}_2\text{TiSi}_6\text{O}_{15}$ has been produced by hydrothermal techniques at SNL, designated SNL-A.[8, 16]

Recent research in the silicotitanate field in the interest of new materials for catalysis and ion exchange has led to numerous new phases. These materials consist of framework tetrahedral silicon atoms, framework 4- 5- or 6- coordinate titanium atoms, interstitial or interlayer exchangeable cations, and removable water molecules. Some examples of silicotitanates include Ti- β [17] and TS-1 [18], TS-2 [19], Ti-ZSM-48 [20], Ti-mordenite[17, 21], ETS-4 [22, 23], ETS-10 [24, 25], a Na silicotitanate analogue of the mineral zorite [26], AM-4 [27], lintersite (naturally-occurring) [28], layered $\text{Na}_2\text{TiSi}_2\text{O}_7 \cdot 2\text{H}_2\text{O}$ [29], a sodium silicotitanate analogue of the mineral penkvilksite [30], nafertisite (naturally-occurring) [31], a Cs silicotitanate analogue to pharmacosiderite [11-13], CST (crystalline silicotitanate) [1-5] and JDF-L1 [32, 33]. Reported information on these phases has greatly aided in the research into the Cs-Ti-Si-O-H₂O phases.

The exploratory hydrothermal syntheses of Cs-Ti-Si-O-H₂O phases carried out at SNL are focused on phase identification, phase stability and interphase relationships, based on synthesis parameters including Cs₂O:SiO₂:TiO₂ precursor ratios, pH, temperature and precursor solution concentration. New Cs-Ti-Si-O-H₂O phases obtained from this study are characterized thoroughly with a focus on properties which affect Cs adsorption/desorption. Phase characterization techniques include powder X-ray diffraction, surface area analysis, FTIR, solid-state ¹³³Cs and ²⁹Si MAS NMR, thermal analyses, SEM/EDS, and electron beam irradiation. Additionally, ongoing calorimetric studies with Navrotsky et al., are providing information on thermodynamic stability of both the new phases and their ion exchanged derivatives.[34]

Experimental

Synthesis of SNL-B. Titanium isopropoxide (TIPT) and tetraethylorthosilicate (TEOS), were combined and added dropwise to a mixture containing 50% CsOH solution and amorphous TiO_2 in a 23 ml Teflon liner for a Parr pressure reactor. After stirring for 30 minutes, H_2O was added and the mixture was stirred for 30 minutes more. The final pH of the mixture is 12.5. The final reactant stoichiometry is $Cs:Ti:Si:H_2O = 4:1:4:383$. The amorphous TiO_2 is 1 – 10 atomic % of the total titanium content of the precursor mixture. The loaded pressure reactor was placed in a 120 °C oven for 14 days. The product was collected by filtration and washed with hot water.

Ion Exchange Experiments. Selectivity experiments for sodium and strontium were performed by combining 0.1 g ion exchanger (SNL-B) with 10 g of 50 ppm metal nitrate ($NaNO_3$, $Sr(NO_3)_2$) solution. The ion exchanger was shaken with the solution overnight at 300 rpm. The solution was filtered and analyzed for sodium or strontium concentration by atomic adsorption spectroscopy (AAS). The distribution coefficients (K_d) were calculated in ml/g where K_d is defined as:

$$K_d = [M_{IE} / g IE] / [M_{sln} / ml sln] \quad (1)$$

where M_{IE} is concentration of the metal for exchange (Na or Sr) in ppm which is adsorbed onto the ion exchanger (IE), "g IE" is weight of the ion exchanger (SNL-B) in grams, M_{sln} is concentration of the metal for ion exchange in ppm which remains in solution after contact with SNL-B, and "ml sln" is volume of the solution for ion exchange in milliliters. For maximum loading experiments, the ion exchanger was stirred overnight at 70 °C in a 100-fold excess, 1 molar solution of a metal salt of the exchange ion ($Sr(NO_3)_2$ or $NaNO_3$). The ion exchanged product was collected by vacuum filtration and thoroughly washed with hot water to remove excess salt. The percent of exchange was determined by AAS, and the ion exchange step was repeated until maximum loading was achieved.

