

**Monte Carlo Simulations of Temperature-Programmed and
Isothermal Desorption from Single-Crystal Surfaces**

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ABSTRACT

The kinetics of temperature-programmed and isothermal desorption have been simulated with a Monte Carlo model. Included in the model are the elementary steps of adsorption, surface diffusion, and desorption. Interactions between adsorbates and the metal as well as interactions between the adsorbates are taken into account with the Bond-Order-Conservation-Morse-Potential method.

The shape, number, and location of the TPD peaks predicted by the simulations is shown to be sensitive to the binding energy, coverage, and coordination of the adsorbates. In addition, the occurrence of lateral interactions between adsorbates is seen to strongly effect the distribution of adsorbates on the surface.

Temperature-programmed desorption spectra of a single type of adsorbate have been simulated for the following adsorbate-metal systems: CO on Pd(100); H₂ on Mo(100); and H₂ on Ni(111). The model predictions are in good agreement with experimental observation. TPD spectra have also been simulated for two species coadsorbed on a surface; the model predictions are in qualitative

agreement with the experimental results for H₂ coadsorbed with strongly bound atomic species on Mo(100) and Fe(100) surfaces as well as for CO and H₂ coadsorbed on Ni(100) and Rh(100) surfaces. Finally, the desorption kinetics of CO from Pd(100) and Ni(100) in the presence of gas-phase CO have been examined. The effect of pressure is seen to lead to an increase in the rate of desorption relative to the rate observed in the absence of gas-phase CO. This increase arises as a consequence of higher coverages and therefore stronger lateral interactions between the adsorbed CO molecules.

Robert F. Bell

To

Mariette,

Michael,

Mom,

and Dad

Table of Contents

Chapter I	Introduction.....	1
Chapter II	A Review of Theoretical Models of Adsorption, Diffusion, Desorption, and Reaction of Gases on Metal Surfaces.....	3
Chapter III	A Monte Carlo Model for the Simulation of Temperature-Programmed Desorption Spectra.....	158
Chapter IV	Monte Carlo Simulation of the Temperature- Programmed Desorption of Coadsorbed Species.....	201
Chapter V	Monte Carlo Simulation of the Effect of Pressure on Isothermal and Temperature-Programmed Desorption Kinetics.....	246
Appendix	FORTRAN Programs for Simulating Temperature- Programmed and Isothermal Desorption.....	278

List of Figures

Chapter II

Fig. 2.1 a. Potential energy surface for molecular adsorption as a function of orientation angle θ and distance from the surface z . The energy contours are in eV.
 b. One-dimensional potential versus reaction coordinate s [36]..... 39

Fig. 2.2 a. Potential energy surface for dissociative adsorption as a function of distance from the surface z and internal bond distance r . The energy contours are in eV. b. One-dimensional potential versus reaction coordinate s [36]..... 40

Fig. 2.3 Variation of $S(1)(\theta)/S(1)(0)$ with θ for adsorption via a precursor mechanism. The precursor parameter $K = k_a^*/k_d^*$ 41

Fig. 2.4 a. Variation in f_d' with θ at $T = 300$ K [26]. The nomenclature (5,6) corresponds to $W_{AA} = 5$ kJ/mol and $z = 6$, respectively. b. Variation in f_d'' with θ [24]. The nomenclature (2,6) corresponds to $W_{AA} = 2k_bT$ and $z = 6$, respectively..... 42

Fig. 2.5 Variation of $v_d(1)(\theta)/v_d(1)(0)$ and $\Delta E_d(\theta)$ with θ for repulsive (top panels) and attractive (bottom panels) interactions [26]. Solid line: $|W_{AA}| = 5$ kJ/mol, $z = 6$; dashed line: $|W_{AA}| = 5$ kJ/mol, $z = 4$; dotted line: $|W_{AA}| = 3$ kJ/mol, $z = 6$ 43

Fig. 2.6 Variation in f_r with θ [25]. The nomenclature (1,0,3) corresponds to $W_{AA} = k_bT$, $W_{AB} = 0$, and $W_{BB} = 3k_bT$, respectively..... 44

Fig. 3.1 Simulated (O) and experimental (\blacktriangle) variation of $S(1)(0)$ versus $E_T \cos \theta_i$ for Xe adsorption on Pt(111) [46]. E_T is the incident kinetic energy and θ_i is the incident

angle measured from the surface normal.....	56
Fig. 3.2 Simulated (●) variation of $S(2)(0)$ versus $E_T \cos^2 \theta$ for H_2 adsorption on $Cu(100)$ [31]. E_T is the incident kinetic energy and θ is the incident angle measured from the surface normal. The solid line is a guide to the eye.....	57
Fig. 3.3 Simulated (solid line) [53] and experimental (circles) [54] variation of $S(2)(0)$ versus incident kinetic energy for N_2 adsorption on $W(110)$	58
Fig. 3.4 a. Potential of mean force versus height above the surface for CO on $Ni(111)$. b. Simulated variation of $S(1)(0)$ versus T for CO adsorption on $Ni(111)$ [36].....	59
Fig. 3.5 Simulated variation of $S(2)(0)$ versus incident kinetic energy for H_2 adsorption on $Ni(100)$ [42].....	60
Fig. 3.6 Variation in $S(2)(\theta)/S(2)(0)$ versus θ for repulsive nearest-neighbor interactions [59].....	61
Fig. 3.7 Variation in $S(1)(\theta)/S(1)(0)$ versus θ when nearest-neighbor interactions occur [60].....	62
Fig. 3.8 a. Variation in $S(2)(\theta)/S(2)(0)$ versus θ when precursor states exist for a fixed value of η . b. Variation in $S(2)(\theta)/S(2)(0)$ versus θ when precursor states exist for a fixed value of L [59].....	63
Fig. 3.9 Experimental and predicted variation in $S(2)(\theta)$ versus coverage for N_2 adsorption on $W(100)$ [59].....	64
Fig. 3.10 Experimental (error bars) [62,63] and predicted (●) [61] variation in $S(1)(\theta)$ versus coverage for N_2 adsorption on $Ru(001)$	65
Fig. 4.1 Experimental [69] and predicted [79] variation in D versus $1/T$ for Rh diffusion on $Rh(100)$	77

Fig. 4.2 Variation in D with 1/T for hydrogen diffusion on W(110). Symbols are the experimental data [90,91] and the lines are the model predictions [89].....	78
Fig. 4.3 a. Predicted variation in D with θ for hydrogen diffusion on a metal surface at 27 K [92]. b. Experimental variation in D with θ for hydrogen diffusion on W(110) at 27 K [91].....	79
Fig. 4.4 Variation in $D(\theta)/D(0)$ with θ for repulsive interactions for $WAA = k_b T^*$, where T^* is the standard temperature [93].....	80
Fig. 4.5 Variation in $D(\theta)/D(0)$ with θ for nearest-(w_1) and next-nearest-(w_2) neighbor interactions [94].....	81
Fig. 4.6 Variation in D with $WAA/k_b T$ [98]. The open circles are from simulations using eq. 4.3 and the closed circles are from simulations using eq. 4.5. The crosses are the rescaled diffusivities from eq. 4.5.....	82
Fig. 4.7 Variation of D with θ for repulsive (■), attractive (▲), and no nearest-neighbor interactions (●) [99].....	83
Fig. 4.8 a. Simulation of the θ dependence of D for repulsive nearest-neighbor and attractive next-nearest neighbor interactions [100]. b. Experimental variation in D with θ for oxygen on W(110) [101].....	84
Fig. 4.9 Variation in $D(\theta)/D(0)$ with θ for equivalent nearest-(w_1) and next-nearest-(w_2) neighbor repulsive interactions [104]. a. Disordered region. b. Ordered region.....	85
Fig. 4.10 a. Variation in $D(\theta_A)/D(0)$ with θ_A in the presence of an immobile adsorbate B for $\theta_B=0$ (□), $\theta_B=0.15$ (○), and $\theta_B=0.36$ (△). b. Variation in $D(\theta_A)/D(0)$ with θ_B .	

The filled symbols represent upper limits [105].....	86
Fig. 5.1 Arrhenius plot for Ar and Xe desorption from Pt(111). Solid and dashed curves are dynamically corrected and uncorrected TST, respectively [29].....	106
Fig. 5.2 Activation energy (a) and preexponential factor (b) profiles versus T for CO desorption from Ni(111) [36]....	107
Fig. 5.3 a. Predicted CO TPD spectra from a W ribbon [124]. b. Experimental CO TPD spectra from a W ribbon [125]. The CO exposures (Pa s) are: a) 1.9×10^{-5} ; b) 6.0×10^{-5} ; c) 1.6×10^{-4} ; d) 1.9×10^{-4} ; e) 2.5×10^{-4} ; and f) 3.1×10^{-3} . c. Predicted activation energy profile versus coverage [124].....	108
Fig. 5.4 Predicted [126] and experimental [127] CO TPD spectra from Ir(110) for $\theta_{CO} = 0.82$	109
Fig. 5.5 Predicted [126] and experimental [128] O ₂ TPD spectra from Ir(110). Curves in both panels correspond to oxygen coverages of 1.0, 0.75, 0.50, and 0.25.....	110
Fig. 5.6 Variation in $\Delta E_d(\theta)$ and $v_d(1)(\theta)/v_d(1)(0)$ versus coverage when surface reconstruction occurs [60].....	111
Fig. 5.7 Comparison of Monte Carlo and continuum models for the simulation of non-associative TPD spectra in the absence of lateral interactions [132].....	112
Fig. 5.8 a. Monte Carlo simulation of TPD spectra [133]. b. Continuum simulation of TPD spectra [134]. Lateral interactions between adsorbates in the continuum simulations are described with the Braggs-Willaims approximation.....	113
Fig. 5.9 Simulation [135] of TPD spectra for both nearest- neighbor interactions w_1 and next-nearest neighbor interactions w_2 . a. Continuum model. b. Monte Carlo	

model. The lateral interactions are in kcal/mol.....	114
Fig. 5.10 Experimental (a) [62,63] and simulated (b) [61] TPD spectra of molecular N ₂ from Ru(001).....	115
Fig. 5.11 Simulated (a) [132] and experimental (b) [136] TPD spectra of H ₂ from Mo(100).....	116
Fig. 5.12 Simulated (a) [132] and experimental (b) [137] variation in E _d with θ _H for H ₂ desorption from Mo(100).....	117
Fig. 5.13 a. Simulated TPD spectra of B and A ₂ [139]. b. Experimental TPD spectra of CO and H ₂ [142]. The dashed curves correspond to desorption of each species adsorbed separately, and the solid curves correspond to desorption of both species coadsorbed.....	118
Fig. 5.14 Variation in E _d with θ _i for the simulations in fig. 5.13a [139]. The dashed curves correspond to desorption of each species adsorbed separately, and the solid curves correspond to desorption of both species coadsorbed.....	119
Fig. 6.1 Theoretical [126] and experimental [160] TPSR spectra for CO oxidation and CO desorption on Ir(110).....	133
Fig. 6.2 a. Predicted variation of E _d versus θ for CO desorption from Ir(110) [126]. b. Predicted variation of E _r versus θ for CO oxidation [126].....	134
Fig. 6.3 TPSR spectra for CO oxidation on Ir(111). a. Experimental spectra as a function of gas exposure [163]. b. Theoretical spectra as a function of coverage [161].....	135
Fig. 6.4 Steady-state rate of CO oxidation versus T on Ir(111). Solid lines are theoretical calculations [162]; circles are experimental data in arbitrary units [163]; and dashed line is experimental data [164]. a. P _{CO} = P _{O₂} = 1.1×10 ⁻⁷ Torr. b. P _{CO} = 5×10 ⁻⁸ Torr and	

$P_o = 1.2 \times 10^{-7}$ Torr.....	136
Fig. 6.5 Steady-state rate of hydrogen oxidation versus T on Pt(111) for $P_H = 1.34 \times 10^{-5}$ Pa and $P_o = 6.7 \times 10^{-6}$ Pa. Solid line is the theoretical calculation [165] and the dashed line is the experimental result [166].....	137
Fig. 6.6 Rate of CO_2 production versus t for the reaction of CO with preadsorbed oxygen for $P_{CO} = 1.0 \times 10^{-8}$ Torr and $\theta_O = 2/3$ monolayer [175]. Solid lines are the experimental results and the dashed lines are the model predictions.....	138
Fig. 6.7 Concentration profiles versus t for CO oxidation for an inlet step change from 2% CO to a 1% CO and 3% O_2 mixture [183]. Circles are the experimental data and the solid line is the model prediction.....	139

Chapter III

Fig. 1 Comparison of Monte Carlo and continuum models for the simulation of first order desorption where E_d is constant. (See Table 1 for parameter values).....	171
Fig. 2 Comparison of Monte Carlo and continuum models for the simulation of second order desorption where E_d is constant. (See Table 1 for parameter values).....	172
Fig. 3 Illustrations of the effects of M-A and M-A plus A-A interactions on the TPD spectra for second order desorption. (See Table 1 for parameter values).....	174
Fig. 4 Illustrations of the effects of M-A and M-A plus A-A interactions on the average E_d for second order desorption.....	175
Fig. 5 Illustrations of the effects of M-A and M-A plus A-A interactions on $f(CN+1)$ for second order desorption.....	178

Fig. 6	Monte Carlo simulations of CO desorption from Pd(100). (See Table 1 for parameter values).....	181
Fig. 7	TPD spectra for CO desorption from Pd(100) [23].....	182
Fig. 8	Variation of the average E_d with θ_{CO} for CO desorption from Pd(100).....	183
Fig. 9	TPD spectra for H ₂ desorption from Mo(100) [26].....	185
Fig. 10	Monte Carlo simulations of H ₂ desorption from Mo(100). (See Table 1 for parameter values).....	187
Fig. 11	Variation of the average E_d with θ_H for H ₂ desorption from Mo(100).....	188
Fig. 12	TPD spectra for H ₂ desorption from Ni(111). Panel a is from ref. [37] and panel b is from ref. [36].....	190
Fig. 13	Monte Carlo simulations of H ₂ desorption from Ni(111). (See Table 1 for parameter values).....	192
Fig. 14	Variation of the average E_d with θ_H for H ₂ desorption from Ni(111).....	193

Chapter IV

Fig. 1	a. Simulation of A ₂ desorption. See table 1 for model parameters. b. Variation of the average E_d with θ_A	215
Fig. 2	Surface map for A(●) at $\theta_A=1.0$ for the simulation in fig. 1a. The open circles are metal atoms.....	216
Fig. 3	a. Simulation of A ₂ desorption with $\theta_B=0.25$. Curve 1 is for $Q_{B,n}=140$ kcal/mol and curve 2 is for $Q_{B,n}=190$ kcal/mol. See table 1 for model parameters. b. Variation of the average E_d with θ_A	218
Fig. 4	Surface map for A(●) and B(■) at $\theta_A=1.0$ and $\theta_B=0.25$	

for the simulations in fig. 3a. The open circles are metal atoms.....	219
Fig. 5 a. Simulation of A ₂ desorption with θ _B =0.50. Curve 1 is for Q _{B,n} =140 kcal/mol and curve 2 is for Q _{B,n} =190 kcal/mol. See table 1 for model parameters. b. Variation of the average E _d with θ _A	220
Fig. 6 a. Simulation of A ₂ desorption with θ _B =0.25 when attractive A-B interactions occur. Curve 1 is for Q _{B,n} =140 kcal/mol and curve 2 is for Q _{B,n} =190 kcal/mol. See table 1 for model parameters. b. Variation of the average E _d with θ _A	222
Fig. 7 a. Simulation of A ₂ desorption with θ _B =0.50 when attractive A-B interactions occur. Curve 1 is for Q _{B,n} =140 kcal/mol and curve 2 is for Q _{B,n} =190 kcal/mol. See table 1 for model parameters. b. Variation of the average E _d with θ _A	223
Fig. 8 Surface map for A(●) and B(■) at θ _A =1.0 and θ _B =0.25 for the simulation in fig. 6a. The open circles are metal atoms.....	224
Fig. 9 a. Simulation of pure component desorption for A ₂ and B. Curve 1 is for Q _{A,n} =55 kcal/mol and curve 2 is for Q _{A,n} =52 kcal/mol. Curve 3 is for bridge-bonded species B with Q _{B,n} =33 kcal/mol. Curve 4 is for on-top bonded species B with Q _{B,n} =33 kcal/mol. See table 2 for model parameters. b. Variation of the average E _d with coverage.....	229
Fig. 10 a. Simulation of A ₂ and B codesorption for case I. See table 2 for model parameters. b. Variation of the average E _d with coverage.....	231
Fig. 11 Surface map for A(■) and B(●) at θ _A =0.51 and θ _B =0.32 for the simulation in fig. 10a. Open circles are metal atoms.....	233

Fig. 12	a. Simulation of A ₂ and B codesorption for case II. See table 2 for model parameters. b. Variation of the average E _d with coverage.....	234
Fig. 13	a. Simulation of A ₂ and B codesorption for case III. See table 2 for model parameters. b. Variation of the average E _d with coverage.....	236
Fig. 14	Surface map for A(■) and B(●) at $\theta_A=0.52$ and $\theta_B=0.50$ for the simulation in fig. 13a. The open circles are metal atoms.....	237
Fig. 15	Experimentally observed TPD spectra for H ₂ and CO desorption from Ni(100) [18]. The dashed lines are for desorption of each species adsorbed separately, and the solid lines are for desorption of both species coadsorbed.....	239

Chapter V

Fig. 1	Variation in CO coverage versus time on polycrystalline Pd at $T_s = 380$ K (taken from Yamada et al. [6]). The open circles represent θ_T and the filled symbols represent the decay of $\theta(C^{18}O)$ for $t > t_1$. a) The isotope switch was made at a) $t_1 = 20$ s (●), 40 s (■), 60 s (◆), and 120 s (▼); b) at $t_1 = 10$ s (●), 20 s (■), 40 s (◆), 60 s (▼), and 120 s (▲); c) at $t_1 = 5$ s (●), 10 s (■), 20 s (◆), and 60 s (▼); d) at $t_1 = 5$ s (●), 10 s (■), 20 s (◆), and 60 s (▼).....	255
Fig. 2	Variation of $k_d P$ versus P. The points designated by (●) are from the analysis of experimental data presented by Yamada et al. [5] and the points designated by (■) are the values obtained from the Monte Carlo simulations.....	257

Fig. 3 Simulated variation in CO coverage versus time on Pd(100) at $T_s = 380$ K. The open circles represent θ_T and the filled symbols represent the decay of $\theta(C^{18}O)$ for $t > t_1$. a) The isotope switch was made at $\theta(C^{18}O) = 0.30$ (●) and 0.60 (◆). b-d) The isotope switch was made at $\theta(C^{18}O) = 0.30$ (●), 0.60 (◆), and 0.85 (▼)..... 260

Fig. 4 Variation in the average value of E_d as a function of θ_T . Curve a is for the $C^{18}O$ molecules which desorb and curve b is for all of the $C^{18}O$ molecules on the surface. The results shown are for simulation presented in fig. 3c for the case where the isotope switch is made at $\theta(C^{18}O) = 0.30$ 262

Fig. 5 Variation in CO coverage versus time at $T_s = 380$ K based on Zhdanov's model [30]. The solid curves represent the numerical solutions of eqs. 10-13 and the symbols represent the results of Monte Carlo simulation using eq. 12 to describe $E_d(\theta)$ 264

Fig. 6 TPD spectra of $C^{18}O$ desorption from Ni(100) reported by Yates and Goodman [4]. a) Desorption in the absence of gas-phase $C^{16}O$. For curve a, the $C^{18}O$ exposure is 0.4 L; for curve b, 0.8 L; for curve c, 1.5 L; for curve d, 3.0 L; and for curve e, 8.0 L. b) Desorption of $C^{18}O$ (4.0 L dose) in the presence of gas-phase $C^{16}O$. For curve a, the $C^{16}O$ flux is 0 $cm^2 s^{-1}$; for curve b, $4.1 \times 10^{14} cm^2 s^{-1}$; for curve c, $8.3 \times 10^{14} cm^2 s^{-1}$; and for curve d, $14.5 \times 10^{14} cm^2 s^{-1}$ 267

Fig. 7 Simulated TPD spectra of C¹⁸O desorption from Ni(100) . a) Desorption in the absence of gas-phase C¹⁶O. For curve a, $\theta^*(C^{18}O) = 0.60$; for curve b, 0.55; and for curve c, 0.25; b) Desorption of C¹⁸O ($\theta^*(C^{18}O) = 0.60$) in the presence of gas-phase C¹⁶O. For curve a, the C¹⁶O flux is $4.1 \times 10^{14} \text{ cm}^2 \text{ s}^{-1}$; for curve b, $7.3 \times 10^{14} \text{ cm}^2 \text{ s}^{-1}$; and for curve c, $14.5 \times 10^{14} \text{ cm}^2 \text{ s}^{-1}$ 269

Fig. 8 a) Variation in θ_T versus T. b) Variation in the average E_d for the desorbing CO molecules versus T. In both panels, curve a corresponds to a C¹⁶O flux of 0 $\text{cm}^2 \text{ s}^{-1}$; curve b, $4.1 \times 10^{14} \text{ cm}^2 \text{ s}^{-1}$; and curve c, $14.5 \times 10^{14} \text{ cm}^2 \text{ s}^{-1}$ 271

List of Tables

Chapter II

Table 2.1 Translational, rotational, and vibrational partitions functions for molecules [37].....	38
Table 2.2 Range of the preexponential factor from absolute rate theory [37].....	38
Table 3.1 Dissociative sticking probability of hydrogen on Cu surfaces at $T_s=1000$ K [33].....	54
Table 3.2 Dissociative sticking probability of hydrogen on Ni surfaces at $T_s=300$ K [33].....	54
Table 3.3 Activation energy (kcal/mol) for dissociative adsorption of hydrogen and deuterium on Ni surfaces [55].....	55
Table 3.4 Values of the parameters used to represent the data in fig. 3.9 [59].....	55
Table 4.1 Arrhenius parameters for atomic self-diffusion [64]....	76
Table 6.1 Activation energy barriers (kcal/mol) for selected surface reactions determined from the BOC-MP method [17].....	131
Table 6.2 Parameters used to represent the CO-NO reaction network [168].....	132

Chapter III

Table 1 Model parameters.....	169
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Chapter IV

Table 1	Model parameters for the TPD simulations of associative desorption in the presence of a fixed coadsorbate.....	214
Table 2	Model parameters for the TPD simulations of associative desorption and molecular desorption.....	228

Chapter I

Introduction

The phenomenon of chemical reactions occurring on surfaces is the basis for a number of technological areas such as catalysis, chemical vapor deposition, and coatings. Each of these processes can be envisioned to proceed by one or more of the following elementary steps: adsorption, surface diffusion, desorption, and surface reaction. Adsorption, desorption, and surface diffusion transport adsorbates to, from, and on the surface, respectively, whereas surface reactions transform reactants to products. In order to correctly predict the kinetics of surface phenomena, an understanding of the individual and collective effects of these four elementary steps is necessary.

The earliest models for describing gas-surface interactions were based on the assumption that adsorbates interact strongly with the surface but not with each other. In many instances, though, kinetic data could not be adequately explained with this assumption. A large body of subsequent work, including present day research, has thus been devoted to studying both adsorbate-surface interactions and the lateral interactions between species on the surface. The objective of this work is to understand kinetic processes occurring on well-defined metal surfaces. Of principal interest is successfully accounting for the kinetics over a wide range of adsorbate surface concentrations.

In chapter II of this work, the theoretical models currently

available for describing adsorption, surface diffusion, desorption, and surface reaction are reviewed. Although a large number of approaches exist for treating surface kinetics, Monte Carlo models are the most appropriate for describing the effects of lateral interactions between adsorbates over a wide range of the surface coverage. For this reason, a Monte Carlo simulation technique has been developed for describing the adsorption, surface diffusion, and desorption of adsorbates. In chapter III, the development of a Monte Carlo model for simulating the temperature-programmed desorption (TPD) of a single type of adsorbates on single-crystal metal surfaces is presented. In chapter IV, the model is extended to simulate the TPD spectra of two species coadsorbed on a surface. Finally, in chapter V, the influence of adsorption kinetics on the isothermal desorption and temperature-programmed desorption of a single type of species is examined. An interesting feature of the model is that it predicts kinetic behavior in good agreement with experimental observation.

Chapter II

A Review of Theoretical Models of Adsorption, Diffusion, Desorption, and Reaction of Gases on Metal Surfaces

ABSTRACT

A review is presented of the theoretical approaches available for describing the kinetics of gas adsorption, diffusion, desorption, and reaction on metal surfaces. The prediction of rate and diffusion coefficients based on molecular dynamics, transition-state theory, stochastic diffusion theory, and quantum mechanics is discussed, and the success of these theoretical approaches in representing experimental observation is examined. Consideration is also given to the effects of lateral interactions between adsorbates and to the ability of lattice-gas models to provide a representation of the dependences of rate and diffusion coefficients on adsorbate coverage. Finally, the utility of continuum and Monte Carlo models for describing the kinetics of complex surface processes in terms of elementary processes is addressed.

1.0 INTRODUCTION

The kinetics of chemical reactions occurring on the surfaces of metals is of continuing interest to researchers working in the areas of catalysis, chemical vapor deposition, etching, and corrosion. While it has been known for some time that the kinetics of such reactions depend on the properties of the metal, as well as those of the adsorbed reactants, intermediates and products, it has only been in the past two decades that it has become possible to understand the dynamics of surface reactions at the atomic and molecular level. Progress has come through careful experimental measurements of the rates of elementary processes such as adsorption, diffusion, desorption, and reaction, and through the development of theoretical methods for predicting the dynamics of such processes. Theoretical studies have also proven useful for interpreting the experimentally observed effects of surface composition/structure and adsorbate coverage on rate and diffusion coefficients, and for describing the kinetics of complex surface processes in terms of a sequence of elementary steps. It is therefore timely to examine the current state of the art of theoretical methods for predicting rate and diffusion coefficients for elementary processes occurring on metal surfaces and the extent to which mechanistic models are successful in representing the overall kinetics of complex surface processes (e.g., catalysis).

Previous reviews on the subjects of adsorption [1,2,3], surface diffusion [4,5], desorption [1,2,3,6,7], and surface reaction [2,3,8] have been primarily concerned with summarizing experimental

results and with presenting techniques for interpreting experimental data. While theoretical methods for predicting rate and diffusion coefficients are discussed in some of these publications, none has included a comprehensive overview of all four elementary processes. In other reviews [9-11], the theoretical approaches used to model gas-surface dynamics are described, but the success or failure of a given approach to produce results consistent with experimental observation is not treated. The purpose of this review is to put into perspective the different theoretical approaches available for describing rate and diffusion coefficients and to evaluate the success of a given approach in providing quantitative agreement with experiment. The simulation of complex surface processes in which the dynamics of individual elementary processes are represented explicitly is also discussed. To limit the scope of this review, attention is restricted to the interactions of gases with metal surfaces.

The remainder of the text is contained in Sections 2-7. Section 2 reviews the theoretical concepts used to represent rate and diffusion coefficients and discusses the relative merits and limitations of each approach. The effects of adsorbate coverage on these coefficients is also examined in this section. Sections 3-6 illustrate the success of different theoretical approaches in predicting the magnitude of the coefficients for adsorption, diffusion, desorption, and reaction, respectively. Sections 5 and 6 also discuss the simulation of temperature-programmed desorption and reaction experiments. A set of concluding remarks summarizing the present state of knowledge in gas-surface dynamics is presented

in Section 7.

2.0 THEORETICAL CONCEPTS

2.1 Adsorbate-Surface Potentials

The dynamics of processes occurring at a metal surface are governed by the adsorbate-surface potential. This function is comprised of contributions due to metal-adsorbate interactions and lateral interactions between adsorbates. Ideally, the adsorbate-surface potential should be obtained by an accurate *ab initio* solution of the electronic Schrödinger equation. However, even with currently available supercomputers, it is not possible to produce complete energy hypersurfaces that are accurate enough for use in dynamical studies. As a consequence, all theoretical analyses of elementary processes occurring at surfaces are based on semi-empirical expressions, the exact form of the potential depending on the nature of the problem being addressed.

Metal-Adsorbate Interactions

For weakly bonding situations between an adsorbate, A, and n metal atoms, the metal-adsorbate interaction can be described by a sum of Lennard-Jones contributions:

$$V = \sum_{i=1}^n 4\epsilon \left[\left(\frac{\sigma}{r_i} \right)^{12} - \left(\frac{\sigma}{r_i} \right)^6 \right] \quad (2.1)$$

where r_i is the separation distance between the adsorbate and the i th metal atom, and ϵ and σ are Lennard-Jones parameters. Where covalent bonding between the adsorbate and surface occurs, the metal-adsorbate interaction can be described by a sum of two-

center Morse potentials

$$V = V_{0A} \sum_{i=1}^n (x_i^2 - 2x_i) \quad (2.2)$$

$$x_i = \exp[-(r_i - r_0)/a] \quad (2.3)$$

where V_{0A} is the equilibrium (minimum) potential energy, x_i is the Pauling bond order, r_i is the bond distance, r_0 is the equilibrium bond distance, and a is a scaling parameter.

More complex forms of the adsorbate-surface potential have also been used. An example is the London, Eyring, Polanyi, Sato (LEPS) potential surface, which has been used in studies of the adsorption of diatomic molecules on metal surfaces [12]. The LEPS potential for the molecule AB is given by

$$V = p_{AS} + p_{BS} + p_{AB} + [\chi_{AB}(\chi_{AB} - \chi_{AS} - \chi_{BS}) + (\chi_{AS} + \chi_{BS})^2]^{1/2} \quad (2.4)$$

where p_{ij} and χ_{ij} are the Coulomb and exchange contributions between i and j, and subscript S denotes the surface. The contributions p_{ij} and χ_{ij} are expressed in the form of modified Morse potentials.

While the adsorbate-surface potential is a multi-dimensional function in general, critical features of the potential energy surface are best characterized by examining the variation of the potential along a trajectory corresponding to the so-called reaction coordinate, s. Illustrations of two-dimensional potential surfaces and their one-dimensional representations corresponding to trajectories along the reaction coordinate are presented in figs. 2.1 and 2.2 for the cases of non-dissociative and dissociative adsorption, respectively.

For non-dissociative adsorption of a molecule AB, the one-

dimensional potential in fig. 2.1b is zero for large distances from the surface. As the adsorbate-metal distance is reduced, the potential goes through a weak minimum and then a weak maximum before it reaches the final minimum corresponding to the adsorbed state. The first minimum is associated with the presence of a weakly-bound precursor state, whereas the maximum corresponds to the activation energy barrier for adsorption. For dissociative adsorption, a qualitatively similar picture is observed. In this instance, the one-dimensional potential in fig. 2.2b is described as a function of s , the reaction coordinate.

Lateral Interactions

While interactions between adsorbates can be ignored at low coverages, experimental evidence indicates that with increasing coverage, such interactions can become significant. Several processes can contribute to lateral interactions. Repulsive interactions can arise from the coupling of mutually aligned dipoles on the surface and from changes in adsorbate-site orbital overlap due to the interaction of multiple adsorbates with a single adsorption site. The direct interactions of orbitals on adjacent species may lead to either attractive or repulsive contributions to the total energy. The classification of an interaction between two species as purely attractive or purely repulsive, however, is not always precise. Quantum models have demonstrated that the nature of an interaction between adsorbates depends upon the separation distance and the crystallographic direction on the surface [13,14].

The description of lateral interactions depends upon the nature of the system considered. For the adsorption of weakly bound

(physisorbed) species, lateral interactions between adsorbates can be handled by the addition of appropriate Lennard-Jones terms to eq. 2.1. For strongly bound (chemisorbed) species, the effects of lateral interactions can be expressed as a perturbation in the heat of adsorption of the adsorbate in the absence of lateral interactions.

The simplest representation of lateral interactions for chemisorbed species is based on the assumption that all contributions are pairwise additive. In such a case, the heat of adsorption of species A is given by

$$Q_A = Q^0 - nw_{AA} \quad (2.5)$$

where Q^0 is the heat of adsorption of A in the absence of lateral interactions, w_{AA} is the energy of each A-A interaction between nearest neighbors (w_{AA} is positive for repulsive interactions and negative for attractive interactions), and n is the number of nearest neighbors. The form of eq. 2.5 is derived empirically and is not based on any physical model for the interactions of nearest-neighbor adsorbates.

Recently, an alternative to eq. 2.5 has been developed based on the Bond-Order-Conservation-Morse-Potential (BOC-MP) approach pioneered by Shustorovich [15-17]. This method describes each two-center interaction between an adsorbate atom A and a surface metal atom M by a Morse potential, and the total heat of adsorption is given by the sum of all two-center interactions. A further assumption of the BOC method is that along a reaction path describing the interactions of a molecular or atomic species with a metal surface, the total bond order, x, is conserved and normalized to unity.

Within the BOC-MP framework, the heat of chemisorption for an isolated atom A on a surface is given by

$$Q_{A,n'} = Q_{oA} (2 - 1/n') \quad (2.6)$$

where Q_{oA} is the heat of chemisorption of A in the on-top position, and n' is the number of metal atoms to which A is coordinated. For an isolated molecular adsorbate AB coordinated via atom A to n' metal atoms, the heat of chemisorption can be approximated by

$$Q_{AB,n'} = \frac{Q_{oA}^2}{D_{AB} + Q_{oA} / n'} \quad (2.7)$$

where D_{AB} is the gas-phase A-B bond energy.

Equations 2.6 and 2.7 are valid for isolated adsorbate atoms or molecules on a surface. For higher coverages, however, situations may arise in which more than one adsorbate is bonded to a metal atom and, furthermore, the adsorbates may interact directly with each other. To account for these metal-adsorbate (M-A) and adsorbate-adsorbate (A-A) interactions, the total binding energy of species A is partitioned as follows:

$$Q_{A,n'} = Q_{A,n'}^{(1)} + Q_{A,n'}^{(2)} \quad (2.8)$$

where $Q_{A,n'}^{(1)}$ is the heat of adsorption due to M-A interactions and $Q_{A,n'}^{(2)}$ is the heat of adsorption due to A-A interactions. Both $Q_{A,n'}^{(1)}$ and $Q_{A,n'}^{(2)}$ can be calculated explicitly as a function of the local occupancy of nearest-neighbor sites, under the constraint that the total bond order of A for both M-A and A-A interactions is conserved to unity.

The value of $Q_{A,n'}^{(2)}$ can be expressed as

$$Q_{A,n}^{(2)} = \sum_{i=1}^{n'} \sum_{l=1}^L 0.5 D_{AA} (2\delta_{il} - \delta_{il}^2) \quad (2.9)$$

where D_{AA} is the A-A bond dissociation energy, and δ_{il} is the bond order for the A-A interaction between the A atom coordinated with metal i and the l -the nearest-neighbor A atom also coordinated with metal atom i . The summation over l in eq. 2.9 is to account for all nearest-neighbor A atoms. The occurrence of A-A interactions weakens the bond order associated with the M-A interactions and, as a consequence, the bond order for each component of an $M_{n'}-A$ bond is given by

$$x_{i,n'} = \frac{1}{n'} - \sum_{l=1}^L \delta_{il} \quad (2.10)$$

The value of $Q_{A,n'}^{(1)}$ when more than one adsorbate is bonded to an individual metal atom is given by

$$Q_{A,n'}^{(1)} = \sum_{i=1}^{n'} \left[\frac{Q_{0A}}{m_i} \left(2 - \frac{1}{m_i} \right) (2x_{i,n'} - x_{i,n'}^2) \right] \quad (2.11)$$

where m_i is the number of adsorbates bonded to the i -the metal atom, and $x_{i,n'}$ is given by eq. 2.10.

Activation Energy Barriers

The characteristic of the potential energy hypersurface having the greatest effect on the dynamics of an elementary process is the height of the activation barrier, E , for movement from the region of the potential surface associated with the reactant state to that associated with the product state. Since the full, multi-dimensional potential surface cannot be described in most instances,

semiempirical models must be used to relate E to the properties of the reactants and to describe the effects of lateral interactions between adsorbates.

The simplest representation of E when lateral interactions occur is based on the assumption that all contributions are pairwise additive. In such a case, E is given by

$$E = E^\circ - nw_{AA} \quad (2.12)$$

where E° is the activation energy in the absence of lateral interactions, w_{AA} is the contribution of nearest-neighbor A-A interactions to E , and n is the number of nearest neighbors. The form of eq. 2.12 is empirical and is not based on any physical model for the interactions of nearest-neighbor adsorbates.

The BOC-MP [17] approach provides an alternative method for estimating the magnitude of E . The activation energy for dissociative adsorption from the gas phase, E_a , can be written as

$$E_a = \frac{1}{2} [D_{AB} - (Q_A + Q_B) + \frac{Q_A Q_B}{Q_A + Q_B} - Q_{AB}] \quad (2.13)$$

For non-associative desorption, E_d is given by

$$E_d = Q_A \quad (2.14)$$

whereas for associative desorption, E_d is given by

$$E_d = Q_A + Q_B - D_{AB} + E_a \quad \text{if } E_a > 0 \quad (2.15)$$

or

$$E_d = Q_A + Q_B - D_{AB} \quad \text{if } E_a < 0 \quad (2.16)$$

Expressions similar to eqs. 2.15 and 2.16 can be used to calculate the activation energy for reactions, E_r . In the BOC-MP framework, the activation energy for diffusion across a bridge site, E_{diff} , is given by

$$E_{\text{diff}} = \frac{(n'-2)}{(4n'-2)} Q_A \quad (2.17)$$

where n' is the coordination number for the adsorbed species in a hollow site (viz., $n' = 3, 4, 5$). The influence of lateral interactions on the activation energy can be determined using eq. 2.8 to calculate Q_i for each species i appearing in eqs. 2.13-2.17.

2.2 Kinetics of Adsorption, Desorption, and Surface Reaction

The relationship between the rate of a surface process and the adsorbate coverage—the kinetics of the process—depends on the nature of the process and the extent to which lateral interactions are important. In the absence of such interactions, and assuming that all adsorbed species are randomly distributed, simple expressions can be written to relate the rate of a process to the surface coverage, θ . The dependence of the rate of elementary processes on θ are much more complex when lateral interactions are included.

Non-Interacting Adsorbates

For non-interacting adsorbates, the rate of adsorption of gas-phase species, r_a , can be written as:

$$r_a = S^{(\alpha)}(\theta)F \quad (2.18)$$

where F is the flux of the adsorbate and $S^{(\alpha)}(\theta)$ is the sticking coefficient. The superscript α on $S^{(\alpha)}(\theta)$ is 1 for non-dissociative and 2 for dissociative adsorption. The form of $S^{(\alpha)}(\theta)$ depends on whether adsorption occurs directly from the gas phase or via a precursor state. For direct adsorption,

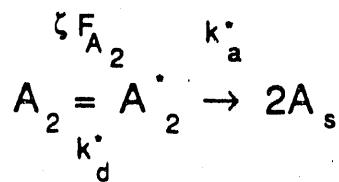
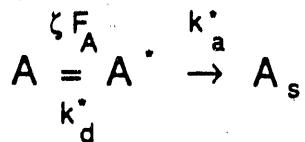
$$S^{(\alpha)}(\theta) = S^{(\alpha)}(0)(1-\theta)^\alpha \quad (2.19)$$

where $S^{(\alpha)}(0)$ is the sticking coefficient at zero coverage. If $S^{(\alpha)}(0)$ is assumed to obey an Arrhenius expression, then

$$S^{(\alpha)}(0) = S_0^{(\alpha)} \exp(-E_a/k_b T) \quad (2.20)$$

where $S_0^{(\alpha)}$ and E_a are the preexponential factor and activation energy for adsorption, respectively, and k_b is the Boltzmann constant.

When adsorption is assumed to proceed via a precursor state, the mechanism of adsorption can be represented by [18]:



where the precursor species for non-dissociative adsorption is denoted by A^* (A_2^* for dissociative adsorption), the chemisorbed species by A_s , ζ is the trapping probability from the gas phase into the precursor state, k_a^* is the rate constant for adsorption from the precursor state into the chemisorbed state, and k_d^* is the rate constant for desorption from the precursor state. The precursor state can be located over an empty site (an intrinsic precursor) or over an occupied site (an extrinsic precursor). Two different approaches have been used to represent the adsorption rate. The

first is based on a continuum description and uses the stationary-state approximation to determine the concentration of precursor species [19,20]. The second approach is based on a successive-site statistical model of the type first proposed by Kisliuk [21,22].

Although conceptually different, both approaches have been shown to lead to equivalent forms of the adsorption rate in many cases [20,23].

If the intrinsic and extrinsic precursors are energetically equivalent and each occupies only a single adsite, then the rates of non-dissociative and dissociative adsorption can be written as [18]:

$$r_a = \frac{\zeta F_A k_a (1-\theta_A)}{k_d + k_a (1-\theta_A)} \quad (2.21)$$

$$r_a = \frac{\zeta F_{A_2} k_a (1-\theta_A)^2}{k_d + k_a ((1-\theta_A))^2} \quad (2.22)$$

Equations 2.21 and 2.22 can be used together with eq. 2.18 to write expressions for $S^{(\alpha)}(\theta)/S^{(\alpha)}(0)$. Thus,

$$S^{(1)}(\theta)/S^{(1)}(0) = \frac{(1+K)(1-\theta_A)}{1+K(1-\theta_A)} \quad (2.23)$$

$$S^{(2)}(\theta)/S^{(2)}(0) = \frac{(1+K)(1-\theta_A)^2}{1+K(1-\theta_A)^2} \quad (2.24)$$

where $K = k_a/k_d$. A plot of $S^{(1)}(\theta)/S^{(1)}(0)$ versus θ is shown in fig. 2.3. When $K \gg 1$, $S^{(1)}(\theta)/S^{(1)}(0) = 1$ and when $K \ll 1$, $S^{(1)}(\theta)/S^{(1)}(0) = (1-\theta)$. Since the value of K is temperature dependent, the shape of $S^{(\alpha)}(\theta)/S^{(\alpha)}(0)$ versus θ will depend on temperature.

The rate of desorption for a randomly distributed adsorbate in the absence of lateral interactions can be written as

$$r_d = k_d^{(\alpha)} \theta^\alpha \quad (2.25)$$

where $k_d^{(\alpha)}$ is the rate coefficient for desorption. If $k_d^{(\alpha)}$ can be described by an Arrhenius expression, then

$$k_d^{(\alpha)} = v_d^{(\alpha)} \exp(-E_d/k_b T) \quad (2.26)$$

where $v_d^{(\alpha)}$ and E_d are the preexponential factor and activation energy for desorption, respectively.

Implicit in the formulation of eq. 2.26 is the assumption that desorption occurs directly from the adsorbed state. This assumption is unnecessarily restrictive since it is conceivable that the adsorbate passes through a weakly bound precursor state before leaving the catalyst surface. Making the same assumptions used in deriving eqs. 2.21 and 2.22, the rates of non-associative and associative desorption can be written as [18]:

$$r_d = \frac{k_d^{(1)} k_d \dot{\theta}_A}{k_a \dot{\theta}_A + k_d} \quad (2.27)$$

$$r_d = \frac{k_d^{(2)} k_d \dot{\theta}_A^2}{k_a \dot{\theta}_A^2 + k_d} \quad (2.28)$$

In the limit $k_a \dot{\theta}_A \ll k_d$, eqs. 2.27 and 2.28 reduce to eq. 2.25.

Surface reactions can be classified into two generic types. The first includes reactions between two adsorbed species or an adsorbed species and a vacant site (Langmuir-Hinshelwood processes). For randomly distributed adsorbates on a surface in the

absence of adsorbate-adsorbate interactions, the rate of reaction is given by:

$$r_r = k_r \theta_A \theta_B \quad (2.29)$$

or

$$r_r = k_r \theta_A \theta_V \quad (2.30)$$

where k_r is the rate coefficient, θ_i is the surface coverage of species i and θ_V is the fraction of vacant sites. If k_r follows an Arrhenius expression, then

$$k_r = v_r \exp(-E_r/k_b T) \quad (2.31)$$

where v_r and E_r are the preexponential factor and the activation energy for reaction, respectively.

The second class of reactions includes the direct interaction of a gas-phase species with an adsorbed species to form a product which may either remain adsorbed or desorb into the gas phase (Eley-Rideal processes). For such processes, the rate of reaction can be written as

$$r_r = k_r \theta_A P_B \quad (2.32)$$

where P_B is the partial pressure of reactant B. The form of eq. 2.32 is similar to that for adsorption and so k_r can be represented as a reactive sticking coefficient, S_o , by the expression

$$k_r = \frac{S_o a_s}{(2\pi m_B k_b T_g)^{1/2}} \exp(-E_r/k_b T) \quad (2.33)$$

where a_s is the area per reaction site, and m_B is the molecular weight of species B.

Interacting Adsorbates

When lateral interactions become significant, the relationships between the rate of an elementary process and the

adsorbate coverage become quite complex and cannot, in general, be written in closed form. An exception to this occurs in the case where a lattice-gas model is used to describe the effects of lateral interactions. In such a model, each adsorbate is assumed to be localized on a two-dimensional array of surface sites, and each site is assumed to be either vacant or occupied by a single adsorbate. A given adsorbate can interact with adsorbates on nearest-neighbor sites, next-nearest-neighbor sites, etc., but in most variants of the lattice-gas model, only nearest-neighbor interactions are taken into account. Using these assumptions, relationships can be derived between the adsorbate coverage and the rate of adsorption, desorption, and surface reaction [24-26].

The rate of non-dissociative adsorption is defined as

$$r_a = FS_0^{(1)} \sum_n P_{V,n} \exp\left[\frac{-(E_a^\circ - E_s)}{k_b T}\right] (1-\theta) \quad (2.34)$$

where $S_0^{(1)}$ is the initial sticking coefficient, $P_{V,n}$ is the probability of finding a vacant site with n occupied nearest-neighbor sites, E_a° is the activation energy for adsorption in the absence of nearest-neighbor interactions, and E_s is the interaction energy of A with its neighbors. E_s can be related to the energy for a given pairwise interaction, WAA , by $E_s = nWAA$. For dissociative adsorption, the rate of adsorption can be expressed as

$$r_a = FS_0^{(2)} \sum_{n,m} P_{V,V;nm} \exp\left[\frac{-(E_a^\circ - E_s)}{k_b T}\right] \quad (2.35)$$

where $P_{V,V;nm}$ is the probability that two nearest-neighbor sites are vacant such that one is surrounded by n nearest neighbors and the other is surrounded by m nearest neighbors.

To complete the description of the lattice-gas model, it is necessary to develop expressions for $P_{V,n}$ and $P_{VV;nm}$ in terms of θ_A and W_{AA} . Since an exact solution to this problem can only be achieved for a coverage of 0.5, an approximation must be made to obtain a solution in closed form. One of the more widely used postulates is the quasi-chemical approximation (QCA) which assumes that the adsorbates maintain an equilibrium distribution on the surface. The QCA treats the probabilities $P_{V,n}$ and $P_{VV;nm}$ in terms of the probabilities P_{AA} , P_{AV} , and P_{VV} , where P_{AA} is the probability that two nearest-neighbor sites are occupied by species A, P_{VV} is the probability that two nearest-neighbor sites are vacant, and P_{AV} is the probability that of two nearest-neighbor sites, one is occupied and the other is vacant. These probabilities are described by the equations:

$$P_{AA} + P_{AV} + P_{VV} = 1 \quad (2.36)$$

$$2P_{AA} + P_{AV} = 2\theta \quad (2.37)$$

$$\frac{P_{AA}P_{VV}}{P_{AV}^2} = \frac{1}{4} \exp\left[\frac{-W_{AA}}{k_b T}\right] \quad (2.38)$$

The solution to eqs. 2.36-2.38 is given by

$$P_{AA} = \theta - \left\{1 - [1 - 2\eta\theta(1-\theta)]^{1/2}\right\} / \eta \quad (2.39)$$

$$P_{AV} = 2\left\{1 - [1 - 2\eta\theta(1-\theta)]^{1/2}\right\} / \eta \quad (2.40)$$

$$P_{VV} = 1 - \theta - \left\{1 - [1 - 2\eta\theta(1-\theta)]^{1/2}\right\} / \eta \quad (2.41)$$

where $\eta = 2[1 - \exp(-W_{AA}/k_b T)]$.

In the QCA, different pairs of sites are considered to be independent and hence

$$P_{V,n} = \frac{z!}{n!(z-n)!} \frac{(0.5P_{AV})^n (P_{VV})^{z-n}}{(0.5P_{AV} + P_{VV})^z} \quad (2.42)$$

where z is the number of nearest-neighbor sites. Using eqs. 2.39-2.42, the following functions can be defined to represent the sums appearing in eqs. 2.34 and 2.35, respectively [25]:

$$\begin{aligned} f_a'(\theta, w_{AA}) &= \sum_n P_{V,n} \exp\left[\frac{nw_{AA}}{k_b T}\right] \\ &= \left(\frac{0.5P_{AV} \exp\left[\frac{w_{AA}}{k_b T}\right] + P_{VV}}{1-\theta} \right)^z \end{aligned} \quad (2.43)$$

$$\begin{aligned} f_a''(\theta, w_{AA}) &= \sum_{n,m} P_{VV;nm} \exp\left[\frac{(n+m)w_{AA}}{k_b T}\right] \\ &= P_{VV} \left(\frac{0.5P_{AV} \exp\left[\frac{w_{AA}}{k_b T}\right] + P_{VV}}{1-\theta} \right)^{2z-2} \end{aligned} \quad (2.44)$$

When $w_{AA} = 0$, $f_a'(\theta_A, w_{AA}) = 1$ and $f_a''(\theta_A, w_{AA}) = (1-\theta)^2$, and eqs. 2.34 and 2.35 become identical to eqs. 2.18-2.20. Thus, f_a' and f_a'' represent corrections to eqs. 2.18-2.20 due to lateral interactions.

The rate of non-associative desorption is defined as

$$r_d = v_d^{(1)} \sum_n P_{A,n} \exp\left[\frac{-(E_d^0 - E_s)}{k_b T}\right] \theta \quad (2.45)$$

where $P_{A,n}$ is the probability of finding an adsorbate A with n nearest-neighbors, E_d^0 is the activation energy for desorption in the absence of nearest-neighbor interactions, and E_s is the interaction energy of A with its neighbors. E_s can be related to the energy for a given pairwise interaction, w_{AA} , by $E_s = nw_{AA}$. For associative desorption, the rate of desorption can be expressed as

$$r_d = v_d^{(2)} \sum_{n,m} P_{AA;n,m} \exp\left[\frac{-(E_d^0 - E_s)}{k_b T}\right] \quad (2.46)$$

where $P_{AA;n,m}$ is the probability that two nearest-neighbor sites are occupied by A atoms such that one is surrounded by n nearest neighbors and the other is surrounded by m nearest neighbors.

Expressions for $P_{A,n}$ and $P_{AA;n,m}$ can be written using the QCA. Thus, for $P_{A,n}$ one can write

$$P_{A,n} = \frac{z!}{n!(z-n)!} \frac{P_{AA}^n (0.5P_{AV})^{z-n}}{(P_{AA} + 0.5P_{AV})^z} \quad (2.47)$$

Using eqs. 2.39-2.41 and 2.47, one can define the following functions [24,26]:

$$\begin{aligned} f_d(\theta, w_{AA}) &= \sum_n P_{A,n} \exp\left[\frac{n w_{AA}}{k_b T}\right] \\ &= \left(\frac{P_{AA} \exp\left[\frac{w_{AA}}{k_b T}\right] + 0.5P_{AV}}{\theta} \right)^z \end{aligned} \quad (2.48)$$

$$\begin{aligned} f_d(\theta, w_{AA}) &= \frac{\sum_{n,m} P_{AA;n,m} \exp\left[\frac{(n+m) w_{AA}}{k_b T}\right]}{\theta^2} \\ &= P_{AA} \left(\frac{P_{AA} \exp\left[\frac{w_{AA}}{k_b T}\right] + 0.5P_{AV}}{\theta} \right)^{2z-2} \end{aligned} \quad (2.49)$$

Substitution of eqs. 2.48 and 2.49 into eqs. 2.45 and 2.46, respectively, gives:

$$r_d = v_d^{(1)} \exp\left[\frac{-E_d^0}{k_b T}\right] \theta f_d(\theta, w_{AA}) \quad (2.50)$$

$$r_d = v_d^{(2)} \exp \left[\frac{-E_d^0}{k_b T} \right] f_d'(\theta, w_{AA}) \quad (2.51)$$

When $w_{AA} = 0$, $f_d'(\theta_A, w_{AA}) = 1.0$ and $f_d''(\theta_A, w_{AA}) = \theta_A^2$, and eqs. 2.50 and 2.51 become identical to eqs. 2.25 and 2.26. Thus, f_d' and f_d'' represent corrections to the Polanyi-Wigner expressions for desorption (eqs. 2.25 and 2.26) due to lateral interactions. The functions f_d' and f_d'' are plotted in fig. 2.4 for attractive and repulsive values of $w_{AA}/k_b T$. It is apparent that even for small values of $w_{AA}/k_b T$, lateral interactions can have a profound effect on f_d' and f_d'' .