General Instrumentation and Characterization Techniques. The powder X-ray diffraction data (XRD) were collected on a Siemens D500 instrument using Cu K_α radiation between $2\theta = 5$ and 60° , a step size of 0.05° and step time of 5 seconds. The powder pattern was indexed for unit cell determination using DICVOL91 software[35]. The ²⁹Si MAS NMR (Magic Angle Spinning Nuclear Magnetic Resonance) spectra were obtained on a Bruker ASX300 at 59.6 MHz using direct single pulse acquisition with 120 second recycle delay, 256 - 512 scans, spinning at 4 kHz, with Q₈M₈ (12.4 ppm) as an external reference. The ²⁹Si CP (cross polarization) MAS NMR spectra were recorded with a 5 millisecond contact time and 10 second recycle delay, and 128 to 256 scans. The ¹³³Cs MAS NMR were recorded on an AMX400 at 52.49 MHz using direct single pulse excitation with 10 second recycle delay, 2 μ s pulse, 512 scans, spinning at 12.5 KHz, with 0.5 M CsCl solution (0.00 ppm) as an external standard. The CP ¹³³Cs experiments were obtained using contact times varying between 1 – 6 ms, 1 s recycle delay for ¹H, and 1024 scans. Elemental analysis for Cs, Na and Sr by atomic adsorption spectroscopy was performed on a Perkin Elmer 5100 PC AAS instrument. Solutions and standards were prepared with 1000 ppm ionization suppressant. An acetylene/air flame was used for Na and Cs analyses and an acetylene/nitrous oxide flame was used for Sr analysis. The DTA-TGA experiments were performed on a STD 2960 TA DTA-TGA instrument with alumina as a standard for DTA. Samples (10 - 15 mg) were heated in air. To obtain the water desorption/readsoption curves, a sample of SNL-B was heated at $10^\circ\text{C}/\text{min}$ up to 500° , cooled to 30° at $1^\circ\text{C}/\text{min}$, and held for 200 minutes. Scanning electron microscopy (SEM) data was collected on a JEOL JSM-T300 SEM with energy dispersive capabilities. Transmission Electron Microscopy (TEM) data was collected on a JEM 2000FX TEM at University of Michigan. Infrared data (FTIR) was obtained from pressed pellet samples containing 1 – 3 wt % sample in a KBr matrix, analyzed on a Perkin Elmer Spectrum GX. Surface areas were determined using a Quantachrome Autosorb with nitrogen as the adsorbant. Samples were outgassed under vacuum at 100°C . Qualitative determination of microporosity was made via inspection of isotherms over the relative pressure range of 10^{-5} and 10^{-2} and qualitative evaluation made using Deboer analysis of t-plots.

Leach Tests. The standard PCT (product consistency test) leach test, a common technique developed to evaluate chemical durability of nuclear waste forms in aqueous environments[36, 37], was performed on SNL-B. A sample of the material (0.2 g) was placed in a hydrothermal reaction vessel with 10 grams of water at 90 °C for 1, 2, 3, 7, and 10 days. After the designated time of heating, each sample was filtered, and the leachate solution was analyzed for Cs concentration by AAS. The solid product was analyzed by XRD to determine if any phase changes occurred as a result of the leach test. Surface area of samples for leach rate calculations were measured, as described above.

Electron Irradiation Studies. Electron irradiation studies of SNL-B were conducted with a JEM 2000FX transmission electron microscope at the University of Michigan. The electron energy used was 200 keV and the sample was irradiated at a dose rate of $\sim 10^{18}$ electrons / s \cdot cm 2 . Temperature dependence of amorphization dose was studied over the range of 20 - 700 °C by using hot stage capabilities of the TEM.