Another way of visualizing the effects of lateral interactions is to look at the coverage dependencies of the apparent activation energy and the apparent preexponential factor. For non-associative desorption, these quantities are defined by [26]:

$$\Delta E_d(\theta) = k_b T^2 \frac{d \ln f_d}{dT} \quad (2.52)$$

$$\frac{v_d^{(1)}(\theta)}{v_d^{(1)}(0)} = f_d' \exp \left[\frac{\Delta E_d(\theta)}{k_b T} \right] \quad (2.53)$$

Figure 2.5 shows that the variation in $\Delta E_d(\theta)$ and $v_d(1)(\theta)/v_d(1)(0)$ with θ is a strong function not only of $w_{AA}/k_b T$, but also of z . The apparent activation energy is seen to vary monotonically with coverage whereas the apparent preexponential factor exhibits a more complex coverage dependence.

The lattice-gas model developed for the description of associative desorption can be extended to describe the kinetics of

Langmuir-Hinshelwood processes. The rate of the bimolecular reaction $A_s + B_s \rightarrow C_s$ or C_g can be written as

$$r_r = v_r \sum_i P_{AB;i} \exp\left[-\frac{(E_r^0 - E_s)}{k_b T}\right] \quad (2.54)$$

where v_r is the preexponential factor, $P_{AB;i}$ is the probability that two nearest sites are occupied by an AB pair and that this pair has the environment i , E_r^0 is the activation energy for the reaction of A and B in the absence of nearest-neighbor interactions, and E_s is the contribution to the activation energy due to nearest-neighbor interactions. The lateral interaction energy of the AB pair is determined by $k, l, k',$ and l' , where k and l (k' and l') are the numbers of A and B neighbors of A_s (B_s), respectively. Using these indices, E_s is written as

$$E_s = kw_{AA} - (l+k')w_{AB} - l'w_{BB} \quad (2.55)$$

In the QCA, the probability $P_{AB;i}$ can be expressed in terms of $P_{AA}, P_{AB}, P_{BB}, P_{AV}, P_{BV}, P_{VV},$ and P_{AB} . These probabilities are described by the equations

$$P_{AA} + P_{AB} + P_{AV} + P_{BB} + P_{BV} + P_{VV} = 1 \quad (2.56)$$

$$2P_{AA} + P_{AB} + P_{AV} = 2\theta_A \quad (2.57)$$

$$2P_{BB} + P_{AB} + P_{BV} = 2\theta_B \quad (2.58)$$

$$\frac{P_{AA}P_{VV}}{P_{AV}^2} = \frac{1}{4} \exp(-w_{AA}/k_b T) \quad (2.59)$$

$$\frac{P_{BB}P_{VV}}{P_{BV}^2} = \frac{1}{4} \exp(-w_{BB}/k_b T) \quad (2.60)$$

$$\frac{P_{AA}P_{BB}}{P_{AB}^2} = \frac{1}{4} \exp\left[-\frac{(w_{AA} + w_{BB} - 2w_{AB})}{k_b T}\right] \quad (2.61)$$

Zhdanov [24,25] has shown that within the QCA, r_r can be written as

$$r_r = v_r \exp \left[\frac{-E_r^o}{k_b T} \right] P_{AB} \left(\frac{P_{AA} \exp \left[\frac{W_{AA}}{k_b T} \right] + 0.5 P_{AB} \exp \left[\frac{W_{AB}}{k_b T} \right] + 0.5 P_{AV}}{\theta_A} \right)^{z-1} \times \\ \left(\frac{P_{BB} \exp \left[\frac{W_{BB}}{k_b T} \right] + 0.5 P_{AB} \exp \left[\frac{W_{AB}}{k_b T} \right] + 0.5 P_{BV}}{\theta_B} \right)^{z-1} \quad (2.62)$$

It is apparent from eq. 2.62 that inclusion of interactions between species leads to a rate expression dependent on the sign and magnitude of these interactions. Moreover, the dependence of r_r on θ_A and θ_B is seen to differ significantly from that based on the assumption of a random adsorbate distribution, as described by eq. 2.29. To establish the extent to which eqs. 2.29 and 2.62 differ, it is useful to define the function [25]

$$f_r = \frac{\sum_i P_{AB,ii} \exp \left[\frac{\Delta E_r(\theta)}{k_b T} \right]}{\theta_A \theta_B} \quad (2.63)$$

Calculations of f_r for a square lattice are given in fig. 2.6. Both the magnitude and sign of the A-A, A-B, and B-B interactions influence f_r .

2.3 Diffusion

The diffusion of an adsorbate on a surface can be characterized by two types of motion. When the activation energy for diffusion,

E_{diff} , is less than $k_b T$, the adsorbate translates freely across the surface. This type of motion, which is termed mobile diffusion, is characteristic of physisorbed species. When E_{diff} is significantly greater than $k_b T$, the adsorbate moves from site to site by hopping over the activation barrier. Diffusion via a hopping mechanism is characteristic of chemisorbed species.

In the presence of a concentration gradient, the diffusive flux, J , of an adsorbate is given by Fick's law:

$$J = -D \nabla C \quad (2.64)$$

where C is the concentration of adsorbate and D is the concentration-independent diffusivity. For a freely mobile adsorbate, the diffusivity can be expressed as:

$$D = \frac{1}{d} \bar{v} \lambda \quad (2.65)$$

where \bar{v} is the mean velocity of the adsorbate, λ is the mean-free path, and d is the number of dimensions in which the motion occurs ($d=2$ for an isotropic surface). For hopping motion, the diffusivity is written as

$$D = \frac{1}{2d} \Gamma \lambda^2 \quad (2.66)$$

where Γ is the hopping frequency and λ now represents the mean-free hopping length. Since hopping is an activated process which is known to depend on temperature, it is customary to express Γ as

$$\Gamma = v_{\text{diff}} \exp(-E_{\text{diff}}/k_b T) \quad (2.67)$$

where v_{diff} is the preexponential factor for hopping. Substitution of eq. 2.67 into eq. 2.66 leads to:

$$D = \frac{1}{2d} \lambda^2 v_{\text{diff}} \exp(-E_{\text{diff}}/k_b T)$$

$$= D_0 \exp(-E_{\text{diff}}/k_b T) \quad (2.68)$$

where D_0 is referred to as the preexponential factor for diffusion.

2.4 Theoretical Descriptions of Surface Dynamics

Theoretical descriptions of the dynamics of elementary processes occurring on metal surfaces have been developed on the basis of molecular dynamics, transition-state theory, absolute rate theory, stochastic diffusion theory, and quantum mechanics. Each of these approaches is reviewed to illustrate the manner in which rate or transport coefficients are related to the adsorbate-surface potential.

Molecular Dynamics

The motion of an adsorbate near or on a surface can be obtained from molecular dynamics (MD) simulations. Such simulations provide a spatial and temporal description of the adsorbate trajectory, which can, in turn, be used to determine the sticking coefficient and the surface diffusion coefficient. For reasons discussed below, MD simulations are not well suited, in most cases, for determining the rate coefficient for desorption.

If the atoms in the metal surface are assumed to be stationary, then the adsorbate trajectory, $\mathbf{Y}(t)$, is given by

$$m\ddot{\mathbf{Y}} = -\partial V(\mathbf{Y})/\partial \mathbf{Y} \quad (2.69)$$

where m is the adsorbate mass and $V(\mathbf{Y})$ is the adsorbate-surface potential. A more realistic description of adsorbate-surface interactions is obtained when the metal atoms are allowed to vibrate. Treating the motion of all the metal atoms, however, is

computationally infeasible. To circumvent this, the problem is formulated in the following manner [27]. The adsorbate is assumed to interact strongly with n (1-10) atoms of the metal. This cluster constitutes the primary zone, and the interactions between the adsorbate and the metal atoms in the primary zone is described by $V_p(x_p, Y)$, where Y describes the position of the adsorbate and x_p describes the positions of the metal atoms in the primary zone. Adsorbate interactions with metal atoms outside the primary zone (i.e., in the secondary zone or heat bath) are described by the potential $V_s(Y)$. Thus the total interaction potential is $V(x_p, Y) = V_s(Y) + V_p(x_p, Y)$. The motion of the adsorbate is given by

$$m\ddot{Y} = -\partial V(x_p, Y)/\partial Y \quad (2.70)$$

whereas the motion of the metal atoms in the primary zone is given by a generalized Langevin equation

$$m'\ddot{x}_p = -\omega_{pp}^2 m' x_p - \partial V(x_p, Y)/\partial x_p + m' \int_0^t \Theta(t-\tau) x_p d\tau + R(t) \quad (2.71)$$

where m' is the mass of metal atom. The parameter ω_{pp} in eq. 2.71 is the characteristic frequency for vibration of metal atoms in the primary zone, and the functions $\Theta(t)$ and $R(t)$ correspond to a memory kernel and a random force. The integral involving $\Theta(t)$ describes the dissipation of energy from the primary zone to the heat bath, whereas $R(t)$ describes the transfer of energy from the heat bath to the primary zone. The last two terms in eq. 2.71 satisfy the second fluctuation-dissipation theorem, so that the surface temperature remains constant.

The solution of eq. 2.69 or eqs. 2.70 and 2.71 gives $Y(t)$, the

trajectory of the adsorbate as a function of time. The probability of adsorption can be determined by examining a collection of trajectories calculated for different initial conditions. Since the probability of adsorption is related to the time scale of the observation (i.e., a species which adsorbs may desorb at a later time), a working definition of adsorption must be adopted. As an example of such a definition, an adsorbate can be assumed to be trapped and equilibrated with the surface if at some time in its history the total energy of the adsorbate on the surface becomes less than $-3k_bT$ [28,29]. Alternatively, one can assume that adsorption occurs when the adsorbate-surface separation distance is less than some prescribed value. Thus, the fraction of all adsorbate trajectories satisfying the working definition condition can be defined as the sticking coefficient for non-dissociative adsorption, $S(1)(T)$. For dissociative adsorption of a molecule A_2 , the fraction of all trajectories in which the A-A bond distance exceeds a specified value at some time is defined as $S(2)(T)$ [30-35].

MD simulations can also be used to predict the diffusion coefficient of an adsorbate on a surface. For mobile diffusion, D is defined by the Einstein relationship as

$$D = \langle [Y(t) - Y(0)]^2 \rangle / 4t \quad (2.72)$$

where $\langle [Y(t) - Y(0)]^2 \rangle$ represents the ensemble average of the mean-square displacement.

The use of MD simulations to describe the dynamics of desorption and surface reactions is limited by technical difficulties. The numerical algorithms used for the solution of eq. 2.69 or eqs. 2.70 and 2.71 require integration step lengths comparable with the

time scale of the fastest motion, usually a vibrational period of 10^{-14} s. While simulation of processes which occur on the 1 to 10^3 picosecond time scale is practical, direct simulation of slow events which occur over microseconds or longer is usually prohibitive. In the case of desorption and surface reaction, the trajectory of the adsorbate must pass through a narrow region of phase space (i.e., the region of the col on the potential surface). Because such events occur infrequently, very long integration times are required if $E_d \gg k_b T$, and consequently, direct simulation of such events is impractical. As discussed below, the description of infrequent events is best handled by a dynamical form of transition-state theory.

Transition-State Theory

In classical transition-state theory (TST), the rate for a species going from state A to state B, $A \rightarrow B$, is defined as the equilibrium flux of adsorbate trajectories across a plane in phase space. The plane, S, which lies between the reactants and products, must be traversed at least once in going from A to B. While the location of the plane is arbitrary, it is often convenient to specify the location of S near or at the col on the potential energy surface (i.e., the transition state). Since some trajectories of species A may make multiple crossings of S before finally going to B whereas other trajectories which cross S correspond to the processes $A \rightarrow A$ or $A \rightarrow B$, simply counting adsorbate trajectories leads to an over-estimation of the rate. Thus, TST provides an upper limit to the true rate. The derivation of the TST formalism presented in the balance of this section is based on the work of Tully and coworkers [28,36].

For the process $A \rightarrow B$, the TST equilibrium rate coefficient, k , is given by [28]

$$k = \frac{\int dp \int dq \int_0^\infty dv_s P(p, q, s_0, v_s) v_s \xi(p, q, v_s)}{\int dp \int dq \int_{-\infty}^\infty dv_s \int_{-\infty}^\infty ds P(p, q, s, v_s)} \quad (2.73)$$

where the vectors q and p denote position and momenta, respectively, s is the coordinate normal to the counting plane located at s_0 , and v_s is the component of velocity in the s direction. Reactant A is located in regions of phase space where $s < s_0$; conversely, product B is located in regions of phase space where $s > s_0$. The equilibrium probability density function in phase space is defined as

$$P(p, q, s, v_s) = N \exp(-H(p, q, s, v_s)/k_b T) \quad (2.74)$$

which is valid for a canonical ensemble. Here, H is the the classical Hamiltonian of the system and N is a normalization constant. The factor $\xi(p, q, v_s)$ in eq. 2.73 is used to correct for trajectories which make multiple crossings of S or which do not contribute to the process $A \rightarrow B$.

It is convenient to factor eq. 2.73 into [28]

$$k = k_{TST} f_s \quad (2.75)$$

where k_{TST} is the uncorrected equilibrium rate constant given by

$$k_{TST} = \frac{\int dp \int dq \int_0^\infty dv_s P(p, q, s_0, v_s) v_s}{\int dp \int dq \int_{-\infty}^\infty dv_s \int_{-\infty}^\infty ds P(p, q, s, v_s)} \quad (2.76)$$

and f_s is the dynamic correction term given by

$$f_s = \frac{\int dp \int dq \int_0^\infty dv_s P(p, q, s_0, v_s) v_s \xi(p, q, v_s)}{\int dp \int dq \int_0^\infty dv_s P(p, q, s, v_s)} \quad (2.77)$$

For a canonical ensemble where H is given as the sum of kinetic and potential energy terms, eq. 2.76 can be reduced to

$$k_{TST} = \left(\frac{k_b T}{2\pi m}\right)^{1/2} \frac{\int dq \exp[-V(q, s)/k_b T]}{\int dq \int_{-\infty}^\infty ds \exp[-V(q, s)/k_b T]} \quad (2.78)$$

where $(k_b T/2\pi m)^{1/2}$ is the mean velocity.

Equation 2.78 can be rewritten in terms of a one-dimensional potential of mean force, $W(s)$, defined as [36]

$$W(s) = -k_b T \ln[g(s)/g(\infty)] \quad (2.79)$$

$$g(s) = G^{-1} \int dq \exp[-V(q, s)/k_b T] \quad (2.80)$$

where G is an arbitrary normalization factor. Substitution of eqs.

2.79 and 2.80 into eq. 2.78 leads to

$$k_{TST} = \left(\frac{k_b T}{2\pi m} \right)^{1/2} \frac{\exp[-W(s)/k_b T]}{\int_{-\infty}^{\infty} ds \exp[-W(s)/k_b T]} \quad (2.81)$$

Thus, the multi-dimensional potential $V(q,s)$ appearing in eq. 2.78 can be represented in the one dimension of the reaction coordinate, s .

In calculating k from eq. 2.73, the product $k_{TST}f_s$ is independent of the location of S , but the individual terms k_{TST} and f_s are not. The most suitable position for S depends on the rate process under investigation. For desorption, placement of S far enough away from the surface so that adsorbate-surface interactions are negligible allows f_s to be equated with the thermally-averaged sticking coefficient. For diffusion, the location for S is specified such that trajectories crossing the bridge sites between two surface atoms are normal to S . An alternative approach for locating S is used in variational transition-state theory (VTST). In this approach, the position of S is varied to obtain the minimum value of k_{TST} , i.e., the closest approximation to k . The principle consequence of using VTST is that the effect of multiple crossings of the potential barrier is minimized. From eq. 2.81, it is evident that the VTST location of S corresponds to the maximum value of $W(s)$.

A number of techniques exist to determine the recrossing factor f_s . An exact approach is to calculate f_s by forward and backward integration of trajectories initiated at the dividing surface S . Approximations to f_s have also been proposed. One

example is the unified statistical model which predicts that f_s is given by [36]

$$f_s = \frac{\exp[W^t/k_b T]}{\{\exp[W^t/k_b T] + \exp[W^{tt}/k_b T] - \exp[W^*/k_b T]\}} \quad (2.82)$$

where W^t is the higher maximum in the potential of mean force, W^{tt} is the lower maximum, and W^* is the minimum that lies between the two maxima.

The unified statistical model also provides an approximation for the sticking coefficient $S^{(\alpha)}(T)$ for adsorption into the chemisorbed state [36]

$$S^{(\alpha)}(T) = \{1 + \exp[W^l/k_b T] - \exp[W^*/k_b T]\}^{-1} \quad (2.83)$$

W^l is the potential of mean force at the inner barrier for adsorption. If a secondary maximum or minimum does not exist, then the corresponding potential of mean force W^l or W^* is simply set equal to zero in eq. 2.83.

Dynamically corrected transition-state theory can also be used to determine the diffusion coefficient for chemisorbed species. In this case, hopping of the adsorbate is treated as if it were a reaction moving the adsorbate from site i to site j. The hopping frequency Γ_{ij} is then given by

$$\Gamma_{ij} = k_{TST} f_s \quad (2.84)$$

and the diffusion coefficient is related to Γ_{ij} by

$$D = \frac{\lambda^2}{4} \Gamma_{ij} \quad (2.85)$$

Absolute Rate Theory

It is evident from the above discussion that the evaluation of

k_{TST} from eq. 2.78 requires a knowledge of $V(q,s)$. If an accurate description of $V(q,s)$ is not available, as is most often the case, then k_{TST} can be estimated using absolute rate theory as

$$k_{TST} = \frac{k_b T}{h} \frac{q_t}{q_o} \exp[-(E_t - E_o)/k_b T] \quad (2.86)$$

where q_t and q_o are the partition functions for the transition state and the reactant, E_t and E_o are the energies of the transition state and the reactant, and h is Planck's constant.

To evaluate k_{TST} using eq. 2.86, estimates must be made of q_t and q_o . An accurate determination of q_o is possible only if the reactant is in the gas phase. For a reactant in the adsorbed state, q_o can be estimated provided some reasonable assumptions of the reactant structure and mobility can be made. Estimation of q_t is much more difficult since the exact nature of the transition-state complex is unknown, and hence, its structure and mobility can not be defined with any accuracy. Nevertheless, it has been found [37-40] that reasonable estimates of the preexponential factor in k_{TST} can be made using eq. 2.86. Formulas for calculating the contributions to q_t and q_o due to translational, rotational, and vibrational degrees of freedom are given in table 2.1. Also indicated in table 2.1 are the magnitudes of each partition function for four species. In table 2.2 are shown the range of values for preexponential factors which can be accounted for by absolute rate theory. It is apparent that for some surface processes, the preexponential factor can fall in a range which spans several orders of magnitude.

In absolute rate theory, the electronic contributions to the partition functions of the reactant and transition-state complex are

factored out and expressed in the exponential factor appearing in eq. 2.86. The difference between the electronic energy states of the transition complex and the reactant, ($E_t - E_0$), is defined as the activation energy, E . Since absolute rate theory does not provide a means for determining E , this parameter must be estimated using one of the methods presented in Section 2.1.

Stochastic Diffusion Theory

Goddard and coworkers [41] have developed a theory for the desorption of atoms and molecules from surfaces based on classical stochastic diffusion theory. This approach uses a one-dimensional representation of the adsorbate-surface interaction. The flux of desorbing atoms or molecules across a plane parallel to the surface and located at s_0 is given by

$$F(s_0, u_0) = \int_{u_0}^{\infty} du P(s_0, u) u \quad (2.87)$$

where $P(s_0, u)$ is the probability of finding the adsorbate at $s=s_0$ with velocity u , and u_0 is the smallest (positive) velocity for which a particle at s_0 will desorb. $P(s_0, u)$ is determined by solving the generalized Liouville equation. The effective force acting on the adsorbate and the friction constant for energy dissipation are deduced from a representation of the adsorbate-surface interactions using a generalized Langevin equation (see eqs. 2.70-2.71). The final expression for $k_d(1)$ in the case of atomic desorption is

$$k_d^{(1)} = \frac{\Omega_0}{2\pi} \exp[-E_d/k_b T] \quad (2.88)$$

where Ω_0 is the frequency of the vibration for the adsorbate at the

bottom of the adsorbate-surface potential well and E_d is the apparent activation energy for desorption

$$E_d = V(s_0) + 1/2\mu\omega_0^2 \quad (2.89)$$

For molecular desorption, $k_d(1)$ is given by

$$k_d^{(1)} = \frac{\Omega_0}{2\pi} \left(\frac{2\mu l^2 \Omega_r^2 \gamma_0^2}{\pi k_b T} \right) \exp[-E_d/k_b T] \quad (2.90)$$

where μ is the reduced mass for frustrated rotational motion of the molecule, l is the effective length of the molecule, Ω_r is the rotational frequency, and γ_0 is the maximum bending angle for the molecule. Since the molecule has rotational modes, E_d for molecules must be modified to:

$$E_d = V(s_0) + 1/2\mu\omega_0^2 + 1/2\mu l^2 \dot{\gamma}^2 + 1/2\mu l^2 \Omega_r^2 \gamma^2 \quad (2.91)$$

where γ is the bending angle.

Quantum Mechanics

Quantum mechanical treatments of the dynamics of adsorption, diffusion, and desorption have, thus far, been limited to atoms and diatomic molecules, and in many instances adsorbate motion has been restricted to one dimension. Different theoretical approaches have been used to determine rate or diffusion coefficients. Thus, for example, dissociative adsorption of a diatomic molecule [42] has been treated by solving the time-dependent Schrödinger equation

$$\psi(r,z;t) = \psi(r,z;0) \exp[-iHt/2\pi/\hbar] \quad (2.92)$$

where $\psi(r,z;t)$ is the wave function for the adsorbate, r is the atom-atom bond distance, z is the distance of the molecule's center of mass above the plane of the adsorption surface, and H is the Hamiltonian operator. Dissociative adsorption is assumed to occur

when r exceeds a prescribed critical value. By contrast, phonon-induced desorption of atoms and molecules has been handled by determining the probability of finding an oscillator in the state n at time t , $P_n(t)$ [43,44]. The oscillators can occupy two types of states: bound states separated by discrete energy intervals or free states of continuous energy. The temporal evolution of $P_n(t)$ is governed by the master equation

$$P_n(t) = P_n(0) \exp[-Wt] \quad (2.93)$$

where W is the matrix of transition probabilities for bound-bound and bound-free transitions. The rate coefficient for desorption is then given by

$$k_d^{(1)} = \left[\sum_{n,m} (W^{-1})_{nm} P_m(0) \right]^{-1} \quad (2.94)$$

Equation 2.94 has also been used to determine the hopping frequency for surface diffusion [45]. In such a case, the elements of W are obtained using the potential function appropriate for diffusion.

Table 2.1

Translational, rotational, and vibrational partition functions for molecules [37]

Molecule	Translation $r^2 2\pi m k_b T / h^2$	Rotation $8\pi^2 k_b T I / h^2 \sigma$	Vibration $1 / (1 - \exp[-hv/k_b T])$
H ₂	33	2.9	1.0
CO	460	180	1.0
Cl ₂	1200	710	1.3
Br ₂	2600	2100	1.7

Area of a unit cell, $r^2 = 10^{-15} \text{ cm}^2$; m , I , v , and σ are the mass, moment of inertia, vibrational frequency, and symmetry number of a molecule; $T = 500 \text{ K}$.

Table 2.2

Range of the preexponential factor from absolute rate theory [37]

Process	Preexponential factor	Units
Molecular adsorption	$10 \cdot 10^{-10} \cdot 10^{-17}$	$\text{cm}^3 \text{ s}^{-1}$
Dissociative adsorption	$10 \cdot 10^{-10} \cdot 10^{-17}$	$\text{cm}^3 \text{ s}^{-1}$
Molecular desorption	$10^{13} \cdot 10^{19}$	s^{-1}
Associative desorption	$10^{-4} \cdot 10^4$	$\text{cm}^2 \text{ s}^{-1}$
Langmuir-Hinshelwood reaction	$10^{-4} \cdot 10^4$	$\text{cm}^2 \text{ s}^{-1}$
Eley-Rideal reaction	$10^{-6} \cdot 10^{-17}$	$\text{cm}^3 \text{ s}^{-1}$
Unimolecular reaction	$10^{12} \cdot 10^{13}$	s^{-1}
Surface diffusion	$10^{-2} \cdot 10^{-4}$	$\text{cm}^2 \text{ s}^{-1}$

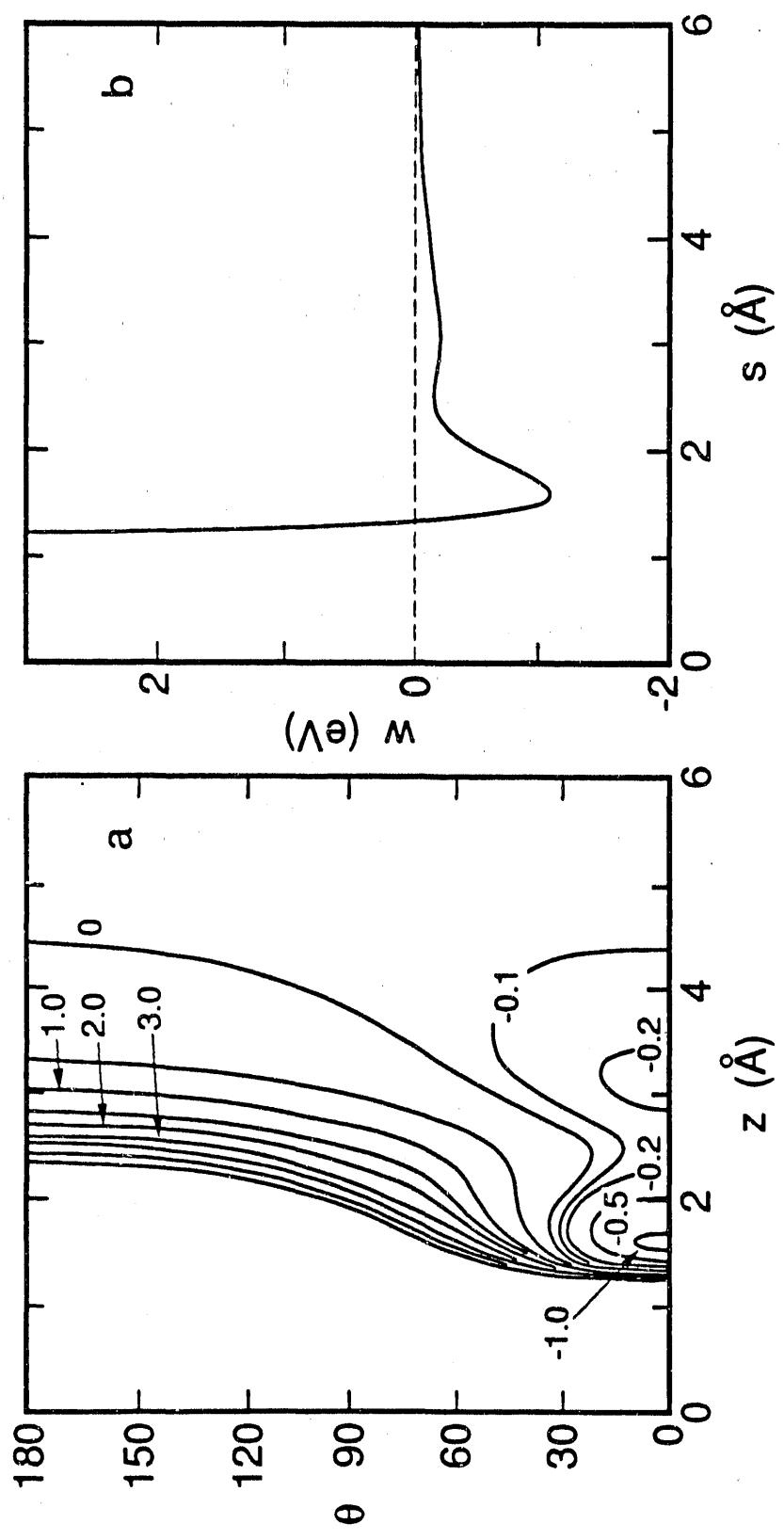


Fig. 2.1 a. Potential energy surface for molecular adsorption as a function of orientation angle θ and distance from the surface z . The energy contours are in eV. b. One-dimensional potential versus reaction coordinate s [36].

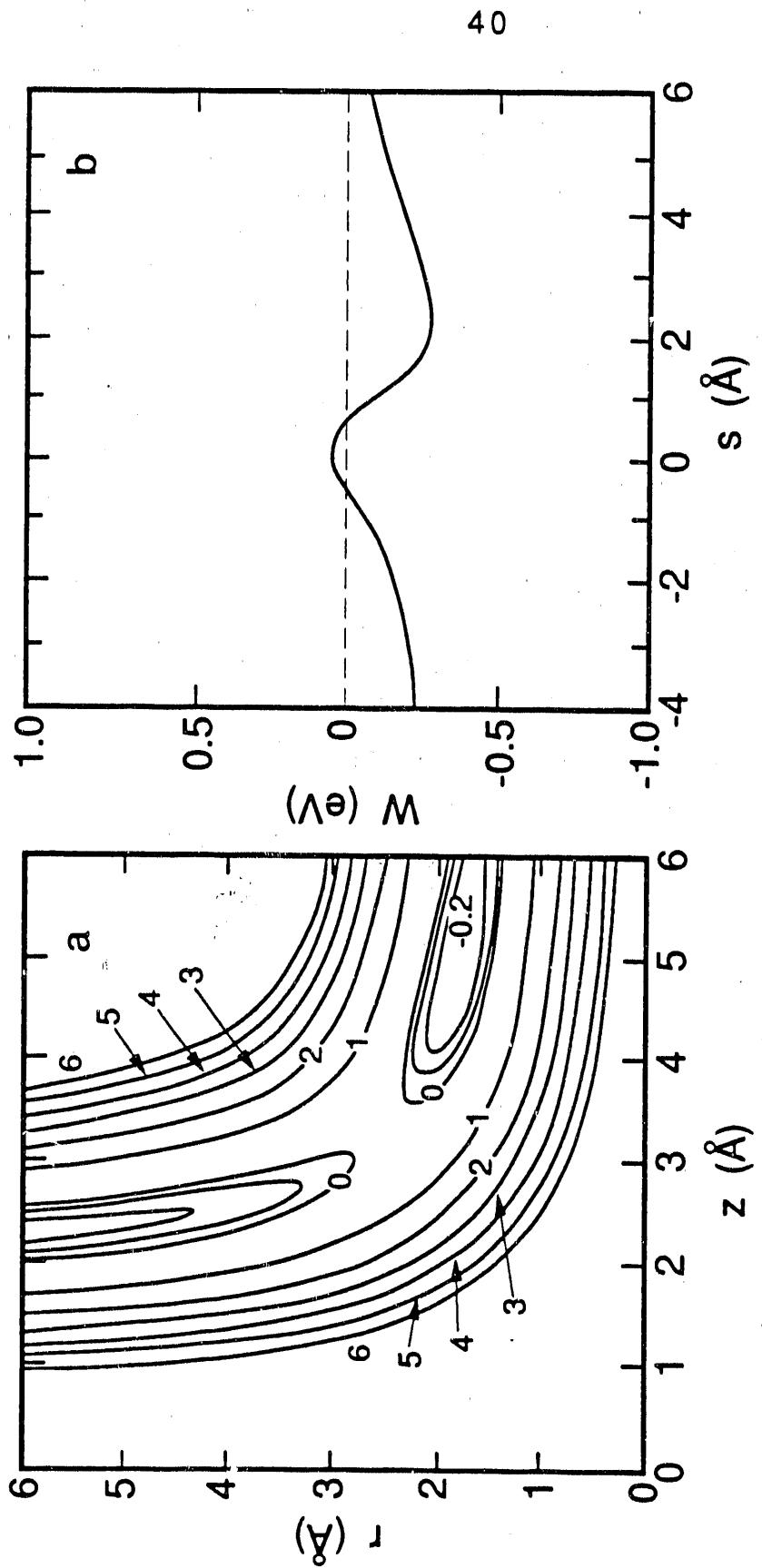


Fig. 2.2 a. Potential energy surface for dissociative adsorption as a function of distance from the surface z and internal bond distance r . The energy contours are in eV. b. One-dimensional potential versus reaction coordinate s [36].

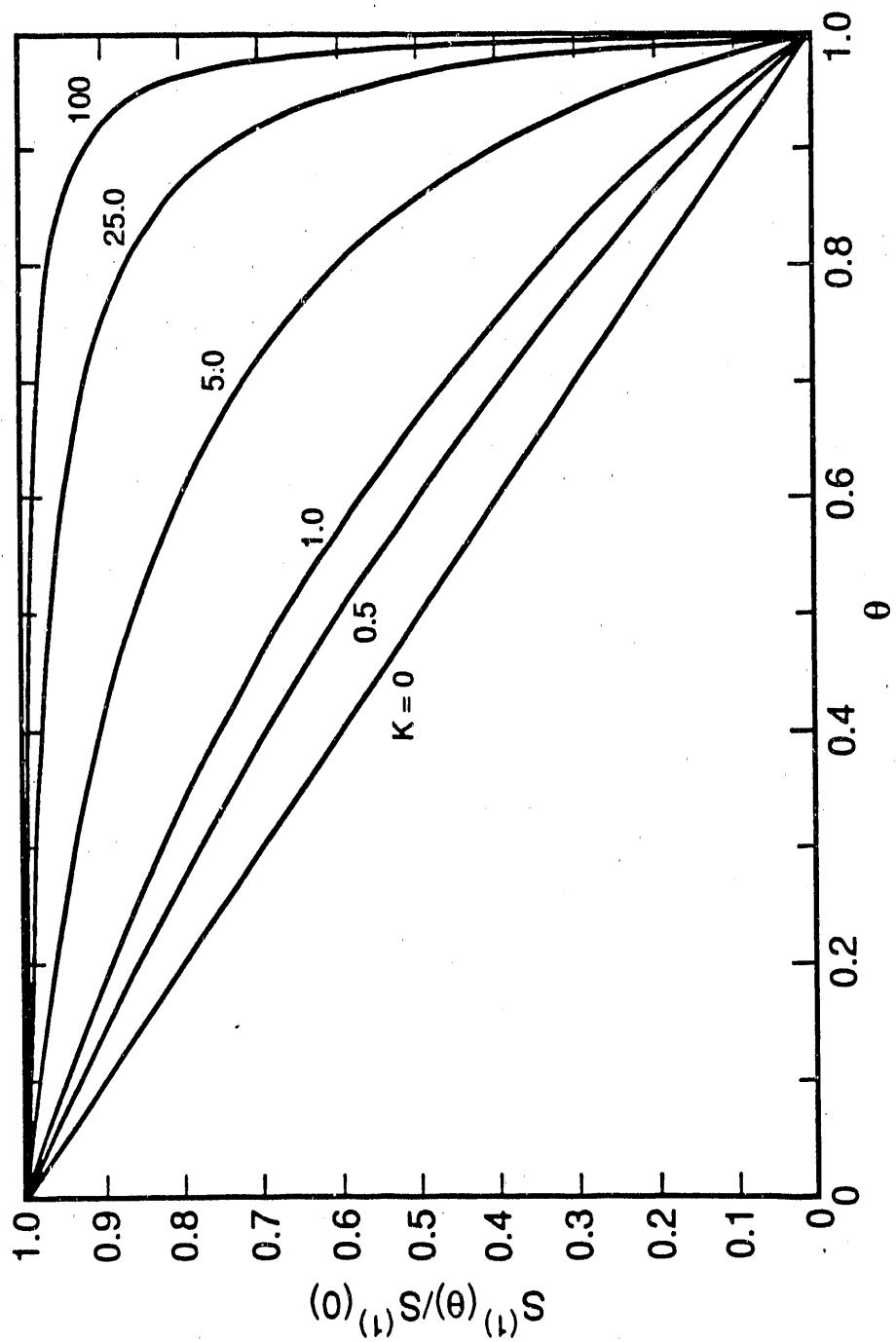


Fig. 2.3 Variation of $S(1)(\theta)/S(1)(0)$ with θ for adsorption via a precursor mechanism. The precursor parameter $K = k_a/k_d$.

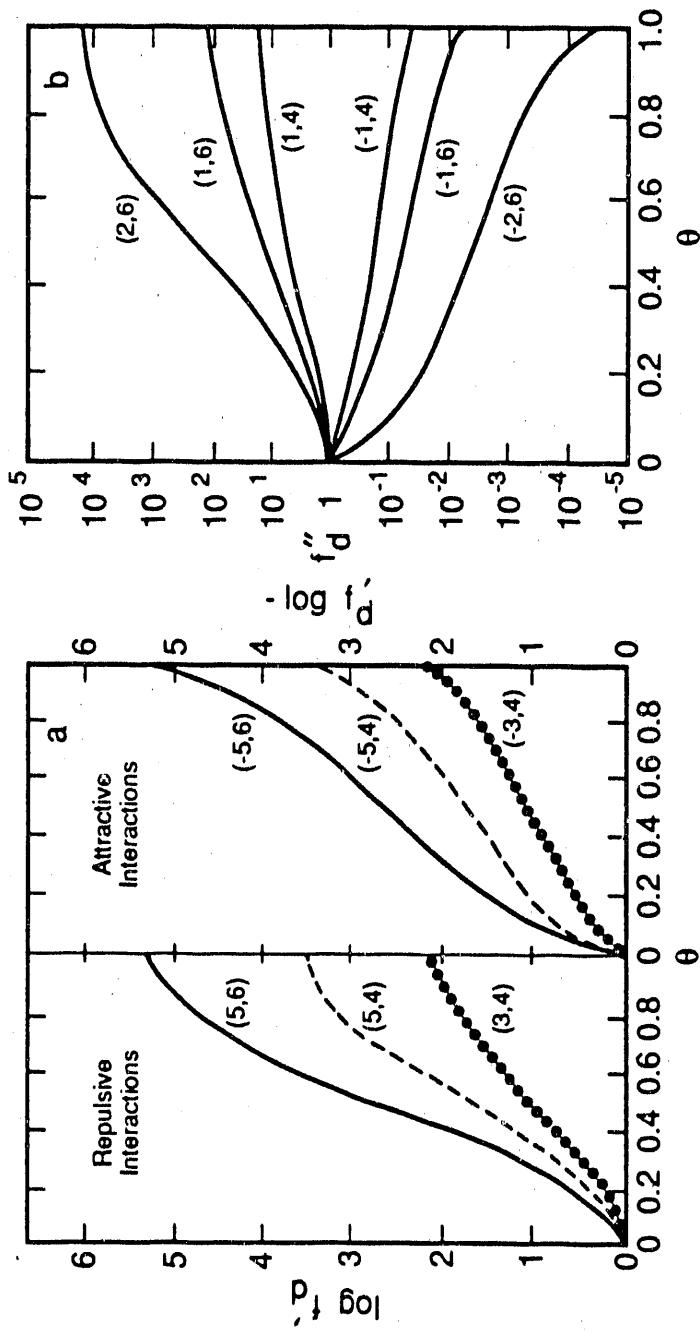


Fig. 2.4 a. Variation in f_d' with θ at $T = 300$ K [26]. The nomenclature (5,6) corresponds to $WAA = 5$ kJ/mol and $z = 6$, respectively. b. Variation in f_d'' with θ [24]. The nomenclature (2,6) corresponds to $WAA = 2k_bT$ and $z = 6$, respectively.

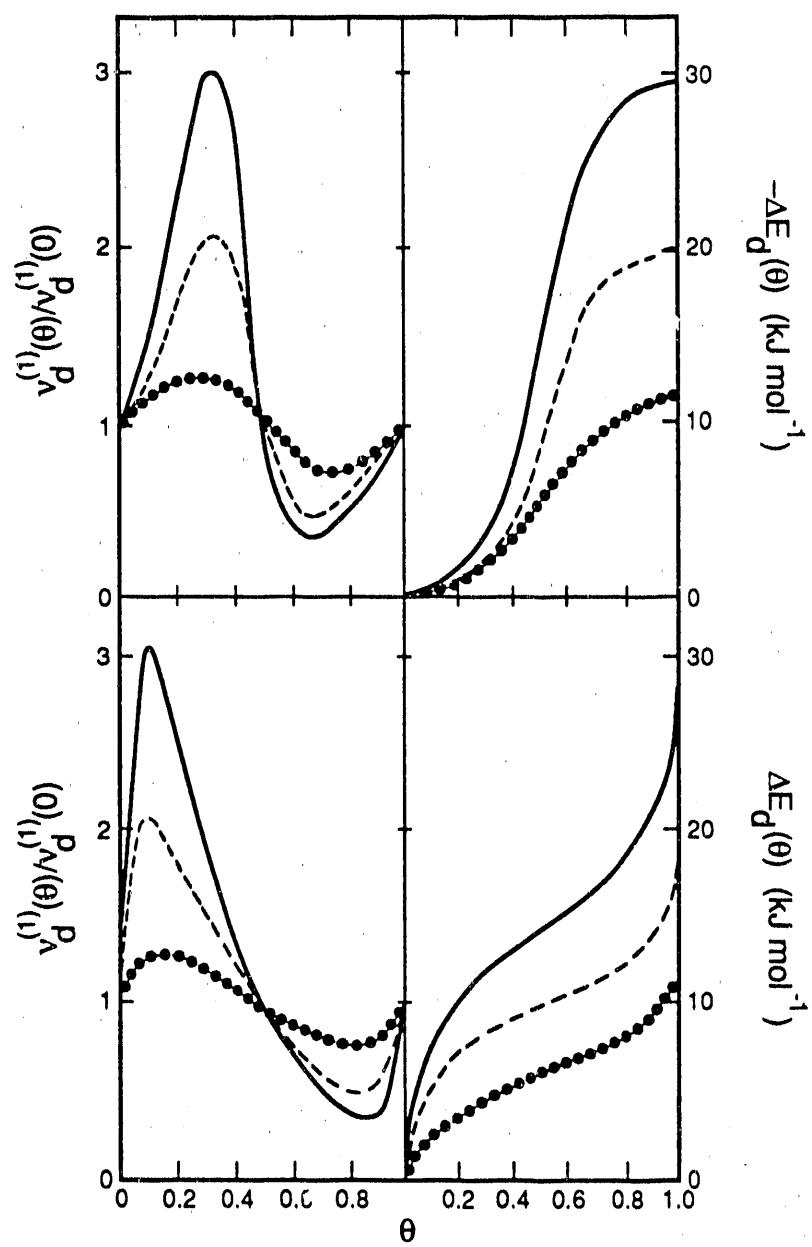


Fig. 2.5 Variation of $v_d(1)(\theta)/v_d(1)(0)$ and $\Delta E_d(\theta)$ with θ for repulsive (top panels) and attractive (bottom panels) interactions [26]. Solid line: $|WAA| = 5 \text{ kJ/mol}$, $z = 6$; dashed line: $|WAA| = 5 \text{ kJ/mol}$, $z = 4$; dotted line: $|WAA| = 3 \text{ kJ/mol}$, $z = 6$.

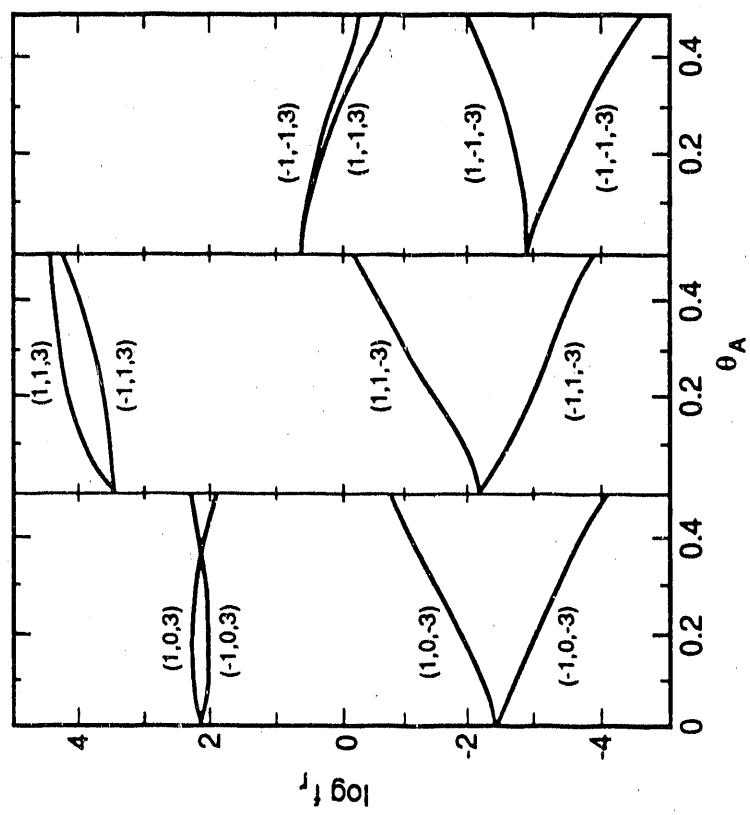


Fig. 2.6 Variation in f_r with θ [25]. The nomenclature $(1,0,3)$ corresponds to $w_{AA} = k_b T$, $w_{AB} = 0$, and $w_{BB} = 3k_b T$, respectively.

3.0 ADSORPTION

3.1 Molecular Dynamics

Arumainayagam et al. [46] have used a Langevin model to calculate $S(1)(0)$ for Xe adsorption on Pt(111). Morse potentials were used to represent the Xe-surface interactions. During the course of a trajectory, adsorbate sticking was assumed to occur when the Xe-surface interactions fell below $-2k_bT$. As seen in fig. 3.1, the level of agreement between the experimental and predicted values of $S(1)(0)$ is quite high. The authors noted that this level of agreement could only be obtained by adjusting the parameters of the Morse potential responsible for the steepness of the repulsive wall and the amount of surface corrugation. The observed reduction in $S(1)(0)$ with increasing adsorbate kinetic energy was attributed to inefficient energy transfer between the adsorbate and the surface.

Using an approach similar to that described above, Muhlhausen et al. [47] have determined $S(1)(0)$ for NO adsorption on the (111) faces of Pt and Ag. The NO-surface potential included a term to account for the orientation of NO relative to the surface. For both surfaces, $S(1)(0)$ was determined to be ~0.7 at 300 K and to decrease monotonically to ~0.08 at 2000 K. The dependence of $S(1)(0)$ on the incident kinetic energy of NO was also examined and $S(1)(0)$ was observed to decrease with increasing kinetic energy of NO.

The sticking coefficient of Si on Si(100) has been examined by NoorBatcha et al. [48]. The adsorbate-surface potential was described by adatom-surface Morse potentials and a Keating

potential which accounts for the bending and stretching modes of the lattice. The criterion for adsorbate sticking was an attractive interaction for a duration of more than five vibrational periods. The value of $S(1)(0)$ was determined to be 0.96 at 1500 K and to be relatively insensitive to temperature.

Molecular dynamics simulations of the dissociative adsorption of H_2 on $Cu(100)$ have been carried out by Gelb and Cardillo [30,31]. A LEPS potential energy surface was used to describe the interactions of H_2 and H with the metal surface. Dissociative adsorption was assumed to occur when the H_2 internuclear separation distance exceeded 3.5 Å. As shown in fig. 3.2, their calculations predict that $S(2)(0)$ passes through a maximum with increasing kinetic energy of the H_2 molecule. The increase in $S(2)(0)$ observed at low kinetic energies is attributed to an increase in the probability of surmounting the activation barrier for dissociative adsorption. For very large values of the kinetic energy, however, the time of interaction of the adsorbate with the surface is so small as to preclude sufficient rearrangement of the H_2 molecules into configurations favorable for dissociation, and hence, the value of $S(2)(0)$ decreases as the kinetic energy increases.

Lee and DePristo [33-35] have reported on the dissociation kinetics of H_2 on Ni and Cu crystal faces. A generalized Langevin model was used to simulate adsorbate trajectories and a LEPS potential energy surface was used to describe the H_2 - and H-metal interactions. Dissociative adsorption was assumed to occur when the H_2 internuclear separation distance exceeded 2.8 Å. A comparison of the predicted and experimental values of $S(2)(0)$ as a

function of incident kinetic energy of the adsorbate is given in tables 3.1 and 3.2. For the Cu surfaces, both the predicted values of $S^{(2)}(0)$ and the dependence of these values on kinetic energy are in fair agreement with the experimental observation [49,50], whereas for the Ni surfaces, the predicted values of $S^{(2)}(0)$ and the dependence on kinetic energy do not agree well with the experimental data [51,52]. Using the same approach, Kara and DePristo [53] have calculated $S^{(2)}(0)$ for N_2 adsorption on W(110). As can be seen in fig. 3.3, the agreement between theory and experiment [54] is very good in this case.

3.2 Transition-State Theory

The influence of precursor states on molecular adsorption has been examined by Doren and Tully [36] using dynamic TST. In their model, adsorbate-surface interactions were described by a one-dimensional potential of mean force $W(z)$, which depends only on z , the distance of the molecular center of mass from the surface. The presence of a precursor state was defined by the occurrence of a secondary minimum in $W(z)$, located between the chemisorbed state and z approaching infinity. As illustrated in 3.4a for the case of CO adsorption on Ni(111), the depth of the well for the precursor state decreases with increasing temperature, whereas the height of the activation barrier for dissociative adsorption increases. Figure 3.4b shows that the sticking coefficient for molecular adsorption of CO first decreases and then gradually increases with increasing temperature. This behavior can be explained as follows. At low

temperatures, the kinetic energy of the impinging adsorbate is small and consequently very little energy transfer is required for the adsorbate to be trapped into the precursor state. Since the barrier for subsequent trapping into the adsorbate state is small at low temperatures, adsorption into that state proceeds rapidly and hence $S(1)(0)$ is large. At intermediate temperatures, the activation energy barrier becomes more pronounced (see fig. 3.4a), and the inability of adsorbates to surmount this higher barrier leads to lower values of the sticking coefficient. At high temperatures, the kinetic energy of the incident molecules is now sufficient to directly overcome the activation energy barrier without initial trapping into the precursor state. The behavior observed at low and high temperatures is referred to in the literature as precursor-mediated adsorption and direct adsorption, respectively.

Truong et al. [55] have used a variational TST approach which included quantum effects to determine the activation energy for dissociative adsorption, E_{diss} , of H₂ and D₂ on Ni(100), (111), and (110) surfaces. The adsorbate-surface interactions were described by a LEPS potential energy surface. To account for the quantum effects of tunneling and reflection, k_{TST} was multiplied by a ground-state transmission coefficient. The apparent activation energies are listed in table 3.3. As can be seen, the value of E_{diss} depends on the crystal face, the temperature, and the molecular weight of the adsorbate. For H₂ adsorption on Ni(100), E_{diss} is negative up to 500 K, indicating that the classical barrier to desorption lies below the reactant energy. On Ni(110), the value of E_{diss} is less than or equal to 0.4 kcal/mol below 500 K, whereas on Ni(111), it is about 1.0

kcal/mol below 500 K. The complex dependence of E_{diss} with temperature for both H₂ and D₂ was partially attributed to the effects of tunneling.

Adams and Doll [56] have also used a dynamic TST model to examine the influence of precursor states on the sticking coefficient for dissociatively adsorbed species. Their calculations suggest that S(2)(0) decreases with increasing temperature and is sensitive to the well depth in the precursor state and the rate of energy dissipation in the adsorbed state.

3.3 Quantum Mechanical Models

A quantum mechanical description of H₂ adsorption on Ni(100) has been developed by Jackson and Metiu [42]. The interactions of H₂ and H with the metal surface were described by a LEPS potential energy surface. The principal features of the energy surface were the barriers for molecular adsorption, E_a , dissociation, E_{diss} , and atomic diffusion, E_{diff} . The adsorption behavior was found to depend on the relative magnitudes of E_a and E_{diss} . When $E_a = 0$ and $E_{diss} < E_{diff}$ (restricted adsorbate mobility), the dissociation probability S(2)(0) was larger than when $E_a = 0$ and $E_{diss} > E_{diff}$ (mobile adsorbate). This behavior is attributed to the fact that restriction of adsorbate mobility increases the probability for H-atom recombination and hence, lowers the value of S(2)(0). As can be seen in fig. 3.5, S(2)(0) increases with increasing kinetic energy and is smaller for the larger values of E_{diss} . The quantum nature of the dissociation process is especially evident for $E_{diss} = -4.423$ eV. Classically,

dissociation should occur with unit probability for all kinetic energies, since E_{diss} is below the zero-point energy of H_2 . The process of quantum reflection upon traversing a potential change, however, leads to much lower values of $S(2)(0)$.

Using an approach similar to the one described above, Chiang and Jackson [32] have investigated the isotope effect for adsorption of H_2 , D_2 and T_2 on $Ni(100)$. Their calculations showed that for a fixed value of the incident kinetic energy, $S(2)(0)$ decreases with increasing molecular weight. This behavior arises from the influence of the zero-point energy on the apparent barrier to dissociation. In qualitative agreement with this prediction, Hamza and Madix [50] have observed larger values of $S(2)(0)$ for H_2 dissociation on $Ni(100)$ than for D_2 .

Asscher et al. [57] have calculated $S(2)(0)$ for N_2 on $Re(0001)$ and $Fe(111)$. The principal feature of the potential energy surface was an activation energy barrier to dissociation $E_{diss} = 73$ kJ/mol. It was shown that $S(2)(0)$ increases exponentially with increasing incident kinetic energy, in agreement with the experimental results for N_2 on $Fe(111)$ [58]. The low values of $S(2)(0)$ observed for low initial kinetic energies were attributed to the large activation energy for dissociation. It is interesting to note that classically, non-zero values of $S(2)(0)$ would not have been expected at low kinetic energies for such a large value of E_{diss} ; the observation of a finite value of $S(2)(0)$ is thus attributable to tunneling of the molecular wave function through the barrier.

3.4 Effects of Lateral Interactions

King and Wells [59] have developed a lattice-gas model to treat the influence of lateral interactions on the rate of dissociative adsorption. For adsorption to occur, it was assumed that a pair of unoccupied nearest-neighbor sites must exist on the surface. The sticking coefficient for adsorption is thus given as $S(2)(\theta) = S(2)(0)P_{vv}$ where P_{vv} is given by eq. 2.41. A plot of $S(2)(\theta)/S(2)(0)$ is shown in fig. 3.6 as a function of the parameter $\eta/2 = [1 - \exp(-W_{AA}/k_b T)]$. At low coverages, $S(2)(\theta)/S(2)(0)$ is practically independent of $\eta/2$, whereas at high coverage, $S(2)(\theta)/S(2)(0)$ decreases more rapidly with increasing $\eta/2$. The values of W_{AA} for $\eta/2$ evaluated at $T = 300$ K are also shown in fig. 3.6 and it can be seen that $S(2)(\theta)/S(2)(0)$ is not a strong function of W_{AA} . It is also interesting to note some of the limiting forms of $S(2)(\theta)/S(2)(0)$. When $W_{AA} = 0$, $S(2)(\theta)/S(2)(0) = (1-\theta)^2$ which is equivalent to the expression for non-activated adsorption for randomly distributed adsorbates. For large repulsive interactions ($W_{AA} \rightarrow \infty$), $S(2)(\theta)/S(2)(0) = 1-2\theta$ for $\theta \leq 0.5$, which represents pseudo first-order dissociative adsorption kinetics.

Zhdanov [60] has examined the influence of lateral interactions on the adsorption rate using a lattice-gas model. The sticking coefficient for adsorption was determined from absolute rate theory. For non-dissociative adsorption, the rate was given by

$$r_a = \frac{k_b T}{h} \sum_n \frac{q_A^\dagger}{q_A} P_{V,n} \exp\left[\frac{-E_a^0}{k_b T}\right] (1-\theta) N_{A,g} \quad (3.1)$$

and the ratio of $S^{(1)}(\theta)/S^{(1)}(0)$ by

$$\frac{S^{(1)}(\theta)}{S^{(1)}(0)} = \sum_n \frac{q_A^t}{q_A} P_{V,n} (1-\theta) \quad (3.2)$$

Figure 3.7 shows curves of $S^{(1)}(\theta)/S^{(1)}(0)$ for attractive, repulsive and no interactions, when $q_A t/q_A = 0.3n$. For a fixed value of θ , $S^{(1)}(\theta)/S^{(1)}(0)$ is largest for attractive interactions. $S^{(1)}(\theta)/S^{(1)}(0)$ is also seen to decrease non-linearly with coverage, independent of the sign of W_{AA} .

The influence of lateral interactions on adsorption kinetics has been investigated for the case of dissociative adsorption via a precursor mechanism. King and Wells [59] have shown that when intrinsic and extrinsic precursor states exist, $S^{(2)}(\theta)/S^{(2)}(0)$ can be written as

$$\frac{S^{(2)}(\theta)}{S^{(2)}(0)} = \frac{1}{1 + L \left(\frac{1}{P_W} - 1 \right)} \quad (3.3)$$

where $L \propto r_d^*/(r_a^{**} + r_d^{**})$. The desorption rates from extrinsic and intrinsic precursor states are denoted by r_d^* and r_d^{**} , respectively, and the adsorption rate from the intrinsic precursor state is denoted by r_a^{**} . If $r_a^{**} \gg r_d^*, r_d^{**}$, then $L \rightarrow 0$, whereas when $r_a^{**} \ll r_d^*, r_d^{**}$, then $L \rightarrow 1$. L is thus a measure of the relative importance of the adsorption and desorption processes. A plot of $S^{(2)}(\theta)/S^{(2)}(0)$ parametric in L for a fixed value of $\eta/2$ is shown in fig. 3.8a. When $L \rightarrow 0$, $S^{(2)}(\theta)/S^{(2)}(0)$ is larger than when $L \rightarrow 1$. Figure 3.8b show a plot of $S^{(2)}(\theta)/S^{(2)}(0)$ parametric in $\eta/2$ for a fixed value of L . At

high coverage, $S(2)(\theta)/S(2)(0)$ is strongly dependent on the magnitude of WAA and decreases with increasingly repulsive values of WAA. It is significant to note that the shape and magnitude of the plots of $S(2)(\theta)/S(2)(0)$ differ considerably depending on whether precursor states do (see figs. 3.8a and 3.8b) or do not exist (see fig. 3.6).

King and Wells [59] have used eq. 3.3 to describe the dissociative adsorption of N_2 on a W(100) surface. Values of the parameters required to fit eq. 3.3 to the experimental data are listed in table 3.4, and a comparison of the fitted and experimental results is shown in fig. 3.9. Good agreement is observed at both low and high temperatures. It is interesting to note that L increases with increasing temperature, a trend which reflects the temperature dependence of the desorption rate.

The adsorption of molecular N_2 on Ru(001) has been simulated by Hood et al. [61] using a Monte Carlo model. Adsorption was assumed to proceed via a precursor mechanism in which both intrinsic and extrinsic precursor sites could be occupied. Repulsive nearest-neighbor interactions between precursor and chemisorbed species were taken to be 0.25 kcal/mol and attractive next-nearest neighbor interactions were taken to be 0.45 kcal/mol. As seen in fig. 3.10, the predicted variation in $S(1)(\theta)$ is in fair agreement with the experimentally observed trend [62,63]. The initial rise in $S(1)(\theta)$ is attributed to the formation of energetically favorable islands in which molecules are arranged in $\sqrt{3} \times \sqrt{3}$ R30° domains, whereas the decrease in $S(1)(\theta)$ at high coverage is attributed to crowding effects.

Table 3.1

Dissociative sticking probability of hydrogen on Cu surfaces for $T_s=1000$ K [33]

Kinetic Energy(eV)	Cu(110)		Cu(100)	
	Calc.	Exp. ^a	Calc.	Exp. ^a
0.15	0.03	0.08	<0.01	0.03
0.20	0.06	0.115	0.01	0.045
0.30	0.11	0.135	0.06	0.095
0.40	0.15	0.14	0.10	0.100
0.60	0.27	-	0.27	-

^aFrom Refs. [49,50].

Table 3.2

Dissociative sticking probability of hydrogen on Ni surfaces for $T_s=300$ K [33]

Kinetic Energy(eV)	Ni(100)		Ni(110)		Ni(111)	
	Calc.	Exp. ^a	Calc.	Exp. ^b	Calc.	Exp. ^b
0.03	0.17	0.4	0.55	0.96	0.09	0.02
0.05	0.18	0.6	0.52	0.96	0.08	0.07
0.07	0.18	0.7	0.52	0.96	0.09	0.10
0.10	0.20	0.8	0.52	-	0.09	0.27
0.14	0.20	-	0.52	-	0.10	-
0.20	0.20	-	0.50	-	0.12	-

^aFrom Refs. [51].^bFrom Refs. [52].

Table 3.3

Activation energy (kcal/mol) for dissociative adsorption of hydrogen and deuterium on Ni surfaces [55]

T(K)	Ni(100)		Ni(110)		Ni(111)	
	H ₂	D ₂	H ₂	D ₂	H ₂	D ₂
100-140	-0.87	-1.31	0.29	0.10	0.99	0.61
140-200	-1.23	-1.02	0.35	0.09	0.92	0.63
200-300	-1.01	-0.67	0.40	0.05	0.86	0.70
300-500	-0.71	-0.30	0.37	-0.05	0.94	1.11
500-800	0.16	0.54	-0.28	-0.29	1.59	2.21

Table 3.4

Values of the parameters used to represent the data in fig. 3.9 [59]

T _s (K)	N _s (10 ¹⁴ cm ⁻²)	L	$\eta/2$	S(2)(0)
300	9.5	0.082	0.989	0.585
433	10.0	0.157	0.987	0.49
663	10.5	0.256	0.977	0.31
773	9.5	0.517	0.986	0.21

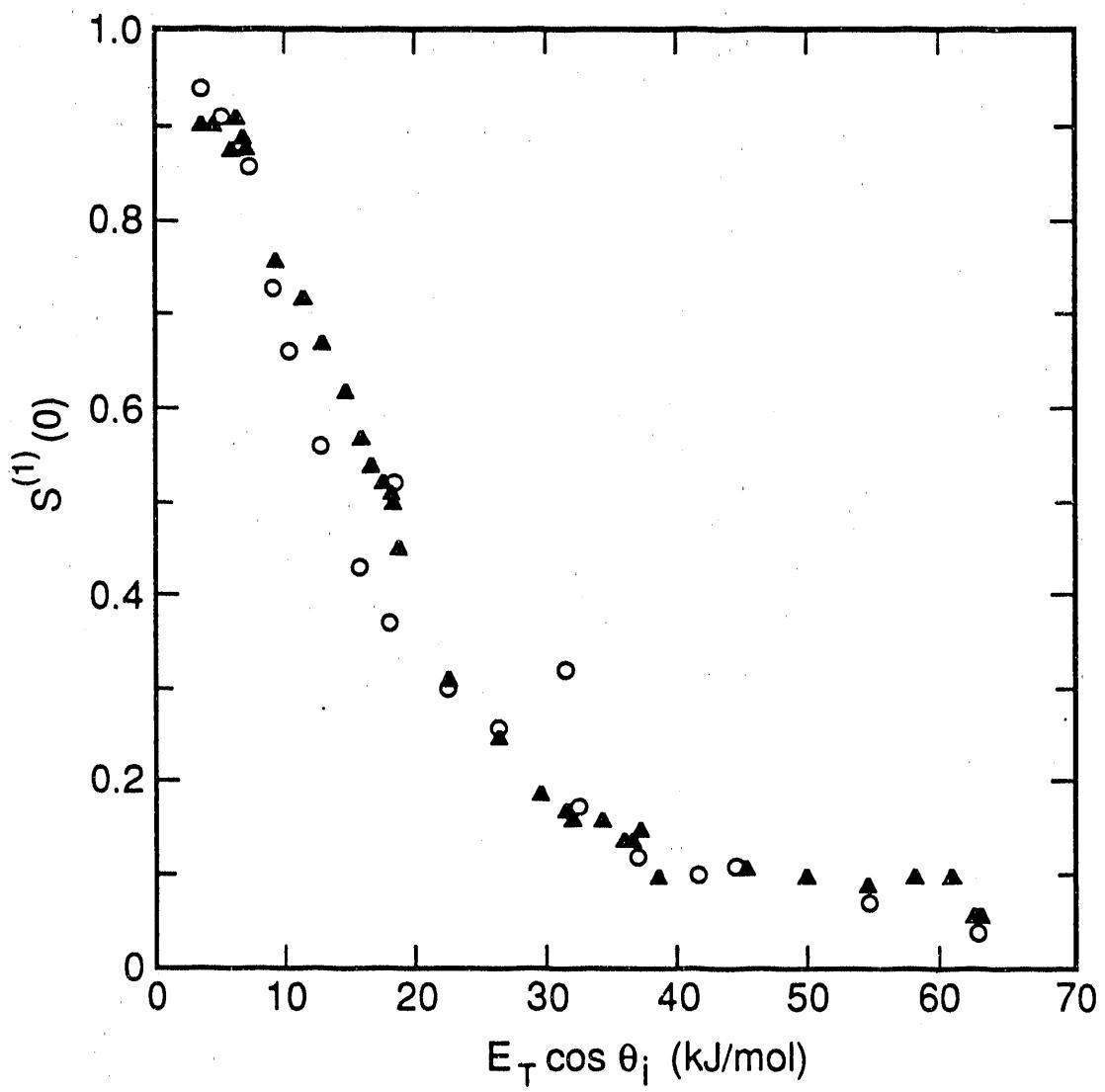


Fig. 3.1 Simulated (O) and experimental (▲) variation of $S^{(1)}(0)$ versus $E_T \cos \theta_i$ for Xe adsorption on Pt(111) [46]. E_T is the incident kinetic energy and θ_i is the incident angle measured from the surface normal.

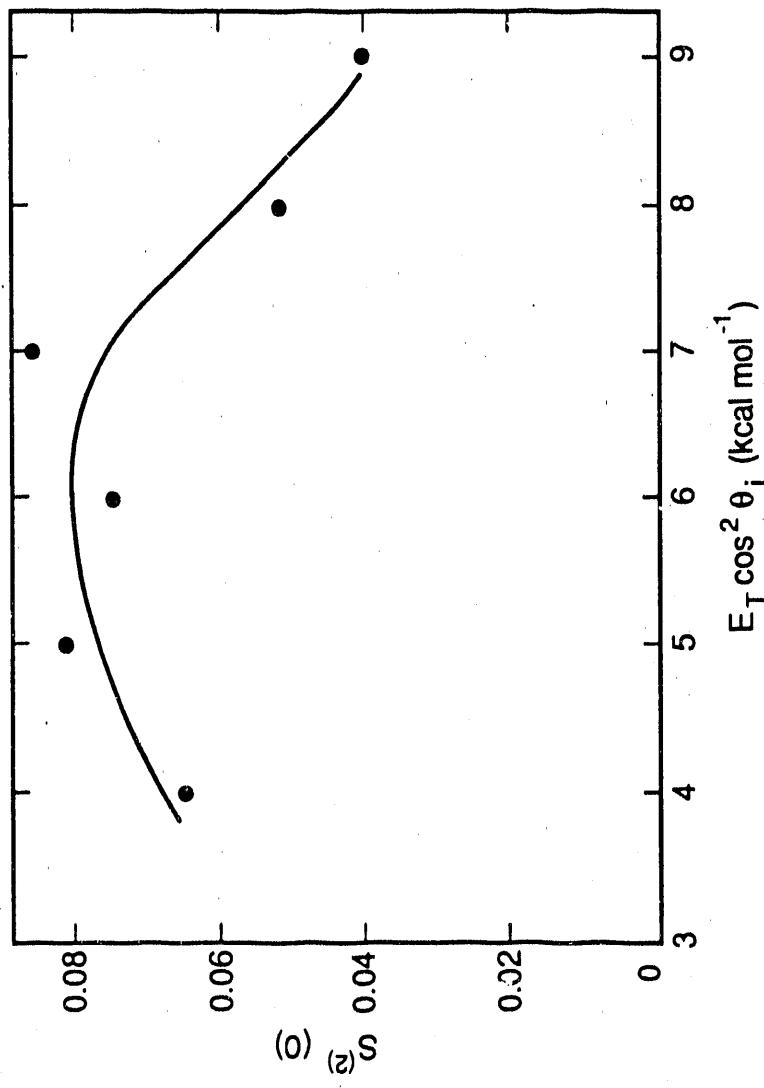


Fig. 3.2 Simulated (●) variation of $S(2)(0)$ versus $E_T \cos^2 \theta_i$ for H_2 adsorption on $Cu(100)$ [31]. E_T is the incident kinetic energy and θ_i is the incident angle measured from the surface normal. The solid line is a guide to the eye.

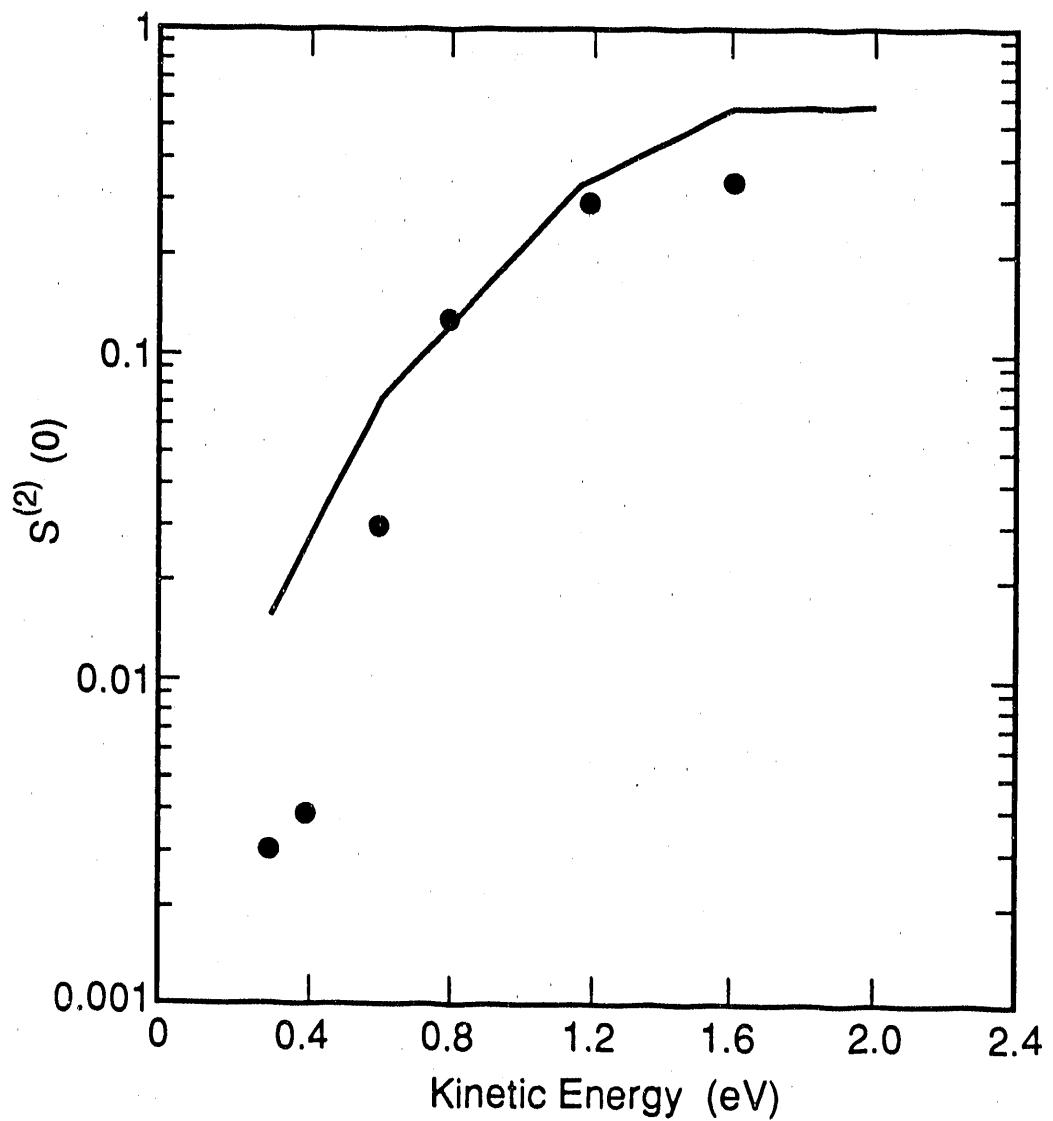


Fig. 3.3 Simulated (solid line) [53] and experimental (circles) [54] variation of $S(2)(0)$ versus incident kinetic energy for N_2 adsorption on $W(110)$.

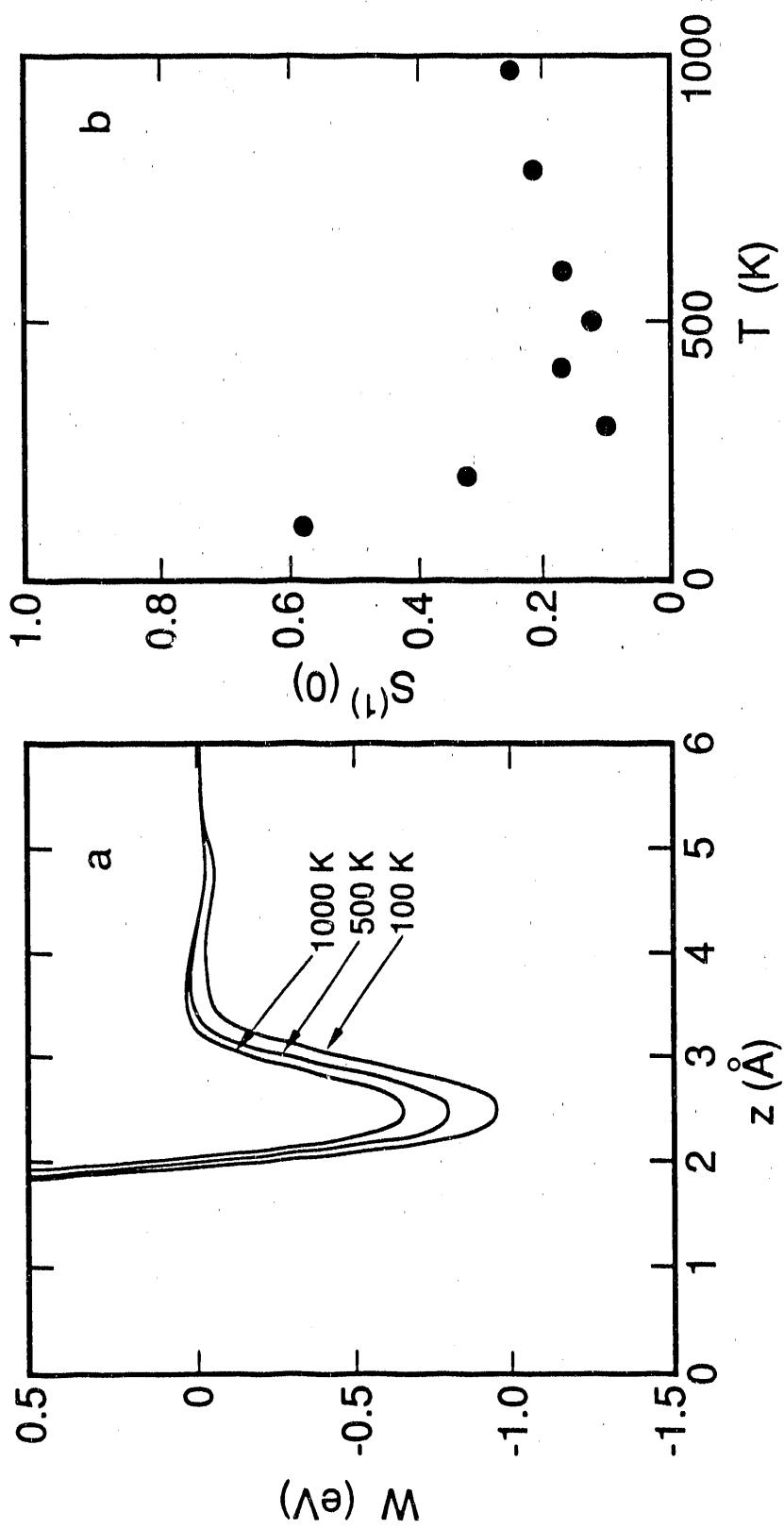


Fig. 3.4 a. Potential of mean force versus height above the surface for CO on Ni(111). b. Simulated variation of $S(1)(0)$ versus T for CO adsorption on Ni(111) [36].

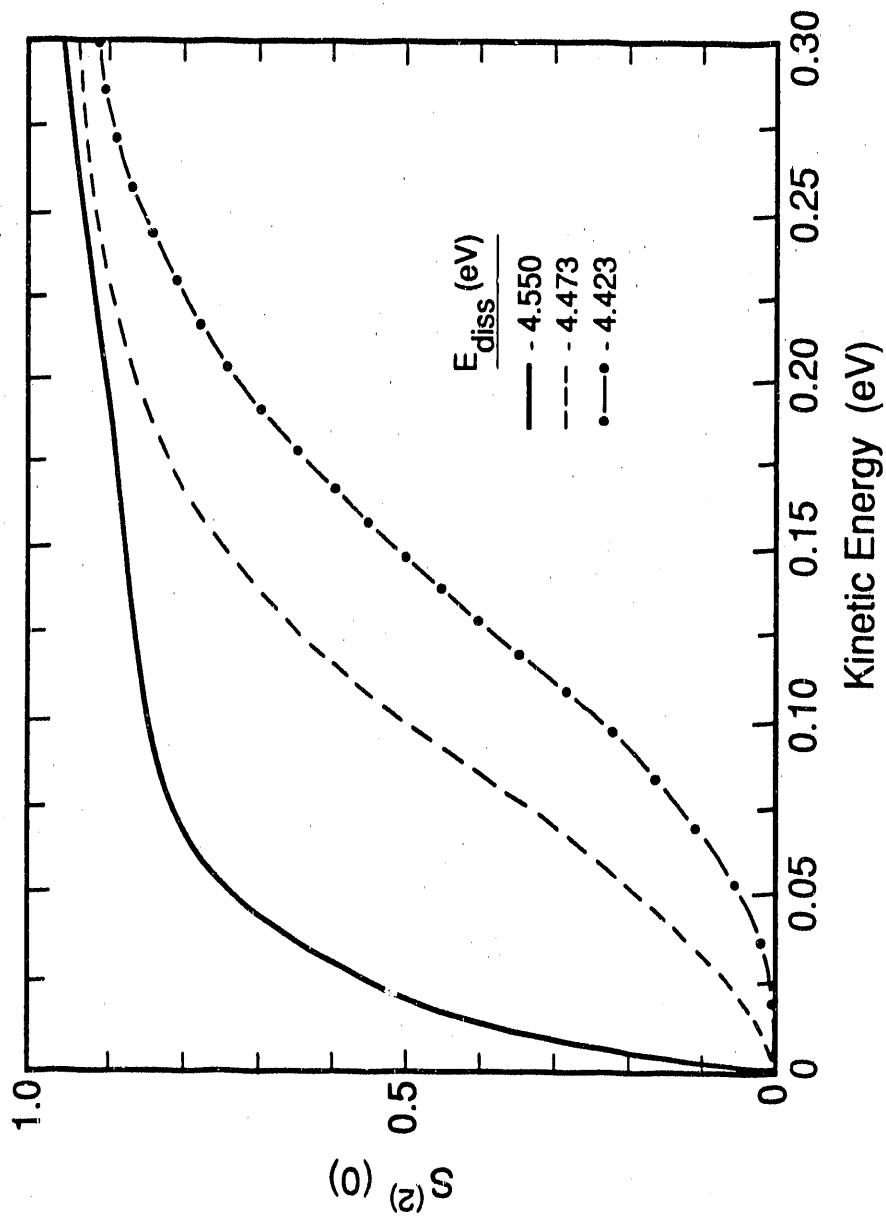


Fig. 3.5 Simulated variation of $S(2)(0)$ versus incident kinetic energy for H_2 adsorption on $Ni(100)$ [42].

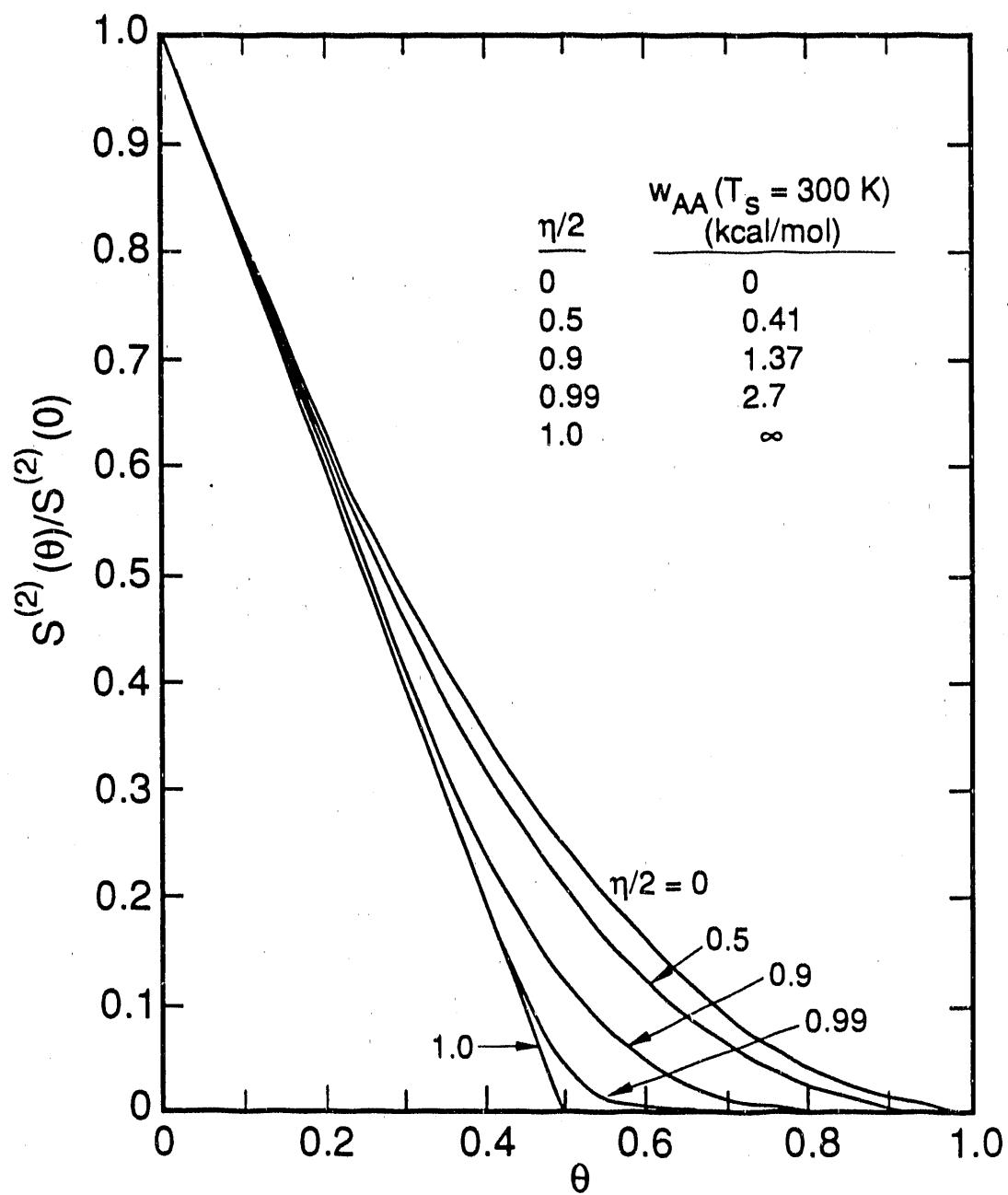


Fig. 3.6 Variation in $S^{(2)}(\theta)/S^{(2)}(0)$ versus θ for repulsive nearest-neighbor interactions [59].

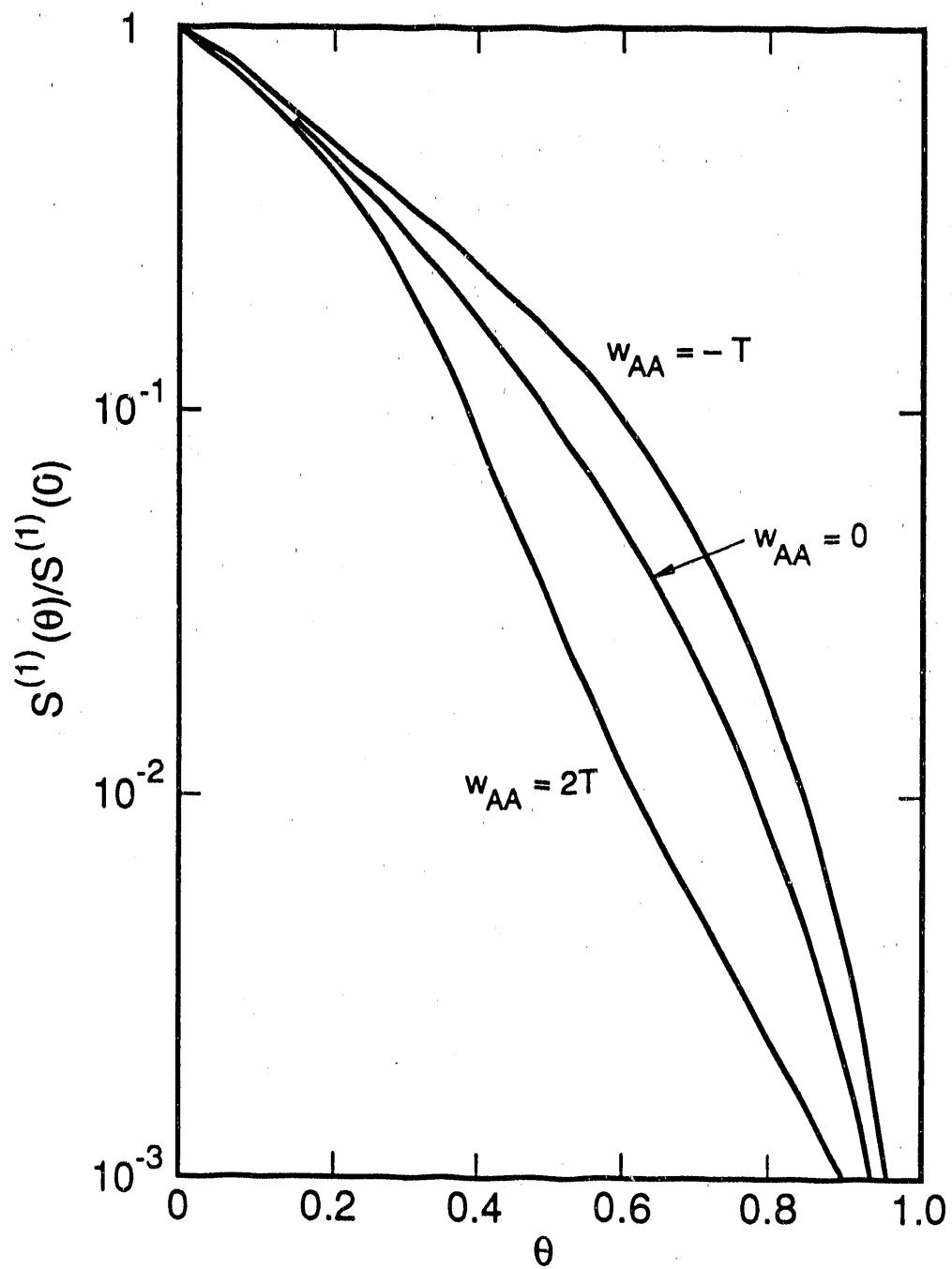


Fig. 3.7 Variation in $S^{(1)}(\theta)/S^{(1)}(0)$ versus θ when nearest-neighbor interactions occur [60].

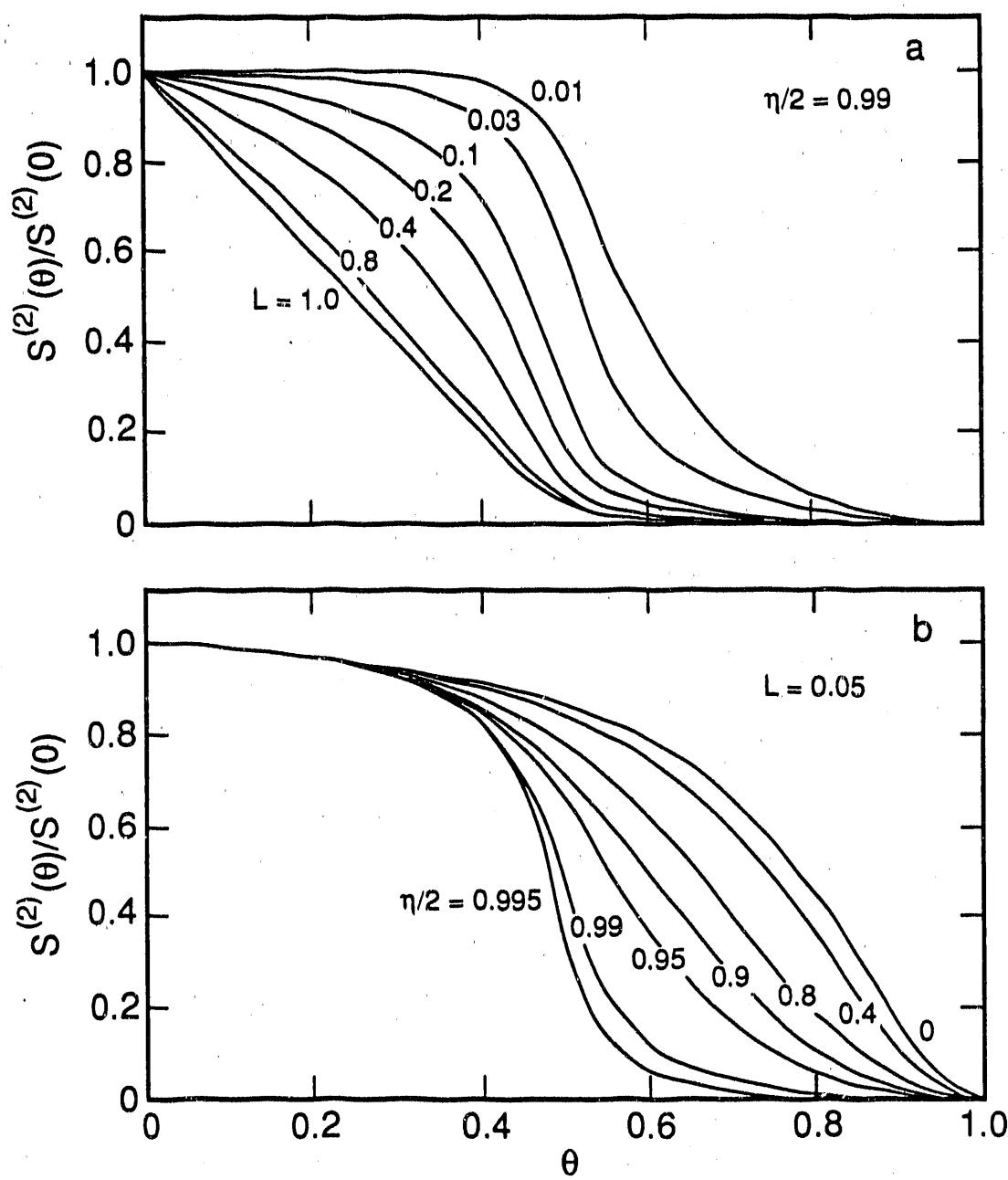


Fig. 3.8 a. Variation in $S^{(2)}(\theta)/S^{(2)}(0)$ versus θ when precursor states exist for a fixed value of η . b. Variation in $S^{(2)}(\theta)/S^{(2)}(0)$ versus θ when precursor states exist for a fixed value of L [59].

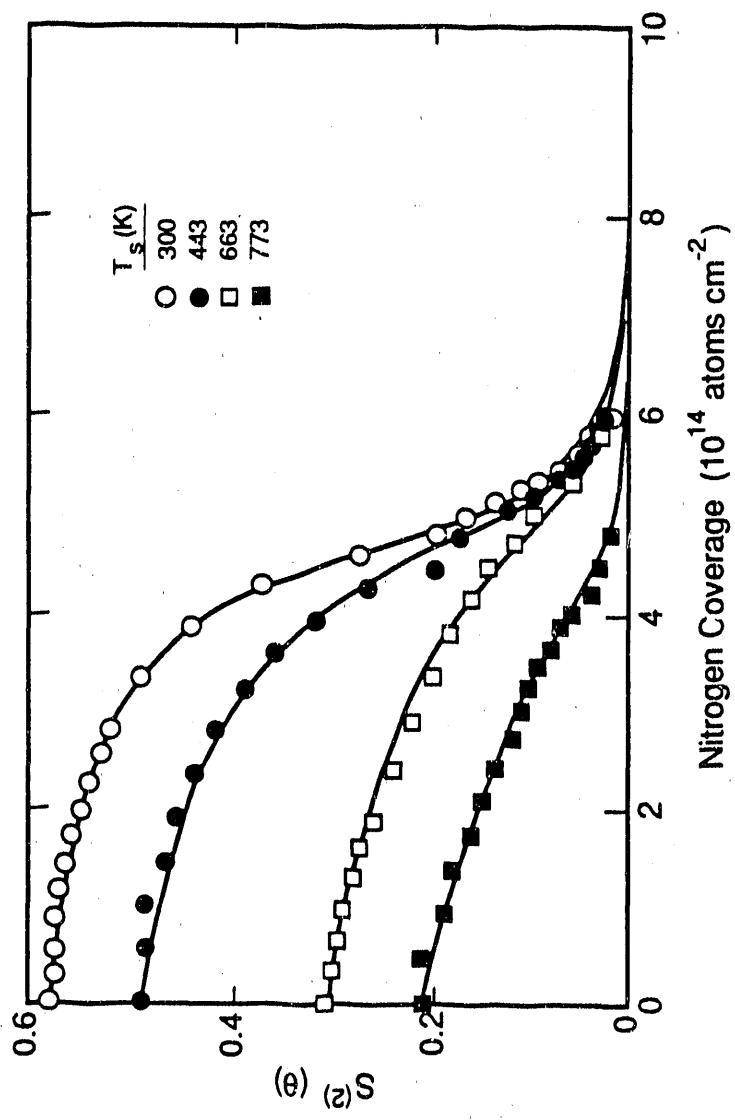


Fig. 3.9 Experimental and predicted variation in $S(2)(\theta)$ versus coverage for N_2 adsorption on $\text{W}(100)$ [59].

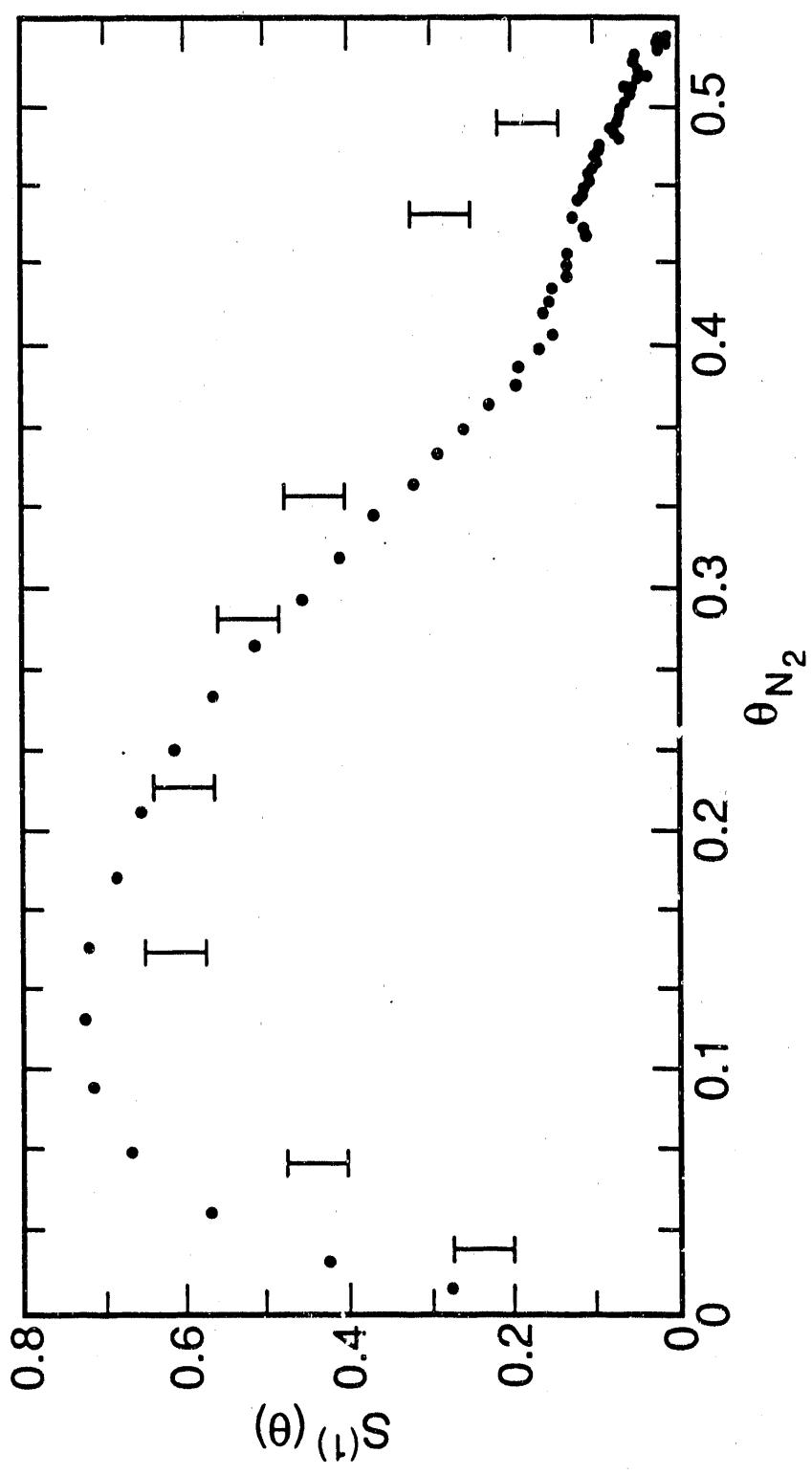


Fig. 3.10 Experimental (error bars) [62,63] and predicted (●) [61] variation in $S(1)(\theta)$ versus coverage for N_2 adsorption on Ru(001).

4.0 DIFFUSION

4.1 Molecular Dynamics

McDowell and Doll [64-67] have used MD simulations to determine the self-diffusivity of W and Rh atoms on well-defined surfaces of these metals. The interactions of the adsorbate and surface atoms in all the simulations were described by Lennard-Jones potentials, and motion of the top 1-2 layers of the surface was included in the calculations. Table 4.1 lists the predicted and experimental values of E_{diff} and D_0 . In general, the predicted values of E_{diff} and D_0 agree well with the experimentally determined quantities. It was also noted that increasing the number of moving layers from 1 to 2 lead to only a small increase in the calculated diffusivities.

The diffusivities of C and O on Pt(111) have been calculated by Doll and Freeman [74] from MD simulations in which the adsorbate-metal interactions were described by Lennard-Jones potentials. The preexponential factor D_0 and the activation energy E_{diff} for carbon were determined to be 3.4×10^{-3} cm²/s and 26.1 kcal/mol whereas for oxygen these quantities were 1.5×10^{-3} cm²/s and 18.1 kcal/mol. The predicted values of E_{diff} are in fair agreement with the experimentally determined values of 31 ± 4 kcal/mol for carbon [75] and 16 kcal/mol for oxygen [76].

Levine and Garofalini [77] have also examined the diffusivity O on Pt(111) with an MD approach. The predicted values of E_{diff} and D_0 from these simulations agree reasonably with the values determined

by Doll and Freeman [74] quoted above.

The self-diffusion of Si on Si(100) has been modelled by NoorBatcha et al. [48]. The adsorbate-metal potential energy surface was described by Morse potentials and the interactions between Si atoms in the lattice were represented by a Keating potential. The values of D_0 and E_{diff} for Si diffusion were determined to be $6.35 \pm 1.44 \times 10^{-4} \text{ cm}^2/\text{s}$ and $3.63 \pm 0.47 \text{ kcal/mol}$. The predicted activation energy is comparable to that determined experimentally, 4.6 kcal/mol for Si on Si(111) [78].

4.2 Transition-State Theory

Voter and Doll [79] have used dynamic TST to determine the self-diffusivity of Rh on Rh(100). The adsorbate-surface interactions were described by Lennard-Jones potentials and the diffusivities were calculated using eq. 2.85. As shown in fig. 4.1, the predicted and experimentally determined self-diffusivities are in good agreement. The predicted values of D_0 and E_{diff} for Rh atoms are $6.6 \pm 0.06 \times 10^{-3} \text{ cm}^2/\text{s}$ and $23.82 \pm 0.05 \text{ kcal/mol}$, whereas the experimentally determined quantities are $10^{-3} \text{ cm}^2/\text{s}$ and $20.2 \pm 1.7 \text{ kcal/mol}$ [69].

The diffusivity of hydrogen on a rigid Cu(100) surface has been investigated by Valone et al. [80] and Lauderdale and Truhlar [81]. Although the approaches used by these authors were based on classical TST, quantum effects were included in both models. Valone et al. [80] treated all motion classically. Quantum effects were incorporated by use of an effective potential which accounts

for the effects of zero-point energies and tunneling. Lauderdale and Truhlar [81], on the other hand, treated motion along the reaction coordinate classically, whereas other modes of motion were treated quantum mechanically. The classical motion in the reaction coordinate, however, was multiplied by a semi-classical transmission coefficient to account for zero-point energies and tunneling. In both studies, the inclusion of quantum effects was shown to lead to a non-Arrhenius dependence of the diffusivity at low temperatures. Lauderdale and Truhlar [81] were able to show that the major source of the non-Arrhenius behavior is tunneling. Both Valone [80] and Lauderdale and Truhlar [81] found the ratio of the diffusion coefficient determined from quantum mechanics to that determined from classical mechanics to be of order 10^3 for H and 10 for D at 120 K. This level of agreement is remarkable in view of the differences in the formalisms used.

The two models discussed above were extended to examine the influence of the lattice motion on the self-diffusivity of hydrogen. In the work reported by Valone et al. [82], only atoms in the topmost layer were allowed to move, whereas in the work reported by Lauderdale and Truhlar [83] and Truong and Truhlar [84,85], both surface and subsurface atom movement was included. Each of these studies showed that movement of the substrate resulted in larger values of the diffusivity for both H and D, the effect being most pronounced at low temperature. At 100 K, the diffusivity for H on a moving surface was 10 to 10^3 larger than that for H on a rigid surface.

The diffusion of H on Ru(001) in the temperature range 260-

330 K has been simulated by two groups. In the work of Mak and George [86], the H-surface potential was described by a sum of pairwise-additive Morse potentials whereas in the study by Truong et al. [87], an ab initio potential was used. The values of D_0 and E_{diff} determined by Mak and George [86] were $2.2 \times 10^{-3} \text{ cm}^2 \text{ s}^{-1}$ and 3.84 kcal/mol, respectively, whereas the corresponding values calculated by Truong et al. [87] were $10^{-3} \text{ cm}^2 \text{ s}^{-1}$ and $4.1 \pm 0.5 \text{ kcal/mol}$. The values of D_0 predicted in both studies are larger than the experimentally determined value of $D_0 = 6.3 \times 10^{-4} \text{ cm}^2 \text{ s}^{-1}$ [88]. The value of E_{diff} determined experimentally is 4.0 kcal/mol [88] and falls between the two predicted values.

4.3 Quantum Mechanical Models

Auerbach et al. [89] have used a quantum mechanical model to treat the diffusion of hydrogen at low coverage on W(110). The model takes into account the effects of phonon-adsorbate interactions on the motion of the adsorbate. At low temperature, the adsorbate motion arises from phonon-assisted tunneling whereas at high temperature, the motion occurs by an activated hopping process. Figure 4.2 shows the predicted and experimentally determined [90,91] diffusivities as a function of temperature. It can be seen that the model successfully represents the experimentally observed trends. The model developed by Auerbach et al. [89] also predicts that in the activated diffusion regime, the preexponential factor for diffusion increases exponentially with increasing mass of the adsorbate, a trend which, although opposite to the classical

isotope dependence, is in accord with the experimentally observed mass dependence of the preexponential factor for hydrogen on W(110) [91].