Results and Discussion

Synthesis of SNL-B. The synthesis method described above results in a typical crystallite size of 0.1 x 2 μ m. Procedure requires addition of small amounts of TiO₂ to form the silicotitanate molecular sieve; exclusion produces only *Cs/Si/Ti-pharmacosiderite*. All samples contain approximately 10 volume % of a known impurity phase [*Cs/Si/Ti-pharmacosiderite*]. However, the major phase of all samples is SNL-B (90%) as confirmed by XRD, SEM/EDS and TGA. Consequently, bulk analyses were carried out with caution and consideration of the impurity phases. SNL-B is classified as a microporous molecular sieve based on both ion exchange behaviors and characterization results, including ion exchange capability, reversible hydration/dehydration of internal water molecules, irradiation response, and other spectroscopic characterizations discussed below.

Characterization of SNL-B. Through SEM-EDS and TGA, the formula of SNL-B has been determined to be Cs₃TiSi₃O_{9.5} \cdot 3H₂O. Other techniques for characterization include

XRD, surface area analysis, FTIR, and solid state ^{29}Si and ^{133}Cs MAS NMR. The results of these characterizations are discussed below.

The *Powder X-ray diffraction pattern* of SNL-B is shown in **figure 1**. A thorough search of the JCPDS data base [38] by comparing the three most intense peaks did not provide a match for SNL-B to any known phase. Additionally, a literature search of silicate, aluminosilicate and silicotitanate phases did not provide a structural analogue to this phase. This phase was indexed as an orthorhombic system with cell parameters $a = 10.83$ Å, $b = 7.43$ Å and $c = 7.11$ Å. The major observed and calculated peaks are listed in **Table I**. Crystallites of SNL-B were observed by SEM, as shown in **figure 2**. The crystals exhibit a blade-like habit, $\sim 0.1 \times 2 \mu\text{m}$ in size, and the crystals grow in radiating aggregates. Due to the extremely small size of these crystallites, single crystal XRD structure determination has not yet proven successful.

The formula of SNL-B was determined by careful *EDS-SEM* analysis on single crystals for Cs:Ti:Si ratios, and the water content was determined by thermogravimetric analysis. Bulk elemental analysis using atomic emission and adsorption techniques to determine composition were not utilized, due to the presence of the known impurity phase. Rather, multiple (20) EDS analyses were carried out on single crystals of SNL-B. Known Cs/Si/Ti phases, $\text{Cs}_2\text{TiSi}_6\text{O}_{15}$ [16] and *Cs/Si/Ti pharmacosiderite* [11], were used as standards for the EDS analysis. The Cs:Ti:Si ratio of SNL-B determined by this technique is 3:1:3.

Thermogravimetric data was obtained on a sample, which (**figure 3**) gave SNL-B a 7.5 weight % volatile component (water) which is consistent with the formula $\text{Cs}_3\text{TiSi}_3\text{O}_{9.5} \cdot 3\text{H}_2\text{O}$. The data on SNL-B was obtained by initially removing the hydrous content of the *Cs/Si/Ti pharmacosiderite* impurity phase by heating. *Cs/Si/Ti pharmacosiderite* undergoes irreversible dehydration and structure collapse by 400 °C, whereas SNL-B exhibits thermal stability upon removal of water and up to 500 °C. The impurity content was confirmed to be $\sim 10\%$ by observation of the percentage of irreversible water loss by TGA of the mixed sample. The remaining sample (SNL-B) was

rehydrolyzed and analyzed for water content. The water is volatilized between 100 – 200 °C, and does not result in any structural change in SNL-B. Furthermore, the water is completely readSORBED upon slow cooling, as observed in **figure 3**. Therefore, the water content is present as occluded H₂O molecules rather than hydroxyl anions.