The coverage dependence of the diffusivity of hydrogen on W(110) has been determined by Whaley et al. [92] using a model which accounts for the band structure of hydrogen atoms in a periodic surface potential. The model differentiates between low- and high-coverage diffusion. At low coverage, the hydrogen motion is described by a random collision process in a uniform potential, whereas at high coverage, the hydrogen motion is affected by the perturbations of nearest-neighbor H-H interactions on the surface potential. The predicted diffusivities for the isotopes of hydrogen are shown in fig. 4.3a and display a complex coverage dependence. For comparison, the experimentally observed diffusivity profiles with coverage are shown in fig. 4.3b [91]. With the exception of the deuterion profile at high coverage, the model predictions agree qualitatively with the experimental results.

4.4 Effects of Lateral Interactions

The occurrence of lateral interactions between species can influence the magnitude of the activation energy barrier for diffusion as well as the distribution of adsorbates on the surface. Lateral interactions between adsorbates are also responsible for ordered adsorbate overlayers, island formation, and order-disorder phase transitions. Both continuum, lattice-gas models and Monte Carlo models have been developed to determine the effects of lateral

interactions on the diffusion of adsorbates on surfaces.

Continuum Models

Reed and Ehrlich [93] have investigated the influence of lateral interactions on the coverage dependence of the diffusivity. In this work, the flux was expressed in terms of the chemical potential gradient:

$$J = -\Gamma(\theta)\lambda^2 \left[\frac{1}{1-\theta} + \left(\frac{\partial \mu / k_b T}{\partial \ln \theta} \right)_T \right] \frac{\partial C}{\partial x} \quad (4.1)$$

where μ is the chemical potential. The coverage-dependent diffusivity was then given by

$$D(\theta) = \Gamma(\theta)\lambda^2 \left[\frac{1}{1-\theta} + \left(\frac{\partial \mu / k_b T}{\partial \ln \theta} \right)_T \right] \quad (4.2)$$

Both $\Gamma(\theta)$ and μ were determined using a lattice-gas model with the quasi-chemical approximation. Calculations of $D(\theta)/D(0)$ were made for selected values of the lateral interaction parameter $w_{AA} = kT^*$, where T^* is the reference temperature. As seen in fig. 4.4 for repulsive interactions, $D(\theta)/D(0)$ increases with increasing θ up to $\theta = 0.8$ and then decreases weakly. Although not shown here, for attractive lateral interactions, $D(\theta)/D(0)$ decreases rapidly from a value 1.0 at $\theta = 0$ to a value of 0.05 at $\theta = 0.5$ and then remains relatively constant at this value to $\theta = 1.0$.

Using an approach similar to that described above, Zhdanov [94] has examined how both nearest- (w_1) and next-nearest-neighbor interactions (w_2) influence the diffusivity. As seen in fig. 4.5, $D(\theta)/D(0)$ goes through a maximum with increasing θ . The location of the maximum is sensitive to the magnitude of the w_2 . In a related

study, Zhdanov [95] has treated the coverage dependence of the diffusion coefficient when two different species, A and B, are coadsorbed. This study demonstrated that $D(\theta)/D(0)$ is sensitive to both the magnitude and sign of the interactions between A and B.

Monte Carlo Models

The diffusive motion of an ensemble of adsorbates can be described using Monte Carlo techniques. The adsorbate atoms or molecules are assumed to be located at fixed positions on a lattice of surface sites. The hopping of adsorbates from site i to site j can be described by P_{ij} , a transition probability which depends on the configuration of nearest-neighbor sites surrounding the initial and final sites. Simulations begin by first choosing an adsorbate at random and then selecting one of the z nearest-neighbor sites. If the neighboring site is occupied, the procedure is begun again. If the neighboring site is unoccupied, the hopping probability P_{ij} is computed and compared to a random number R ($0 \leq R \leq 1$). The adsorbate is moved from site i to site j if $R \leq P_{ij}$, otherwise it remains in its original position. Each time an adsorbate is selected, the time is increased by an amount $\Delta t = \tau/N$, where $1/\tau$ is the attempt frequency of hops and N is the number of adsorbates in the ensemble. The surface diffusion coefficient, D , is calculated from the ensemble average of the mean-square displacement of adsorbates using eq. 2.72.

The transition probability P_{ij} discussed above can be described in terms of an activation energy for diffusion, E_{diff} , or an energy difference between sites i and j , $\Delta E = (Q_j - Q_i)$. In the presence of lateral interactions, both E_{diff} and ΔE depend on the configuration of

nearest-neighbor sites surrounding the initial and final sites. If adsorbate diffusion is characterized by E_{diff} , P_{ij} is given by

$$P_{ij} = \exp[-E_{\text{diff}}/k_b T] \quad (4.3)$$

If the representation of adsorbate diffusion is in terms of the energy difference ΔE , P_{ij} is described by either Metropolis [96] or Kawasaki [97] dynamics. For Metropolis dynamics,

$$P_{ij} = \exp\left[\frac{\Delta E}{k_b T}\right] \quad \text{for } \Delta E < 0$$

$$= 1 \quad \text{for } \Delta E \geq 0 \quad (4.4)$$

For Kawasaki dynamics,

$$P_{ij} = \frac{\exp[\Delta E/k_b T]}{1 + \exp[\Delta E/k_b T]} \quad (4.5)$$

Kang and Weinberg [98] have recently discussed the proper specification of P_{ij} and τ required to correctly represent the dynamics of diffusion. Their analysis shows that P_{ij} must be written as given in eq. 4.3 and $\tau = 1/v_{\text{diff}}$, where v_{diff} is the frequency of frustrated translation motion parallel to the surface. It was also shown that specification of P_{ij} using either Kawasaki or Metropolis dynamics does not produce a proper description of the surface dynamics. This is demonstrated in fig. 4.6 which shows plots of D versus $1/T$ for two representations of P_{ij} . In one case, P_{ij} is given by the energy barrier model (eq. 4.3) whereas in the other, P_{ij} is given by Kawasaki dynamics (eq. 4.5). As seen in fig. 4.6, the two models lead to different dependences of the diffusivity with temperature. Kang and Weinberg [98] indicate, however, that the diffusivities determined from Kawasaki dynamics can be brought into agreement with those determined from the energy barrier model by rescaling τ .

The diffusivities determined from the rescaled Kawasaki model are also shown in fig. 4.6 and are seen to be in good agreement with the diffusivities predicted from the energy barrier model. In general, though, the rescaling factor is not known and thus accurate diffusivities can only be obtained from the energy barrier model.

Models in which eqs. 4.3-4.5 are used to determine the coverage dependence of the diffusivity are discussed next. Bowker and King [99] have used a Monte Carlo model to investigate the influence of lateral interactions on the diffusivity. Lateral interactions were described in a pairwise additive manner. Concentration gradients were established on a two-dimensional array and jumps of adsorbates between sites were monitored as a function of coverage and time. Coverage-dependent diffusivities were then obtained for the three cases of repulsive, attractive, and no interactions between adsorbates. As seen in fig. 4.7, D increases with increasing θ for repulsive interactions, whereas it decreases with increasing θ for attractive interactions. In the absence of lateral interactions, D is coverage independent. Bowker and King [100] have also examined the influence of repulsive nearest-neighbor and attractive next-nearest neighbor interactions on D . Figure 4.8a shows that the predicted profile of D goes through a maximum near $\theta = 0.45$. This type of coverage dependence is qualitatively similar to that observed experimentally for oxygen diffusion on W(110), which is shown in fig. 4.8b [101].

Tringides and Gomer [102] have also simulated the diffusion of oxygen on W(110) using a Monte Carlo model. Both pairwise and triplet interactions between adsorbates were considered. The

experimentally observed [103] increase in the activation energy and preexponential factor with increasing coverage were accounted for with nearest-neighbor interactions of 2.1 kcal/mol, next-nearest-neighbor interactions of -0.7 kcal/mol, and triplet interactions of -1.2 or 0 kcal/mol.

The effects of adsorbate coverage on the diffusion coefficient for ordered or disordered adsorbates have been modelled by Sadiq and Binder [104]. Figure 4.9a shows that for disordered adsorbates, the diffusivity goes through a maximum as a function of coverage. Simulations were also conducted for a case in which the adsorbates can form an ordered overlayer at $\theta = 0.5$. As seen in fig. 4.9b, the diffusion coefficient exhibits maxima at $\theta = 0.33$ and $\theta = 0.67$. The minimum at $\theta = 0.5$ is attributed to the formation of an ordered surface overlayer which causes a reduction in the diffusivity.

Mak et al. [105] have examined the effect of an immobile species B on the diffusivity of a mobile species A assuming no lateral interactions. Figure 4.10a displays the dependence of $D(\theta_A)/D(0)$ on θ_B . $D(\theta_A)/D(0)$ is seen to decrease with increasing coverage of B. On the other hand, $D(\theta_A)/D(0)$ decreases linearly with θ_B , as shown in fig. 4.10b. The intercept with the abscissa at $\theta_B = 0.408$ is in agreement with percolation estimates for the coverage of B above which no further diffusion of A can occur.

Table 4.1
Arrhenius parameters for atomic self-diffusion [64]

Atom/Surface	D ₀ (cm ² /s)		E _{diff} (kcal/mol)	
	Exp.	Calc.	Exp.	Calc.
W/W(110)	^a 2.1×10 ⁻³		^a 19.8	
	^b 2.6×10 ⁻³		^b 21.2	
	^c 6.2×10 ⁻³		^c 20.8	
	Average 3.6×10 ⁻³	3.59±1.77×10 ⁻³	20.6	22.7±3.0
W/W(211)	^d 3.0×10 ⁻⁴		^d 17.5	
	^e 1.2×10 ⁻⁴		^e 16.8	
	^f 2.0×10 ⁻²		^f 19.7	
	^g 2.0×10 ⁻³		^g 17.8	
Average	5.6×10 ⁻³	2.18±2.72×10 ⁻³	18.0	20.5±5.2
Rh/Rh(111)	^b 2×10 ⁻⁴	7.10±1.28×10 ⁻⁴	^b 3.6	6.2±0.6
Rh/Rh(100)	^b 1×10 ⁻³	4.06±2.34×10 ⁻³	^b 20.2	20.8±3.7

^aFrom Ref. [68]. ^bFrom Ref. [69]. ^cFrom Ref. [70]. ^dFrom Ref. [71]. ^eFrom Ref. [72].
^fFrom Ref. [73].

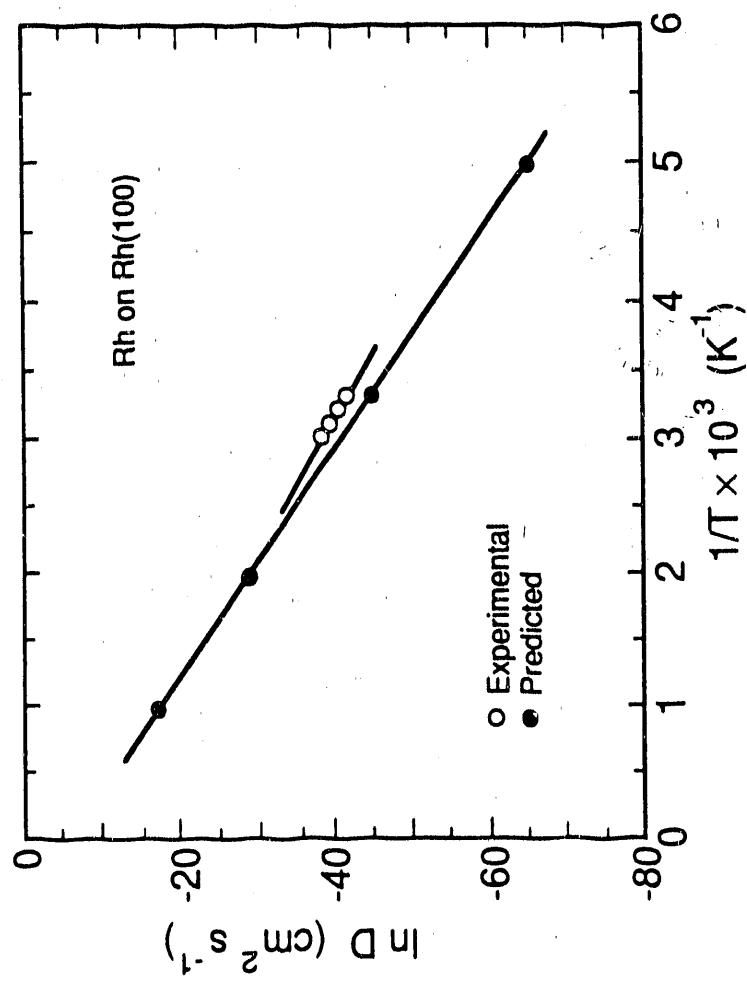


Fig. 4.1 Experimental [69] and predicted [79] variation in D versus $1/T$ for Rh diffusion on Rh(100).

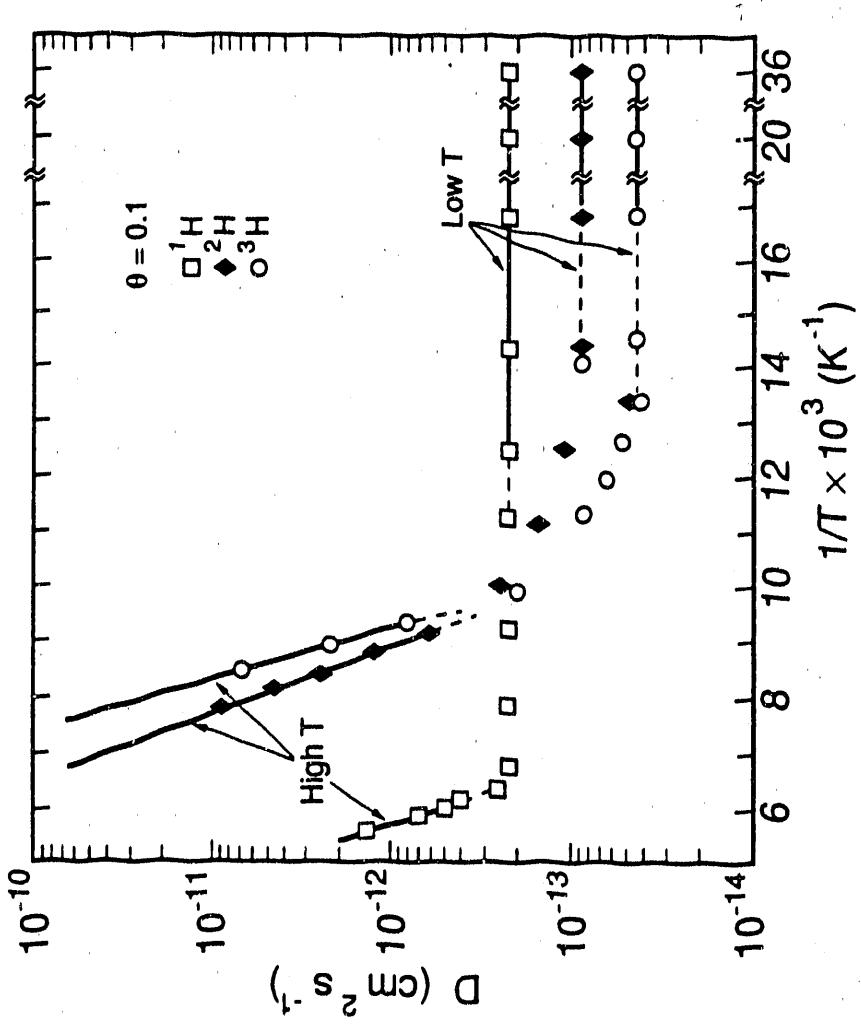


Fig. 4.2 Variation in D with $1/T$ for hydrogen diffusion on $\text{W}(110)$. Symbols are the experimental data [90,91] and the lines are the model predictions [89].

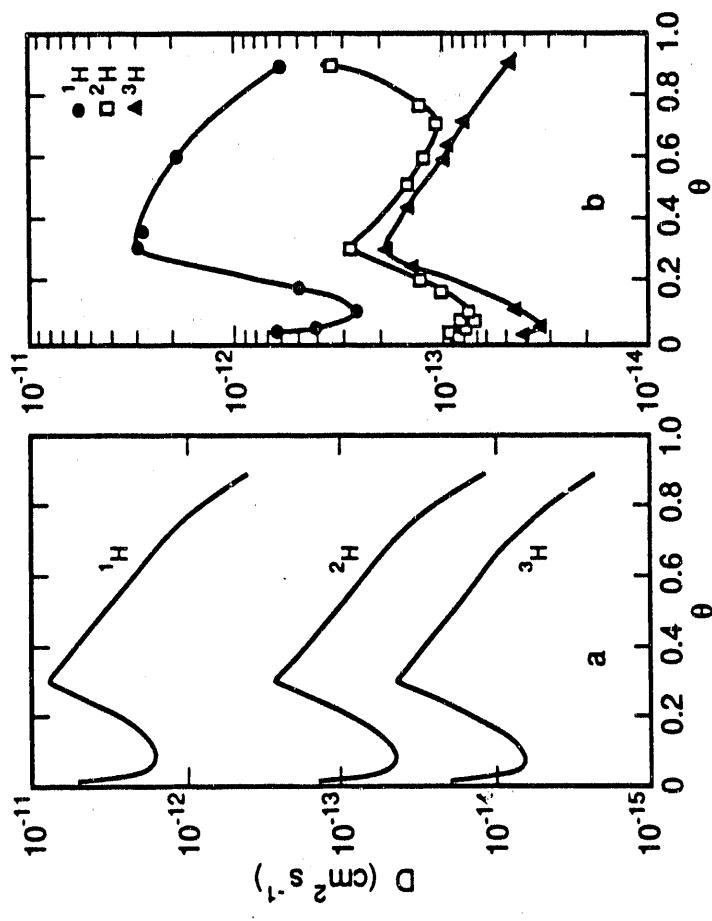


Fig. 4.3
 a. Predicted variation in D with θ for hydrogen diffusion on a metal surface at 27 K [92].
 b. Experimental variation in D with θ for hydrogen diffusion on W(110) at 27 K [91].

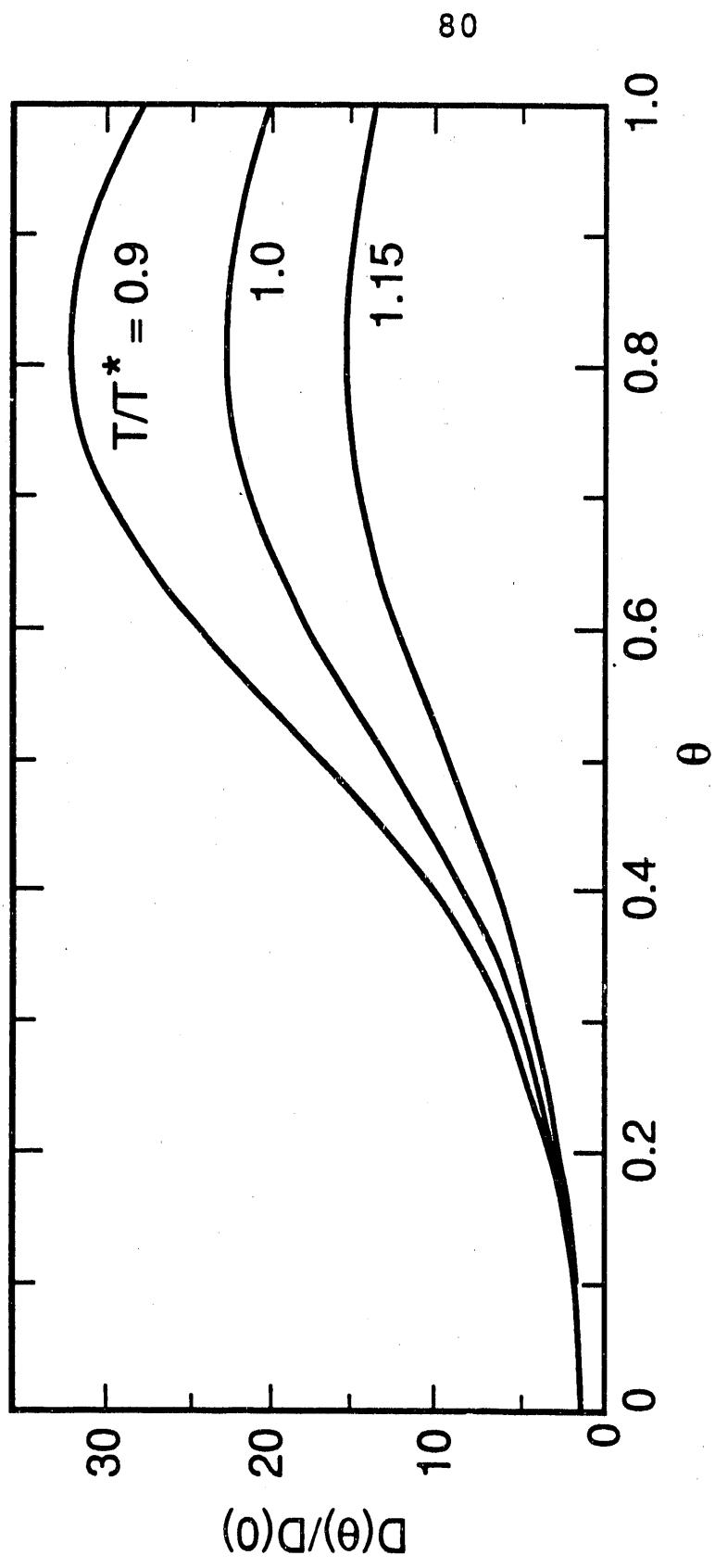


Fig. 4.4 Variation in $D(\theta)/D(0)$ with θ for repulsive interactions for $WAA = k_b T^*$, where T^* is the standard temperature [93].

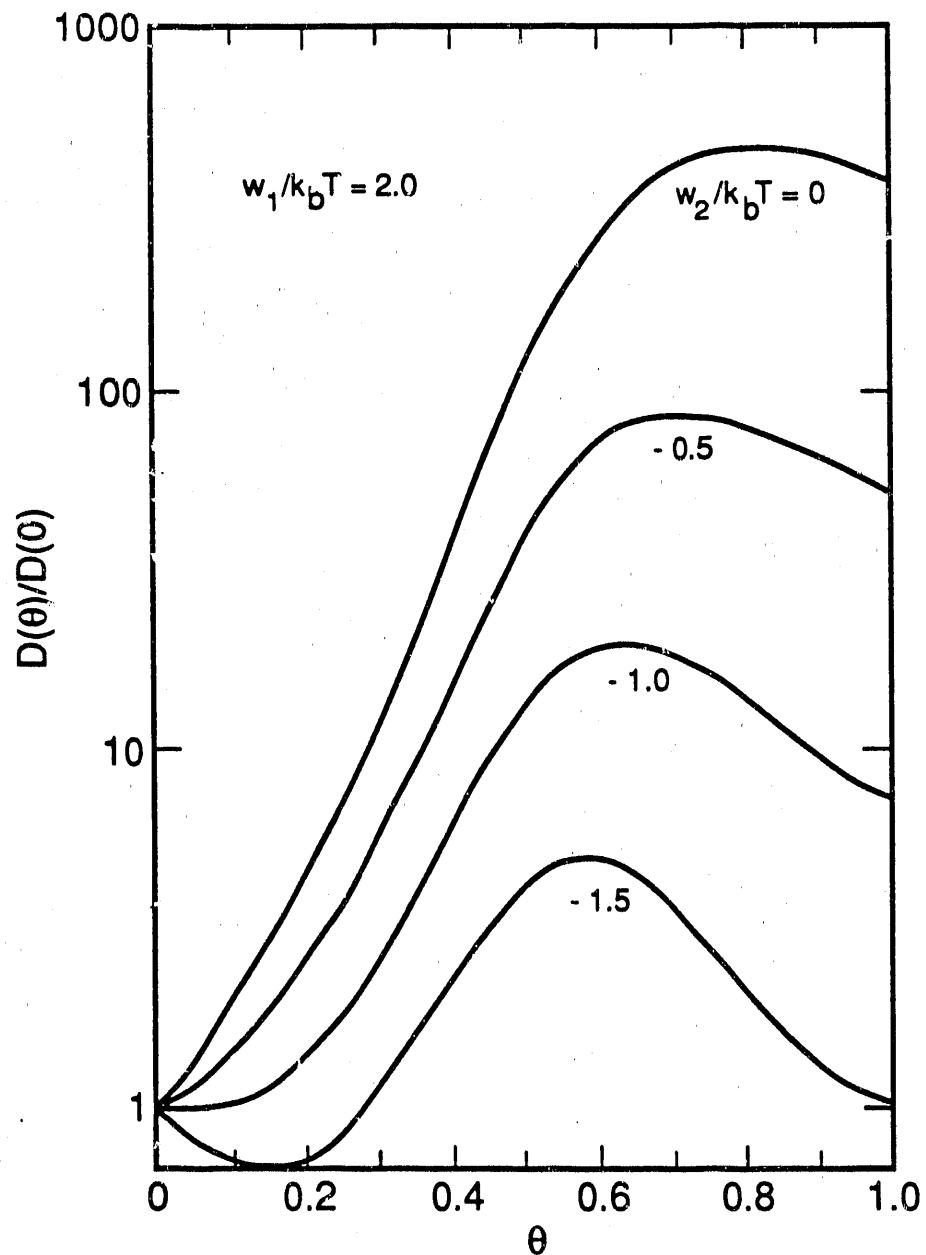


Fig. 4.5 Variation in $D(\theta)/D(0)$ with θ for nearest- (w_1) and next-nearest- (w_2) neighbor interactions [94].

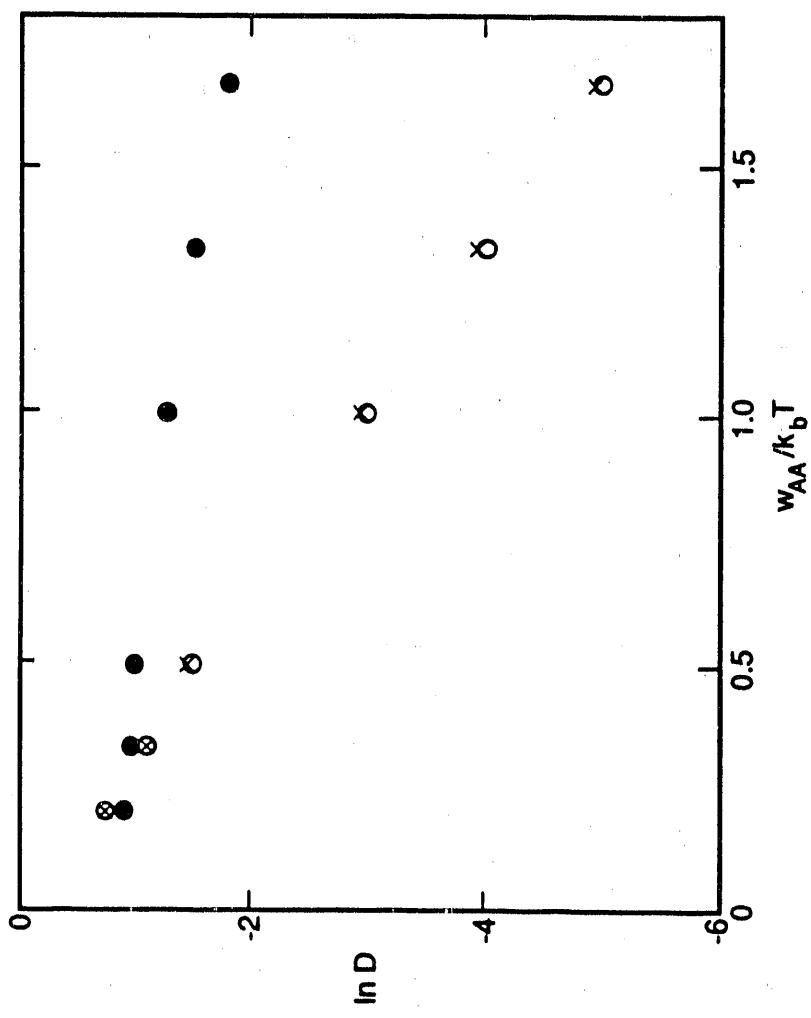


Fig. 4.6 Variation in D with $w_{AA}/k_b T$ [98]. The open circles are from simulations using eq. 4.3 and the closed circles are from simulations using eq. 4.5. The crosses are the rescaled diffusivities from eq. 4.5.

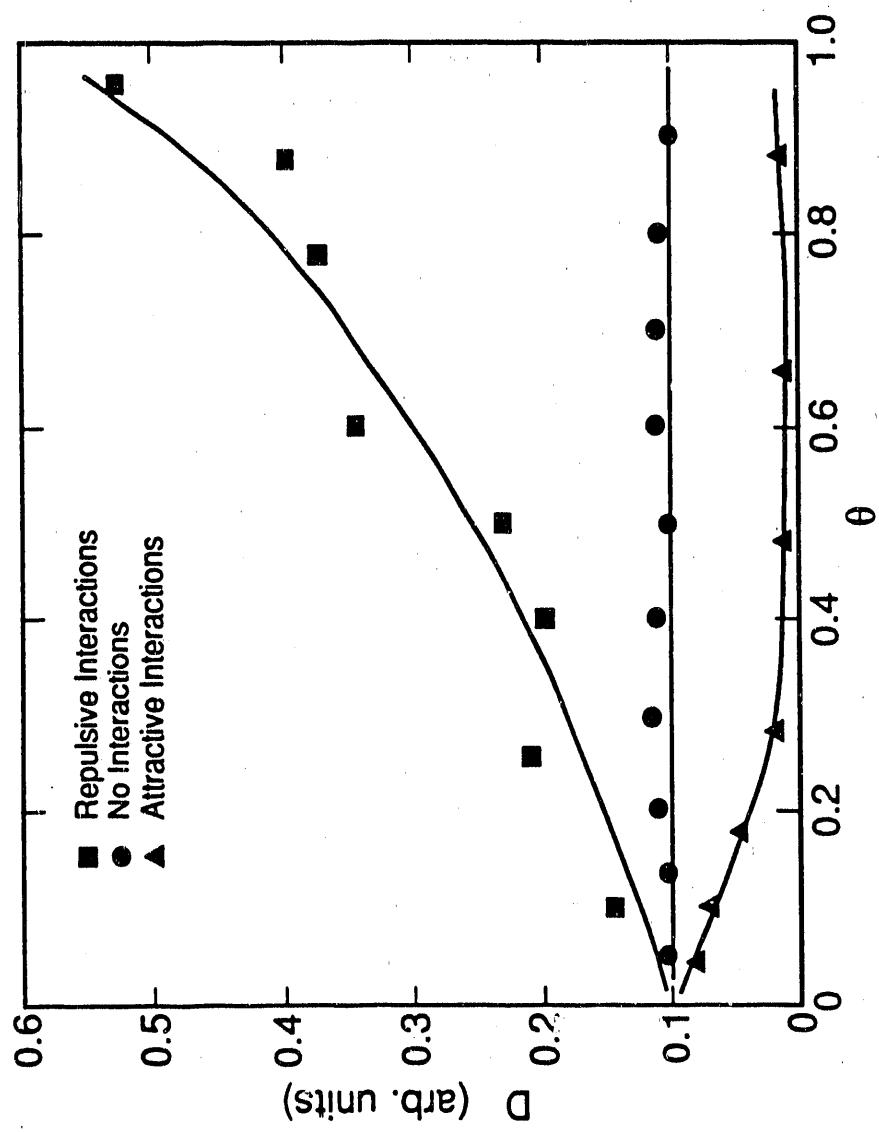


Fig. 4.7 Variation of D with θ for repulsive (■), attractive (▲), and no nearest-neighbor interactions (●) [99].

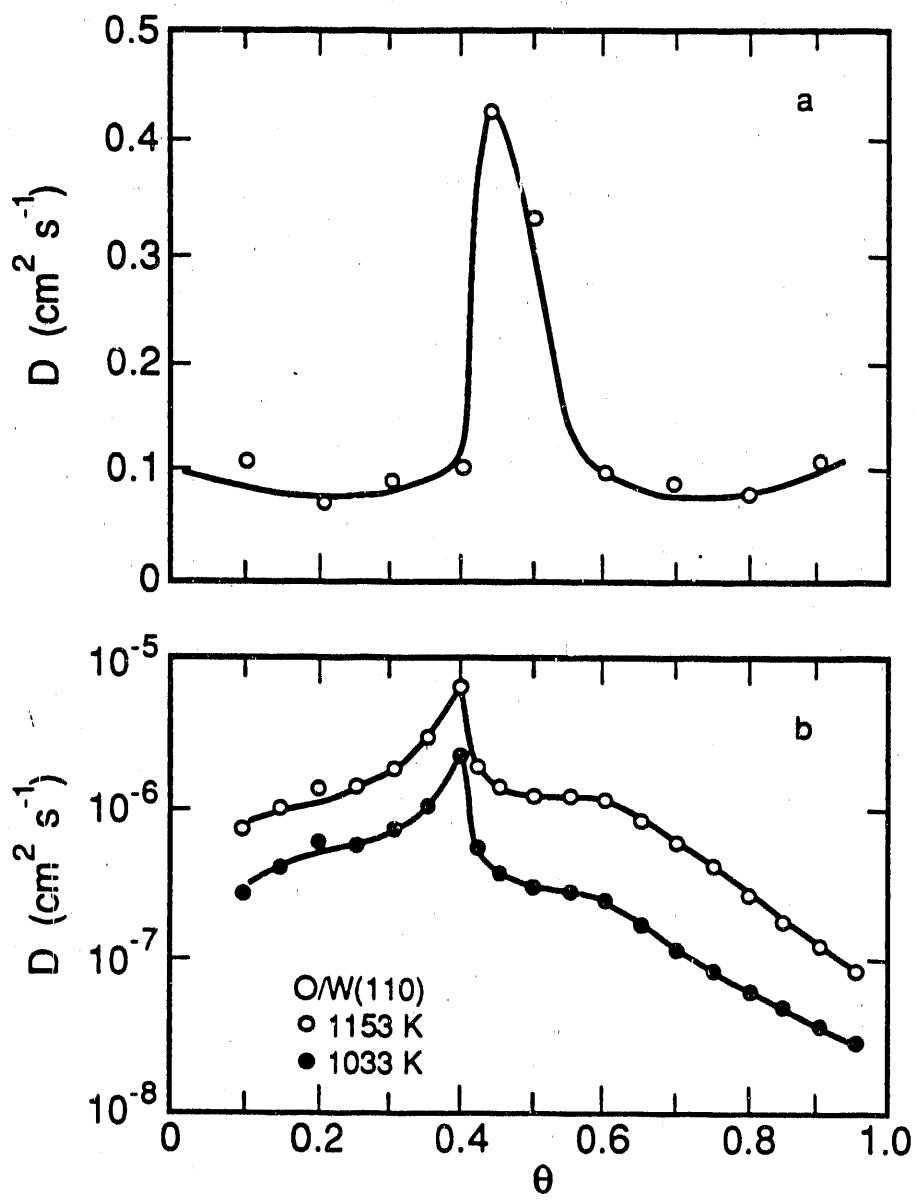


Fig. 4.8 a. Simulation of the θ dependence of D for repulsive nearest-neighbor and attractive next-nearest neighbor interactions [100]. b. Experimental variation in D with θ for oxygen on W(110) [101].

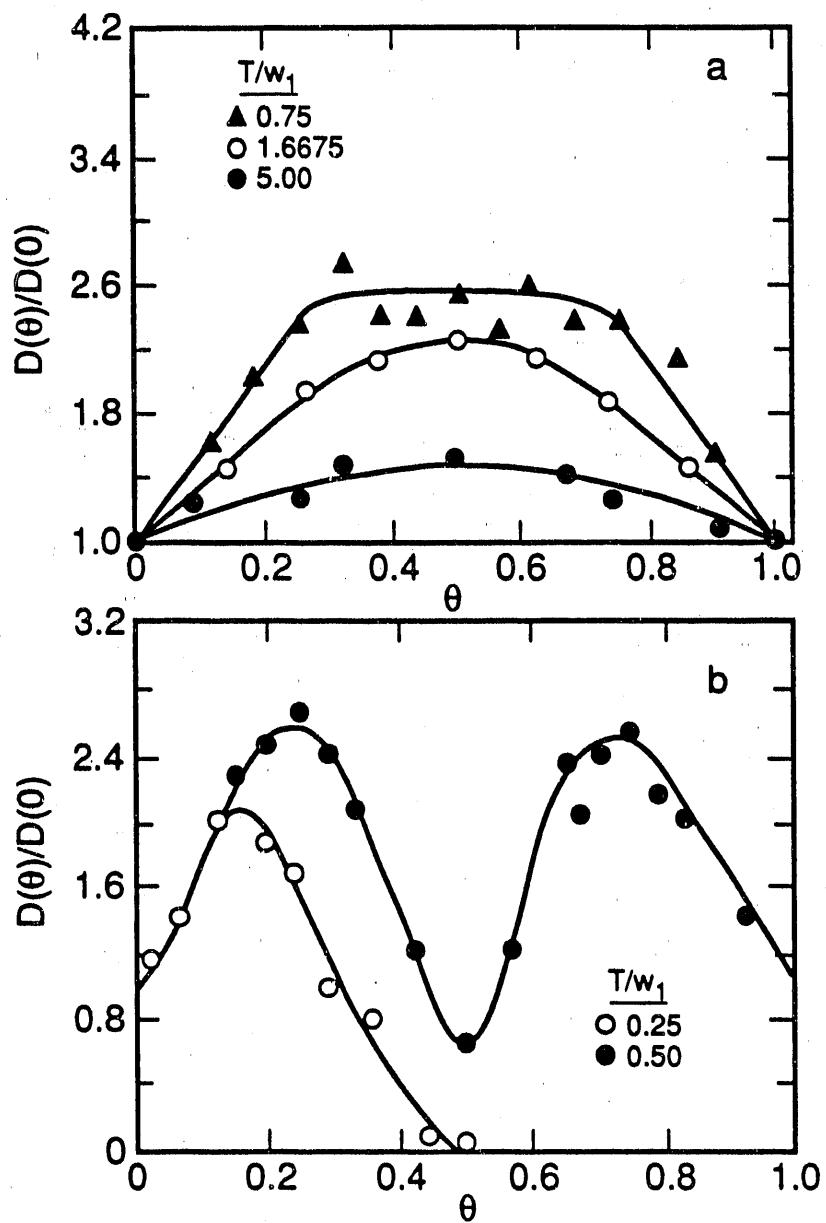


Fig. 4.9 Variation in $D(\theta)/D(0)$ with θ for equivalent nearest- (w_1) and next-nearest- (w_2) neighbor repulsive interactions [104]. a. Disordered region. b. Ordered region.

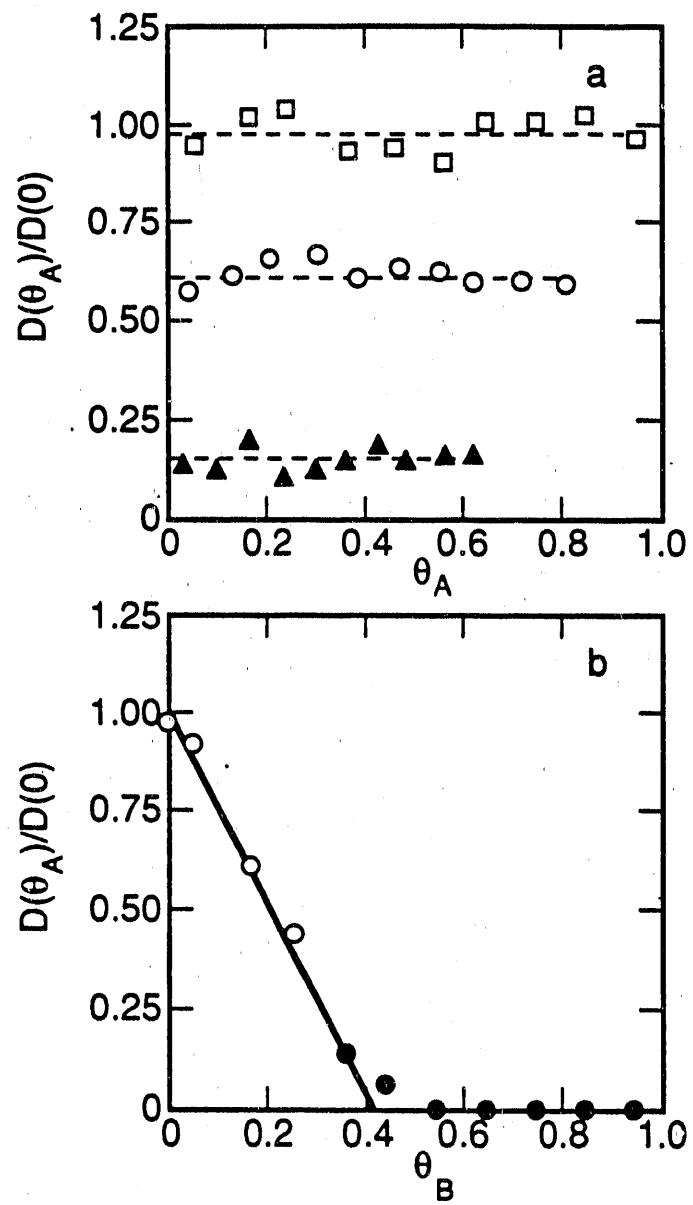


Fig. 4.10 a. Variation in $D(\theta_A)/D(0)$ with θ_A in the presence of an immobile adsorbate B for $\theta_B=0$ (□), $\theta_B=0.15$ (○), and $\theta_B=0.36$ (△). b. Variation in $D(\theta_A)/D(0)$ with θ_B . The filled symbols represent upper limits [105].

5.0 DESORPTION

5.1 Transition-State and Stochastic Diffusion Theory

Tully and coworkers [28,29] have calculated the rate coefficients for Xe and Ar desorption from Pt(111) using both dynamically corrected (eq. 2.75) and classical (eq. 2.76) TST. The results of these calculations are illustrated in fig. 5.1. At low temperatures, both approaches give identical results, indicating that f_s goes to 1.0 as the temperature decreases. It is also observed that at low temperatures, $k_d(1)$ exhibits Arrhenius behavior. The apparent activation energies for Ar and Xe desorption are 8.8 and 28.8 kJ/mol, respectively, in good agreement with the experimental values of 9.2 and 29.3 kJ/mol [106,107]. In the limit of low temperatures, the apparent preexponential factors for Ar and Xe are 5.2×10^{11} and $6.8 \times 10^{11} \text{ s}^{-1}$, respectively, which agree approximately with the vibrational frequencies of the rare-gas atoms on the surface: $9.8 \times 10^{11} \text{ s}^{-1}$ for Ar and $9.2 \times 10^{11} \text{ s}^{-1}$ for Xe. Figure 5.1 shows that with increasing temperature above 200 K, the value of $k_d(1)$ predicted by dynamically corrected TST becomes progressively smaller than that calculated by classical TST. This deviation is a reflection of the decreasing value of $f_s(T)$ [i.e., $S(1)(0)$] as T increases. The non-Arrhenius behavior of $k_d(1)$ is explained in the following way. At high temperatures, the mean energy transferred upon impact of the rare gas atom with the surface is relatively small and so consequently equilibrium can no longer be maintained among adsorbates located near the top of the potential well. This causes a

depletion in the population of the precursor levels due to desorption into the gas phase and hence, to a reduction in the rate of desorption.

Using an approach similar to that described above, Muhlhausen et al. [47] have examined the molecular desorption of NO from Pt(111). The rate of NO desorption exhibited Arrhenius behavior at low temperature, but at temperatures above 1000 K, curvature was observed. This high temperature behavior was explained in the same manner as that for Xe and Ar desorption, discussed above. The preexponential factor determined from the low temperature portion of the Arrhenius plot was $10^{16} \pm 0.4$ s⁻¹, in excellent agreement with the experimental value of 10¹⁶ s⁻¹ [108]. This value of $v_d(1)$, which is larger than the value of 10¹³ s⁻¹ often assumed, is ascribed by the authors to the reduced entropy of chemisorption associated with the alignment of the molecule perpendicular to the surface.

Doren and Tully [36] have calculated $k_d(1)$ for CO desorption from Ni(111) using dynamic TST. Figure 5.2 illustrates plots of the apparent Arrhenius parameters E_d and $v_d(1)$ versus temperature. Both the activation energy and the preexponential factor are seen to be temperature dependent and to decrease with increasing temperature. For temperatures between 400 and 600 K, $v_d(1)$ is in the range of 10¹⁴ to 10¹⁵ s⁻¹. The explanation for this large value of $v_d(1)$ for CO is identical to that presented above for NO.

The associative desorption of H₂ from Si(111) has been modelled by Raff et al. [109] with a variational TST approach. The potential energy surface was defined by the sum of three terms corresponding to the lattice potential, adatom-lattice interactions, and adatom-adatom interactions. The rate coefficient $k_d(2)$ exhibited

Arrhenius behavior over the temperature range investigated. The apparent activation energy and preexponential factor were 55 kcal/mol and $0.20 \text{ cm}^2 \text{ s}^{-1}$, respectively. These values lie within the range of the experimentally observed values of 41-59 kcal/mol for the activation energy and $0.03-136 \text{ cm}^2 \text{ s}^{-1}$ for the preexponential factor [110-112].

Stochastic diffusion theory has been used by Zeiri et al. [41] to describe the desorption of K and Xe from a W(111) surface and the molecular desorption of CO from a Ni(110) surface. The values of the parameters appearing in eqs. 2.88 and 2.90 were specified on the basis of independent experimental results. The temperature dependence of the rate was in excellent agreement with experimental observation [113-116] for the three adsorbate-metal systems studied. The preexponential factor for adsorbed CO was found to be a factor of 100 larger than the value of 10^{13} s^{-1} for adsorbed K or Xe. This difference was attributed to the frustrated surface rotational modes of chemisorbed CO.

5.2 Quantum Mechanical Models

The earliest quantum models, developed by Lennard-Jones, Strachan, and Devonshire [117-121], treated the one-dimensional motion of an adsorbate bonded to a single surface atom. The motion of the adsorbate was only considered in the direction perpendicular to the surface, and only single-phonon exchanges between the substrate and adsorbate were taken into account. As a consequence of the one-dimensional nature of the adsorbate motion, all exchange

of energy between the adsorbate and the substrate leads to motion perpendicular to the surface, with the net result that the transition probabilities from bound to continuum states are overestimated [122].

De et al. [123] have used a one-dimensional, multiphonon model to represent the desorption of K and Xe from a W surface. The phonon spectrum of the substrate was described by a Debye model, and either a harmonic or a Morse potentials was used to describe the interactions between the adsorbate and the substrate. Although the rate of desorption predicted by the model was found to be sensitive to the type of potential used in the calculations, both harmonic and Morse potentials resulted in an Arrhenius dependence of the desorption rate. Good agreement between theory and experiment [113] was obtained for the desorption rates of K from a W surface when the interactions between the adsorbate and the surface were described by a Morse potential. On the other hand, for the desorption of Xe from a W surface, the predicted rates of desorption, using either a harmonic or a Morse potential, were smaller by a factor of 100 than the rates determined experimentally [115]. No explanation for this discrepancy was given by the authors.

Hood et al. [44] have used a one-dimensional quantum model with a Morse potential to analyze the desorption of Ar from a W surface. Included in the model were multiphonon exchange processes between the Ar atoms and the substrate. The rates of desorption predicted by the model exhibited weakly non-Arrhenius behavior. The apparent preexponential factor increased from 10^{10} s^{-1} to 10^{11} s^{-1} with increasing temperature, and the apparent activation energy

increased from 1.3 kcal/mol to 1.5 kcal/mol. These authors noted that the apparent activation energy was 25-35% lower than the binding energy of 1.9 kcal/mol used in the Morse potential. This was ascribed to the high desorption probability of adsorbates promoted to bound states of intermediate energy. As a consequence, the apparent desorption activation energy barrier is the energy necessary to reach intermediate bound states rather than that required to reach the top of the potential well.

Hood et al. [44] have also examined the desorption kinetics of CO from a Cu surface using the model discussed above. The CO molecule was treated as a quasi-atom, i.e., no account was made for the CO bending and stretching modes of vibration. The calculated rates of desorption were observed to exhibit strongly non-Arrhenius behavior as a function of temperature and to vary over approximately one order of magnitude depending on the mode of phonon relaxation used in the calculations. The apparent preexponential factor for desorption ranged from 10^{13} to 10^{17} s⁻¹, with the smaller values being observed at lower temperatures. The activation energies for desorption were seen to be 5-20% smaller than the binding energy value of 16.6 kcal/mol used in the Morse potential. This observation was explained in the same manner as that for the case of Ar desorption from W, discussed above.

5.3 Simulation of TPD Spectra

Both continuum lattice-gas models and Monte Carlo (stochastic) models have been used to simulate TPD spectra of

adsorbates from well-defined surfaces. In the continuum approach, both the explicit coverage dependence of the desorption rate and the dependence of the rate coefficient, $k_d(\alpha)$, are described in terms of the average coverage, θ . By contrast, Monte Carlo models of desorption represent the desorption process in terms of a probability which is specific to each site and its configuration of nearest neighbors. As will be shown below, the influence of finite rates of surface diffusion and the presence of coadsorbates can be incorporated into either type of model.

Simulations Based on Continuum Models

To simulate TPD spectra within the continuum framework, a balance is written between the rate of desorption and the rate at which the surface coverage decreases. Thus,

$$r_d = -\beta \frac{d\theta}{dT} \quad (5.1)$$

where $\beta = dT/dt$ is the rate at which the surface is heated. If lateral interactions between adsorbates can be neglected, $r_d = v_d(\alpha) \exp(-E_d/k_b T) \theta^\alpha$. Integration of eq. 5.1 then yields $\theta(T)$, from which it is possible to determine r_d as a function of T , for an initial adsorbate coverage, θ_0 .

Experience has shown that the assumption of coverage-independent rate parameters does not adequately represent the full shape of TPD spectra observed experimentally. Substantially better agreement between theory and experiment can be achieved using a lattice-gas model of adsorbate-adsorbate interactions. Several examples of the use of lattice-gas models to simulate TPD spectra are presented below. In all but one case, the distribution of

adsorbates on the surface is based on the quasi-chemical approximation (QCA).

Goymour and King [124] have simulated the associative desorption of the high-temperature states of dissociated CO from a tungsten ribbon. The adsorbed C and O atoms were assumed to have equivalent binding energies and the activation energy for associative desorption was set equal to the differential heat of adsorption. The values of the preexponential factor ($v_d(2) = 3 \times 10^{14} \text{ s}^{-1}$), the activation energy at low coverage ($E_d^\circ = 100 \text{ kcal/mol}$), and the nearest-neighbor interaction energies ($w_{CO} = w_{OO} = w_{CC} = 4.8 \text{ kcal/mol}$) were determined by fitting the TPD spectrum predicted by the model to the experimentally observed TPD spectrum. As illustrated in figs. 5.3a and 5.3b, the predicted CO TPD spectrum agrees with that observed experimentally [125]. The coverage dependence of the activation energy for CO desorption was also determined from the model and, as seen in fig. 5.3c, decreases with increasing CO coverage as a consequence of the repulsive C-O, O-O, and C-C interactions.

Zhdanov [126] has also evaluated the applicability of a lattice-gas model for describing desorption from well-defined metal surfaces. Figure 5.4 shows a comparison of the theoretical and experimental [127] TPD spectra for CO desorption from Ir(110). The shape and location of the predicted spectrum is in fair agreement with the experimental results. To obtain this level of agreement, the following parameters were used: $E_d^\circ = 33 \text{ kcal/mol}$, $w_{CO-CO} = 2.2 \text{ kcal/mol}$, and $v_d(1) = 10^{11} \text{ s}^{-1}$. It should be noted that the value of $v_d(1)$ used to fit the data is considerably smaller than values

reported in the experimental literature, which range from 10^{14} - 10^{16} s^{-1} .