Pore surface area of SNL-B was studied by nitrogen adsorption. A pure, sample of the 3D molecular sieve *Cs/Si/Ti pharmacosiderite* was also analyzed for comparison (contains Cs cations and water molecules within channels). The isotherms for both samples displayed significant adsorption at low relative pressures, qualitatively indicating the presence of micropores. Total surface area calculated for SNL-B and *Cs/Si/Ti-pharmacosiderite* are similar; 89 m²/g and 86 m²/g, respectively. The apparent low BET-determined surface areas of SNL-B and *Cs/Si/Ti-pharmacosiderite* (with 5 Å pores, based on crystallographic data [13]) are likely a result of Cs and H₂O within the microporous void space which hinder accessibility. Therefore, the majority of the quantified porosity is actually external surface area. [However, the combination of experimental results indicate that SNL-B has microporosity. Further indication of microporosity is seen in the ion exchange behavior of SNL-B, which is discussed later.]

FTIR: The mid-IR spectrum of SNL-B is shown in **figure 4**. The spectrum has peaks in three distinct regions; 400 - 750 cm⁻¹(I), 850 – 1200 cm⁻¹ (II), and 2800 – 3500 cm⁻¹ (III). The lower frequency region I corresponds with Si-O-Ti vibrational modes and the Si-O-Si bending modes.[33, 39, 40] The higher frequency region II corresponds with Si-O-Si stretching modes.[40] The large broad peak centered at 3500 cm⁻¹ (region III) is assigned as O-H stretching modes.

Solid-state ²⁹Si and ¹³³Cs MAS NMR and CP spectroscopies provide insight into the bonding environments of framework and interstitial atoms of SNL-B. The ²⁹Si and ¹³³Cs spectra are shown in **figures 5a** and **5b**, respectively.

The ²⁹Si MAS NMR chemical shifts are consistent with other reported SiO₄ tetrahedral shifts for silicotitanates.[41-43] The ²⁹Si chemical shifts are assigned Q^x designations, for

which x is the number of Si in the second coordination sphere of the silicon tetrahedron, and 4-x is the number of non-silicon atoms (Ti). The peak at -87 ppm is the *Cs/Si/Ti pharmacosiderite* impurity phase. The broadness of this peak in conjunction with the broad peaks observed in the diffraction pattern indicate poor crystallinity. By comparing the two remaining peaks of SNL-B to those of structurally characterized silicotitanates, the following assignments are made: -97 ppm is a Q² or Q³ peak and -105 ppm is a Q³ or Q⁴ peak. The corollary **cross polarization (CP)** experiments gave only the peak at -97 ppm which suggests this silicon site is in closer proximity to the internal H₂O.

The ¹³³Cs **MAS NMR** spectrum shows that SNL-B has two unique exchange sites occupied by Cs. The peak at 57 ppm is the *Cs/Si/Ti pharmacosiderite* impurity phase. The corollary ¹³³Cs **CP-MAS** experiment gave the peak at 8 ppm, which suggests the Cs in this site is more hydrated than the Cs in the second site (23 ppm). The resolution of the peaks at room temperature suggests that there is no rapid exchange between the Cs sites, as is reported for some Cs containing zeolites.[44] Upfield shifts of Cs peaks in general correspond with "less shielding".[44, 45] Shielding is increased by more cations per cage, smaller cage size, more waters per cage, or increased donation of electrons from the framework oxygen atoms to the unoccupied orbitals of Cs.[44, 45]

Ion Exchange and Selectivity of SNL-B. Ion exchange and selectivity experiments were performed to determine if: 1) SNL-B can exchange its Cs cations and 2) SNL-B demonstrates any selectivity for specific cations. In fact, greater than 90 atomic % of the Cs of SNL-B can be exchanged for Na or Sr by repeated contact with hot solutions containing an excess of the exchange cation. This exchange capacity is comparable to that of *Cs/Si/Ti pharmacosiderite*.[11] The ion exchanges of SNL-B take place without any framework distortion, as indicated by the X-ray diffraction spectra of ion exchanged SNL-B. No peak shifts are observed, only relative peak intensity changes.