A second example from the work of Zhdanov [126] is shown in fig. 5.5. In this instance, the TPD spectra are for the associative desorption of O₂ from an Ir(110) surface. The parameters $E_d^\circ = 67 \text{ kcal/mol}$, $w_{00} = 3.6 \text{ kcal/mol}$, and $v_d(2) = 3.5 \times 10^{11} \text{ s}^{-1}$ ($3.5 \times 10^{-4} \text{ cm}^2/\text{s}$) were used to represent the data, and reasonable agreement between the theoretical and experimental [128] results is observed. The activation energy for O₂ desorption determined from the model was seen to decrease as a function of increasing coverage as a consequence of the repulsive O-O interactions.

While most applications of lattice-gas models have been restricted to a rigid lattice of adsorption sites, Zhdanov [60] has demonstrated the extension of this model to cases where surface reconstruction occurs. The rate of desorption was written in terms of a chemical potential which included contributions from adsorbate-adsorbate interactions, adsorbate-surface interactions, and the surface-free energy. Illustrations of the variations of $\Delta E_d(\theta)$ and $v_d(1)(\theta)/v_d(1)(0)$ determined from the model are shown in fig 5.6. Both the preexponential factor and the activation energy are seen to decrease with increasing coverage. The variation in the apparent preexponential factor over 8 orders of magnitude is in accord with what has been determined for the associative desorption of H₂ from a W(110) surface [129]. More recently, Myshlyavtsev and Zhdanov [130] have reexamined the the role of surface reconstruction with a more accurate approach. In this study, the variation of the preexponential factor with coverage was much smaller, spanning

only two orders of magnitude.

Implicit in the QCA is the assumption that surface diffusion is much more rapid than desorption so that an equilibrium configuration of adsorbates is maintained at all times. Sundaresan and Kaza [131] have investigated the effects of limited adspecies mobility on the rates of desorption when two species are coadsorbed on a surface. A lattice-gas model was used to describe the effects of lateral interactions. To account for the mobility of the adsorbate, a set of three differential equations was used to describe the change in the occupation probabilities of pairs of sites with time. It was demonstrated that the TPD spectra for coadsorbed species are sensitive to the sequence in which the adsorbates are dosed on the surface as well as to the relative coverages and mobilities of the two species, but no direct comparison was made between theory and experiment.

Simulations Based on Monte Carlo Models

In the Monte Carlo approach for simulating TPD spectra, adsorbates are assumed to occupy well-defined sites on a fixed lattice. The occupancy of a given site can change as a consequence of either desorption or hopping (diffusion) of the adsorbate to an adjacent site. The probability of desorption from site i is defined by P_i and the probability for diffusion from site i to site j is defined by P_{ij} .

The probability of desorption of a given species from site i in the time interval Δt can be defined as

$$P_i = v_{d,i} \exp(-E_{d,i}/k_b T) \Delta t \quad (5.2)$$

where $v_{d,i}$ and $E_{d,i}$ are the preexponential factor and the activation

energy for the local environment i , respectively. The value of Δt in eq. 5.2 is chosen so that P_i goes to unity at a temperature sufficiently high to guarantee virtually complete desorption from all sites of type i .

The rate of desorption, r_d , expressed as a turnover frequency based on the number of surface atoms, N_s , is given by

$$r_d = \frac{N_d}{\alpha N_s \Delta t} \quad (5.3)$$

where N_d is the number of adsorbates atoms or molecules desorbing in the time interval Δt . The parameter α is 1 for atomic or molecular desorption and 2 for associative desorption. Consequently, r_d is the rate of desorption as observed from the gas phase.

The diffusion of adsorbates can be handled using the Monte Carlo techniques described in Section 4.4. If the rate of diffusion is of comparable magnitude to the rate of desorption, then the dynamics of diffusion are represented with an energy barrier model and P_{ij} is given by

$$P_{ij} = \exp[-E_{\text{diff}}/k_b T] \quad (5.4)$$

If diffusion occurs very rapidly relative to diffusion, the spatial distribution of adsorbates on the surface will remain very close to equilibrium. Under such circumstances, the probability of moving an adsorbate can be represented by Kawasaki dynamics:

$$P_{ij} = \frac{\exp[-(Q_i - Q_j)/k_b T]}{1 + \exp[-(Q_i - Q_j)/k_b T]} \quad (5.5)$$

where Q_i and Q_j are the heats of adsorption for an adsorbate at sites i and j , respectively. Application of eq. 5.5 for a large number of

hops leads to an equilibrium distribution of adsorbates on surface.

The probability and rate formulations described above are incorporated into a Monte Carlo algorithm for simulating the temperature-programmed desorption of adsorbates. The algorithm presented below is for simulations in which adsorbate diffusion is described by Kawasaki dynamics, the case most frequently treated in the literature.

The surface is represented by an array of numbered sites, and periodic boundary conditions are used to eliminate edge effects. Adsorbate atoms or molecules are placed on the surface lattice to achieve a desired initial coverage, θ_0 . The temperature is initialized at T_0 and taken to be constant at this value for the time interval Δt . During this time interval, the desorption of adsorbates from occupied sites is determined by the outcome of a comparison between a random number, R , and the probability condition for desorption P_i . When $R \leq P_i$, the adsorbate (or pair of adsorbates for associative desorption) is removed from the lattice, and N_d is incremented by one (two for associative desorption). If, on the other hand, $R > P_i$, the adsorbate remains on the lattice. After sampling of the surface is completed, the rate of desorption for this time interval is calculated from eq. 5.3.

Following the calculation of the rate of desorption, the remaining adsorbates are redistributed on the lattice to account for the effects of surface diffusion. Redistribution of the adsorbates to achieve a new equilibrium configuration is achieved as follows. Movement of an adsorbate at site i to a vacant site j is determined by comparing a random number R with the value of P_{ij} given by eq.

5.5. If $R \leq P_{ij}$, the adsorbate is moved from site i to site j , whereas if $R > P_{ij}$, the adsorbate is not moved. Application of this test to the ensemble of adsorbates for a sufficiently large number of hops leads to a minimization of the surface energy and a reestablishment of equilibrium.

The desorption/diffusion calculations described above constitute a Monte Carlo Step (MCS). As noted above, the time interval associated with an MCS is Δt , and over this interval the temperature is constant. At the end of an MCS, the temperature is increased by the increment $\beta \Delta t$, where β is the heating rate. A new MCS is then carried out at the next temperature. This process is repeated until a temperature is reached for which the surface is depleted of adsorbate. A plot of the desorption rate as a function of temperature then yields a TPD spectrum. Several examples of the simulation of TPD spectra using Monte Carlo methods are presented below.

In the absence of lateral interactions, Lombardo and Bell [132] have demonstrated that the continuum and stochastic representations of TPD lead to identical spectra. An illustration of this point for non-associative desorption is given in fig. 5.7. The slight deviations of the Monte Carlo simulations from the curve for the continuum model are attributable to the combined effects of finite lattice size, finite step size, and statistical, random-number fluctuations.

Sales and Zgrablich [133] have utilized a Monte Carlo approach to model the influence of lateral interactions between adsorbates on TPD spectra. Pairwise-additive interactions were used to account

for the lateral interactions between adsorbates. Figure 5.8a shows the simulated TPD spectra for non-associative desorption from a one-dimensional surface for selected values of the interaction parameter W_{AA} . The predicted TPD spectra based on a continuum lattice-gas model using the Bragg-Williams approximation ($E_d = E_d^\circ - zW_{AA}\theta$, where z is the number of nearest-neighbor sites around a given site) [134] are presented in fig. 5.8b. A comparison of the two figures demonstrates that, for $W_{AA}=0$ (no interactions) and for $W_{AA} < 0$ (attractive interactions), the TPD spectra determined by the Monte Carlo and continuum lattice-gas models are in good agreement. For $W_{AA} > 0$ (repulsive interactions), however, the Monte Carlo model predicts two peaks whereas the continuum lattice-gas model predicts one. The authors indicate that the difference between the two models arises from the incorrect assumption of a random distribution of adsorbates in the Bragg-Williams approximation. The influence of the number and distribution of surface binding sites with different energies on TPD spectra was also examined by Monte Carlo simulations, and it was shown that the shape and location of simulated TPD spectra are sensitive to the distribution of adsorbates on the surface as well as to the magnitude and sign of the energetic interactions between species.

In another study, Sales et al. [135] have compared the desorption kinetics predicted from a Monte Carlo model using pairwise-additive energetics for nearest- and next-nearest neighbor interactions with the desorption kinetics predicted from a continuum lattice-gas model using the quasi-chemical approximation for nearest-neighbor interactions and the Bragg-

Williams approximation for next-nearest neighbor interactions. As seen in fig. 5.9, the two models predict many of the same features for selected values of the lateral interaction parameters. The number of peaks, the peak locations, and the coverage dependence of the peak positions predicted by the two models, however, are different. The differences in the spectra predicted from the two models are especially pronounced for the simulations incorporating repulsive nearest-neighbor interactions with attractive next-nearest neighbor interactions.

Hood et al. [61] have combined Monte Carlo simulations and continuum rate expressions to describe molecular desorption of N₂ from a Ru(001) surface. The Monte Carlo segment of the algorithm was used to determine the local environment of each adsorbate and hence the local activation energy for desorption. The coverage and activation energy for each type of surface environment were then used in a continuum rate expression such as eq. 5.1 to solve for the rate of desorption. The parameters used to simulate the TPD spectra were repulsive nearest-neighbor interactions of 0.25 kcal/mol, attractive next-nearest neighbor interactions of 0.45 kcal/mol, and a preexponential factor of 10¹² s⁻¹. In addition, the preexponential factor was postulated to have an increasing exponential dependence with increasing global surface coverage of N₂. As illustrated by figs. 5.10a and 5.10b, the simulated TPD spectra agree with experimental observation [62,63]. The low-coverage peak was attributed to desorption of N₂ molecules located at the perimeter of N₂ islands in which the molecules were arranged in ($\sqrt{3} \times \sqrt{3}$) R30° domains, whereas the high-coverage peak was attributed to

desorption of N₂ from antiphase ($\sqrt{3} \times \sqrt{3}$) R30° domains located within the N₂ islands.

Lombardo and Bell [132] have developed a Monte Carlo algorithm in which the BOC-MP method [15,16] was used to calculate the activation energy for desorption. For non-associative desorption, E_d is given by

$$E_d = \dot{Q}_{A,n} \quad (5.6)$$

and for associative desorption of A₂, E_d is given by

$$E_d = \dot{Q}_{A,n} \dot{Q}_{A',n'} / (\dot{Q}_{A,n} + \dot{Q}_{A',n'}) \quad (5.7)$$

where Q_{A,n} and Q_{A',n'} represent the heats of adsorption of A and A', respectively. The two recombining atoms are designated A and A' to denote that the local environments of each may be different. In order to account for coverage effects, Q_{A,n} and Q_{A',n'} are determined from eq. 2.8.

When metal-adsorbate and adsorbate-adsorbate interactions between adsorbates were included, Lombardo and Bell [132] observed multiple peaks in the simulated TPD spectra, and the activation energy profile of the desorbing species was found to decrease in a non-linear fashion with increasing coverage. It was also shown that the activation energy of the desorbing species is less than or equal to the average value for the entire adlayer. Physically, this means that for a given coverage, adsorbates which are in less stable configurations have a lower activation energy for desorption and hence desorb preferentially.

In the same study, Lombardo and Bell [132] reported simulations of the associative desorption of H₂ from Mo(100). As

illustrated by figs. 5.11a and 5.11b, the simulated TPD spectra for H₂ desorption from a Mo(100) surface agree with the experimentally observed spectra [136]. Figure 5.12a and 5.12b shows a comparison of the simulated and experimentally observed [137] profiles of the activation energy as a function of coverage. Both plots exhibit a step-wise decrease in activation energy as the coverage increases. The observed decrease in the activation energy with increasing coverage is caused by a decrease in the heat of adsorption of H atoms when more than one atom is bonded to a given metal atom.

Monte Carlo models have also been used to simulate the desorption of coadsorbed species. Gupta and Hirtzel [138] have investigated the effects of lateral interactions on the non-associative desorption of coadsorbed species A and B. Both nearest- and next-nearest neighbor interactions between species were represented in a pairwise additive fashion. It was shown that the number of peaks in the TPD spectrum and the spectrum shape and location are sensitive to the lateral interactions and to the relative coverage of each species. These authors also demonstrated that A-B interactions only affect the spectrum of the species desorbing at lower temperatures whereas A-A and B-B interactions only affect the spectra of the respective A and B species.

More recently, Lombardo and Bell [139] have simulated TPD spectra for coadsorbed species using the BOC-MP approach to describe the effects of adsorbate coverage on the energetics of desorption. Nearest-neighbor interactions between the adsorbates and the metal surface as well as nearest-neighbor interactions between the coadsorbates were taken into account. The presence of

a strongly bound coadsorbate on a bcc(100) surface was shown to shift the associative desorption spectrum of adsorbed atoms to lower temperatures. The simulated TPD spectra were found to be in qualitative agreement with experimental results for H₂ coadsorbed with strongly bound atomic species on Mo(100) [140] and Fe(100) [141] surfaces.

TPD spectra were also simulated for the concurrent desorption of B molecules and the associative desorption of A atoms from an fcc(100) surface [139]. Two types of behavior were observed: in one case, both species exhibited new low-temperature features not present in the TPD spectra of A and B when each species was adsorbed alone. In the second case, only the more weakly bound species displayed new spectral features. These types of behavior are in qualitative agreement with what has been observed for the codesorption of CO and H₂ from Ni(100) [142] and Rh(100) [143] surfaces, respectively. Figure 5.13a shows TPD spectra for A₂ and B when each species is adsorbed separately, and when both species are coadsorbed. The experimentally observed TPD spectra for CO and H₂ on a Ni(100) are shown in fig. 5.13b [142]. A comparison of the two figures demonstrates the qualitative agreement between the model predictions and the experimental observations. In fig 5.14 are shown the activation energy profiles versus coverage for both the pure component and codesorption simulations. For the A₂ and B activation energy profiles when both species are coadsorbed, the activation energy profiles are lower than for the respective pure component activation energies and show larger variation with coverage.

5.4 Compensation Effect

Analysis of experimental TPD spectra using empirical rate expressions of the form

$$r_d = v_d^{(\alpha)}(\theta) \exp[-E_d(\theta)/k_b T] \theta^\alpha \quad (5.8)$$

have shown evidence of a compensation effect, namely that $v_d^{(\alpha)}(\theta)$ and $E_d(\theta)$ vary in the same manner for changes in coverage and temperature [144-149]. In particular, it has often been observed that both $v_d^{(\alpha)}(\theta)$ and $E_d(\theta)$ decrease with increasing adsorbate coverage. While the decrease in $E_d(\theta)$ is readily attributable to repulsive lateral interactions, the large decreases in $v_d^{(\alpha)}(\theta)$ (e.g., up to nine orders in magnitude) have not been explained satisfactorily. Seebauer et al. [146] have recently reviewed different theoretical representations for the preexponential factor. Although several of the theories predict compensation behavior, none are able to account for the large variations such as those determined from the analysis of experimental data. The authors suggested, however, that models which account for the phonon vibrational modes of the solid and for surface reconstruction may be able to account for the large variations observed in the preexponential factor as a function of coverage.

The apparent inconsistency between the large variations in $v_d^{(\alpha)}(\theta)$ deduced from experimental data and the significantly smaller variations suggested by various physical models can be reconciled in the following manner. To begin with, it must be recognized that eq.

5.8 is written intuitively as a generalization of eqs. 2.25 and 2.26 rather than on the basis of a physical theory of desorption. For example, reference to eqs. 2.48-2.51 shows that the rate of desorption cannot be factored naturally into a preexponential factor, an exponential factor, and a factor describing the coverage dependence. This indicates that eq. 5.8 is not a valid physical representation of the kinetics of desorption. As a consequence, while eq. 5.8 can be fitted to experimental desorption data, the values of $v_d(\alpha)(\theta)$ and $E_d(\theta)$ determined by this means must be regarded as apparent values. Consistent with this interpretation, fig. 2.5 shows that the apparent preexponential factor deduced from a lattice-gas model in the quasi-chemical approximation exhibits a strong dependence on θ , even though the preexponential factor incorporated into the model is, in fact, assumed to be constant. Inspection of eqs. 2.48-2.51 reveals further that the observed coverage dependence of the apparent preexponential factor can be ascribed to $f_d'(\theta, \text{WAA})$ and $f_d''(\theta, \text{WAA})$ which are each comprised of weighted terms of the probabilities of finding an adsorbate in a specific configuration of nearest neighbors. Viewed in this fashion, the large variations in the apparent preexponential factor can be attributed to the configurational and energetic effects of nearest-neighbor interactions, rather than to an explicit dependence of the preexponential factor on coverage.

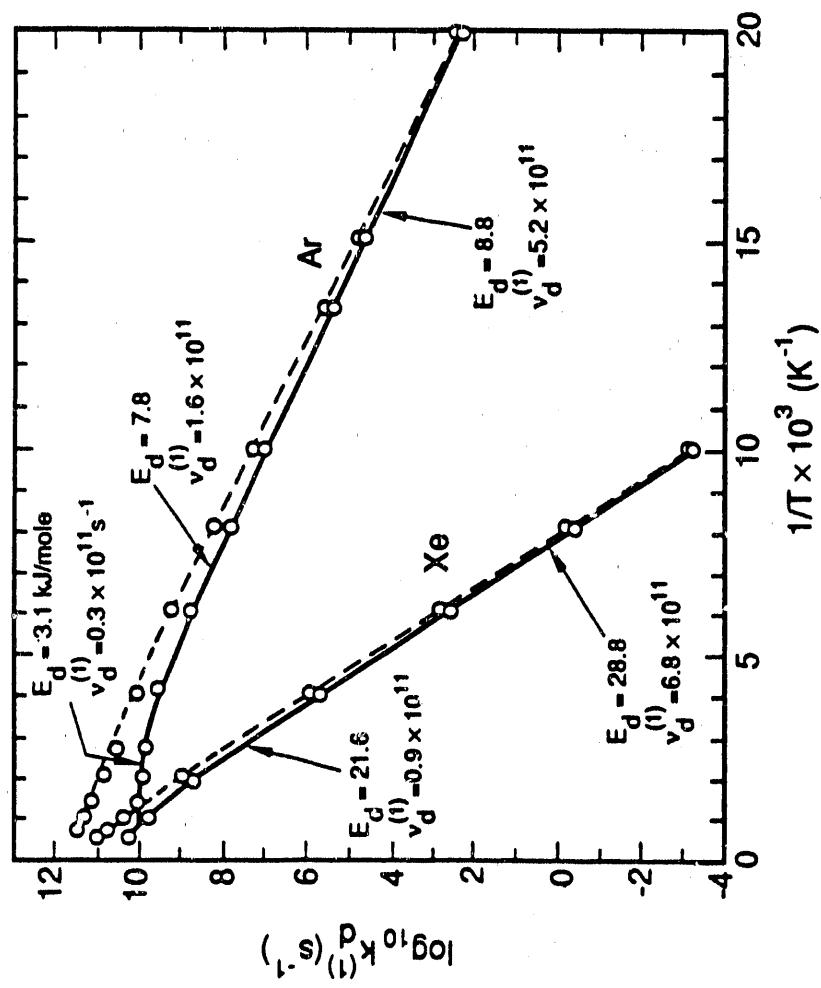


Fig. 5.1 Arrhenius plot for Ar and Xe desorption from Pt(111). Solid and dashed curves are dynamically corrected and uncorrected TST, respectively [29].

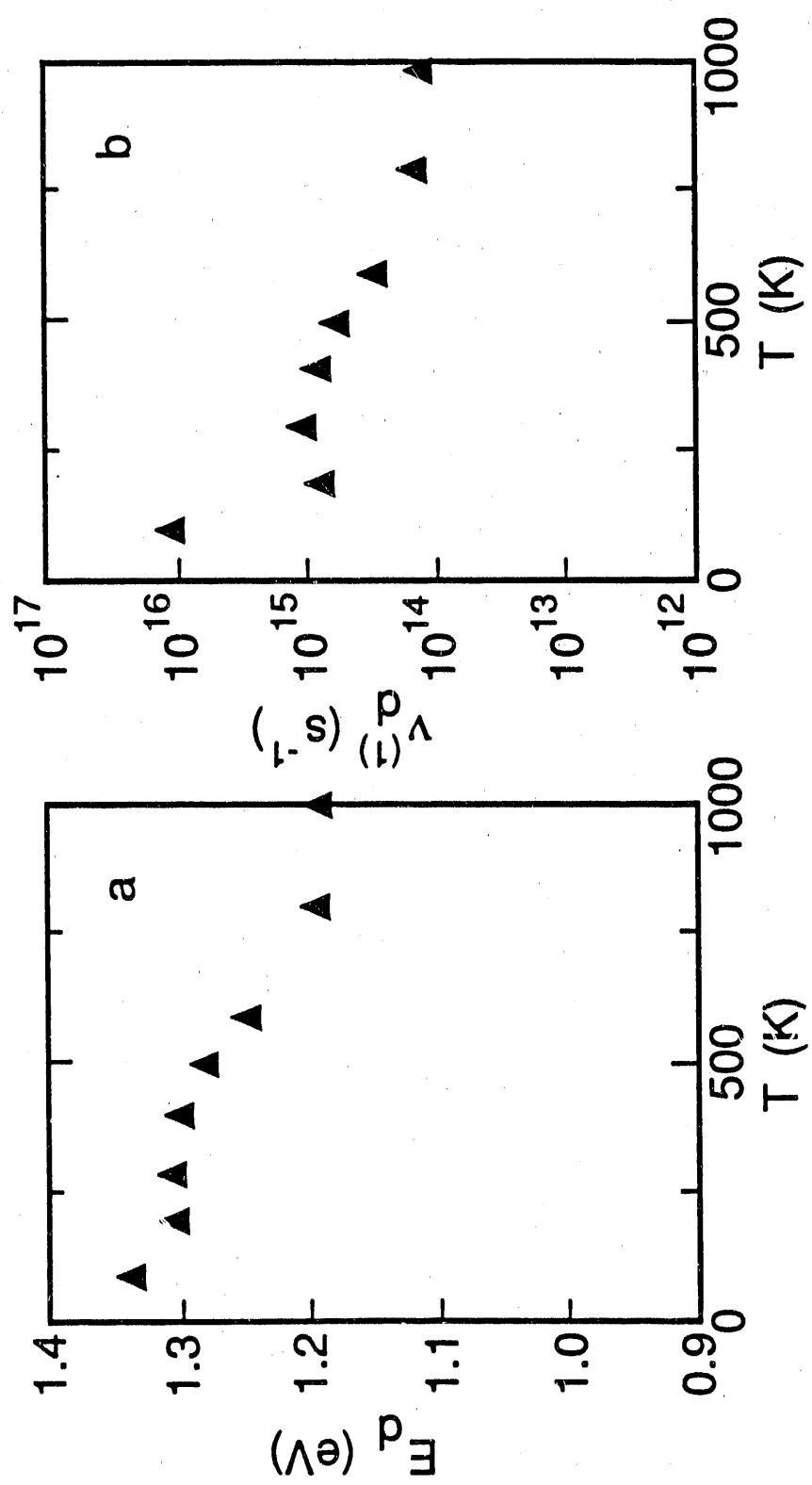


Fig. 5.2 Activation energy (a) and preexponential factor (b) profiles versus T for CO desorption from Ni(111) [36].

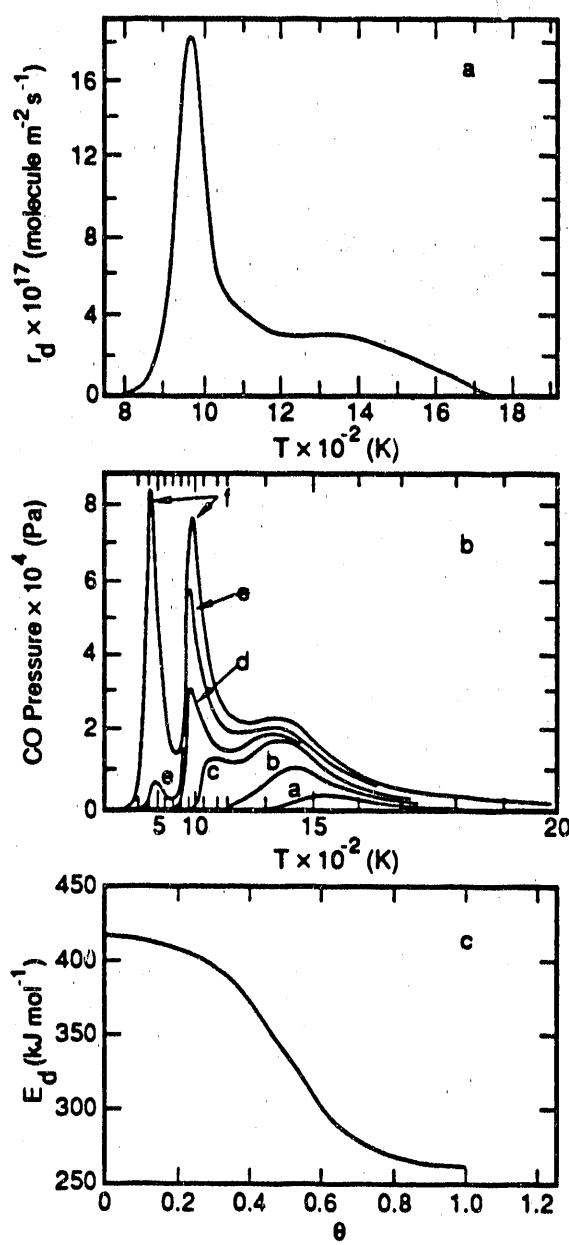


Fig. 5.3 a. Predicted CO TPD spectra from a W ribbon [124]. b. Experimental CO TPD spectra from a W ribbon [125]. The CO exposures (Pa s) are: a) 1.9×10^{-5} ; b) 6.0×10^{-5} ; c) 1.6×10^{-4} ; d) 1.9×10^{-4} ; e) 2.5×10^{-4} ; and f) 3.1×10^{-3} . c. Predicted activation energy profile versus coverage [124].

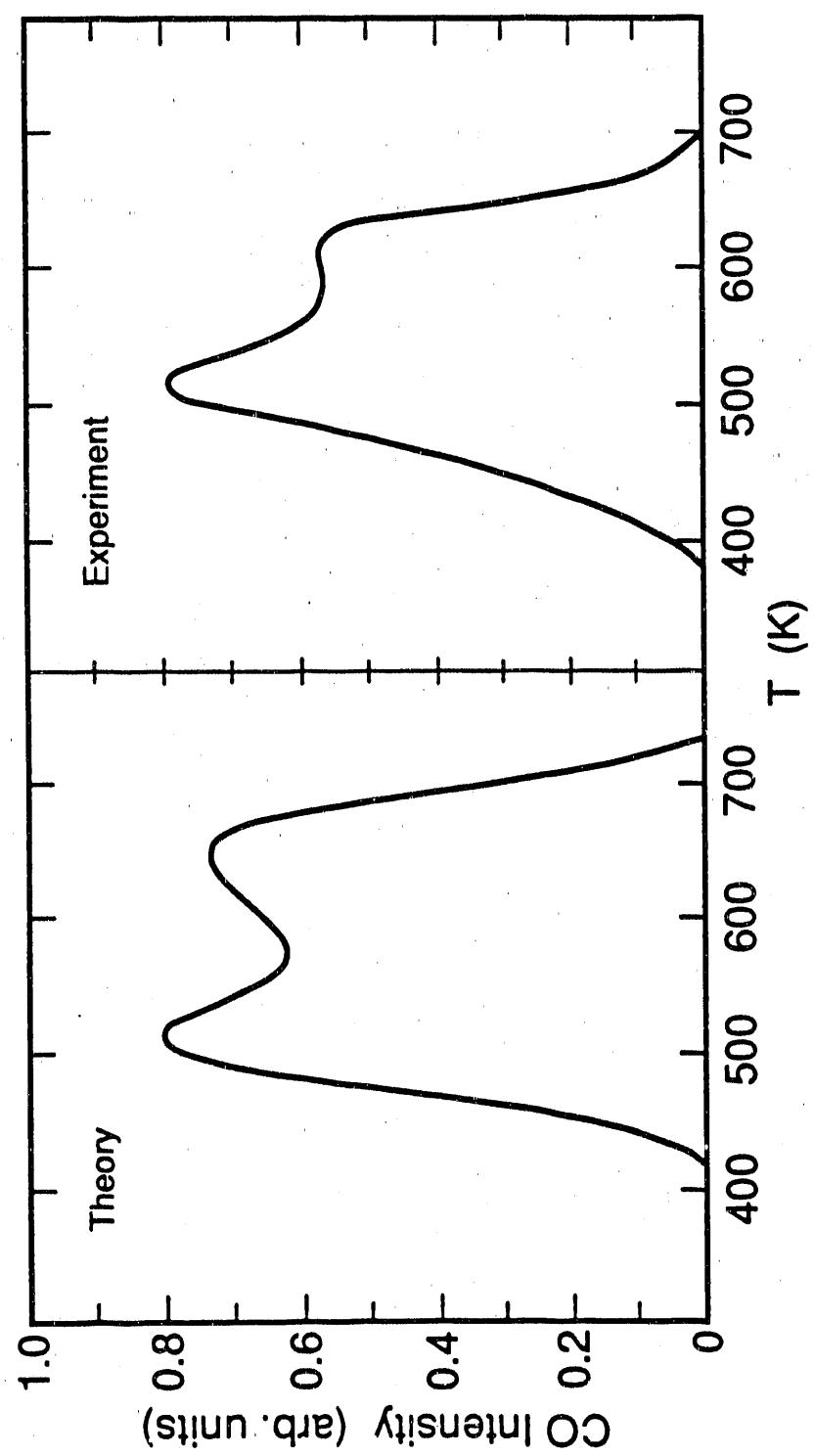


Fig. 5.4 Predicted [126] and experimental [127] CO TPD spectra from Ir(110) for $\theta_{CO} = 0.82$.

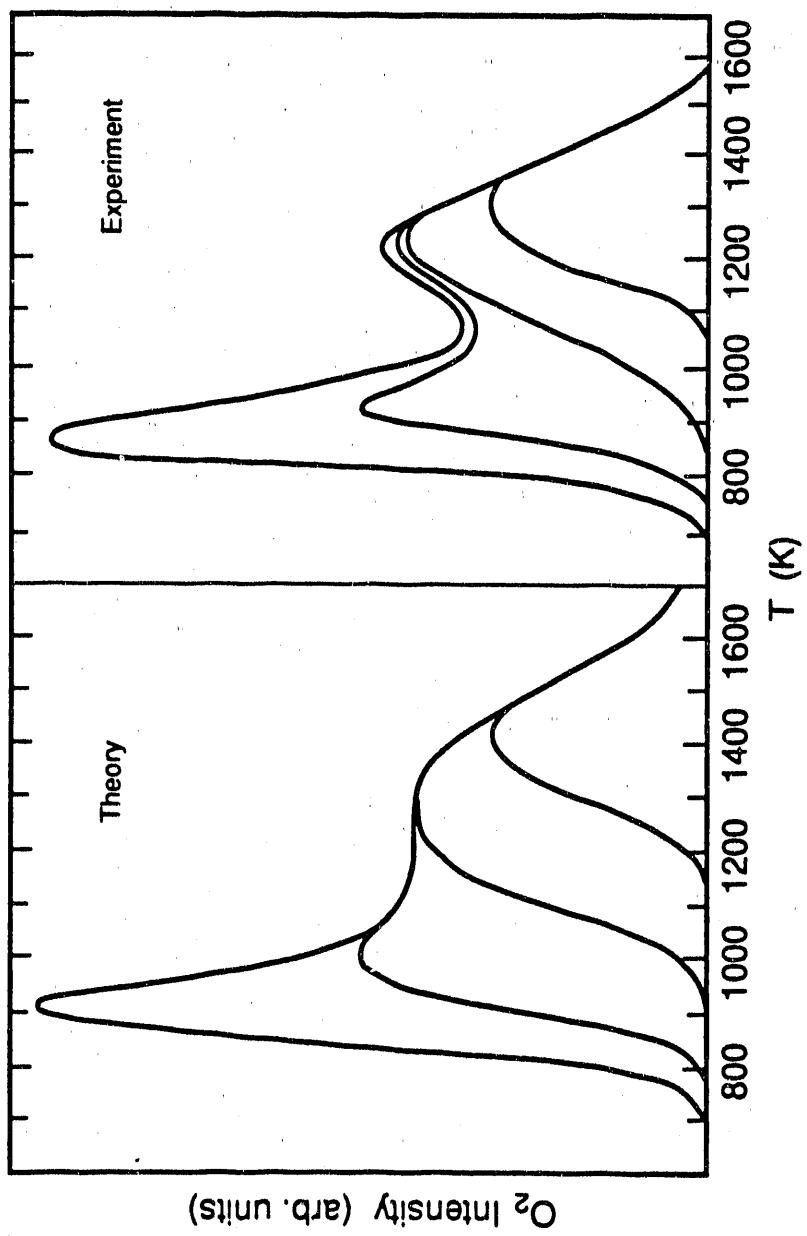


Fig. 5.5 Predicted [126] and experimental [128] O_2 TPD spectra from Ir(110). Curves in both panels correspond to oxygen coverages of 1.0, 0.75, 0.50, and 0.25.

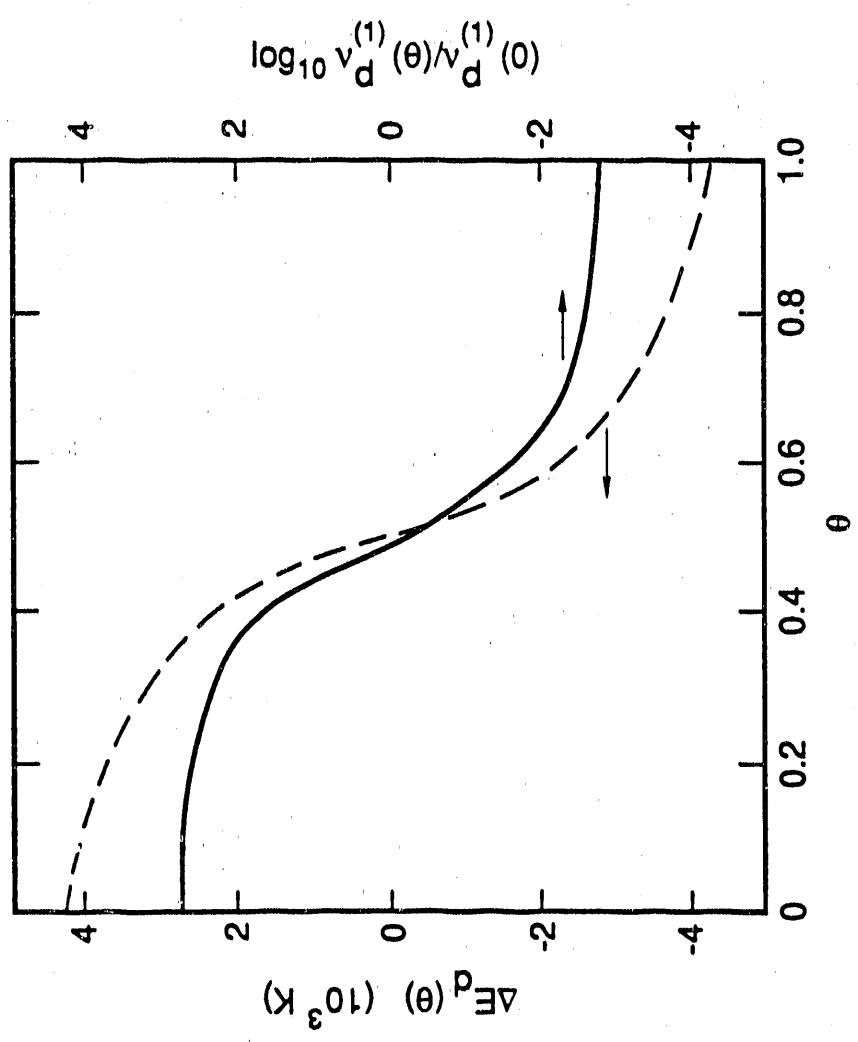


Fig. 5.6 Variation in $\Delta E_d(\theta)$ and $v_d^{(1)}(\theta)/v_d^{(1)}(0)$ versus coverage when surface reconstruction occurs [60].

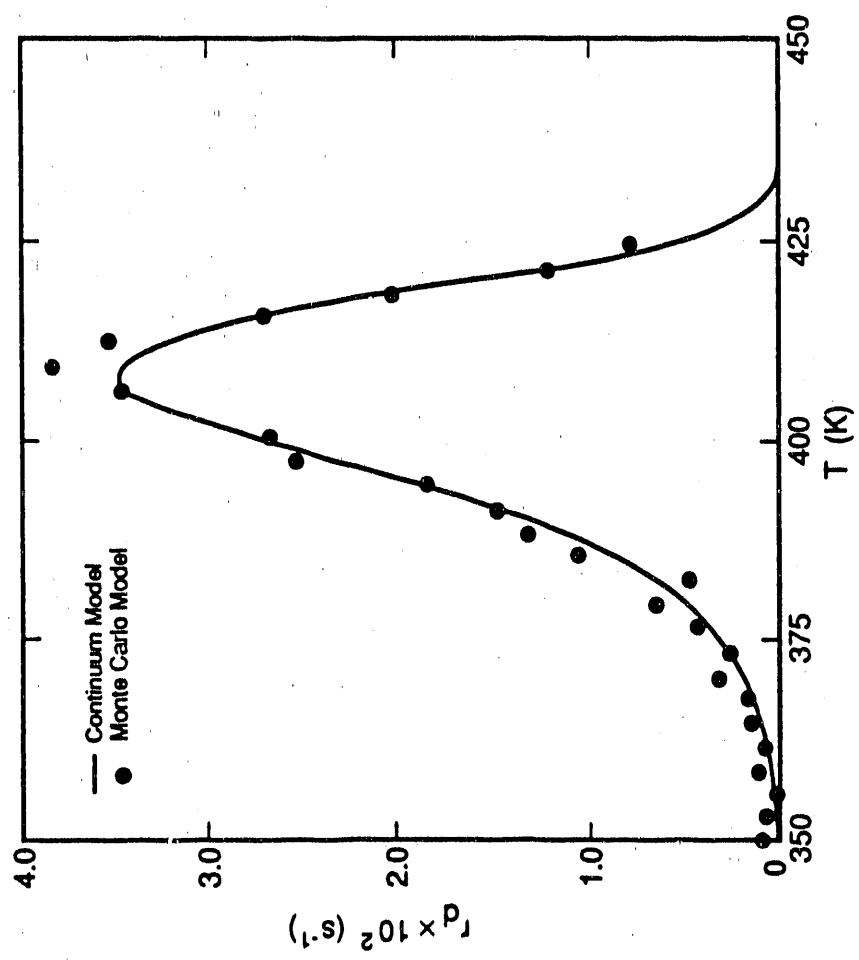


Fig. 5.7 Comparison of Monte Carlo and continuum models for the simulation of non-associative TPD spectra in the absence of lateral interactions [132].

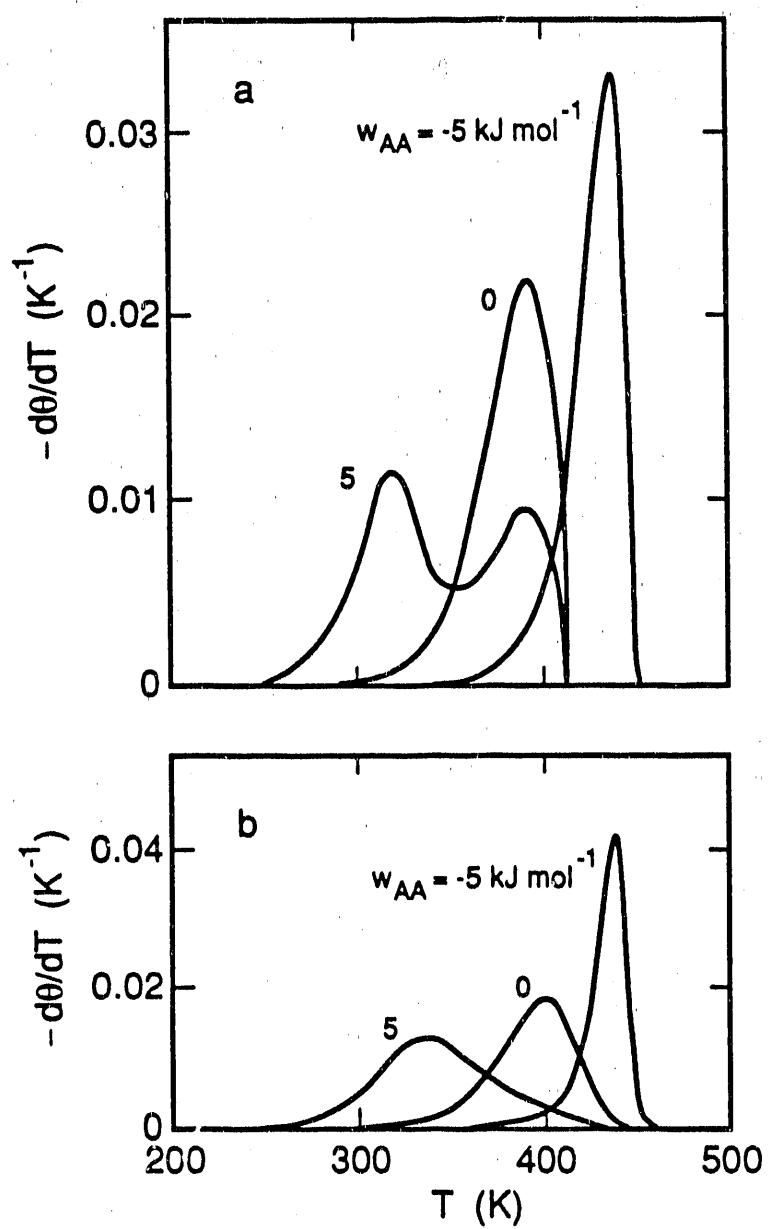


Fig. 5.8 a. Monte Carlo simulation of TPD spectra [133]. b. Continuum simulation of TPD spectra [134]. Lateral interactions between adsorbates in the continuum simulations are described with the Braggs-Willaims approximation.

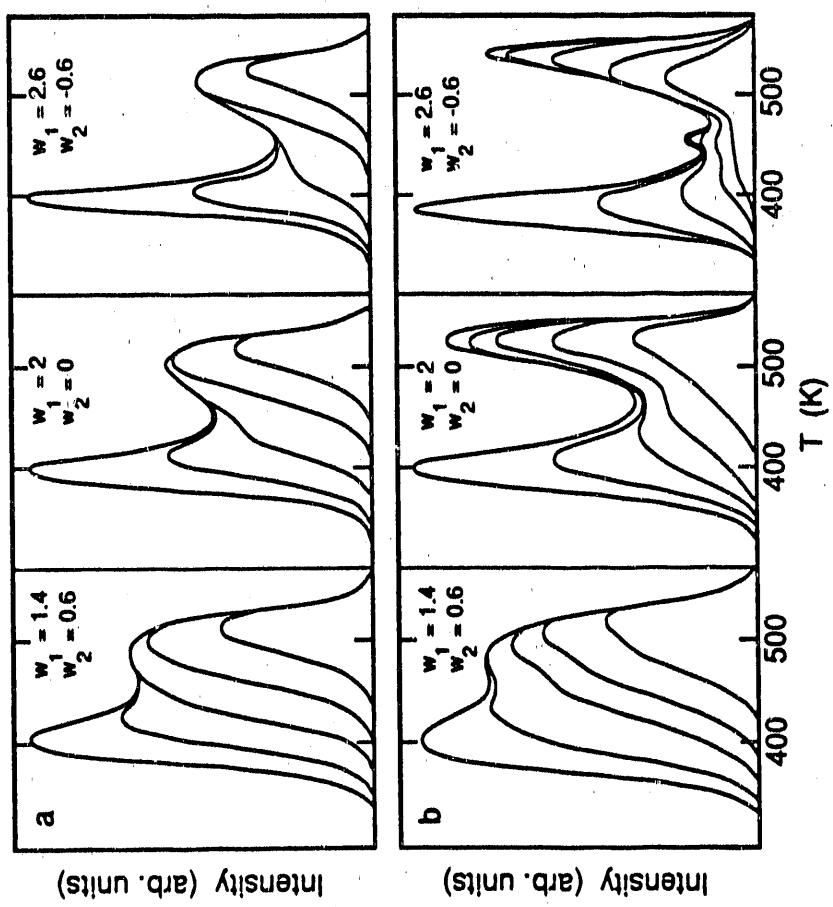


Fig. 5.9 Simulation [135] of TPD spectra for both nearest-neighbor interactions w_1 and next-nearest neighbor interactions w_2 . a. Continuum model. b. Monte Carlo model. The lateral interactions are in kcal/mol.

Fig. 5.9

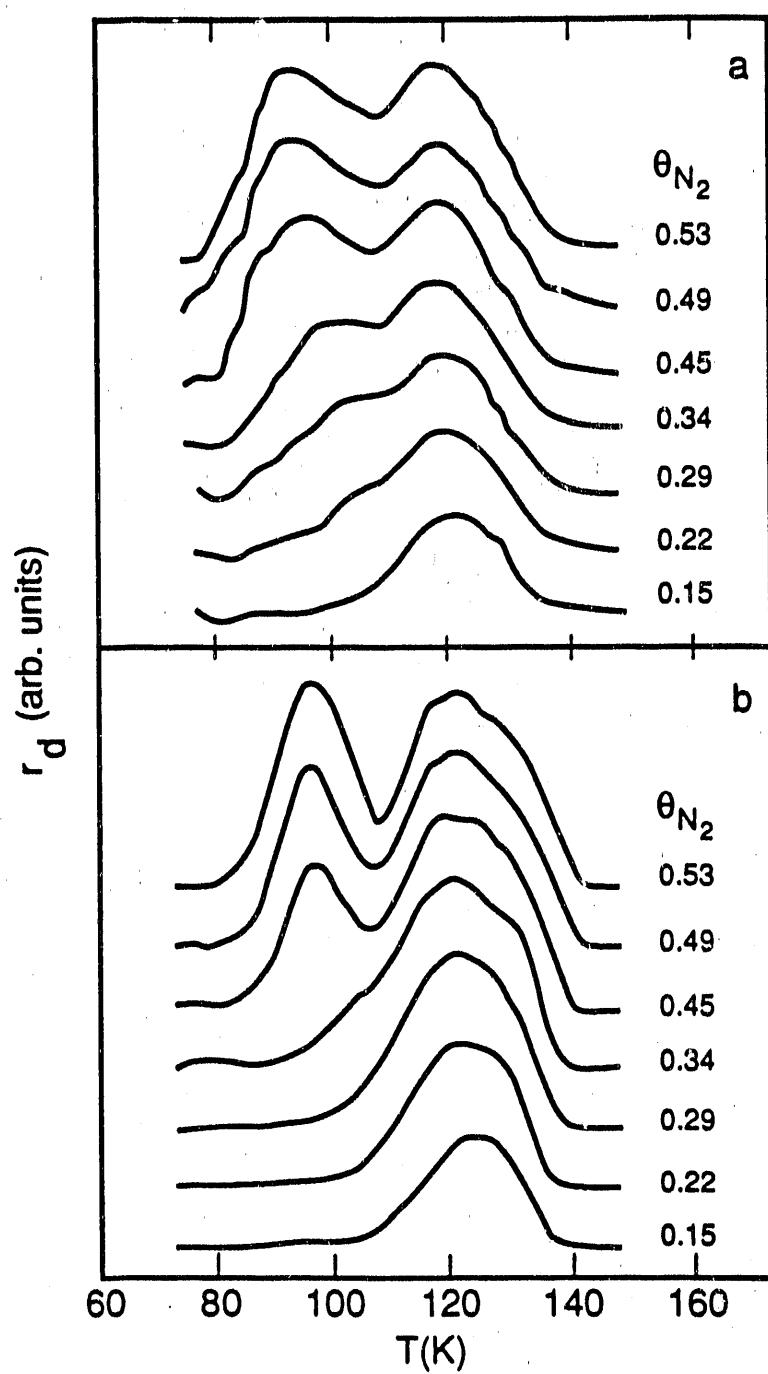


Fig. 5.10 Experimental (a) [62,63] and simulated (b) [61] TPD spectra of molecular N_2 from Ru(001).

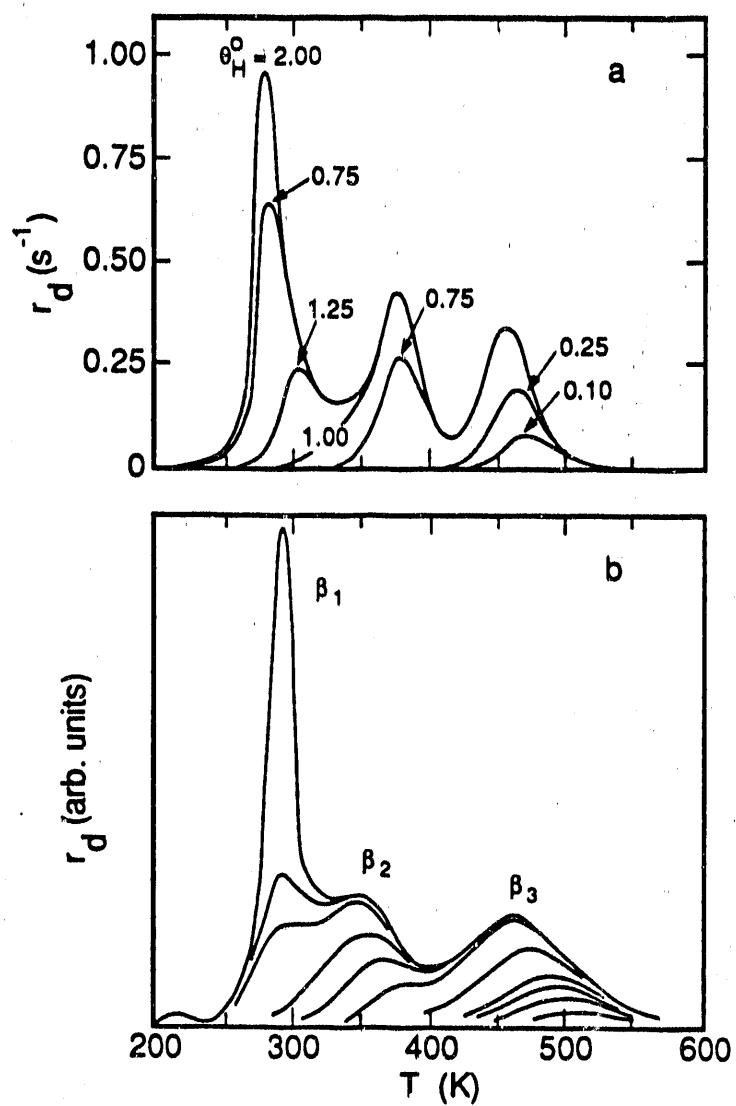


Fig. 5.11 Simulated (a) [132] and experimental (b) [136] TPD spectra of H_2 from $\text{Mo}(100)$.