Additionally, SNL-B exhibits high selectivity for Sr with $K_d > 100,000$ ml/g (i.e. no Sr detected in solution after contact with SNL-B). On the other hand, a $K_d = 19$ ml/g was measured for adsorption of Na onto SNL-B. In these experiments, the selectivity behavior

of the *Cs/Si/Ti pharmacosiderite* impurity phase needs to be taken into account. For comparable experiments, K_d values for Sr and Na on *Cs/Si/Ti pharmacosiderite* of 1800 and 30 ml/g respectively, have been reported.[11] Given the K_d for SNL-B is much greater than that reported for *Cs/Si/Ti pharmacosiderite*, it can be assumed the Sr selectively exhibited by our sample is due to the major SNL-B phase rather than the minor *Cs/Si/Ti pharmacosiderite* impurity phase. However, both SNL-B and *Cs/Si/Ti pharmacosiderite* can arguably contribute to the Na K_d value. Therefore, the K_d value of 19 ml/g of Na on SNL-B is a maximum distribution coefficient. In either case, SNL-B definitely shows extremely good selectivity for Sr and extremely poor selectivity for Na. Ongoing selectivity experiments are underway to determine if this observed disparity in selectivity is due to the preference of SNL-B for divalent cations over monovalent cations.

Chemical durability of SNL-B. SNL-B is metastable relative to the other hydrothermally synthesized *Cs/Si/Ti* ternary oxides, SNL-A and *Cs/Si/Ti pharmacosiderite*. Alteration to one of these phase is observed as a result of extensive hydrothermal treatment of SNL-B. Note also, partial alteration to *Cs/Si/Ti pharmacosiderite* was also observed for the ion exchange experiments in which SNL-B was treated with hot (80 – 90 °C) aqueous solutions.

A standard PCT leach test was carried out to determine the leachability of Cs from SNL-B. The leach rates are plotted in **figure 6**, based on a BET determined external surface area of 89 m²/g. The leach rates for SNL-B are approximately 6 orders of magnitude higher than those observed for *Cs/Si/Ti-pharmacosiderite*, which has a similar external surface area of 86 m²/g. [8] Analysis of the post-leached samples by XRD reveals relative peak intensities of SNL-B are diminished, and broad peaks of *Cs/Si/Ti-pharmacosiderite* are apparent. **Figure 7** shows TEM images of SNL-B before and after leach tests. The leached SNL-B has a “mottled” appearance and a fine precipitate is observable, consistent with high leach rates and a dissolution/reprecipitation texture.

From the synthesis precursor mixture described, SNL-B is the stable phase formed at 120 °C for up to 28 days of heating; at which time SNL-A ($Cs_2TiSi_6O_{15}$) is formed, which is

extremely durable and stable.[16] Alternatively, a heating temperature of 170 °C results in SNL-B after 3 - 4 days, and alteration to the condensed phase (SNL-A) occurs between 5 - 7 days. This transformation is shown in the X-ray diffraction pattern in **figure 8**, for 2.5, 7 and 10 days heating time at 170 °C.

Electron Irradiation Studies. Radiation stability of all Cs-Ti-Si-O-H₂O phases synthesized in this study is important in order to assess the durability of the phases for ion exchange with radionuclides or radioactive waste form applications. Electron irradiation with *in situ* TEM studies have proven to be useful in simulating the effects of ionizing radiation damage caused by fission products.[46, 47] The results from our studies indicate that SNL-B behaves as traditional zeolites.[46, 47]

The amorphization dose due to electron irradiation of SNL-B was studied as a function of temperature. The electron dose on the sample was determined by:

$$Dose = i \times M^2 \times t \text{ (electrons/cm}^2\text{)} \quad (2)$$

where i is the screen current density measured through an area where no sample is present to obscure the electron beam, M is magnification of the microscope and t is the time for the sample to achieve full amorphization. The irradiation-induced transformation from the crystalline-to-amorphous material is observed by the fading of the Bragg-diffraction spots in the electron diffraction pattern (**figure 9**). The dose for amorphization at room temperature is 2×10^{20} electrons/cm². The temperature dependence of irradiation-induced amorphization is shown in **figure 10**. The dose needed for amorphization is constant up to 500 °C, and then decreases sharply as a result of thermally-induced amorphization. This observation also agrees with differential thermal analysis (DTA) of SNL-B for which an exothermic phase change occurs at 550 °C. Analcime, an aluminosilicate zeolite was reported to exhibit similar response to electron irradiation upon heating, as a result of thermally induced structure collapse.[46, 47]

Summary and Conclusions

In the course of our ongoing studies of ternary phases in the system Cs-Ti-Si-O-H₂O, we have produced Cs₃TiSi₃O_{9.5}•3H₂O (SNL-B) a novel molecular sieve ion exchanger. In addition to a *Cs/Si/Ti-pharmacosiderite* analogue (HCs₃Ti₄Si₃O₁₆•4H₂O), SNL-B is the only other reported molecular sieve silicotitanate which has been synthesized with Cs as its charge balancing cation. Spectroscopic characterizations of SNL-B show a complex phase with multiple silicon coordination sites (2) and Cs cation exchange sites (2). SNL-B exhibits a very high leach rate for Cs as a result of its metastability. Additionally, SNL-B is not stable under electron irradiation; it behaves in a fashion similar to other zeolitic materials when exposed to electron irradiation. SNL-B also undergoes alteration to a condensed leach resistant Cs/Si/Ti phase, SNL-A that is the topic of forthcoming publication.[16] Ongoing work includes crystal growth experiments to obtain structural data on this unique phase.

Acknowledgements

This work was supported by the U.S. DOE under contract DE-AC04-94AL85000 (Sandia National Laboratories).

All authors thank the DOE Environmental Management Science Program (EMSP) for funding for this work under separate grants for University of Michigan and Sandia National Laboratories (EMSP grant DE-FG07-97ER45652 at University of Michigan, project # 27601 at Sandia National Laboratories).

We thank Dr. Todd Alam at Sandia National Laboratories for the NMR spectroscopy experiments and Steve Thoma for the surface area analyses.

Sandia is a multiprogram laboratory operated by Sandia Corporation, a Lockheed Martin Company, for the United States Department of Energy under contract DE-AC04-94AL85000.

References

- [1] R. G. Anthony, C. V. Phillip, R. G. Dosch, Waste Management 13 (1993) 503.
- [2] R. G. Anthony, R. G. Dosch, D. Gu, C. V. Philip, Ind. Eng. Chem. Res. 33 (1994) 2702.
- [3] D. M. Poojary, R. A. Cahill, A. Clearfield, Chem. Mater. 6 (1994) 2364.
- [4] T. M. Nenoff, S. G. Thoma, J. L. Krumhansl, "The Stability and Selectivity of TAM5: A Silicotitanate Molecular Sieve for Radwaste Clean-up", SAND96-2580, Sandia National Laboratories, Albuquerque, November, 1996.
- [5] T. M. Nenoff, J. L. Krumhansl, "The Structure of TAM5: A Silicotitanate Molecular Sieve for Radwaste Clean-up", SAND96-2578, Sandia National Laboratories, Albuquerque, November 1996.
- [6] Y. Su, M. L. Balmer, L. Wang, B. C. Bunker, T. M. Nenoff, M. Nyman, A. Navrotsky, MRS Conference Proceedings (Boston), "Evaluation of Thermally Converted Silicotitanate Waste Forms (1998).
- [7] Y. Su, M. L. Balmer, B. C. Bunker, MRS Conference Proceedings (Boston), "Evaluation of Cesium Silicotitanates as an Alternative Waste Form", p. 457 (1997).
- [8] M. Nyman, T. M. Nenoff, Y. Su, M. L. Balmer, A. Navrotsky, H. Xu, MRS Conference Proceedings (Boston), "New Crystalline Silicotitanate (CST) Waste Forms: Hydrothermal Synthesis and Characterization of Cs-Si-Ti-O Phases (1998).