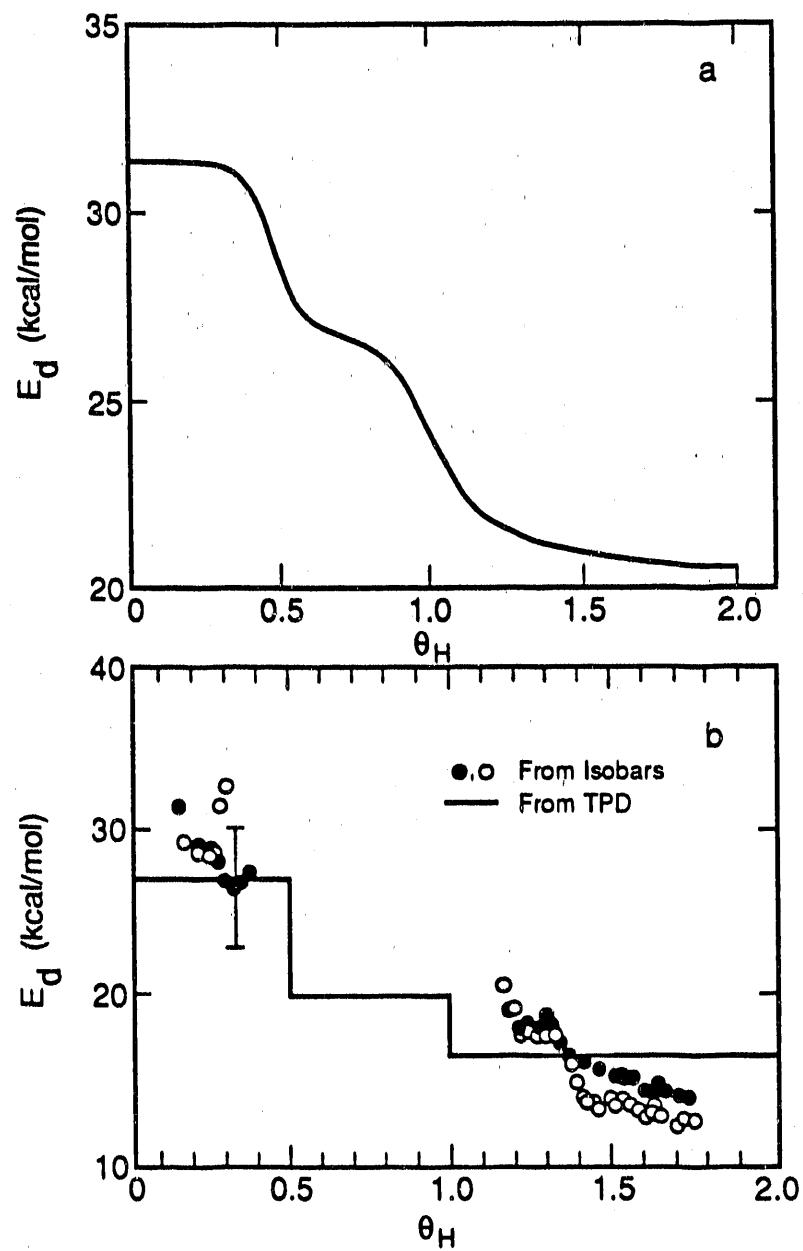


Fig. 5.12 Simulated (a) [132] and experimental (b) [137] variation in E_d with θ_H for H₂ desorption from Mo(100).

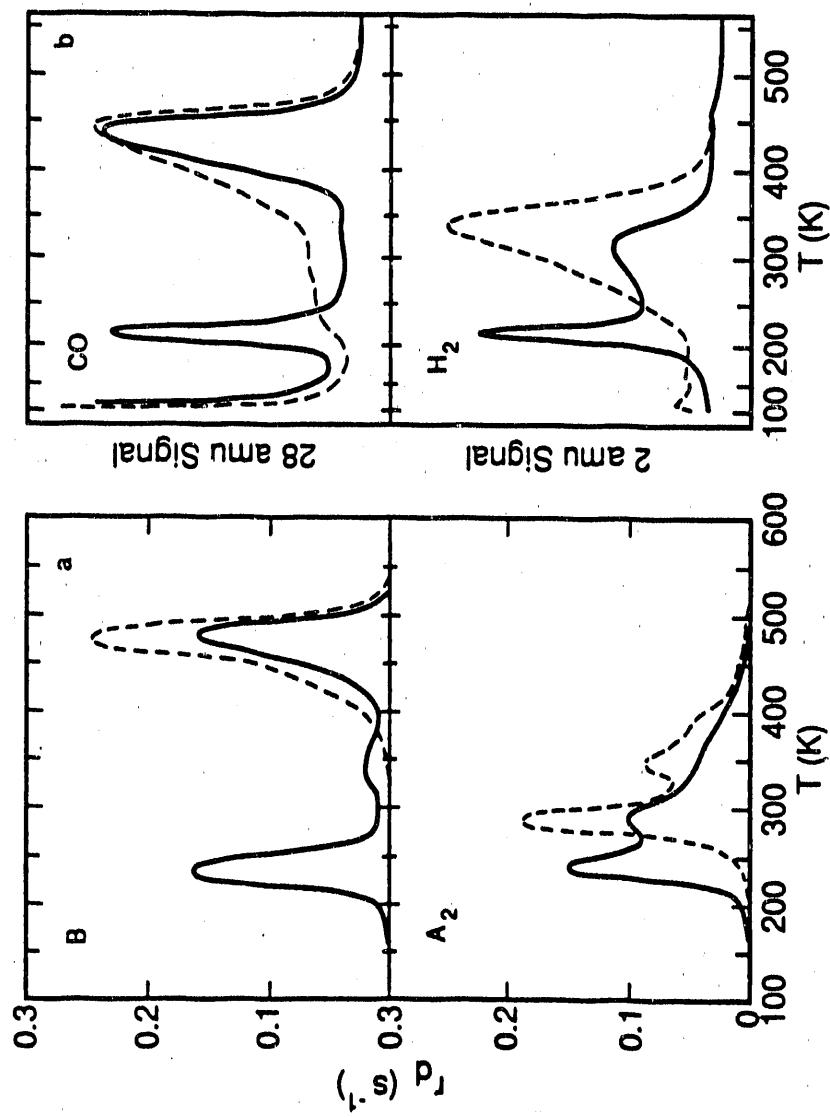


Fig. 5.13 a. Simulated TPD spectra of B and A₂ [139]; b. Experimental TPD spectra of CO and H₂ [142]. The dashed curves correspond to desorption of each species adsorbed separately, and the solid curves correspond to desorption of both species coadsorbed.

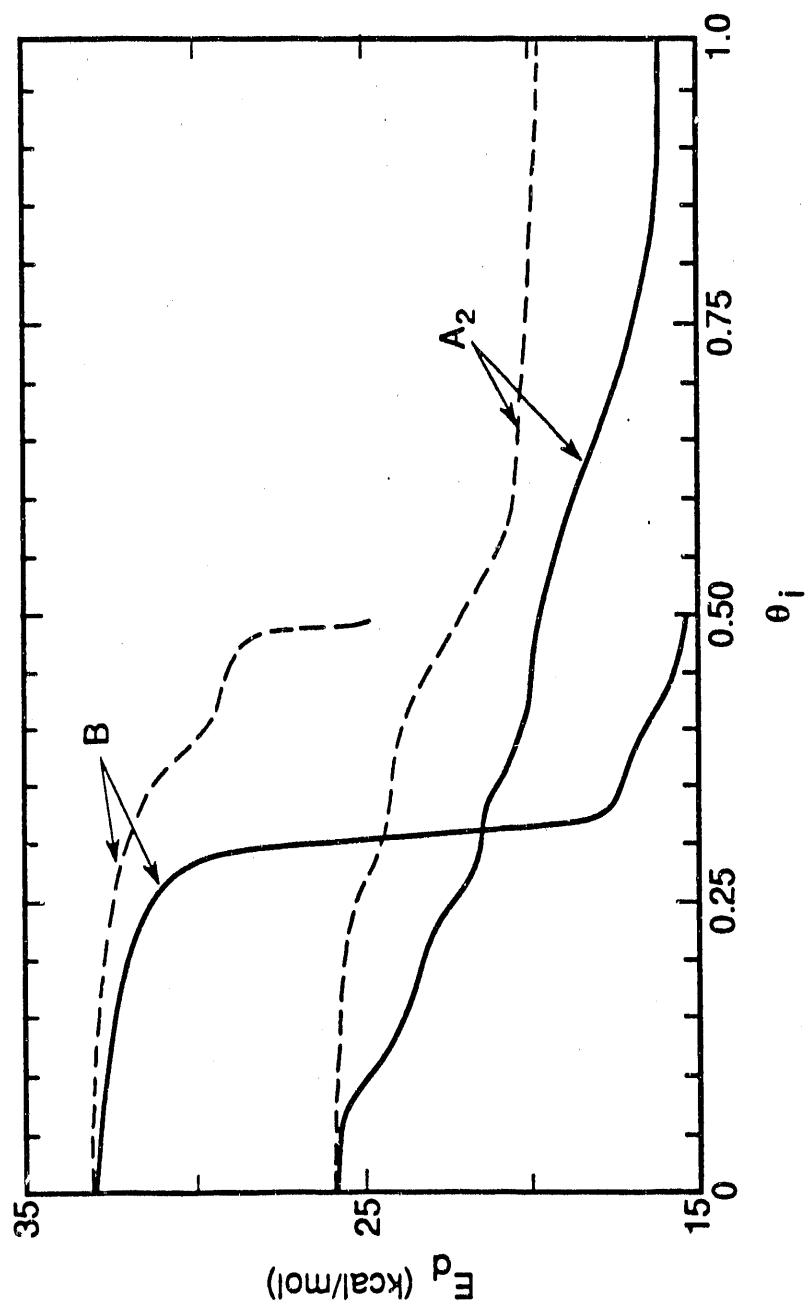


Fig. 5.14 Variation in E_d with θ_i for the simulations in fig. 5.13a [139]. The dashed curves correspond to desorption of each species adsorbed separately, and the solid curves correspond to desorption of both species coadsorbed.

6.0 REACTIONS

6.1 Theoretical Prediction of k_r

With the exception of absolute rate theory, none of the theoretical approaches described in Section 2 have been used to predict rate coefficients for reactions occurring on metal surfaces. Estimates from absolute rate theory of the preexponential factor v_r for Langmuir-Hinshelwood reactions range from 10^4 - 10^4 $\text{cm}^2 \text{ s}^{-1}$ [37]. This approach suffers, however, from the need to make a number of ad hoc assumptions regarding the structure and molecular properties of the transition state. Estimates of the activation energy for surface reactions can be made using the BOC-MP approach. Shustorovich [17] has shown that where comparison can be made with experiment, the BOC-MP method provides good estimates of E_r . This conclusion is illustrated by the examples presented in Table 6.1.

6.2 Simulation of TPSR Spectra and Steady-State Dynamics

Numerous authors have reported theoretical descriptions of reactions occurring on well-defined metal surfaces and compared the results of such simulations with experiment. Two types of reaction studies have been considered. The first involves coadsorption of both reactants followed by a progressive heating of the surface to raise the temperature. The rates at which the reactants and products desorb from the surface are then followed as

a function of time or temperature. In the second type of experiment, reactants are passed over a surface maintained at a constant temperature, and the steady-state rates of product formation are observed. Simulations of both types of experiments have been carried out using continuum models.

Bridge and Lambert [157] have used a lattice-gas model with the QCA to treat the associative recombination of N_s and O_s produced by dissociative adsorption of NO on Pt, Ni, and Ru surfaces. The associative desorption kinetics were simulated for two cases of interactions between the atomic adsorbates. In the first case, repulsive interactions between all atomic species (N-N, O-O, and N-O) were taken into account. In the second case, only repulsive N-N interactions were considered. For both cases of interactions, N₂ desorbed at low temperature, O₂ desorbed at high temperature, and no appreciable NO was formed. This selectivity to N₂ and O₂ products was explained as follows. At low temperatures, formation of N₂ is energetically favored, and atomic nitrogen is depleted from the surface. At intermediate temperatures for which the formation of NO is energetically favorable, the amount of nitrogen remaining on the surface is too small to form appreciable NO. At still higher temperatures, atomic oxygen is the only species remaining on the surface and it desorbs as O₂. The negligible selectivity to NO product formation is thus seen to result from the lack of atomic nitrogen on the surface at temperatures favorable for NO formation. Comparison of the simulated TPD spectra with those observed experimentally [156,158,159] led to the conclusion that O-O repulsions are significant for all three metals. In contrast, the N-N

interactions were found to be either zero or slightly attractive on Pt and Ni, but definitely repulsive on Ru.

The oxidation of CO to CO₂ on an Ir(110) surface has been examined by Zhdanov [126] with a lattice-gas/QCA model. Since the oxidation of CO is accompanied by CO desorption, both elementary processes were included in the model. The values of $v_{d,CO}(1) = 1011$ s⁻¹, $E_{d,CO}^\circ = 33$ kcal/mol, and $w_{CO-CO} = 2.2$ kcal/mol used in these simulations were chosen to obtain a successful representation of the TPD spectrum of CO in the absence of adsorbed oxygen. Likewise, the values of $v_{d,O_2}(2) = 3.5 \times 10^{-4}$ cm² s⁻¹, $E_{d,O_2}^\circ = 67$ kcal/mol, and $w_{OO} = 3.6$ kcal/mol were chosen to obtain a good representation of the TPD spectrum of O₂ in the absence of adsorbed CO. The remaining parameters, $v_r = 1013$ s⁻¹, $E_r^\circ = 37$ kcal/mol, and $w_{CO-O} = 1.7$ kcal/mol, were chosen by trial and error to obtain the best representation of the TPSR spectra. A comparison between theory and experiment [160] is shown in fig. 6.1. It is seen that most features of the experimentally observed TPSR spectra are reproduced in the simulation.

As an additional part of his study, Zhdanov [126] determined the apparent activation energies for CO desorption and oxidation using the following relationships:

$$E_{d,CO} = -k_b T \ln \left[\frac{k_{d,CO}^{(1)}}{v_{d,CO}^{(1)}} \right] \quad (6.1)$$

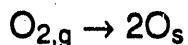
$$E_r = -k_b T \ln \left[\frac{k_r}{v_r 2 \theta_{CO} \theta_O} \right] \quad (6.2)$$

Figure 6.2 shows that both $E_{d,CO}$ and E_r decrease with increasing θ_{CO}

and θ_0 which is a consequence of the repulsive CO-CO, CO-O and O-O interactions.

Zhdanov has also reported on the oxidation of CO over Ir(111) under both TPSR [161] and steady-state conditions [162]. The parameters used to simulate the TPSR spectra were $v_r = 10^{13} \text{ s}^{-1}$, $E_r^\circ = 31 \text{ kcal/mol}$, and $w_{\text{CO-CO}} = 0.7 \text{ kcal/mol}$, $w_{\text{OO}} = 0.8 \text{ kcal/mol}$, and $w_{\text{CO-O}} = 0.5 \text{ kcal/mol}$. The experimental TPSR spectra [163] are shown in fig. 6.3a and the corresponding simulations of these spectra are presented in fig. 6.3b. Comparison of these figures indicates that the TPSR spectra determined from the lattice-gas model are in semi-quantitative agreement with those observed experimentally. Of particular interest is the fact that the lattice-gas model correctly predicts the strong upscale shift in the position of the CO_2 peak with increasing initial oxygen coverage. This trend cannot be represented successfully if lateral interactions are neglected.

The steady-state oxidation of CO over Ir(111) was described by Zhdanov [162] in terms of the following mechanism:



At steady state, the kinetics of the reaction are given by

$$S_{\text{CO}}^{(1)}(\theta) F_{\text{CO}} = r_{\text{d,CO}} + r_r \quad (6.3)$$

$$2S_{\text{O}_2}^{(2)}(\theta) F_{\text{O}_2} = r_r \quad (6.4)$$

where $S_i^{(\alpha)}(\theta)$ and F_i are the sticking coefficient and the flux of component i , respectively. The rate parameters appearing in the

expressions for $r_{d,CO}$ and r_r are identical to those used for describing the oxidation of CO during TPSR. The dependencies of the CO and O₂ sticking coefficients on coverage were approximated by:

$$S_{CO}^{(1)}(\theta) = S_{CO}^{(1)}(0)(1 - \theta_{CO} - \theta_O) \quad (6.5)$$

$$S_{O_2}^{(2)}(\theta) = S_{O_2}^{(2)}(0) P_{vv} \quad (6.6)$$

where $S_{CO}^{(1)}(0) = 1.0$, $S_{O_2}^{(2)}(0) = 0.5$, and P_{vv} is the probability that two nearest-neighbor sites are vacant. A plot of the steady-state oxidation rate as a function of temperature and pressure is shown in fig. 6.4. It is evident that the theoretical model provides a good description of the experimental data [163,164], and in particular, the maximum in the rate at 600 K.

The steady-state oxidation of H₂ over Pt(111) has been analyzed by Zhdanov [165] using an approach similar to that used for the description of CO oxidation over Ir(111). The reaction mechanism is assumed to be:

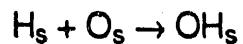
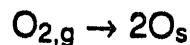


Figure 6.5 compares the steady-state reaction rates observed experimentally [166] with those predicted on the basis of the lattice-gas model. The theoretical model provides a qualitatively correct description of the experimental data.

Sundaresan and Kaza [167] have explored the effects of limited adsorbate mobility on the rates of surface reactions. The formulation of the lattice-gas model used for this study was based

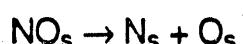
on that used to treat the effects of adsorbate mobility on non-associative desorption described in Section 5.3 [131]. It was demonstrated that limited adsorbate mobilities can significantly alter the rates of surface reactions depending on the lateral interactions and mobilities of the adsorbed species.

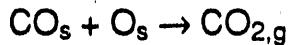
An analysis of the steady-state kinetics of the CO-O₂ and CO-NO reactions over single-crystal Rh(111) and supported Rh/Al₂O₃ has been reported by Oh et al. [168]. The oxidation of CO is represented by the mechanism described above. In the limit of high CO coverage, the authors show that the rate of CO₂ formation can be written as

$$r_r = \frac{2 k_{a,O_2}}{k_{a,CO}} \frac{P_{O_2}}{P_{CO}} \left\{ v_{d,CO}^{(1)} \exp \left[\frac{-(E_{d,CO}^0 - \phi_{CO}\theta_{CO})}{k_b T} \right] \right\} \quad (6.7)$$

In this expression, the activation energy for CO desorption is seen to depend linearly on the coverage of CO. With the values of E_{d,CO}⁰ = 31.6 kcal/mol, ϕ_{CO} = 4.5 kcal/mol, and $v_{d,CO}^{(1)} = 1.6 \times 10^{14} \text{ s}^{-1}$, a quantitative fit of rate data versus 1/T could be achieved for both single crystal and supported Rh catalysts over four orders of magnitude in the rate.

In the same study, Oh et al. [168] modelled the kinetics of the CO-NO reaction for the same Rh catalysts using the following mechanism:





The rate of CO desorption was assumed to decrease linearly with both CO and N coverage and the rate of N₂ desorption was assumed to decrease linearly with N coverage. Although an analytical expression for the rate of N₂ formation could not be obtained in this case, a numerical solution of the problem was achieved. Using the parameter values listed in table 6.2, good agreement was obtained between the measured and predicted rates of NO reduction by CO over a Rh(111) surface. A similar level of agreement was achieved for simulations of NO reduction over Rh/Al₂O₃; however, in this instance a smaller value of the rate coefficient for NO dissociation was required than for Rh(111). The difference in the rate coefficients for Rh(111) and Rh/Al₂O₃ was attributed to the structure sensitivity of the NO dissociation process.

6.3 Effects of Adsorbate Islanding

Lateral interactions between adsorbed species favor their organization into non-random, two-dimensional distributions. Under appropriate conditions, an adsorbate may even form islands of macroscopic dimensions. Such a situation can arise for coadsorbed reactants A and B when $w_{AA} < 0$, $w_{BB} = w_{AB} = 0$, and $T < T_c$, where T_c is the critical temperature of A, below which islands of A form. If the reactant B cannot penetrate the A islands, reaction will only occur at the island boundaries. Consequently, the rate of product formation will depend on the number of A islands, the length and shape of their boundaries, and on related dynamical properties such

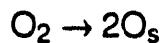
as the rates of island growth and diffusion of B. These characteristics depend in turn on initial conditions such as initial coverages, order, and time delay between A and B adsorption, and temperature. A well-known system exhibiting the behavior just described is $O_s + CO_s \rightarrow CO_{2,g}$ on various metal surfaces [2,175]. In this case, the oxygen atoms are the aggregating species and CO is the mobile reactant. Both Monte Carlo and continuum models have been used to simulate the effects of island formation on reaction kinetics.

Silverberg and Ben-Shaul [176] have used a Monte Carlo model to examine the consequences of the issues raised in the preceding paragraph. The reaction considered was $As + B_s \rightarrow AB_g$ with only attractive interactions between A adsorbates being taken into account. In the initial segment of the simulation, species A was adsorbed on the surface and then allowed to diffuse in order to either fully or partially reach an equilibrium adsorbate configuration. Species B was subsequently adsorbed randomly on unoccupied surface sites. As a consequence of the attractive interactions between A adsorbates, islands of A form, and the reaction occurs primarily at the ramified perimeters of the A islands. The power law dependence of the reaction on the coverage of A was determined to be 0.60, which is larger than the value of 0.50 assumed for perfectly uniform islands. The rate of reaction was further shown to be sensitive to the initial coverage of reactants and to the time allowed for the A species to form islands.

In a series of related studies, Silverberg and Ben-Shaul [177-180] have applied a combination of Monte Carlo and lattice-gas

models to simulate the TPSR of $A_s + B_s \rightarrow AB_g$. Interactions between all species (A-A, A-B, and B-B) were included in the model. As a consequence of attractive A-A interactions, species A formed islands on the surface. The Monte Carlo segment of the algorithm was used to determine the non-equilibrium, non-uniform distribution of A atoms and the rate of AB formation. The lattice-gas models were used to determine the equilibrium distribution of B on surface sites not occupied by A. The shape and location of the simulated TPSR spectra were sensitive to the magnitude and sign of the A-B and B-B interactions as well as to the amount of time allowed for the formation of A islands. It was observed that the two peaks in the AB TPSR spectrum occurred from the reaction of AB species from two types of surface environments. The low-temperature peak corresponded to the reaction of isolated AB pairs whereas the high-temperature peak corresponded to desorption of AB pairs located in the vicinity of other A species.

The effects of lateral interactions and phase separation on the oxidation of SO_2 to SO_3 over Pt have been investigated by Kaza and Sundaresan [181]. The reaction system was analyzed using a lattice-gas model with the Bragg-Williams approximation. The reaction was postulated to proceed via the following steps:



It was observed that the assumption of sufficiently strong attractive interactions between adsorbed SO_3 molecules resulted in a phase separation into a condensed (liquid-like) and a dilute (gas-

like) phase. The inclusion of attractive SO_3 - SO_3 interactions and SO_3 phase separation resulted in a qualitatively correct prediction of the dependence of the rate of SO_2 oxidation on reactant partial pressures [182]. Similar agreement could not be achieved if a random distribution of SO_3 was assumed.

While lattice-gas models provide a sound theoretical basis for predicting the conditions under which island formation can occur, such models require knowledge of the strength of adsorbate-adsorbate interactions. Several authors have shown that many of the consequences of island formation can be explained using Langmuir-Hinshelwood models in which the presence of islands is postulated a priori. An example of this approach is given by Barteau et al. [175], who simulated the isothermal oxidation of CO on a reconstructed $\text{Pt}(100)-(5\times20)$ surface. The rate of reaction between CO and preadsorbed O_2 was written as

$$r_r = k_r \theta_{\text{CO}} \theta_{\text{O}}^{1/2} \quad (6.8)$$

on the assumption that the reaction occurs solely at the perimeter of oxygen islands. An analytical solution to eq. 6.8 was obtained by writing species conservation equations and by assuming that the rates of CO adsorption and desorption were much greater than the rate of reaction. As can be seen in fig. 6.6, it was found that for a range of temperatures, the Langmuir-Hinshelwood model provides an adequate description of the reaction dynamics.

Mukesh et al. [183] have also investigated the oxidation of CO at the perimeter of islands. Two models were examined. In the first, it was assumed that COs aggregates into islands and is surrounded by Os. If the number of CO islands is assumed to be

constant, independent of CO coverage, then $r_r \propto \theta_O(\theta_{CO})^{1/2}$. In the second model, CO_s and O_s are assumed to form separate islands, and hence, $r_r \propto (\theta_O\theta_{CO})^{1/2}$. A comparison of the agreement between the CO islands model and experimental data obtained for supported Pt catalysts is shown in fig. 6.7. The authors note that a similar level of agreement could be achieved if both reactants were assumed to form islands.

The kinetics of H₂ reacting with preadsorbed oxygen on a Pt(111) surface have been modelled by Gland et al. [184]. Oxygen was assumed to form a regular array of islands. The rate of water formation was shown to depend on the size and shape of the oxygen islands as well as on the availability of atomic hydrogen in the peripheral region around each oxygen island. The model provides a qualitatively correct description of the high reaction rates observed at low oxygen coverages, as well as the first-order dependence of the reaction rate on the hydrogen flux to the surface.

TPSR spectra of CO and NO coadsorbed on Pt(100) have been simulated by Fink et al. [185]. The reaction mechanism was identical to that given earlier for NO reduction by CO, with the rate limiting step assumed to be the dissociation of NO. A further assumption of the model was that CO and NO adsorb into mixed islands. The simulated TPSR spectra were in good agreement with the experimentally observed spectra. In particular, the model accurately predicted the narrowness of the CO₂ peak and the insensitivity of the peak position to equal coverages of coadsorbed CO and NO.

Table 6.1

Activation energy barriers (kcal/mol) for selected surface reactions determined from the BOC-MP method [17]

Reaction	Surface	Activation Energy Calc.	Energy Exp.	Ref.
$\text{CO}_{2,\text{g}} \rightarrow \text{CO}_{\text{s}} + \text{O}_{\text{s}}$	Rh(111)	17	17	150
	Re(001)	-5	≤0	151
$\text{CO}_{\text{s}} + \text{O}_{\text{s}} \rightarrow \text{CO}_{2,\text{g}}$	Rh(111)	24	27	152
	Pd(111)	24	25	2
	Pt(111)	23	25	2
	Ag(110)	6.0	5.3	153
	Rh(111)	22	21	154
$\text{NO}_{\text{s}} + \text{N}_{\text{s}} \rightarrow \text{N}_2\text{O}_{\text{s}}$	Rh(100)	21	21	155
	Pt(111)	22	20	156
	Rh(111)	-63	-	-
$\text{N}_2\text{O}_{\text{s}} \rightarrow \text{N}_{2,\text{s}} + \text{O}_{\text{s}}$	Pt(111)	-46	-	-

Table 6.2
Parameters used to represent the CO-NO reaction network [168]

Process	Oh et al.	Literature Values	Refs.
CO adsorption			
$S_{CO}(1)(\theta)$	0.5	0.5	169,170
CO desorption			
$v_d(1)$ (s ⁻¹)	1.6×10^{14}	$1 \times 10^{13.6 \pm 0.3}$	171
E_d° (kcal/mol)	31.6	31.6 ± 1	171
ϕ_{CO} (kcal/mol)	4.5	-	-
ϕ_{NO} (kcal/mol)	10	10	172
CO ₂ formation			
$v_d(2)$ (s ⁻¹)	1×10^{12}	3×10^5	169
E_d (kcal/mol)	14.3	14.3	169
NO adsorption			
$S_{NO}(1)(\theta)$	0.5	-1	173
NO desorption			
$v_d(1)$ (s ⁻¹)	5×10^{13}	2×10^{12}	154
E_d (kcal/mol)	26	26	154
NO dissociation			
v_r (s ⁻¹)	6×10^{13}	6×10^{13}	174
E_r (kcal/mol)	19	19	174
δ -N ₂ formation			
v_r (s ⁻¹)	2×10^9	2×10^9	154
E_r (kcal/mol)	21	21	154
β -N ₂ formation			
$v_d(2)$ (s ⁻¹)	3×10^{10}	3×10^{10}	154
E_d° (kcal/mol)	31	31	154
ϕ_N (kcal/mol)	4	-	-

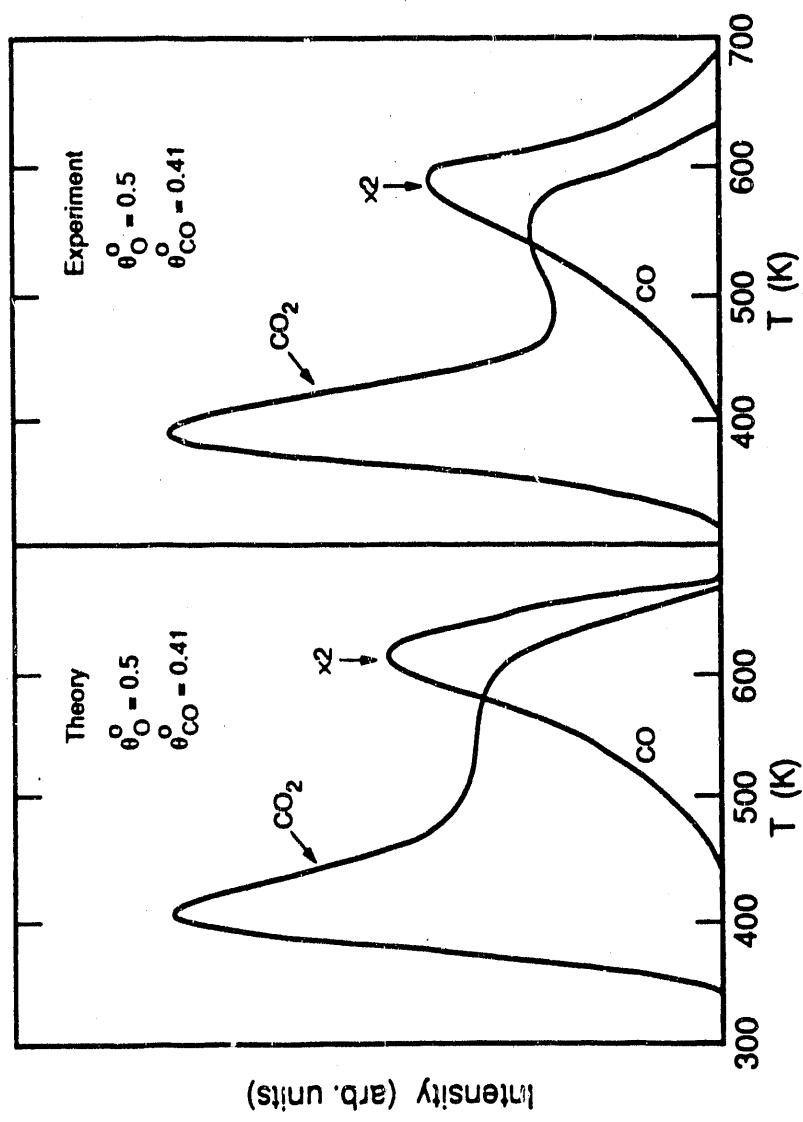


Fig. 6.1 Theoretical [126] and experimental [1160] TPSR spectra for CO oxidation and CO desorption on $\text{Ir}(110)$.

Fig. 6.1

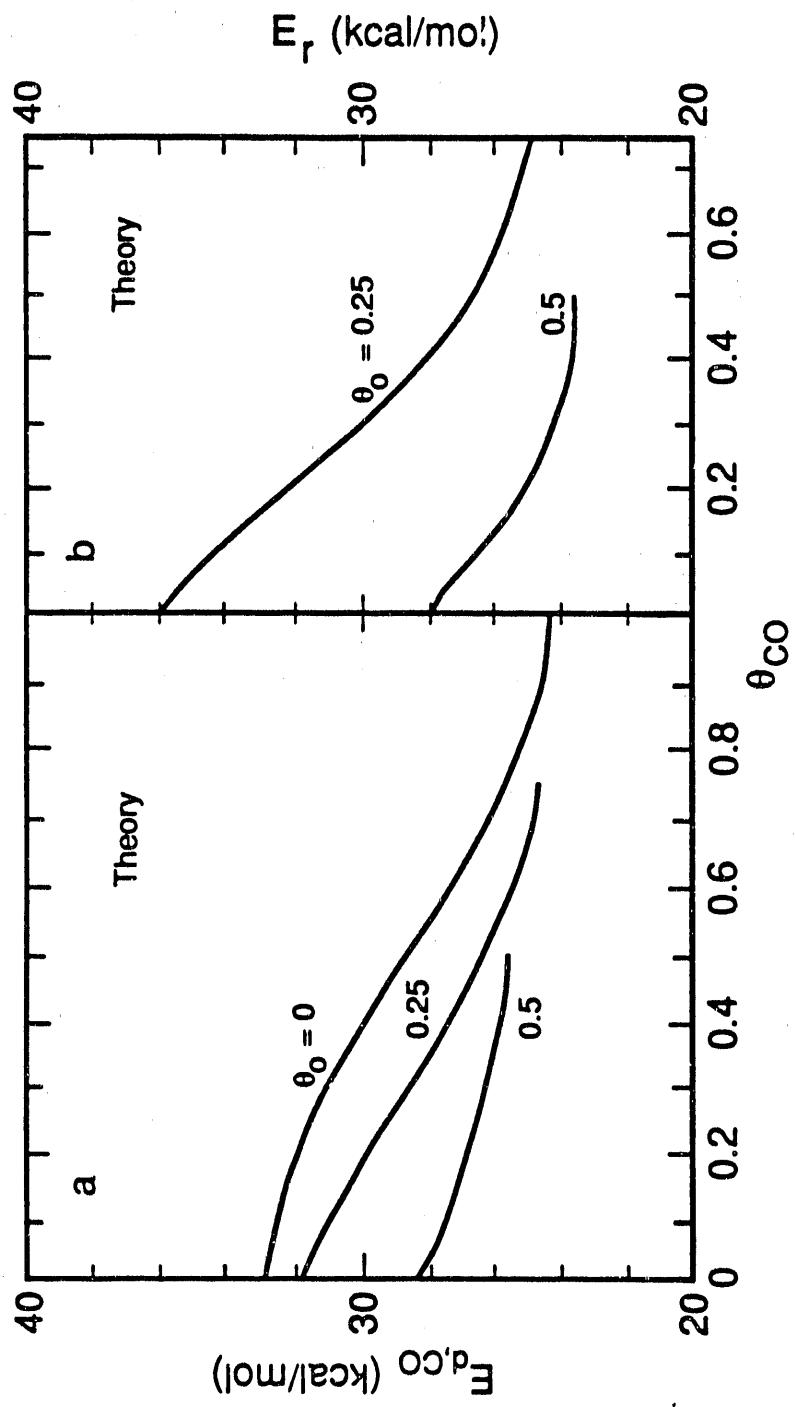


Fig. 6.2 a. Predicted variation of E_d versus θ for CO desorption from Ir(110) [126]. b. Predicted variation of E_r versus θ for CO oxidation [126].

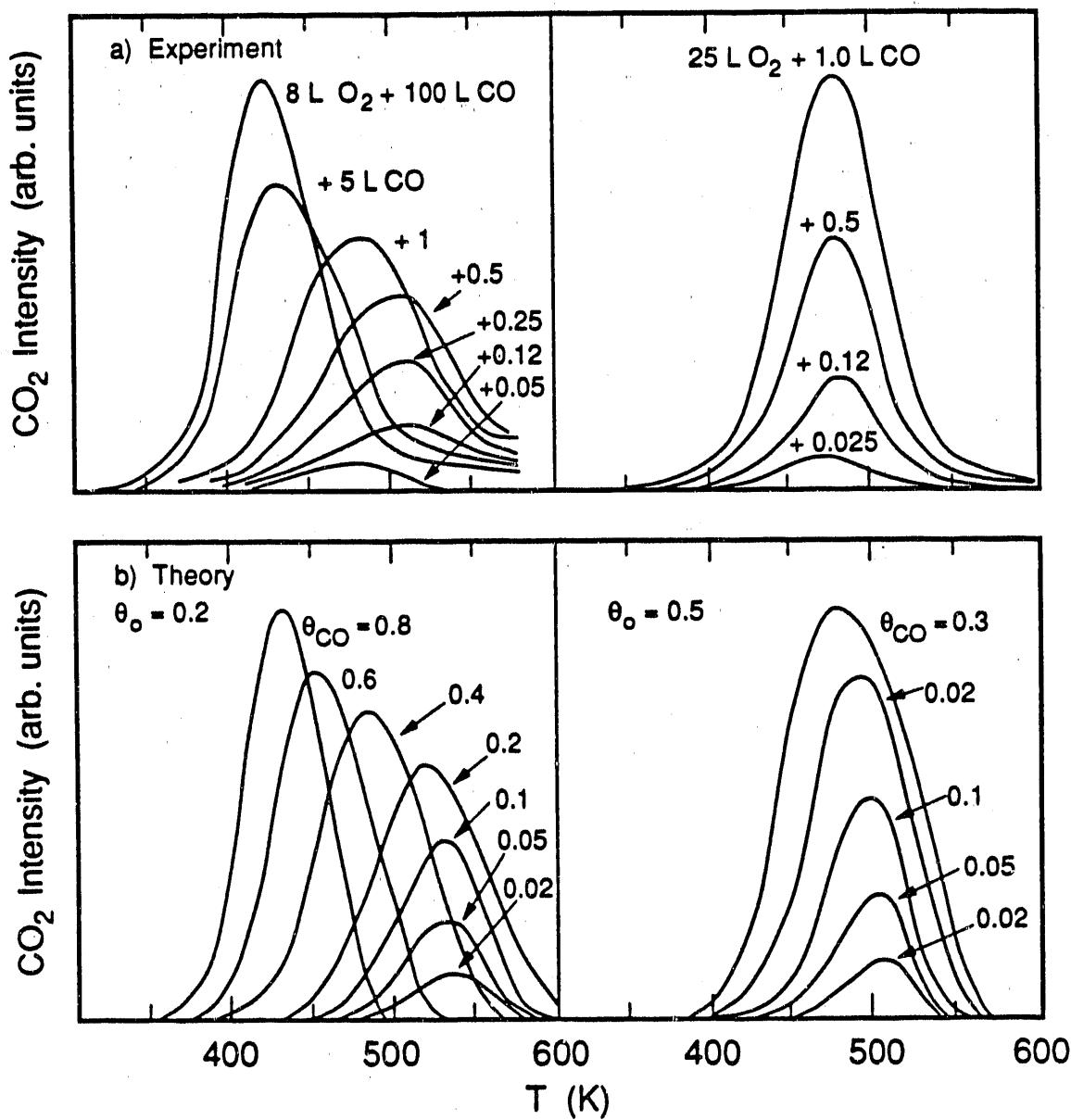


Fig. 6.3 TPSR spectra for CO oxidation on Ir(111). a. Experimental spectra as a function of gas exposure [163]. b. Theoretical spectra as a function of coverage [161].

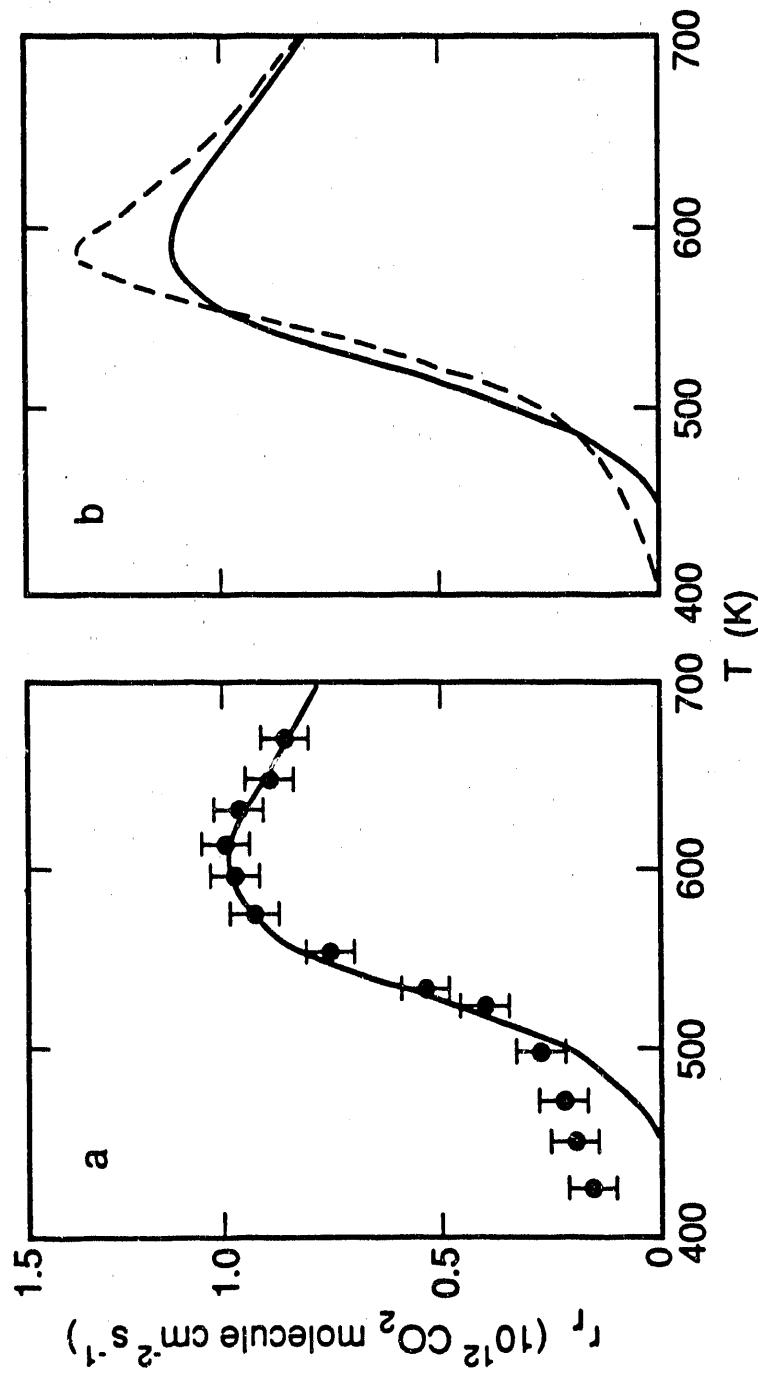


Fig. 6.4 Steady-state rate of CO oxidation versus T on Ir(111). Solid lines are theoretical calculations [162]; circles are experimental data in arbitrary units [163]; and dashed line is experimental data [164]. a. $P_{\text{CO}} = P_{\text{O}_2} = 1.1 \times 10^{-7}$ Torr. b. $P_{\text{CO}} = P_{\text{O}_2} = 5 \times 10^{-8}$ Torr and $P_{\text{O}_2} = 1.2 \times 10^{-7}$ Torr.

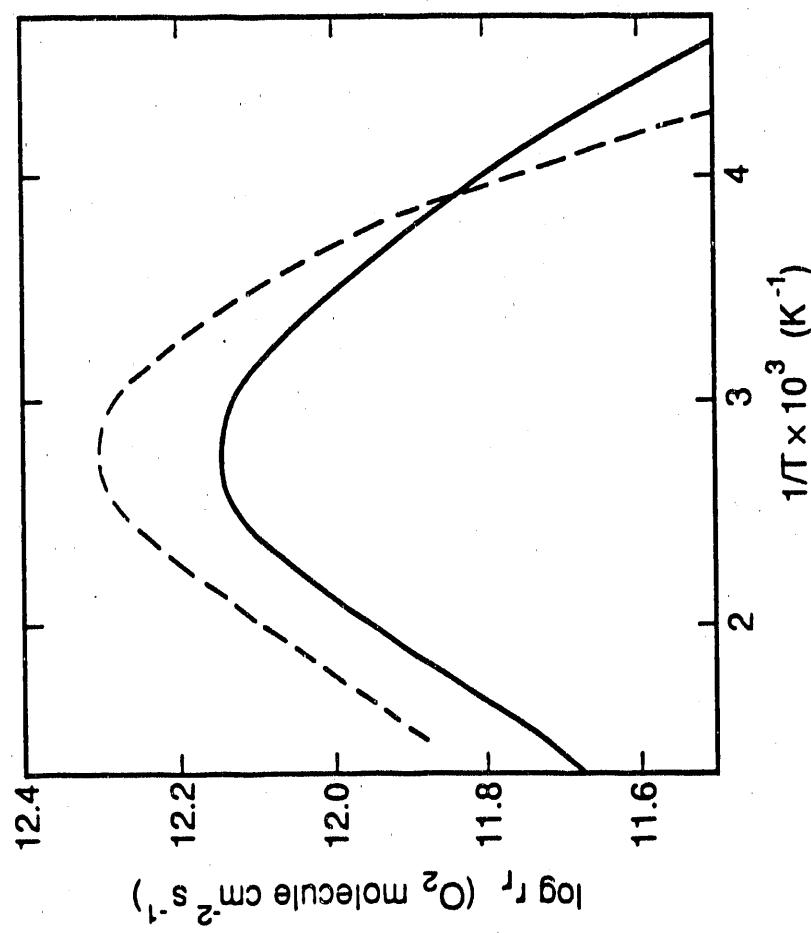


Fig. 6.5 Steady-state rate of hydrogen oxidation versus T on Pt(111) for $P_H = 1.34 \times 10^{-5}$ Pa and $P_O = 6.7 \times 10^{-6}$ Pa. Solid line is the theoretical calculation [165] and the dashed line is the experimental result [166].

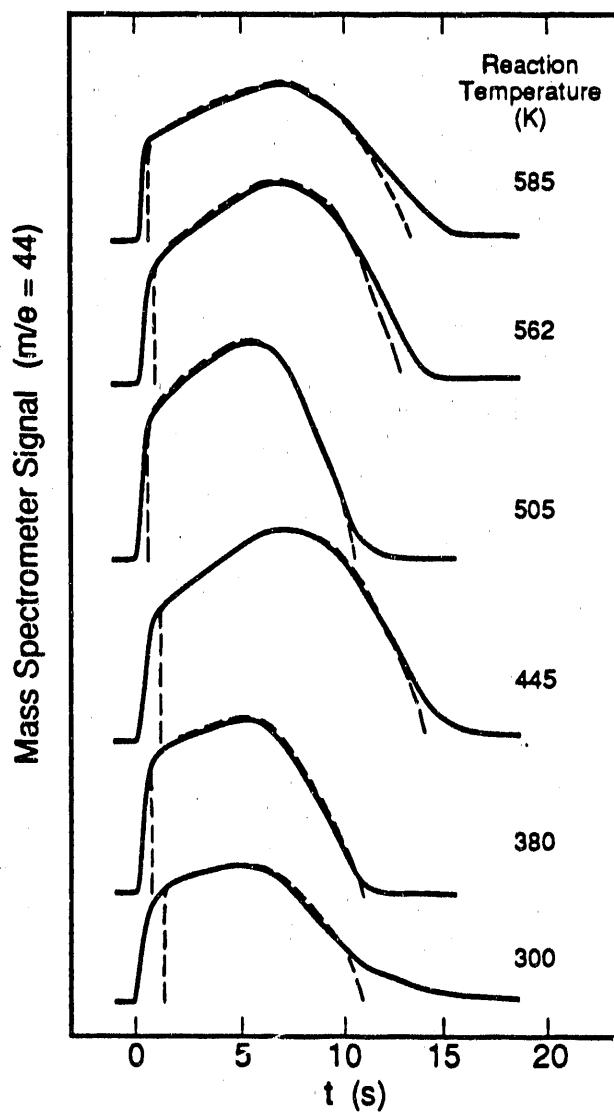


Fig. 6.6 Rate of CO_2 production versus t for the reaction of CO with preadsorbed oxygen for $P_{\text{CO}}=1.0 \times 10^{-8}$ Torr and $\theta_{\text{O}}=2/3$ monolayer [175]. Solid lines are the experimental results and the dashed lines are the model predictions.

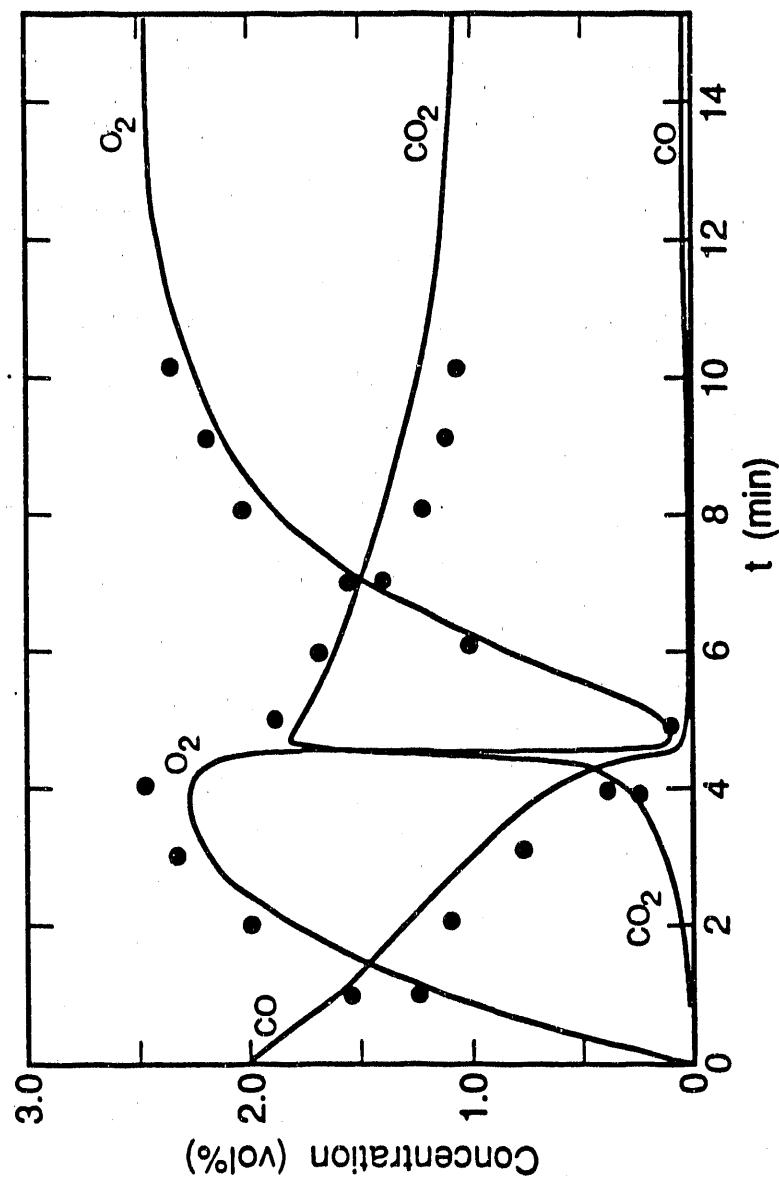


Fig. 6.7 Concentration profiles versus t for CO oxidation for an inlet step change from 2% CO to a 1% CO and 3% O₂ mixture [183]. Circles are the experimental data and the solid line is the model prediction.

7.0 CONCLUDING REMARKS

Considerable progress has been made in developing theoretical methods for predicting the dynamics of elementary processes occurring on metal surfaces. The starting point in all cases is a description of the potential governing the interactions between gas molecules and the atoms at the surface of the metal. At high adsorbate coverages, additional information must be supplied to describe the effects of adsorbate-adsorbate interactions. Accurate potential functions based on a quantum treatment of gas-metal interactions are not yet available and consequently, resort must be made to empirical potential functions. While such functions can be constructed in an ad hoc fashion, there do not appear to be generally accepted criteria for establishing the suitability of such functions for dynamic calculations. At a minimum, it would seem that the potential function chosen should give the experimentally observed heat of adsorption and the vibrational frequency for adsorbate-metal vibrations in the ground state.

As noted in Section 2, both classical and quantum descriptions of the dynamics of gas-metal interactions have been developed. At this time, only the classical approaches offer a way of predicting rate and diffusion coefficients for a wide range of processes and adsorbed species. Direct simulation of molecular dynamics is practical provided that the process of interest is $\sim 10^3$ times slower than the time constant for the fastest mode of motion (e.g., vibration). Experience has shown that molecular dynamics simulation provides estimates of the diffusion coefficient and of

the sticking coefficient for dissociative adsorption of diatomic molecules that are in good quantitative agreement with experiment. By use of either MD or stochastic (Langevin) dynamics, it is also possible to include the effects of lattice vibrations. Where this has been done, it has been found that lattice vibrations do affect the calculated values of the sticking and diffusion coefficients.

The constraints of molecular dynamics with respect to describing the dynamics of infrequent events for which the time constants are $> 10^{-9}$ s can be overcome through the use of dynamically corrected TST. This approach provides a rigorously correct theoretical framework for calculating rate and diffusion coefficients. Dynamically corrected TST provides a rational basis for identifying conditions under which precursor states affect the adsorption and desorption of adsorbates. Calculations of diffusion coefficients and rate coefficients for desorption obtained by means of dynamically corrected TST show good quantitative agreement with experimental measurements. Moreover, this theoretical approach explains why desorption rate coefficients exhibit a deviation from Arrhenius behavior at high temperatures, and why the rate coefficients for the desorption of molecular species are a factor of 100-1000 larger than those for the desorption of atomic species.

Absolute rate theory is useful only for crude estimation of preexponential factors, in as much as ad hoc assumptions regarding the transition state structure are necessary in order to estimate partition functions. This represents a severe limitation which limits the accuracy of the estimates of the preexponential factor to

within a factor of 10 to 10³.

Several theoretical models based on quantum mechanics have been developed. Such models can include the exchange of energy between the adsorbate and the surface as well as the effects of barrier tunneling and reflection. For small molecules and atoms (e.g., H₂ and H), quantum effects such as barrier tunneling and reflection are observed. For more massive adsorbates, quantum effects are projected to be less important.