- [9] M. L. Balmer, Y. Su, I. E. Grey, A. Santoro, R. S. Roth, Q. Huang, N. Hess, B. C. Bunker, MRS Conference Proceedings, "The Structure and Properties of Two New Silicotitanate Zeolites", p. 449 (1997).
- [10] M. L. Balmer, Q. Huang, W. Wong-Ng, R. S. Roth, A. Santoro, J. Solid State Chem. 130 (1997) 97.
- [11] E. A. Behrens, D. M. Poojary, A. Clearfield, Chem. Mater. 8 (1996) 1236.
- [12] E. A. Behrens, A. Clearfield, Microporous Mater. 11 (1997) 65.
- [13] W. T. A. Harrison, T. E. Gier, G. D. Stucky, Zeolites 15 (1995) 408.
- [14] D. E. McCready, M. L. Balmer, K. D. Keefer, Powder Diffr. 12 (1997) 40.
- [15] I. E. Grey, R. S. Roth, M. L. Balmer, J. Solid State Chem. 131 (1997) 38.
- [16] M. Nyman, F. Bonhomme, D. Teter, T. M. Nenoff, R. C. Ewing, L. M. Wang, B. X. Gu, Chem. Mater. submitted (2000).
- [17] X. Wang, X. Guo, Catal. Today 51 (1999) 177.
- [18] C. B. Dartt, M. E. Davis, Appl. Catal. A 143 (1996) 53.
- [19] J. S. Reddy, R. Kumar, P. Ratnasamy, Appl. Catal. 58 (1990) L1.
- [20] D. P. Serrano, H. X. Li, M. E. Davis, J. Chem. Soc., Chem. Commun. (1992) 745.
- [21] B. Kraushaar-Czernezki, J. H. C. v. Hooff, Catal. Lett. 2 (1989) 43.
- [22] S. M. Kuznicki, Patent # 4,853,202, U.S.A. 1989.
- [23] X. Liu, J. K. Thomas, J. Chem. Soc., Chem. Commun. (1996) 1435.
- [24] S. M. Kuznicki, Patent # 4,938,939, U.S.A. 1990.
- [25] T. K. Das, A. J. Chandwadkar, S. Sivasanker, J. Chem. Soc., Chem. Commun. (1996) 1105.
- [26] D. M. Chapman, A. L. Roe, Zeolites 10 (1990) 730.
- [27] M. S. Dadachov, J. Rocha, A. Ferreira, Z. Lin, M. W. Anderson, J. Chem. Soc., Chem. Commun. (1997) 2371.
- [28] S. Merlino, M. Pasero, A. Khomyakov, Zeit. Kristallog. 193 (1990) 137.
- [29] A. Clearfield, A. I. Bortun, L. N. Bortun, R. A. Cahill, Solvent Extr. Ion Exch. 15 (1997) 285.
- [30] Y. Liu, H. Du, F. Zhou, W. Pang, J. Chem. Soc., Chem. Commun. (1997) 1467.