The effects of lateral interactions can be neglected at low adsorbate coverages but can become significant at higher coverages. Such interactions can alter both the spatial distribution of adsorbates on a metal surface and the apparent activation energy barriers. To date, only empirical or semiempirical representations of lateral interactions have been developed which are capable of describing rate or diffusion coefficients for a wide range of surface processes and adsorbate coverages. The simplest and the most frequently used approach is to treat lateral interactions by a sum of pairwise additive contributions to the activation energy. With this approach, both the magnitude and sign of the interaction are treated as adjustable parameters. An alternative technique for describing lateral interactions is the BOC-MP method. In this case, interactions between adsorbates arise as a consequence of through-metal and direct adsorbate-adsorbate interactions.

Analytical expressions for the kinetics of adsorption, diffusion, desorption, and reaction in the presence of lateral interactions can be written using a lattice-gas model with the quasi-chemical approximation. While this approach captures the effects of

lateral interactions, the lattice-gas model requires two rather restrictive assumptions. The first is that lateral interactions can be represented by a sum of pairwise contributions of equivalent strength and the second is that adsorbates always maintain an equilibrium configuration on the metal surface.

An alternative approach to account for the effects of lateral interactions is to incorporate the energetics predicted either from the sum of pairwise additive contributions or from the BOC-MP method into Monte Carlo simulations. The Monte Carlo models treat the kinetics of surface processes in terms of probabilities which are specific to each site and its local environment, and thus require no assumptions about the distribution of adsorbates on the surface.

The present review has shown that both continuum and Monte Carlo models can be used to account for the experimentally observed kinetics of surface processes. In addition, the temperature-programmed desorption and reaction of adsorbates on metal surfaces can also be described with these techniques. Proper representation of lateral interactions has led to the conclusion that the appearance of multiple peaks in TPD and TPSR spectra are a consequence of such interactions.

Both continuum and Monte Carlo models have been used to describe the effects of island formation on the dynamics of surface reactions. Continuum models require assumptions to be made about the number, size, and shape of the islands. Such models have been successful, though, in representing the effects of adsorbate islands on reaction dynamics. By contrast, Monte Carlo models make no assumptions about the concentration of islands or their size and

shape. To date, though, kinetics predicted with Monte Carlo models have not been compared extensively with experimental observation.

Theoretical methods for predicting rate coefficients for other than very simple surface reactions have yet to be developed. The principal difficulty is the absence of accurate methods for generating the relevant potential energy hypersurface. Should it become possible to generate such functions, then the desired rate coefficients could be calculated using dynamically corrected TST. For reactions involving the removal or addition of a hydrogen atom, quantum effects might be expected to be important.

It is evident from this review that theoretical methods are now available for explaining many of the phenomena observed when gases interact with metal surfaces. One area for future research is the development of realistic potential hypersurfaces for describing surface reactions. Such calculations should preferably be based on ab initio quantum chemical methods, to the extent possible. Another area that should be considered is the influence of adsorbate coverage on the reconstruction of metal surfaces and the effects of surface reconstruction on the dynamics of elementary processes occurring on metal surfaces. Finally, the suitability of stochastic models for describing reaction kinetics over a wide range of conditions should be explored more fully.

NOMENCLATURE

A,A ₂ ,B,AB	Adsorbates
A ⁺ ,A ₂ ⁺	Precursor species
A _s	Chemisorbed species
A-A	Adsorbate-adsorbate interaction
a	Morse potential parameter
a _s	Area per reaction site
C	Concentration
D	Diffusivity
D(θ)	Coverage-dependent diffusivity
D _{AB}	Dissociation energy of AB
D ₀	Preexponential factor for diffusion
d	Dimensionality
E	Activation energy
E _a	Activation energy for adsorption
E _d	Activation energy for desorption
E _{diff}	Activation energy for diffusion
E _{diss}	Activation energy for dissociation
E _r	Activation energy for reaction
E _s	Lateral interaction energy
E _{a°}	Activation energy for adsorption in the absence of lateral interactions
E _{d°}	Activation energy for desorption in the absence of lateral interactions
E _{r°}	Activation energy for reaction in the absence of lateral interactions
E _t	Energy of the transition state
E ₀	Energy of the reactant
E _T	Incident kinetic translational energy
F	Flux of adsorbate
f _r	Correction factor for reaction due to lateral interactions
f _s	Correction term in TST
f _{a',fa"}	Correction factors for adsorption due to lateral interactions
f _{d',fd"}	Correction factors for desorption due to lateral interactions
G	Normalization constant
H	Classical Hamiltonian

H	Quantum Hamiltonian operator
h	Planck's constant
J	Diffusive flux
K	k_a^*/k_d^*
k	rate coefficient
k_b	Boltzmann constant
k_a	Rate coefficient for adsorption
$k_d(\alpha)$	Rate coefficient for desorption
k_{TST}	Rate coefficient from TST
k_a^*	Rate coefficient for adsorption from the precursor state
k_d^*	Rate coefficient for desorption from the precursor state
L	$r_d^*/(r_a^{**} + r_d^{**})$
LEPS	London-Eyring-Polanyi-Sato potential energy surface
l	Effective length of an adsorbate
M-A	Metal-adsorbate interaction
m	Molecular weight or mass of an adsorbate
m_B	Molecular weight of species B
m_i	Number of adsorbates bonded to i th metal atom
m'	Mass of a surface atom
N	Normalization constant
N_d	Number of species which desorb
N_s	Number of surface metal atoms
n	Number of nearest neighbors
n'	Number of metal atoms bonded to an adsorbate
P	Probability
$P(p,q)$	Probability density in phase space
P_B	Pressure of species B
$P_n(t)$	Probability of finding an adsorbate in state n at time t
P _{AA}	Probability that two nearest-neighbor sites are occupied by A
P _{AV}	Probability that of two nearest-neighbor sites are occupied, one is occupied by A and the other is vacant
P _V	Probability that two nearest-neighbor sites are vacant
$P_{A,n}$	Probability that a site occupied by A has n nearest neighbors
$P_{AB;i}$	Probability that a pair of sites occupied by A and B has the environment i
P_i	Probability of desorption from the i th site
P_{ij}	Probability of diffusion from site i to site j

p	Position
Q, Q_i	Heat of adsorption
Q°	Heat of adsorption in the absence of lateral interactions
Q_{oA}	Heat of adsorption of A in the on-top position
$Q_{A,n}$	Heat of adsorption of A as a function of coordination
$Q_{A,n}(1)$	Heat of adsorption of A due to M-A interactions
$Q_{A,n}(2)$	Heat of adsorption of A due to A-A interactions
$Q_{A,n}^*$	Total heat of adsorption of A
q	Momentum
q_o	Partition function for adsorbates in the transition state
q_t	Partition function for adsorbates in the transition state
R	Random number
R	Random force
r, r_i	Distance
r_a	Rate of adsorption
r_d	Rate of desorption
r_o	Equilibrium bond distance
r_r	Rate of reaction
S	Dividing plane
$S^{(\alpha)}(\theta)$	Sticking coefficient
$S^{(\alpha)}(0)$	Sticking coefficient at zero coverage
$S_0^{(\alpha)}$	Preexponential factor for adsorption
S_0	Reactive sticking coefficient
s	Reaction coordinate
s_o	Location of the dividing plane
T	Temperature
T_c	Critical temperature
T_s	Temperature of the surface
TST	Transition state theory
t	time
u_o	Velocity required for an adsorbate located at s_o to desorb
V	Potential
V_{oA}	Equilibrium energy
V_p	Interaction potential in the primary zone
V_s	Interaction potential in the secondary zone
v	Mean velocity
v_s	Velocity in the s direction
W	Potential of mean force

W	Matrix of transition probabilities
W_{AA}	Interaction energy between an AA pair
X, X_i	Bond order
X_p	Position in the primary zone
Y	Adsorbate trajectory
z	Number of nearest-neighbor sites or distance
α	Order of a rate process
β	Heating rate
δ_{ij}	A-A bond order
ϵ	Lennard-Jones parameter
ϕ	Coefficient for describing how E varies with total coverage
Γ, Γ_{ij}	Hopping frequency
γ	Bending angle
γ_0	Maximum bending angle
η	$2[1-\exp(-W_{AA})/k_b T]$
λ	Mean-free path or mean-free hopping length
μ	Chemical potential
μ_i	Reduced mass of species i
μ_r	Reduced mass for rotation
$v_d(\alpha)$	Preexponential factor for desorption
v_{diff}	Preexponential factor for hopping
v_r	Preexponential factor for reaction
Θ	Memory kernel
θ	Coverage
θ_A°	Initial coverage of A
θ_v	Coverage of vacant sites
θ_i	Angle of incidence from the surface normal
χ_{ij}	Exchange integral
ρ_{ij}	Coulomb integral
σ	Lennard-Jones parameter
τ	Time
Ω_0	Vibration frequency of an adsorbate at the bottom of the potential well
Ω_r	Rotational frequency
ω_{pp}	Vibration frequency for the metal atoms in the primary zone

ξ Correction factor in TST
 ψ Wave function
 ζ Trapping probability

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Chapter III

A Monte Carlo Model for the Simulation of Temperature-Programmed Desorption Spectra**ABSTRACT**

A Monte Carlo model has been developed for describing the temperature-programmed desorption of adsorbates from single crystal surfaces. This model takes into account the effects of surface diffusion and the influence of metal-adsorbate (M-A) and adsorbate-adsorbate (A-A) interactions on the coverage dependence of the activation energy for desorption. The inclusion of M-A and A-A interactions has a pronounced effect on the shape of the predicted TPD spectrum. Where only a single peak is observed in the absence of M-A and A-A interactions, multiple peaks are found when these interactions are included. The inclusion of M-A and A-A interactions is also shown to produce a nonlinear decline in the activation energy for desorption as a function of increasing adsorbate coverage. Simulated TPD spectra for CO desorption from a Pd(100) surface and for H₂ desorption from Mo(100) and Ni(111) surfaces, obtained using the Monte Carlo model, are in satisfactory agreement with those observed experimentally. An important feature of the model is that it describes correctly the observed dependence of the activation energy on adsorbate coverage.

1.0 INTRODUCTION

Temperature-programmed desorption (TPD) spectroscopy is widely used to characterize the interactions of adsorbates with metal surfaces [1-4]. The analysis and interpretation of TPD spectra have usually been carried out using rate expressions which assume the surface and adlayer to be homogeneous for a given adsorbate coverage. In such models, the rate of desorption is given by the product of a rate coefficient and an integer power dependence on the surface coverage, θ . The rate coefficient is written as the product of a preexponential factor, k_d° , and an exponential function, the argument of which contains the activation energy for desorption, E_d . While the earliest modeling efforts assumed k_d° and E_d to be independent of θ , more recent studies have demonstrated that both parameters can be coverage dependent.

Analyses of experimental TPD spectra indicate that E_d generally decreases with increasing adsorbate coverage. Most attempts to account for this trend have been based on a lattice-gas model with pairwise-additive energetics [5-10]. In this model, adsorbed species are considered to be localized on a two-dimensional surface and interactions between nearest-neighbor adsorbates are assigned either an attractive or repulsive contribution to the total binding energy of the adsorbate. Simulations of TPD spectra based on pairwise-additive energetics with either the quasichemical approximation or with Monte Carlo methods have been successful in accounting qualitatively (and in a few instances, quantitatively) for deviations from the assumption of non-interacting adsorbates [6-10]. It should be noted, though, that

the use of pairwise-additive energetics is not physically satisfying, since no a priori basis exists for assigning the relative magnitude and sign of an interaction or for assuming linearly additive contributions from nearest-neighbor adsorbates. Furthermore, the assumption of a homogeneous surface in the quasichemical approximation is not warranted. For a given global coverage, different local configurations of adsorbates may exist on the surface and hence the value of E_d for each environment would not be equivalent.

Different dependencies of k_d° on θ have been reported by various authors, some indicating that k_d° increases with θ , and others, that k_d° decreases with θ [11-17]. Attempts to explain such trends have been made using absolute rate theory. While this approach does permit for a rationalization of why k_d° should be coverage dependent, it has not proven to be adequate for making accurate predictions of k_d° .

The objective of this study is to examine the influence of adsorbate-metal interactions, adsorbate-adsorbate interactions, and surface mobility on the rates of adsorbate desorption from well-defined surfaces. A lattice-gas model approach is used. Rather than rely on average or mean-field coverage effects, however, the influence of local environment is taken into account explicitly. The local binding energy for an adsorbate is determined using the Bond-Order-Conservation (BOC) method [18]. This method allows one to account for the coordination of the adsorbate with the surface and the effects of nearest-neighbor adsorbates. In the latter case, both through-metal and adsorbate-adsorbate interactions are included.

BOC calculations of the local heat of adsorption and activation energy for desorption are then incorporated into a Monte Carlo simulation of temperature-programmed desorption. The method is used to examine the desorption of CO from a Pd(100) surface and H₂ desorption from Mo(100) and Ni(111) surfaces. An interesting feature of the model is that it describes not only the shape of the TPD spectrum, but also the dependence of the desorption activation energy and the distribution of local surface site occupancy on the average surface coverage.

2.0 THEORY

We assume that at any moment in time adsorbate atoms or molecules occupy well-defined sites on a single-crystal metal surface, the coordination of the adsorbate being specific to the metal and the adsorbate. The number of metal atoms coordinated to a single adsorbate is 1 for an on-top site, 2 for a bridge site, 3 for a three-fold hollow, and 4 for a four-fold hollow. The coverage is defined as the number of adsorbate atoms or molecules, N, divided by the number of atoms at the surface of the metal, N_s:

$$\theta = N/N_s \quad (1)$$

Adsorbed atoms or molecules are assumed to participate in only two rate processes: surface diffusion and desorption. Since the local coverage varies with position on the surface, it is anticipated that the dynamics of surface diffusion and desorption should be site specific. This specificity will be accounted for through the local heat of adsorption and activation energy for desorption, as discussed below.

The probability of desorption from the local environment i in the time interval Δt can be defined as

$$P_i^{(1)} = k_{d,i}^0 \exp[-E_{d,i} / RT] \Delta t \quad (2)$$

where $k_{d,i}^0$ is the frequency factor for desorption and $E_{d,i}$ is the activation energy for desorption. The value of Δt in eq. 2 must be sufficiently small so that $P_i^{(1)}$ goes to unity at a temperature sufficiently high to guarantee virtually complete desorption from all sites of type i . This point is discussed more fully in the next section.

The rate of desorption, expressed as a turnover frequency based on N_s , is

$$r_d = N_d / (\alpha N_s \Delta t) \quad (3)$$

where N_d is the number of adsorbate atoms or molecules desorbing in the time interval Δt . The parameter α is 1 for molecular desorption and 2 for associative desorption of atoms. Consequently, r_d is the rate of desorption observed from the gas phase. The total rate of desorption from all local environments is simply

$$\begin{aligned} r_d &= \sum_i N_{d,i} / (\alpha N_s \Delta t) \\ &= N_d / (\alpha N_s \Delta t) \end{aligned} \quad (4)$$

Since the activation energy for surface diffusion is 10-15% of that for desorption, whereas the frequency factors for the two processes are comparable, surface diffusion is expected to be a much more rapid process than desorption. As a consequence, we assume that adsorbate atoms or molecules will reposition themselves nearly instantaneously to achieve an equilibrium state. The probability that an adsorbate in site i moves to an adjacent

vacant site j is given by [19]:

$$P_{ij}^{(2)} = \frac{\exp [-(Q_i - Q_j) / RT]}{1 + \exp [-(Q_i - Q_j) / RT]} \quad (5)$$

where Q_i and Q_j are the heats of adsorption for an adsorbate at sites i and j , respectively. Equation 5 weights diffusion jumps according to the magnitude of $(Q_i - Q_j)$. It should be noted that if $Q_i = Q_j$, $P_{ij}^{(2)} = 0.5$, which indicates that two sites of equal energy have equal probability of occupancy. For a sufficiently large number of jumps, application of eq. 5 yields an equilibrium distribution of adsorbates on the metal surface.

Having developed expressions for the probabilities of adsorbate desorption and surface diffusion, we proceed next to indicate how $E_{d,i}$ and Q_i , appearing in these expressions, are calculated. Our approach is to use the Bond-Order-Conservation (BOC) method developed by Shustorovich, since it allows us to account for the effects of local site occupancy without the introduction of locally assigned energy parameters [18].

A principal assumption of the BOC method is that the two-center interaction between an adsorbate atom A and a surface metal atom M can be described by a Morse potential

$$Q(x) = Q_0(2x - x^2) \quad (6)$$

where x is the bond order and Q_0 is the equilibrium value of the M-A bond energy for $x=1$. The bond order in eq. 6 is defined as

$$x = \exp[-(r-r_0)/a] \quad (7)$$

where r is the M-A bond length, r_0 is the equilibrium bond length, and a is a scaling parameter [18]. When A interacts with more than one metal atom, the total heat of adsorption is given by the sum of all

two-center interactions. A second assumption central to the BOC method is that along a reaction path describing the interactions of a molecular or atomic species with a metal surface, the total bond order is conserved and normalized to unity.

Using the above assumptions, Shustorovich [18] has shown that the heat of chemisorption of an atom A interacting with n metal atoms is given by

$$Q_{A,n} = Q_{oA} (2 - 1/n) \quad (8)$$

The heat of adsorption for a molecular adsorbate AB coordinated via atom A to n metal atoms can be approximated by

$$Q_{AB,n} = \frac{Q_{oA}^2}{D_{AB} + Q_{oA} / n} \quad (9)$$

where D_{AB} is the gas-phase dissociation energy for the A-B bond.

Equations 8 and 9 are appropriate for isolated adsorbate atoms or molecules. When the adsorbate coverage increases, however, situations arise in which more than one adsorbate is coordinated with a single metal atom. In this case, the heat of adsorption is given by

$$\dot{Q}_{A,n} = \sum_{i=1}^n \frac{Q_{A,n}}{nm_i} (2 - 1/m_i) \quad (10)$$

where m_i is the number of adsorbates bonded to the i th metal atom.

At high coverages, direct A-A interactions can become significant and must be taken into account in calculating $Q_{A,n}$. To do so, the total bond order associated with A is still normalized to unity but partitioned between the A-A and M-A interactions. The resulting expression for $Q_{A,n}$ can be written as

$$Q_{A,n} = Q_{A,n}^{(1)} + Q_{A,n}^{(2)} \quad (11)$$

where $Q_{A,n}^{(1)}$ is the heat of adsorption due to M-A interactions and $Q_{A,n}^{(2)}$ is the heat of adsorption due to A-A interactions.

The value of $Q_{A,n}^{(2)}$ can be expressed as

$$Q_{A,n}^{(2)} = \sum_{i=1}^n \sum_{l=1}^L 0.5 D_{AA} (2\delta_{il} - \delta_{il}^2) \quad (12)$$

where D_{AA} is the bond dissociation energy for the A-A bond and δ_{il} is the bond order for the A-A interaction between the A atom coordinated with metal atom i and the l th nearest-neighbor atom which is also coordinated with metal atom i . The summation over l in eq. 12 is to account for all nearest-neighbor A atoms. The occurrence of A-A interactions weakens the bond order associated with M-A interactions and as a consequence, the bond order for each component of an M_n -A bond is given by

$$x_{i,n} = \frac{1}{n} - \sum_{l=1}^L \delta_{il} \quad (13)$$

Equation 13 is then used to calculate $Q_{A,n}^{(1)}$, which is given by

$$Q_{A,n}^{(1)} = \sum_{i=1}^n \left[\frac{Q_{oA}}{m_i} \left(2 - \frac{1}{m_i} \right) (2x_{i,n} - x_{i,n}^2) \right] \quad (14)$$

The activation energy for desorption depends on the mode of desorption. For the desorption of a single atom or an adsorbed molecule, the activation energy for desorption is identical to the heat of adsorption. For homonuclear associative desorption, the activation energy, $E_{d,AA}$ is given by [18]

$$E_{d,AA} = \left(\dot{Q}_{A,n} \dot{Q}_{A',n} \right) / \left(\dot{Q}_{A,n} + \dot{Q}_{A',n} \right) \quad (15)$$

where $Q_{A,n}^*$ and $Q_{A',n}^*$ are calculated using eq. 10 or 11. The two recombining atoms are designated as A and A' to denote that the local environments of the two atoms may be different.

The magnitude of the preexponential factor, $k_{d,i}^*$, can be estimated from absolute rate theory [11-17]. The accuracy of these estimates is, in general not high, since numerous assumptions must be made to evaluate the partition functions for the adsorbate. As a consequence, $k_{d,i}^*$ was treated as an adjustable constant in the present work. The values of $k_{d,i}^*$ used for individual simulations are justified in the Results and Discussion Section.

The temperature-programmed desorption of adsorbates was simulated using a Monte Carlo algorithm. The metal surface was represented by a 100 by 100 array of numbered sites, and periodic boundary conditions were used to eliminate edge effects. The number of metal atoms associated with an adsorbate, n , was chosen on the basis of information taken from the literature for a given metal-adsorbate system. Experimental observations reported in the literature were also used to determine the initial structure of the adsorbate overlayer.

The simulation of a TPD spectrum was carried out in the following fashion. Adsorbate atoms or molecules were placed on the metal surface lattice in either an ordered or random fashion to achieve a desired initial coverage, θ_0 . The temperature was initialized at T_0 and taken to be constant at this value for the time interval Δt . A surface site (or pair of adjacent surface sites for associative desorption) was then selected in a random fashion. If the site (or both sites for associative desorption) was occupied, the

probability of desorption was calculated using eq. 2. The local activation energy for desorption was calculated using either eq. 10 or eq. 15, depending on whether the desorption process considered was first or second order, respectively. The calculated desorption probability was then compared with a random number, R , such that $0 < R < 1$. If $R < P_i(1)$, the adsorbate (or pair of adsorbates for associative desorption) was removed from the lattice, and N_d in eq. 4 was incremented by one (two for associative desorption). If $R \geq P_i(1)$, the site (or pair of sites) remained occupied. The preceding steps were repeated N_{s0} times for the case of first order desorption and $0.5 N_{s0}^2$ times for the case of second order desorption. After completing the sampling of the surface, the rate of desorption for this time interval was calculated using eq. 4.

A redistribution of the remaining adsorbates was carried out next to account for the effects of surface diffusion. A surface site and an adjacent site were chosen at random. If the surface site was occupied and the nearest-neighbor site was vacant, a probability of diffusion was calculated using eq. 5. Q_i was taken as the heat of adsorption in the initial site and Q_j was taken as the heat of adsorption in the final site. The calculated value of $P_{ij}(2)$ was then compared with a random number, R , ($0 < R < 1$). If $R < P_{ij}(2)$, the adsorbate was moved from the initial site, i , to the adjacent site, j . Conversely, if $R \geq P_{ij}(2)$, movement of the adsorbate was not allowed. A sufficient number of surface visitations was allowed to attain an equilibrium distribution of adsorbates on the surface.

The desorption and diffusion calculations described above constitute a Monte Carlo Step (MCS). As noted above, the time

interval associated with an MCS is Δt , and over this interval the temperature is constant. At the end of an MCS the temperature was increased by the increment $\beta\Delta t$, where β is the desired heating rate. A new MCS was then carried out at the next temperature. This process was repeated until a temperature was reached for which the surface was depleted of adsorbate. A plot of the desorption rate as a function of temperature then yielded a TPD spectrum. In addition, the average activation energy of desorption could also be determined as a function of the total surface coverage. For both the TPD spectra and the average activation energy profiles, best-fit curves were drawn by eye through the data points. All of the calculations described above were carried out on an IBM 3090 computer. Typical run times were 2-10 CPU minutes. Random numbers were generated by the IMSL linear congruential number generator GGUBFS [20].

3.0 RESULTS AND DISCUSSION

Comparison of Monte Carlo and Continuum Models

Since the Monte Carlo model for TPD presented in the preceding section is new, it is useful to compare the results obtained with this model with those obtained from classical continuum representations. We begin by considering first order desorption for the case where both the activation energy and preexponential factor are constant. The values of these parameters are listed in table 1. To ensure that the probability of desorption given by eq. 2 is normalized to unity at a temperature sufficiently high to guarantee complete desorption, the time interval Δt was determined according to:

Table 1
Model parameters^a

Lattice	Case	$Q_{A,n}$ (kcal/mol)	D_o (kcal/mol)	k_d^* (s ⁻¹) ^c	β (K/s)	θ_o	n	Δt	Fig.
<i>1st order desorption</i>									
fcc(100)	No interactions	30.0	-	1x10 ¹⁵	1.0	1.0	1	0.37	1
fcc(100)	CO on Pd(100)	36.8	-	2x10 ¹⁶	14.0	0.1-0.5	2	0.10	6
<i>2nd order desorption</i>									
fcc(100)	No interactions	60.0	-	1x10 ¹⁵	1.0	1.0	4	3.0	2
fcc(100)	M-A interactions	68.6 ^b	-	1x10 ¹⁵	1.0	1.0	4	3.0	3
fcc(100)	M-A and A-A interactions	60.5 ^b	225.0	1x10 ¹⁵	1.0	1.0	4	3.0	3
bcc(100)	H ₂ on Mo(100)	63.0	109.5	1x10 ¹⁶	74.0	0.1-2.0	2	0.027	10
fcc(111)	H ₂ on Ni(111)	58.3	109.5	1x10 ¹⁶	2.0	0.25-0.8	3	2.0	13

a) Each simulation was 1 run except for H₂ on Ni(111) which was the average of 8 runs.
 b) $Q_{A,n}$ was adjusted so that E_d ($\theta \rightarrow 0$) is the same as in the case of no M-A or A-A interactions.
 c) For second order desorption, k_d^* represents the product of the second order preexponential factor times the concentration of adsorption sites.

$$\Delta t = (k_{d,i}^0 \exp[-E_{d,i} / RT_{\max}])^{-1} \quad (16)$$

where T_{\max} is selected to be a temperature at which the surface is depleted of adsorbate. For the case considered, T_{\max} was chosen to be 450 K, and consequently Δt is 0.37 s. Similar reasoning for determining Δt was used in all the simulations discussed in this paper.

Figure 1 shows that the shape and location of the TPD peak predicted by the Monte Carlo model are in good agreement with those obtained from the continuum model. The scatter in the Monte Carlo calculations is attributable to the combined effects of finite lattice size, finite step size, and statistical random number fluctuations. As demonstrated in the Appendix, the good agreement between the numerical solution and the Monte Carlo simulation is due to the equivalence of the probabilistic and continuum descriptions of desorption.

A comparison between the continuum and Monte Carlo models was also made for second order desorption kinetics with a constant activation energy. For the Monte Carlo simulation, the atomic adsorbates were allowed to diffuse on a fcc(100) lattice with probabilities of diffusion calculated using eq. 5. Only nearest-neighbor species were considered as desorption partners. The parameters used are listed in table 1 and the results are shown in fig. 2. While the Monte Carlo simulation shows some scatter, it agrees well with the solution obtained from the continuum model. When a similar simulation was conducted in the absence of diffusion, a residual coverage of 9% of a monolayer remained on the surface at T_{\max} in the form of isolated adsorbates with no nearest

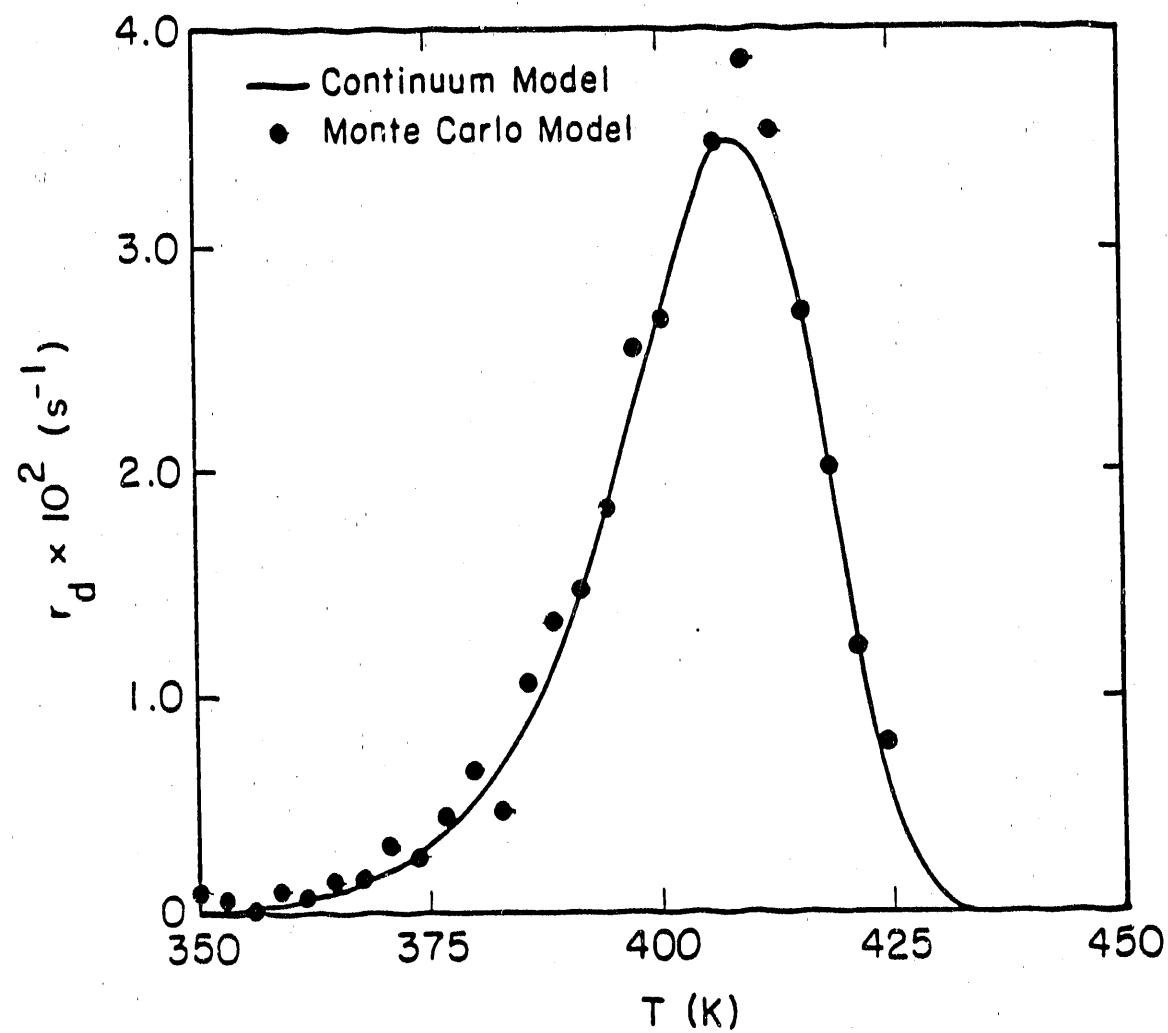


Fig. 1 Comparison of Monte Carlo and continuum models for the simulation of first order desorption where E_d is constant. (See Table 1 for parameter values).

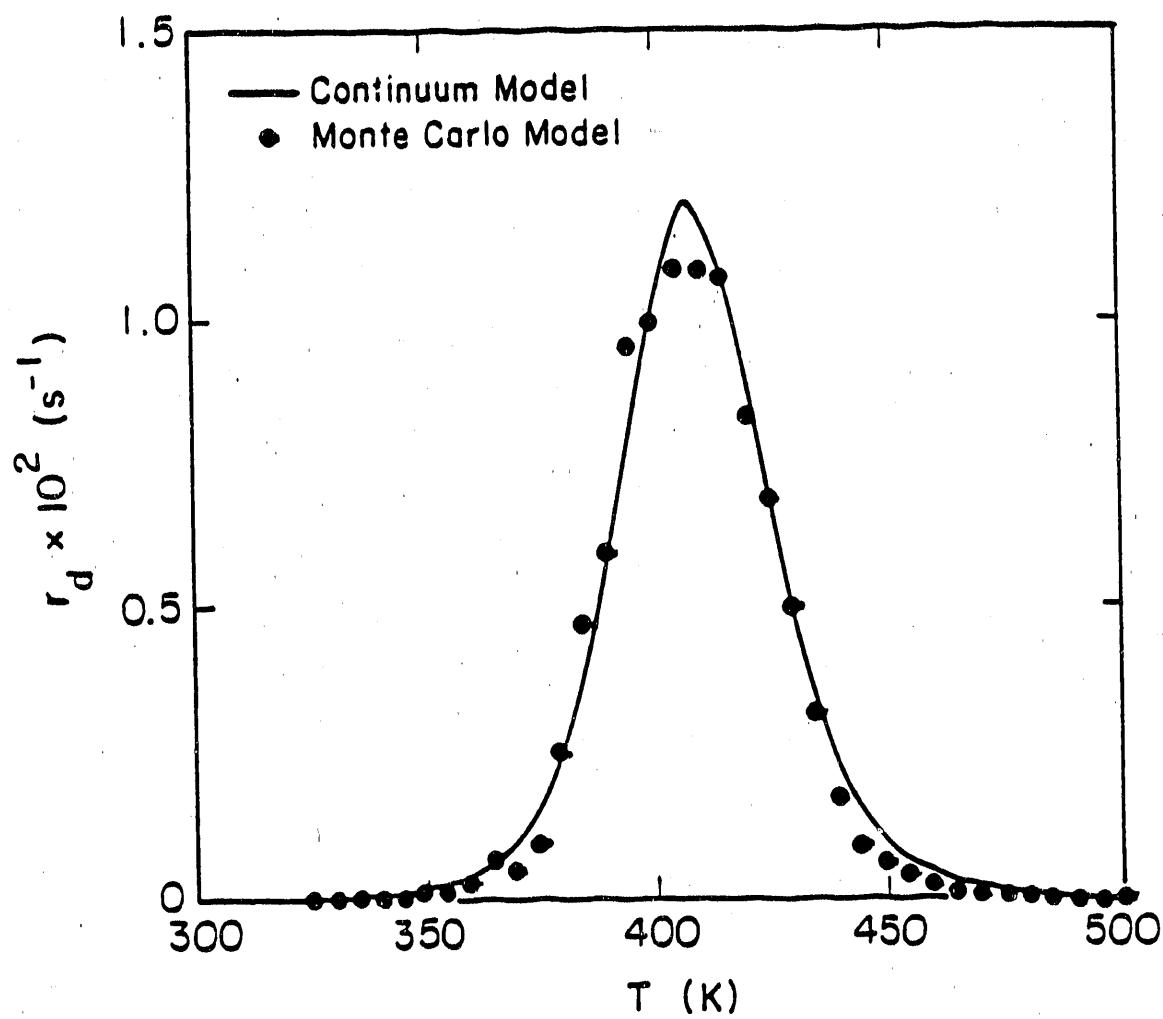


Fig. 2 Comparison of Monte Carlo and continuum models for the simulation of second order desorption where E_d is constant.
(See Table 1 for parameter values).

neighbors. The consequence of this residual coverage on the spectrum shape was an abrupt decrease in the desorption rate at 430 K.

The second order case was next extended to include metal-adsorbate (M-A) interactions. The desorption activation energy for a pair of adjacent adsorbates was calculated using eqs. 10 and 15. The resulting TPD spectrum, shown in fig. 3, exhibits peaks at 201 K, 255 K, and 350 K in marked contrast to the single peak observed at 406 K for the case of constant activation energy. The three peaks arise from the non-linear manner in which the average activation energy for the desorbing species varies with coverage. As shown in fig. 4, the activation energy for desorbing pairs of atoms decreases with increasing increasing adsorbate coverage, the decreases being most pronounced at coverages of 0.25, 0.50, and 0.75. The coverages at which rapid changes in the activation energy occur reflect the (100) lattice geometry chosen for this example. Figure 4 also shows the coverage dependence of the average activation energy for all pairs of adsorbates. This curve lies above the curve for the desorbing species, which indicates that at a given coverage, the desorbing atoms are, on the average, less strongly bound to the surface than those which remain adsorbed.

A comparison of the activation energy profile with the TPD spectrum shows that the flat regions of the activation energy curve correspond to changes in coverage associated with the peaks while the steep regions correspond to rearrangement of the surface by diffusion and stabilization of the adsorbates. It should also be noted that at the highest temperature shown, 0.22 monolayer of adsorbate

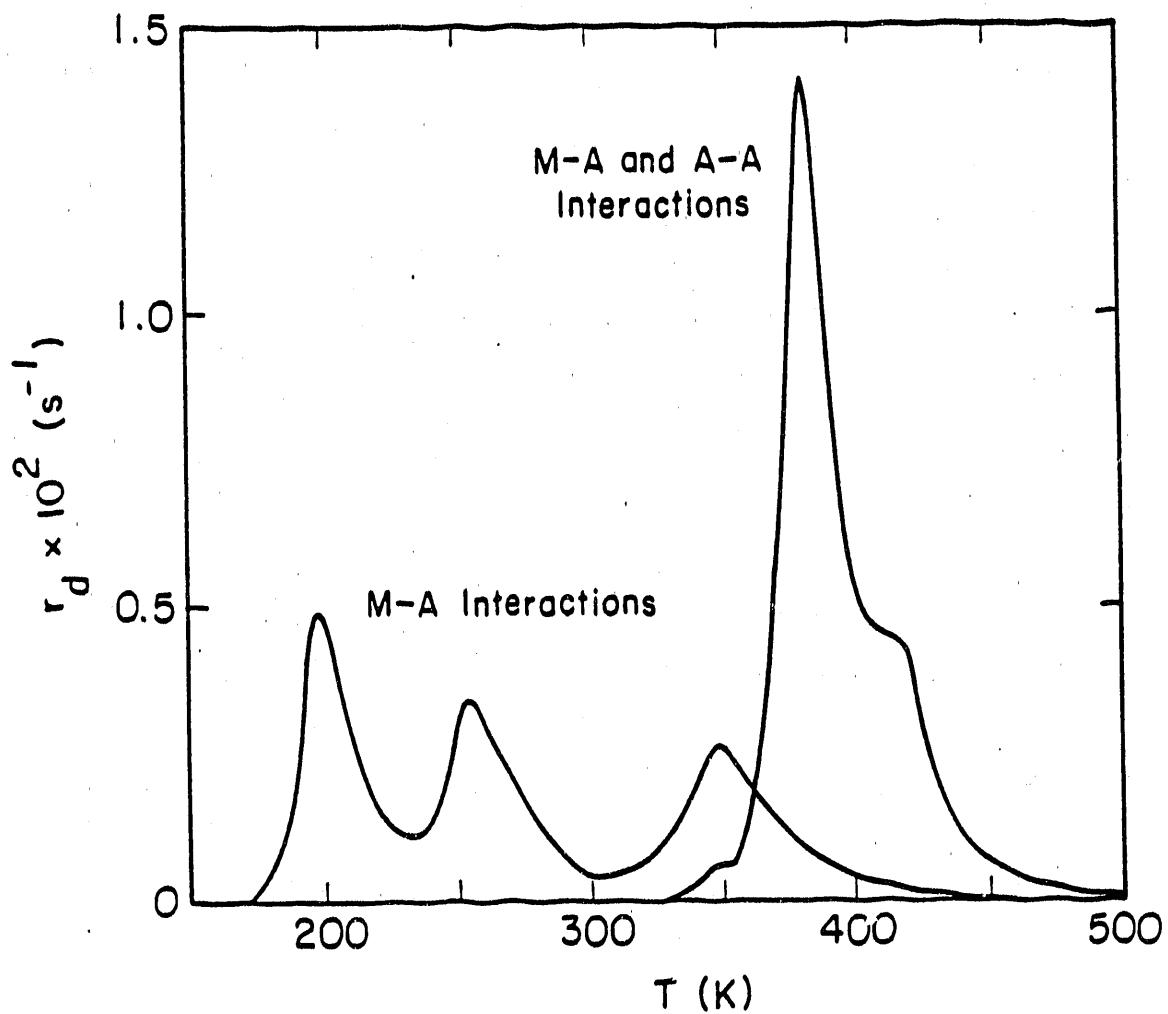


Fig. 3 Illustrations of the effects of M-A and M-A plus A-A interactions on the TPD spectra for second order desorption. (See Table 1 for parameter values).

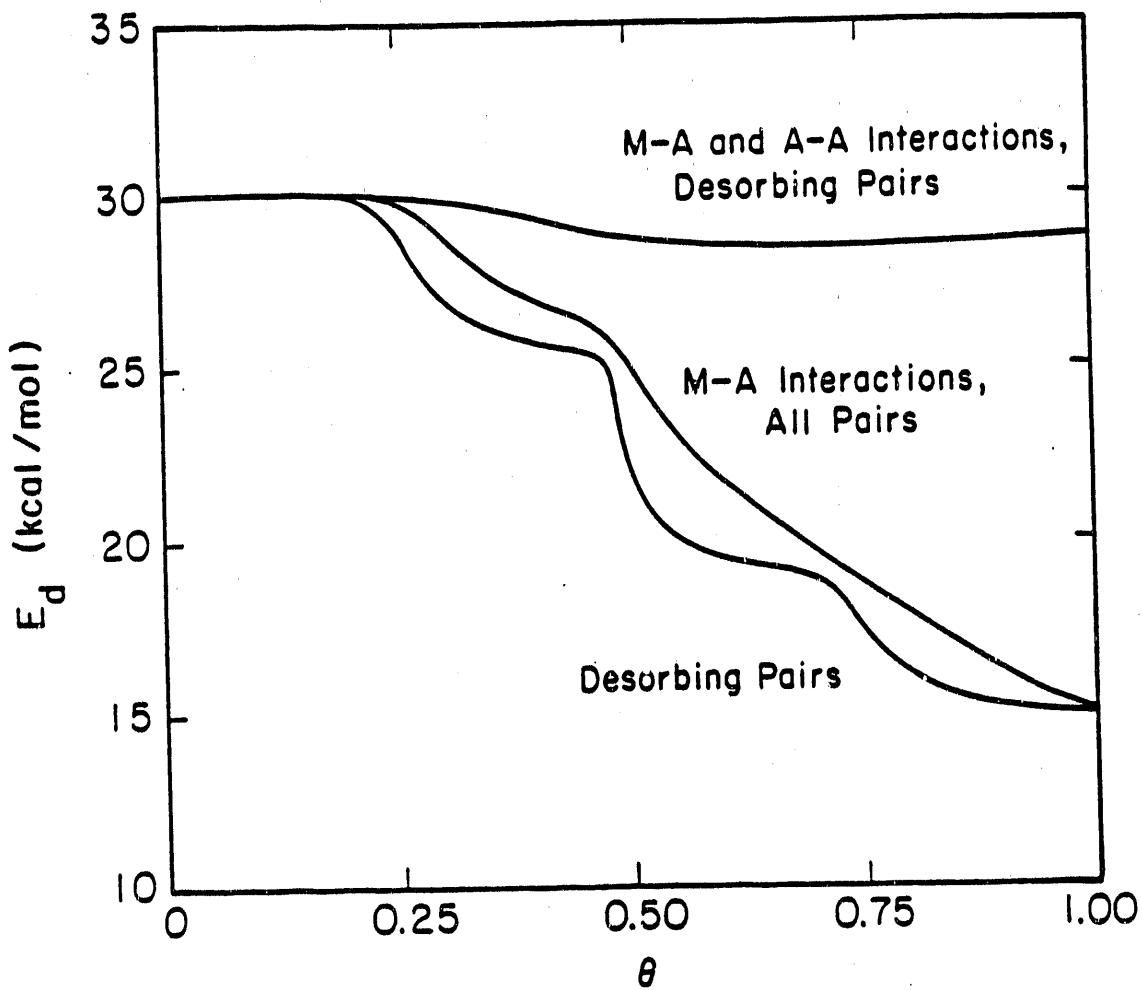


Fig. 4 Illustrations of the effects of M-A and M-A plus A-A interactions on the average E_d for second order desorption.

remained on the surface in the form of isolated atoms. This situation arises because the heat of adsorption of an isolated atom is higher than that for an atom located adjacent to another atom, and as a consequence, an isolated adsorbate is unlikely to diffuse to a lattice location where it can form an A-A pair.

Finally, second order desorption was considered for the case where both metal-adsorbate (M-A) and direct adsorbate-adsorbate (A-A) interactions occur. Activation energies for desorption in this case were calculated from eqs. 11 and 15. The results are shown in fig. 3. A single peak is observed at 384 K with a small low temperature shoulder at 350 K and a high temperature shoulder at 415 K. The large difference between this TPD spectrum and that corresponding to the case for M-A interactions alone arises from the stabilizing influence of the attractive A-A interactions. As seen from the activation energy profile in fig. 4, the high coverage activation energy is 12 kcal/mol greater than that for the case where only M-A interactions are considered. Consequently, the peak shifts to a higher temperature and is intermediate in shape and location between the cases for M-A interactions and no interaction (i.e. constant E_d). The shoulders which are observed arise from the non-linear variation of the activation energy with adsorbate coverage. When A-A interactions are taken into account, the residual coverage at 500 K decreases to 0.03 of a monolayer. This low residual coverage is a consequence of the attractive nature of the A-A interactions which makes diffusion of an A atom to an adjacent lattice site favorable energetically.

To illustrate further the effects of M-A and A-A interactions,

determinations were made of the distribution of local coverages for a fixed average coverage. Figure 5 illustrates the fraction of adsorbate pairs having 1 to 8 nearest- and/or next-nearest neighbors surrounding one member of the pair. Distributions are shown for the desorbing atoms and for all atoms residing on the surface at a given average coverage. Three cases are considered: no M-A or A-A interactions; M-A interactions; and M-A plus A-A interactions. It is evident that the distribution of local coordination numbers is sensitive to the presence or absence of M-A and A-A interactions. M-A interactions cause a narrowing in the distribution whereas A-A interactions have the opposite effect. M-A interactions, which are repulsive in nature, skew the distributions to lower coordination numbers. When M-A and A-A interactions occur, the complex tradeoff between repulsive M-A and attractive A-A effects skews the distribution to high coordination numbers at high coverage ($\theta > 0.5$) and to low coordination numbers at low coverage ($\theta < 0.5$). In the absence of M-A and A-A interactions the distribution of coordination numbers for all pairs, shown in fig. 5, is in excellent agreement with the relationship

$$f(CN + 1) = \frac{7!}{CN! (7-CN)!} \theta^{CN} (1-\theta)^{7-CN}, \text{ where } 0 \leq CN \leq 1 \quad (17)$$

derived for a random placement of adsorbate atoms on the surface sites.

CO Desorption from Pd(100)

The structure of carbon monoxide adsorbed on the (100) plane of palladium is well established. Early investigations have

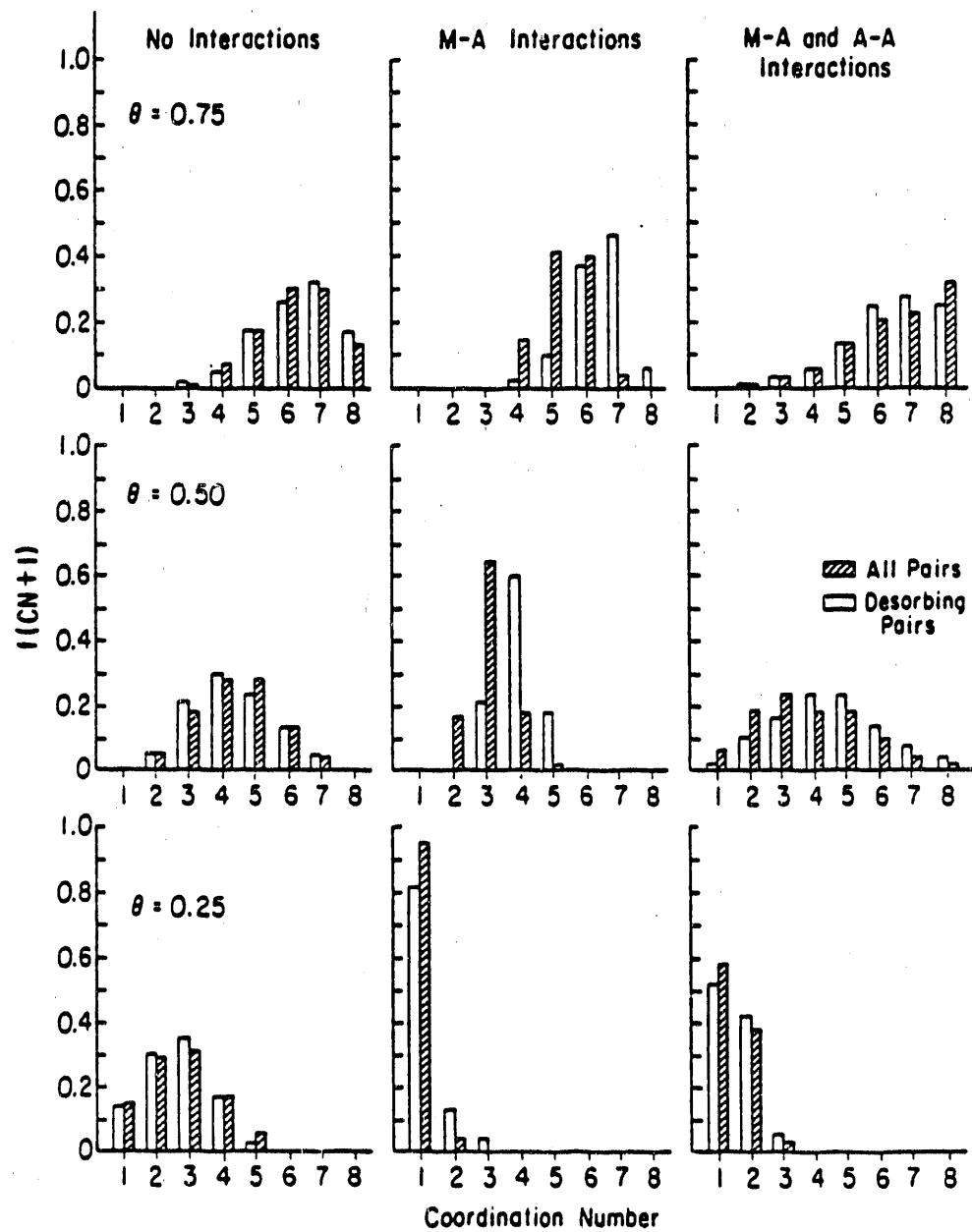


Fig. 5 Illustrations of the effects of M-A and M-A plus A-A interactions on $f(CN+l)$ for second order desorption.

demonstrated that CO adsorption occurs molecularly and is completely reversible [21,22]. LEED analysis indicates that the CO molecules occupy both types of bridge sites and form a $c(2\sqrt{2} \times \sqrt{2})R45$ overlayer at one-half of a monolayer [21-24]. Vibrational spectroscopic measurements by EELS and IR further support the bridge-bonding coordination for all coverages up to one-half monolayer [24,25]. Analysis of equilibrium isosteres and isobars has shown that the isosteric heat of adsorption (which is approximately equal to the activation energy for desorption in the case of nonactivated adsorption) for CO is 38.5 kcal/mol at low coverages and decreases with increasing coverage [22,23]. The activation energy for desorption and the preexponential factor are reported to be 36.8 kcal/mol and $2 \times 10^{16} \text{ s}^{-1}$, respectively, in the limit of low coverage [23].

Simulation of CO desorption from a Pd(100) surface was conducted in the following manner. A carbon monoxide coverage of one-half was placed on a Pd(100) lattice in a $c(2\sqrt{2} \times \sqrt{2})R45$ overlayer. The activation energy and the preexponential factor at zero coverage were taken from the experimental results presented in ref. [23] and are listed in table 1. The activation energy as a function of local CO coverage was calculated from eq. 10. Carbon monoxide molecules were allowed to diffuse on the surface with probabilities for individual diffusion jumps calculated from eq. 5. Because of steric constraints, a maximum of two CO molecules was allowed to be bridge bonded to a single Pd metal atom; furthermore, the adsorbates had to be located 180 degrees from each other on a metal atom. Exclusion of adsorbates bonded 90 degrees to

each other seems reasonable as this distance on the Pd surface is 1.94 Å and the CO hard sphere diameter is approximately 2.7 Å [22].

Results of simulations for coverages equal to and less than 0.5 are shown in fig. 6. The spectra exhibit first order desorption kinetics with a peak temperature at 490 K. Some low temperature broadening of the desorption peak is evident at the highest coverages. The extent of the low temperature broadening predicted by the simulation is not as large as is observed experimentally (see fig. 7). This is due most likely to the use of a constant preexponential factor in the simulations. Interpretation of the experimental data using a continuum model [23] led to the conclusion that the preexponential factor increases steeply with increasing coverage for coverages greater than 0.4. This variation in the preexponential factor was not included in the simulations because it could not be predicted on the basis of local environment without introducing an arbitrary parameter.