- [31] G. Ferraris, G. Ivaldi, A. P. Khomyakov, S. V. Soboleva, E. Belluso, A. Pavese, European J. Miner. 8 (1996) 241.
- [32] M. A. Roberts, G. Sankar, J. M. Thomas, R. H. Jones, H. Du, J. Chen, W. Pang, R. Su, Nature 381 (1996) 401.
- [33] H. Du, M. Fang, J. Chen, W. Pang, J. Mater. Chem. 6 (1996) 1827.
- [34] H. Xu, A. Navrotsky, M. Nyman, T. M. Nenoff, M. L. Balmer, Y. Su, J. Mater. Chem. in press (2000).
- [35] A. Boultif, D. Louer, J. Appl. Crystallography 24 (1991) 987.
- [36] M. G. Mesko, D. E. Day, J. Nucl. Mater. 273 (1999) 27.
- [37] H. Li, M. Tomozawa, J. non-Crystalline Sol. 195 (1996) 188.
- [38] JCPDS Database, Swarthmore 1992.
- [39] H. Du, F. Zhou, W. Pang, Y. Yue, Microporous Mater. 7 (1996) 73.
- [40] P. Launer, Infrared analysis of organosilicon compounds, Petrarch systems 1987.
- [41] M. L. Balmer, B. C. Bunker, L. Q. Wang, C. H. F. Peden, Y. Su, J. Phys. Chem. B 101 (1997) 9170.
- [42] A. Labouriau, T. J. Higley, W. L. Earl, J. Phys. Chem. 102 (1998) 2897.
- [43] P. J. Dirken, M. E. Smith, H. J. Whitfield, J. Phys. Chem. 99 (1995) 395.
- [44] T. Tokuhiro, M. Mattingly, L. E. Iton, M. K. Ahn, J. Phys. Chem. 93 (1989) 5584.
- [45] F. Yagi, N. Kanuka, H. Tsuji, S. Nakata, H. Kita, H. Hattori, Microporous Mater. 9 (1997) 229.
- [46] L. M. Wang, S. X. Wang, R. C. Ewing, 9th Annual International High-Level Radioactive Waste Management Conference (Las Vegas, NV), "Radiation Effects in Zeolites: Relevance to Near-Field Contamination", p. 772 (1997).
- [47] B. X. Gu, L. M. Wang, R. C. Ewing, J. Nucl. Mater. 278 (2000) 64.

Table I. Calculated and Observed Peaks of SNL-B

h	k	l	d_{obs} (Å)	d_{calc} (Å)
1	0	0	10.80	10.83
0	1	0	7.42	7.43
2	0	0	5.41	5.42
0	1	1	5.14	5.14
2	1	0	4.38	4.38
2	1	1	3.73	3.73
0	0	2	3.56	3.56
1	2	0	3.52	3.52
0	1	2	3.21	3.21
2	2	0	3.07	3.07
2	0	2	2.98	2.98
0	2	2	2.57	2.57

Figure Captions

Figure 1. Powder X-ray diffraction patterns of SNL-B (top) and the major peak positions indexed (bottom).

Figure 2. Scanning electron micrographs showing the radiating, blade-like habit of SNL-B crystallites.

Figure 3. Thermogravimetric analysis of SNL-B showing desorption and subsequent adsorption cycle.

Figure 4. Infrared spectrum of SNL-B.

Figure 5. Solid-state MAS NMR spectra of SNL-B: a) ^{29}Si spectrum (top) b) ^{133}Cs spectrum (bottom).

Figure 6. a) PCT leach rates of Cs from SNL-B as a function of time (0.1 g sample heated in 10 ml H_2O at 90 °C).

Figure 7. a) TEM bright field image of SNL-B, prior to leaching. b) TEM bright field image of SNL-B, post-leaching.

Figure 8. X-ray diffraction spectra showing hydrothermal transformation of SNL-B to SNL-A, Cs-silicotitanate condensed phase.[16]

Figure 9. Sequence of selected-area diffraction patterns of SNL-B during 200 KeV electron beam irradiation: A. nonirradiated. B. intermediate state, 1.2×10^{20} electrons/cm 2 irradiation dose. C. fully amorphized, 2.0×10^{20} electrons/cm 2 irradiation dose.

Figure 10. Temperature dependence of amorphization dose for SNL-B under 200 KeV electron irradiation.



