Values of the average activation energy for desorption as a function of coverage are shown in fig. 8. The activation energy exhibits a relatively constant value until a coverage of 0.40 at which point it begins to decrease sharply. Also shown in fig. 8 are values for the activation energy deduced from experimentally determined isobars [23]. It is evident that the coverage dependence of the activation energy determined from the Monte Carlo simulation is roughly similar to that observed experimentally.

H₂ Desorption from Mo(100)

Hydrogen adsorbed on the (100) plane of molybdenum exhibits

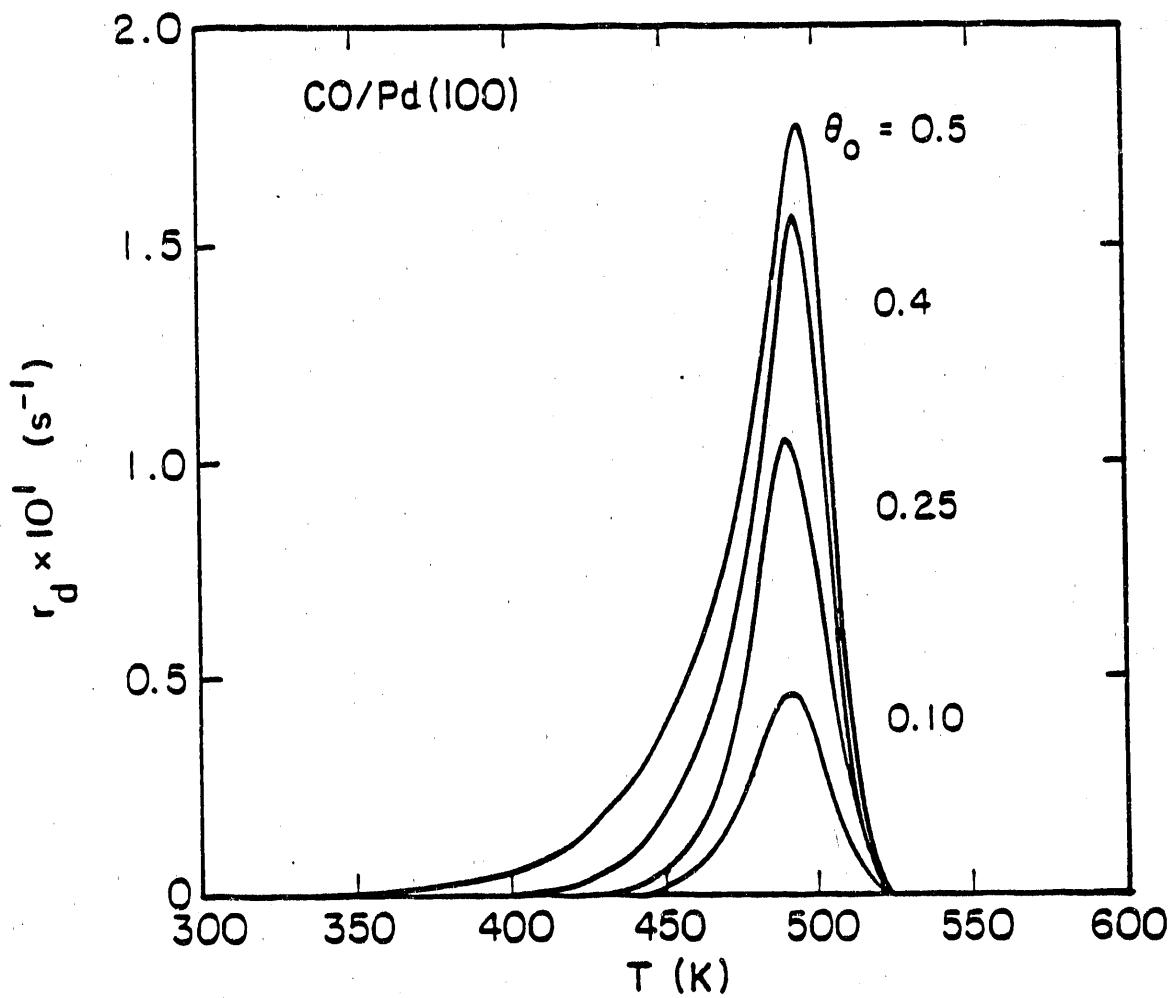


Fig. 6 Monte Carlo simulations of CO desorption from Pd(100). (See Table 1 for parameter values).

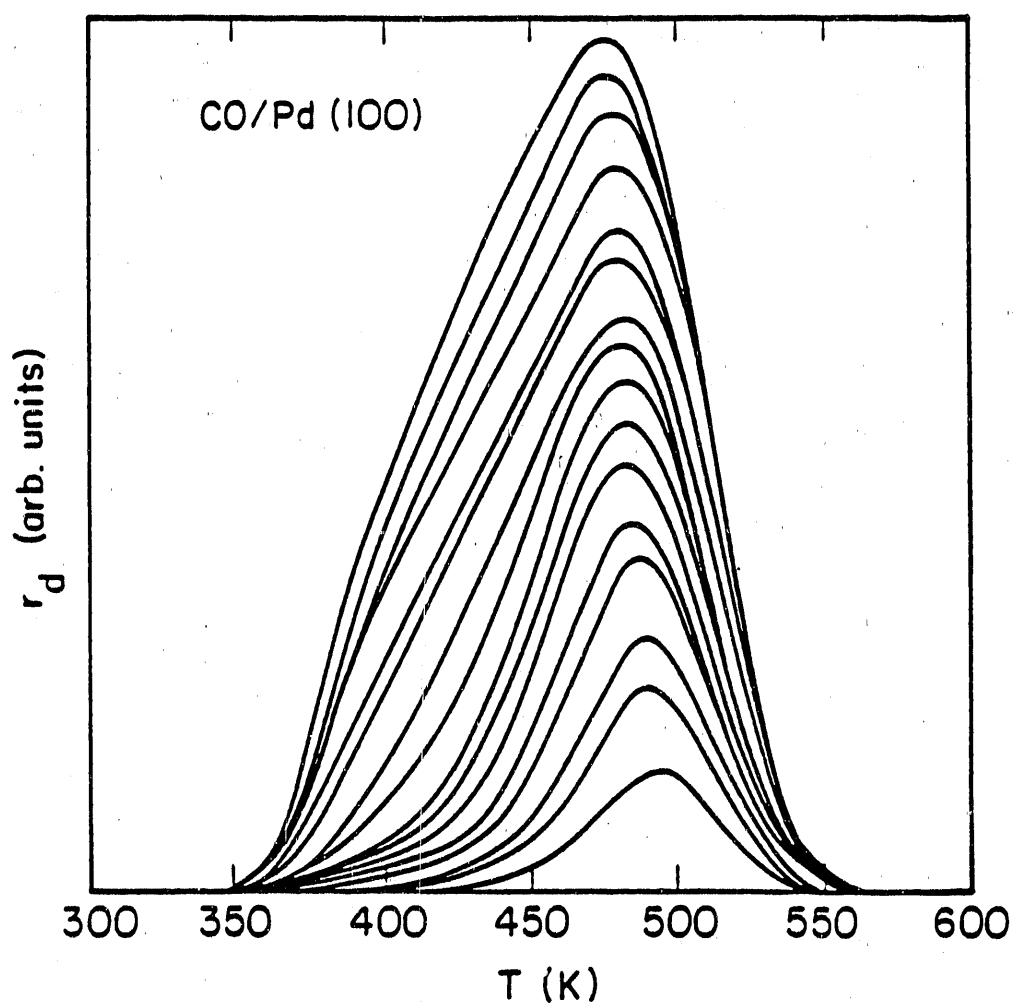


Fig. 7 TPD spectra for CO desorption from Pd(100) [23].

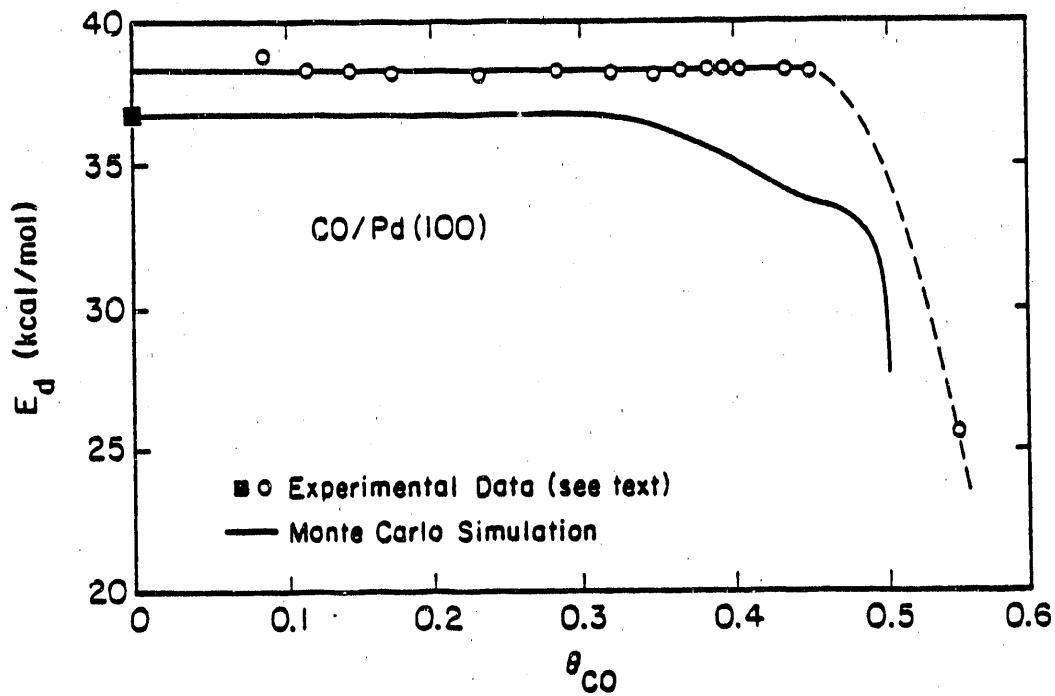


Fig. 8 Variation of the average E_d with θ_{CO} for CO desorption from Pd(100).

TPD spectra such as those shown in fig. 9 [26]. From hydrogen isotope experiments, it is now agreed that the three states, which occur in 2:1:1 ratio, all arise from dissociatively adsorbed hydrogen and that saturation coverage corresponds to two [26,27].

Vibrational spectroscopy suggests that the atomic hydrogen is located in bridge sites but that the nature of these bridge sites changes with coverage [27-29]. It is also believed that the molybdenum surface undergoes a surface reconstruction which consists of periodic displacements of molybdenum atoms by approximately 0.2 Å from their unreconstructed bcc location [27-33]. Although not fully characterized, this reconstruction occurs primarily at temperatures below 300 K and at low adsorbate coverages [27-33].

Simulations were conducted by assigning hydrogen atoms to bridge sites at all coverages. Consideration of atomic hydrogen distribution on the surface required including both nearest- and next-nearest-neighbor hydrogen atoms as potential desorption partners. Values of the activation energy and the preexponential factor used to initialize the model were obtained from a Redhead analysis of the low coverage experimental spectra shown in fig. 9 and are listed in table 1. No attempt was made to account for the effect of molybdenum displacements due to the surface reconstruction. Direct H-H interactions were limited to nearest-neighbor hydrogen atoms and accounted for in the following manner. The bond order for H-H interactions can, in principle, be calculated from eq. 7. It is known, though, that at long H-H distances, the Morse potential is a poor approximation to the bond dissociation characteristics. To

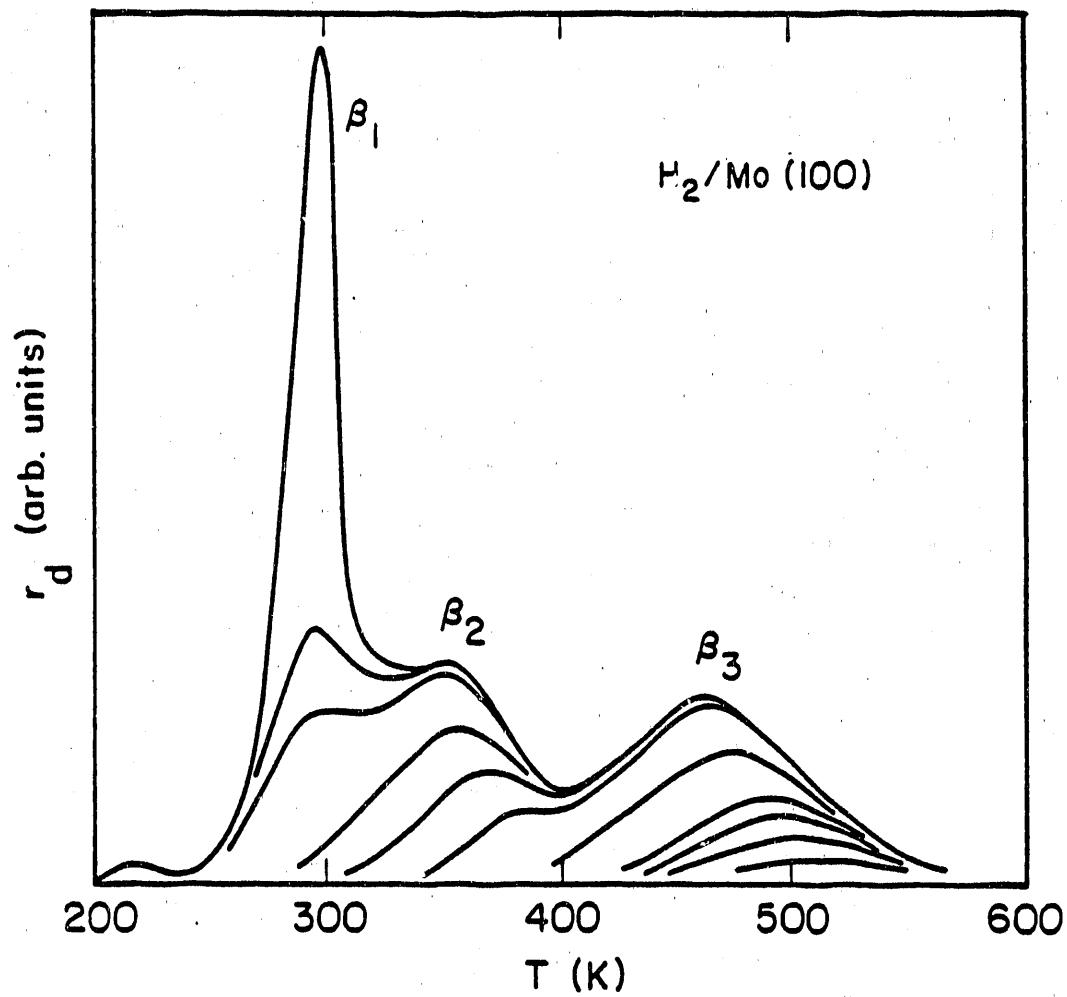


Fig. 9 TPD spectra for H_2 desorption from Mo(100) [26].

obtain better agreement with experimental data for gaseous hydrogen, the Morse potential parameter a in eq. 7 can be replaced by a' which is a three parameter function of distance given by:

$$a' = \{a_0 [1 + \gamma(r - r_0) + \lambda(r - r_0)^2]\}^{-1} \quad (18)$$

Values of the parameters a_0 , γ and λ in eq. 18 were taken from ref. [34]. The activation energy for H_2 desorption was then calculated as a function of coverage using eqs. 11 and 15.

Results of TPD simulations for different initial coverages randomly distributed on the surface are shown in fig. 10. Three peaks are observed at 285, 381, and 459 K. These features are referred to as β_1 , β_2 , and β_3 and have areas which are in the ratio of 2:1:1. As the initial coverage increases, first the β_3 , then the β_2 , and finally the β_1 peak is populated. Many features of the simulated spectra agree quite well with those of the experimentally observed spectra shown in fig. 9. It is noted that the predicted locations of the β_1 and β_3 peaks are almost identical to those observed experimentally, and that the ratios of all three peak areas are identical to those found experimentally. Figure 10 does show, however, that the predicted position of the β_2 peak is 30 K higher than that found experimentally (see fig. 9). Nevertheless, the level of agreement between the spectra shown in figs. 9 and 10 is sufficiently high to suggest that the three β peaks do not arise from three distinct types of the sites but rather from the non-linear manner in which the activation energy varies with coverage.

The variation in the activation energy with coverage predicted from the Monte Carlo model is shown in fig. 11. The activation

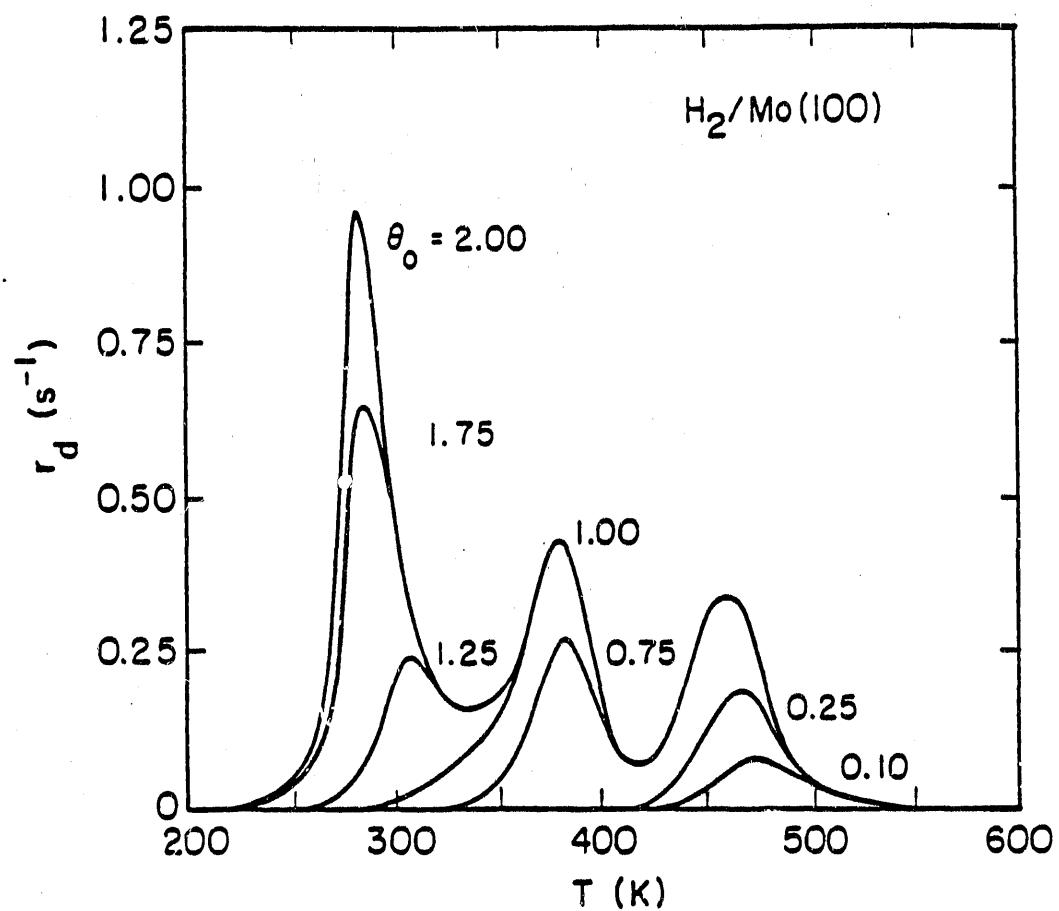


Fig. 10 Monte Carlo simulations of H_2 desorption from $Mo(100)$.
(See Table 1 for parameter values).

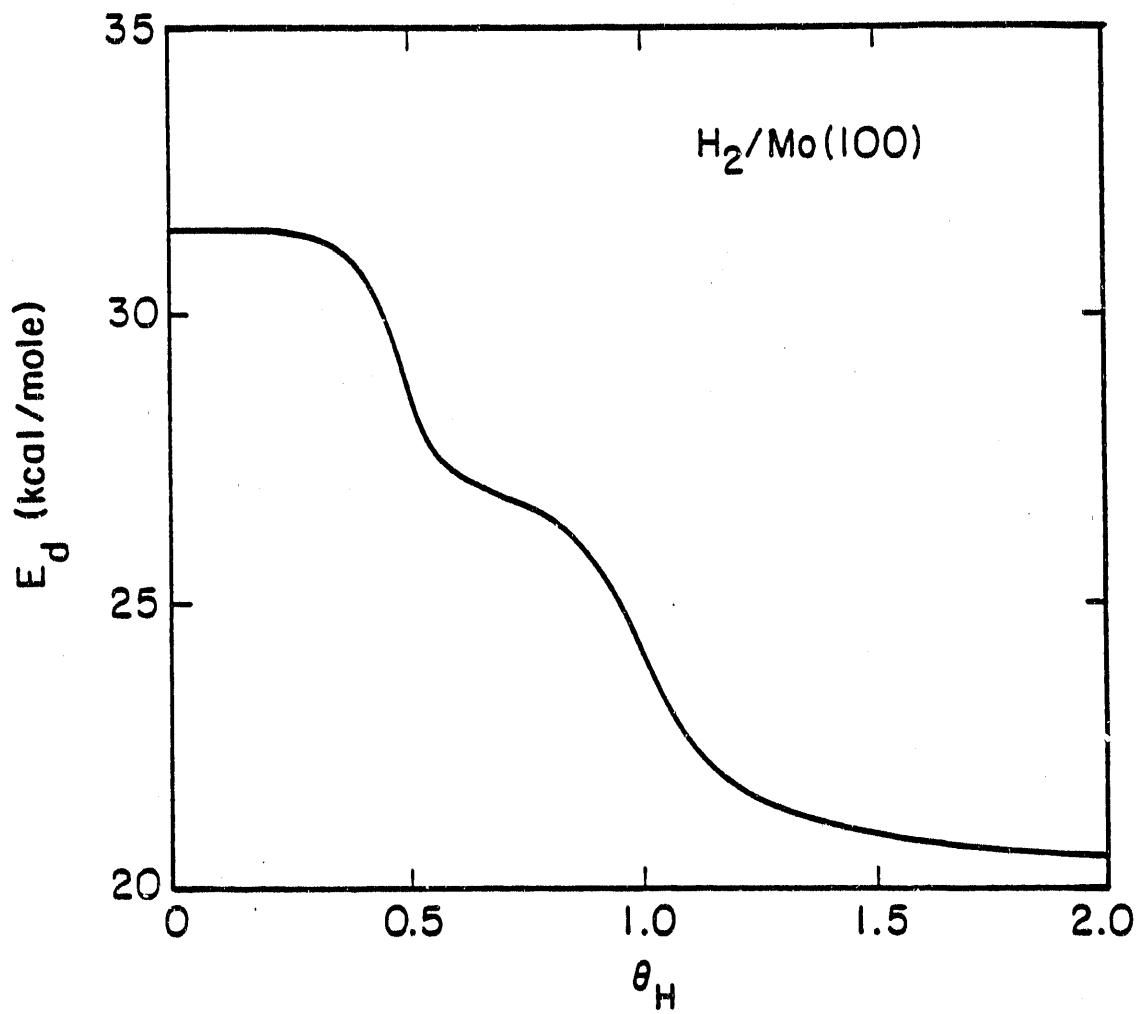


Fig. 11 Variation of the average E_d with θ_H for H_2 desorption from $Mo(100)$.

energy is relatively constant up to a coverage of 0.5 at which point the activation energy decreases rapidly to a new plateau. The second plateau is maintained up to a coverage of about 1.0. Thereafter, the activation energy steadily decreases down to its minimum value at a coverage of 2.0. The pattern shown in fig. 11 closely resembles that observed experimentally [35].

H₂ Desorption from Ni(111)

Desorption of H₂ from the (111) plane of Ni produces TPD spectra such as those shown in fig. 12. Both sets of experimental data exhibit two peaks with maxima separated by 30-80 K [36,37]. Isotopic labelling experiments have shown that the two peaks arise from dissociatively adsorbed hydrogen [36-38]. Unlike Mo(100), there is no evidence for reconstruction of the Ni(111) surface [36]. Dynamic LEED IV analysis [36] and EELS [39] indicate that hydrogen occupies both the fcc and hcp three-fold hollows. LEED analysis further suggests that at temperatures above 270 K, ordering of the hydrogen overlayer does not occur and that island formation is unlikely [36]. The maximum coverage observed is $\theta = 0.8 \pm 0.2$. This is less than a coverage of 2.0 which would correspond to full occupation of both types of three-fold sites [36].

Simulations were conducted by assigning hydrogen atoms to both types of three-fold sites at all coverages. Occupation of adjacent fcc and hcp type hollows was not allowed based on the LEED observations mentioned previously [36]. The preexponential factor and the activation energy at zero coverage were obtained from a Redhead analysis of the TPD spectra reported in ref. [37].

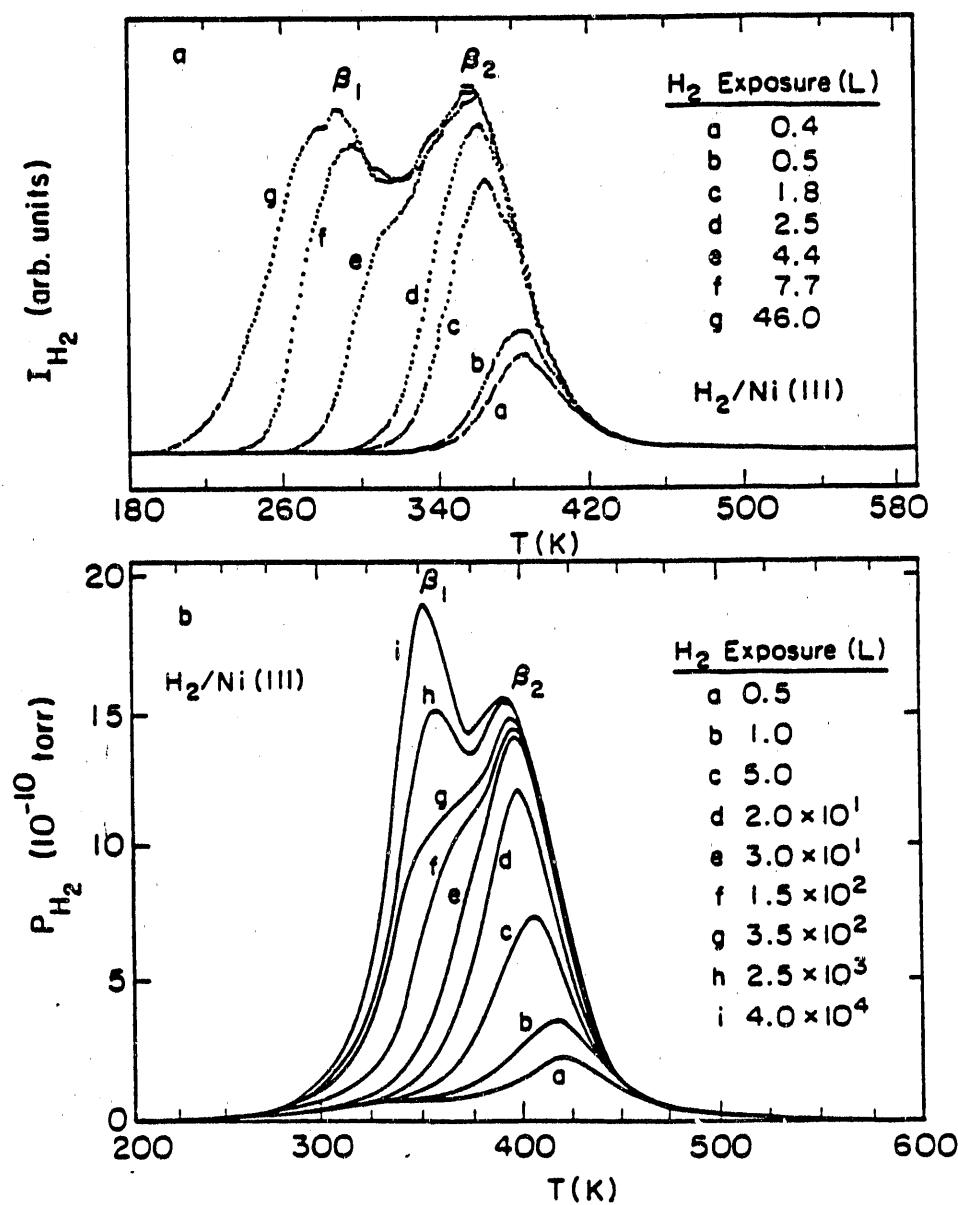


Fig. 12 TPD spectra for H_2 desorption from $\text{Ni}(111)$. Panel a is from ref. [37] and panel b is from ref. [36].

The results of simulations for different initial coverages are shown in fig. 13 and the variation of E_d for desorbing H₂ molecules with coverage is given in fig. 14. Figure 13 shows that for $\theta_0 = 0.80$, three overlapping peaks occur at 275, 300, and 340 K. The first two features fall in the range of temperatures for the β_1 peak seen in fig. 12a, whereas the third feature coincides with the β_2 peak of fig. 12a. Careful inspection of fig. 12a suggests that the peak designated β_1 may in fact be comprised of two components. A similar conclusion can be drawn from the TPD spectrum for D₂ desorption presented in ref. [37]. It should be noted that the relative peak heights in the spectrum computed for $\theta_0 = 0.80$, seen in fig. 13, differ from those shown in fig. 12a, and are closer to those shown in fig. 12b. The reason for this is that an initial coverage of 0.8 may be greater than the highest initial coverage that could be achieved by an H₂ exposure of 46L. This is certainly consistent with the observation that as the maximum exposure is increased, the β_1 feature becomes more intense than the β_2 feature (see fig. 12b).

4.0 CONCLUSIONS

A Monte Carlo model has been developed to describe temperature-programmed desorption of adsorbates from single crystal surfaces. The model accounts for the dependence of the activation energy for desorption on metal-adsorbate (M-A) and adsorbate-adsorbate (A-A) interactions and for the diffusion of adsorbates over the metal surface. The model correctly predicts not only the number and location of the peaks in a TPD spectrum, but also the dependence of the activation energy for desorption on the

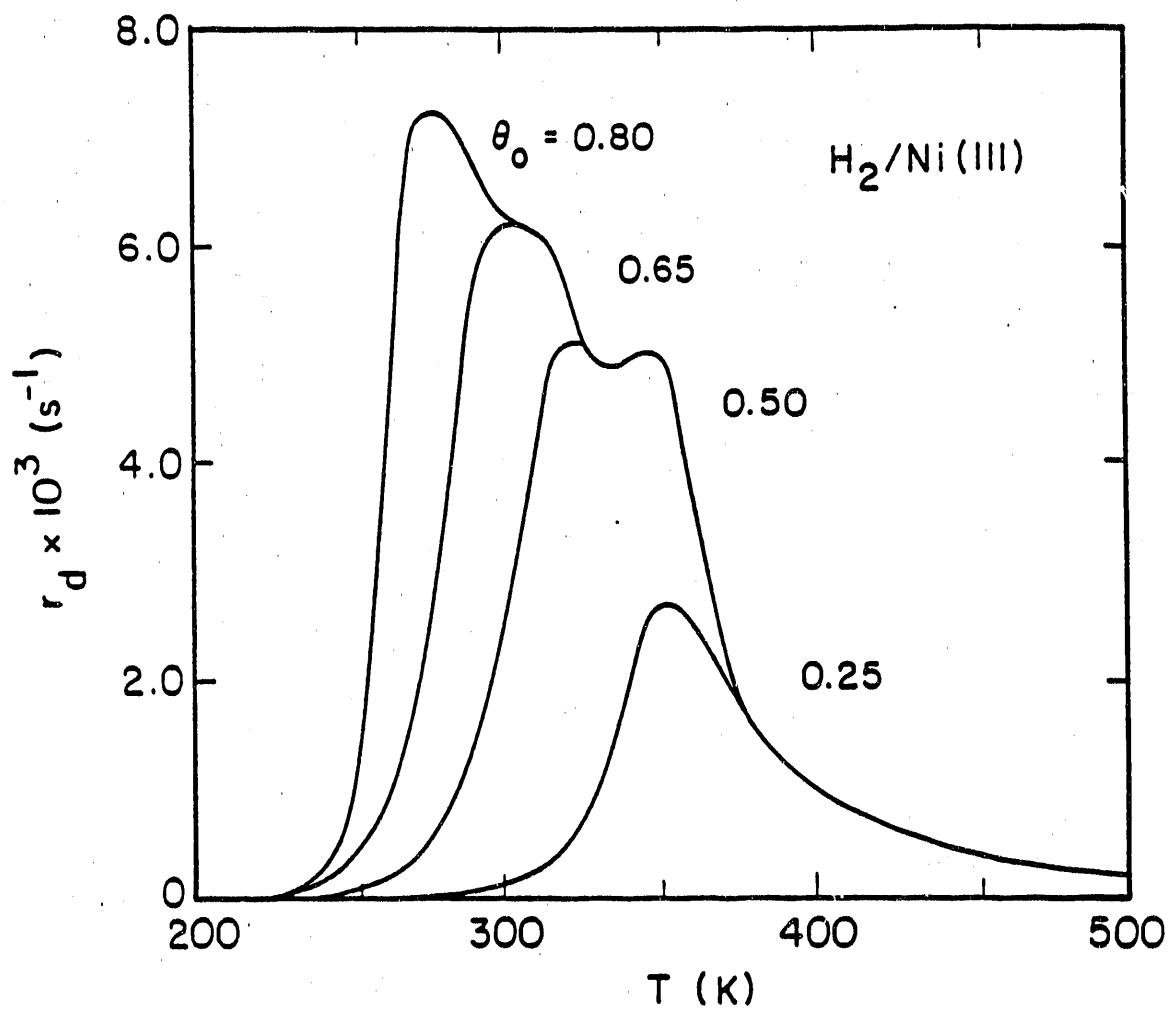


Fig. 13 Monte Carlo simulations of H_2 desorption from $\text{Ni}(111)$.
(See Table 1 for parameter values).

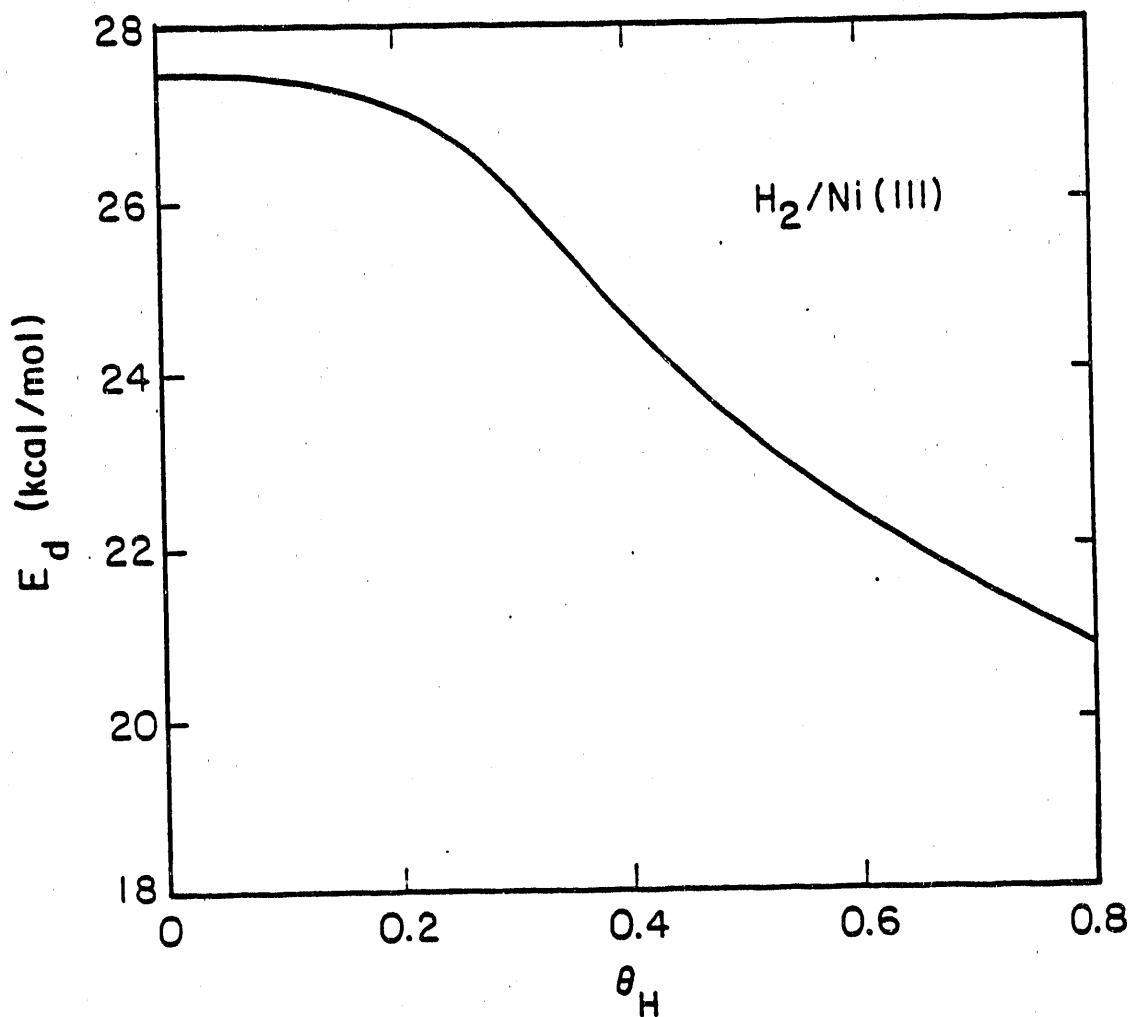


Fig. 14 Variation of the average E_d with θ_H for H_2 desorption from $Ni(111)$.

average surface coverage and the distribution of local coverages for a given average coverage.

The inclusion of M-A and A-A interactions in the model has a pronounced effect on the shape of the predicted TPD spectrum. Where only a single peak is observed in the absence of M-A and A-A interactions, multiple peaks are found when these interactions are included. Moreover, the introduction of M-A and A-A interactions causes the activation energy to decrease with increasing coverage, the shape of this function depending on the strength of these interactions. The calculations also show that the activation energy of the desorbing species is typically less than or equal to the average value for the entire adlayer. In physical terms, this means that for a given coverage, only the most weakly bound species desorb.

The method described in this paper has been used to simulate the desorption of CO from Pd(100) and the desorption of H₂ from Mo(100) and Ni(111). The simulated spectra for CO desorption from Pd(100) are narrower than those observed experimentally, most likely due to the neglect of the variation in the preexponential factor with coverage. However, the predicted variation in activation energy with coverage agrees reasonably well with that found experimentally. Monte Carlo simulation of H₂ desorption from Mo(100) and Ni(111) predicts three TPD peaks in both cases. The number of peaks, the peak temperatures for each peak, and the relative peak intensities are in good agreement with experimental observation. These calculations demonstrate that the appearance of multiple peaks is due to the non-linear nature of the M-A and A-A

interactions, rather than to the occurrence of distinct binding states.

ACKNOWLEDGMENT

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APPENDIX

We derive here probabilistic rate expressions which are equivalent to the continuum rate expressions. For any α -order kinetic rate process, the change in surface coverage can be represented by:

$$d\theta_i / dt = k_{d,i}^0 \exp[-E_{d,i} / RT] \theta_i^\alpha \quad (A1)$$

where subscript i represents a single type of environment. The rate of desorption as observed in the gas phase is then:

$$r_d = \frac{d\theta_i}{\alpha dt} = \frac{1}{\alpha} k_{d,i}^0 \exp\left[\frac{-E_{d,i}}{RT}\right] \theta_i^\alpha \quad (A2)$$

The probability of desorption in a given time interval can be expressed as

$$P_i^{(1)} = \int_{t_1}^{t_2} k_{d,i} \theta_i^\alpha dt / \int_0^{\infty} k_{d,i} \theta_i^\alpha dt \quad (A3)$$

The denominator in eq. A3 is θ_{i0} , the initial coverage in the i th environment. For sufficiently small time intervals, eq. A3 can be approximated by

$$P_i^{(1)} \approx k_{d,i} \theta_i^\alpha \Delta t / \theta_{i0} \quad (A4)$$

The number of adsorbates that have desorbed in the interval Δt is:

$$N_{d,i} = P_i^{(1)} N_s \theta_{i0} \quad (A5)$$

where N_s is the number of adsorption sites. The turnover frequency per surface atom as observed in the gas phase is then

$$r_{d,i} = N_{d,i} / (\alpha \Delta t N_s) \quad (A6)$$

Substitution of eqs. A4 and A5 into eq. A6 yields eq. A2 which demonstrates the equivalence between probabilistic formulations

and the more familiar continuum rate expressions. When many adsorbate environments coexist on the surface, the total rate of desorption can be expressed as

$$r_{d,i} = \sum_i N_{d,i} / (\alpha \Delta t N_s) \quad (A7)$$

NOMENCLATURE

A	Adsorbate
a	Morse potential constant (\AA)
a'	Morse potential function (\AA)
a_0	Morse potential constant
CN	Coordination number
D_{AB}	Bond dissociation energy for A-B bond (kcal/mol)
E_d	Activation energy for desorption (kcal/mol)
$E_{d,i}$	Activation energy for desorption from the i th environment (kcal/mol)
$E_{d,AA'}$	Activation energy for desorption for a diatomic pair A and A' (kcal/mol)
$f(CN+1)$	Fraction of adsorbate pairs having a given number of nearest-neighbors surrounding one member of the pair
k_d°	Preexponential factor for desorption (s^{-1})
$k_{d,i}^\circ$	Preexponential factor for desorption from the i th environment (s^{-1})
$k_{d,i}$	Rate coefficient for desorption from i th environment (s^{-1})
m_i	Number of adsorbates bonded to i th metal atom
N	Number of adsorbate atoms or molecules (mol)
N_d	Number of adsorbate atoms or molecules desorbing (mol)
$N_{d,i}$	Number of adsorbate atoms or molecules desorbing from the i th environment (mol)
N_s	Number of surface metal atoms (mol)
n	Number of metal atoms bonded to an adsorbate
$P_{i(1)}$	Probability of desorption from the i th environment
$P_{ij(2)}$	Probability of diffusion from site i to site j
$Q(x)$	Bond energy (kcal/mol)
Q_b	Equilibrium bond energy (kcal/mol)
Q_{oA}	Heat of adsorption of A in the on-top position (kcal/mol)
Q_i, Q_j	Heats of adsorption in sites i and j (kcal/mol)
$Q_{A,n}$	Heat of adsorption due to M-A interactions (kcal/mol)

$Q_{A,n}$	Heat of adsorption as a function of coordination (kcal/mol)
$Q_{A,n}(1)$	Heat of adsorption due to M-A interactions (kcal/mol)
$Q_{A,n}(2)$	Heat of adsorption due to A-A interactions (kcal/mol)
R	Gas constant (kcal/mol K)
r	Bond length (\AA)
r_0	Equilibrium bond length (\AA)
T	Temperature (K)
T_{\max}	Temperature at which surface is depleted of adsorbate (K)
t	Time (s)
Δt	Time interval (s)
x	Bond order
α	Desorption order
β	Heating rate (K/s)
β_i	Denotes the i th peak in a TPD spectrum
δ	A-A bond order
θ	Coverage
θ_i	Coverage in the i th environment
θ_{i0}	Initial coverage in i th environment
γ, λ	Morse potential constants

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Chapter IV

Monte Carlo Simulation of Temperature-Programmed Desorption of Coadsorbed Species**ABSTRACT**

A Monte Carlo model is presented for describing the temperature-programmed desorption of coadsorbed species from single-crystal surfaces. Interactions between the adsorbates and the metal surface as well as interactions between the adsorbates are taken into account using the bond-order-conservation-Morse-potential (BOC-MP) approach. The number, shape, and location of the peaks is shown to be sensitive to the binding energy, coverage, and coordination of each coadsorbed species. The presence of a strongly bound coadsorbate on a bcc(100) surface is shown to shift the desorption spectrum for associative desorption of adsorbed atoms to lower temperatures. TPD spectra for the concurrent associative desorption of A atoms and the desorption of B molecules from an fcc(100) surface are of two types: in one case, both species exhibit new low-temperature features far removed from their pure component spectra; in the second case, only the species undergoing associative desorption displays new spectral features. The simulated TPD spectra are in qualitative agreement with experimental results for H₂ coadsorbed with strongly bound atomic species on Mo(100) and Fe(100) surfaces as well as for CO and H₂ coadsorbed on Ni(100) and Rh(100) surfaces.

1.0 INTRODUCTION

The interactions of single adsorbates with metal surfaces have been examined extensively using temperature-programmed desorption (TPD) spectroscopy [1-6]. More recently, TPD spectroscopy has also been used to characterize the behavior of coadsorbed species [7-27]. Such studies can be classified into two types: those studies in which one species is strongly bound to the metal surface and does not desorb during a TPD experiment and those studies in which both species are of comparable binding energies and hence desorb in the same temperature range. The presence of a coadsorbed species can alter significantly the TPD spectrum of the adsorbed species. Strongly bound species tend to shift the TPD peaks for the desorbing species to lower temperature and may even cause individual peaks to merge. Coadsorbed species with comparable heats of adsorption may also produce downscale shifts in the position of TPD peaks and, in addition, give rise to new low-temperature features.

Theoretical descriptions of TPD spectra for coadsorbed species have been reported, using both continuum and stochastic models. In the continuum approach, the local heat of adsorption is usually described by pairwise-additive energetics. The interactions between adsorbates are assigned either an attractive or repulsive contribution to the local binding energy of the adsorbate, and the magnitude of the adsorbate-adsorbate interaction is treated as an adjustable parameter. The local occupancy of adsorption sites in the continuum models is specified in terms of a distribution function which depends on the average coverage and on the nature of the

interactions between adsorbates. It should be noted, though, that no a priori basis exists for assigning the relative magnitude and sign of an interaction or for assuming linearly-additive contributions from nearest-neighbor adsorbates. Stochastic models may also use pairwise-additive interactions to describe local energetics but make no assumptions about the distribution of local site occupancies.

Benziger and Schoofs [27] and Sundaresan and Kaza [28] have demonstrated using a continuum approach that the nature and strength of adsorbate-adsorbate interactions as well as the surface mobility of adsorbates can influence the number and shape of TPD peaks for coadsorbed species. Gupta and Hirtzel [29] have used a stochastic model to show that attractive interactions lead to sharper peaks, whereas repulsive interactions lead to broader peaks and, in some instances, multiple peaks. It was also found that the appearance or loss of spectral features was sensitive to the nature and magnitude of the energetic interactions.

This paper describes the simulation of TPD spectra for coadsorbed species using a Monte Carlo model similar to that developed recently [30] for the simulation of TPD spectra for single adsorbates. The model accounts for the effects of surface diffusion and surface coordination, and the effects of local coverage on the activation energy for desorption. Both metal-adsorbate (M-A) and adsorbate-adsorbate (A-A) interactions are described using the bond-order-conservation-Morse-potential (BOC-MP) approach [31]. The energetics determined with the BOC-MP method differ from the pairwise-additive energetics in two ways. First, the BOC-MP

method does not assume linearly-additive nearest-neighbor contributions to the total energetics. Rather, the functional form of the contributions of the M-A and A-A interactions to the total energetics are derived relationships within the model framework. Second, the activation energy for associative desorption is a model prediction rather than a model assumption as in the lattice-gas models. In this study, our interest has been the examination of how the coordination, coverage, and strength of bonding of each coadsorbed species influence TPD spectra. In addition, we show that the BOC-MP method can be used to account for the experimentally observed desorption kinetics when two species are coadsorbed on metal surfaces.

2.0 THEORY

Since the basic formulation of the model used for this work has already been discussed [30], we will briefly review the essential features of the model and then demonstrate how it can be extended to include the effects of two species adsorbed on a surface. In the model, each adsorbate is assumed to occupy a fixed type of site and to participate in only two surface processes: desorption and surface diffusion. Since the coverage in the vicinity of each adsorbate may be different, the local heat of adsorption and activation energy for desorption will depend upon the local environment. As described below, the BOC-MP [31] method can be used to determine the energetics associated with each adsorbed species.

The probability of desorption of a given species from site i in the time interval Δt can be defined as

$$P_i^{(1)} = k_{d,i}^0 \exp(-E_{d,i} / RT) \Delta t \quad (1)$$

where $k_{d,i}^0$ is the frequency factor for desorption and $E_{d,i}$ is the activation energy for desorption. The value of Δt in eq. 1 is chosen so that $P_i^{(1)}$ goes to unity at a temperature sufficiently high to guarantee virtually complete desorption from all sites of type i.

The rate of desorption, r_d , expressed as a turnover frequency based on the number of surface metal atoms, N_s , is

$$r_d = N_d / (\alpha N_s \Delta t) \quad (2)$$

where N_d is the number of adsorbate atoms or molecules desorbing in the time interval Δt . The parameter α is 1 for atomic or molecular desorption and 2 for associative desorption. Consequently, r_d is the rate of desorption as observed from the gas phase.

Since the activation energy for surface diffusion is 10-15% of that for desorption, whereas the frequency factors for the two processes are comparable, surface diffusion is expected to be a much more rapid process than desorption. As a consequence, diffusion is not treated as a rate process, but rather, we assume that adsorbate atoms or molecules will reposition themselves nearly instantaneously to achieve an equilibrium state. The equilibrium distribution of adsorbates can be determined from diffusion probabilities where the probability that an adsorbate in site i moves to an adjacent site j is given by [32]:

$$P_{ij}^{(2)} = \frac{\exp[-(Q_i - Q_j) / RT]}{1 + \exp[-(Q_i - Q_j) / RT]} \quad (3)$$

where Q_i and Q_j are the heats of adsorption for an adsorbate at sites i and j, respectively. Eq. 3 weights diffusion jumps according to the

magnitude of $(Q_i - Q_j)$. It should be noted that if $Q_i = Q_j$, then $P_{ij}(2) = 0.5$, which indicates that two sites of equal energy have equal probability of occupancy. For a sufficiently large number of jumps, application of eq. 3 yields an equilibrium distribution of adsorbates on a metal surface.

The values of the energetic terms $E_{d,i}$ and Q_i appearing in eqs. 1 and 3 are calculated using the BOC-MP approach [31]. With this method, each two-center interaction between an adsorbate atom A and a surface metal atom M is described by a Morse potential

$$Q(x) = Q_0 (2x - x^2) \quad (4)$$

where x is the bond order, and Q_0 is the equilibrium value of the M-A bond energy when $x=1$. The bond order in eq. 4 is defined by

$$x = \exp [- (r - r_0) / a] \quad (5)$$

where r is the M-A bond length, r_0 is the equilibrium bond length, and a is a scaling parameter. When A interacts with more than one metal atom, the total heat of adsorption is given by the sum of all two-center interactions. A further assumption of the BOC method is that along a reaction path describing the interactions of a molecular or atomic species with a metal surface, the total bond order is conserved and normalized to unity.

Within the BOC framework, the heat of chemisorption for an isolated atom A on a surface is given by:

$$Q_{A,n} = Q_{0A} (2 - 1/n) \quad (6)$$

where Q_{0A} is the heat of chemisorption of A in the on-top position, and n is the number of metal atoms to which A is coordinated. For an isolated molecular adsorbate AB coordinated via atom A to n metal atoms, the heat of chemisorption can be approximated by

$$Q_{AB,n} = \frac{Q_{oA}^2}{D_{AB} + Q_{oA}/n} \quad (7)$$

where D_{AB} is the gas phase A-B bond energy.

Eqs. 6 and 7 are valid for isolated adsorbate atoms or molecules on a surface. For higher coverages, however, situations may arise in which more than one adsorbate is bonded to a metal atom and, furthermore, the adsorbates may begin to interact directly with each other. To account for these metal-adsorbate (M-A) and adsorbate-adsorbate (A-A) interactions, the total binding energy of species A is partitioned as follows:

$$Q_{A,n} = Q_{A,n}^{(1)} + Q_{A,n}^{(2)} \quad (8)$$

where $Q_{A,n}^{(1)}$ is the heat of adsorption due to M-A interactions and $Q_{A,n}^{(2)}$ is the heat of adsorption due to A-A interactions. Both $Q_{A,n}^{(1)}$ and $Q_{A,n}^{(2)}$ can be calculated explicitly as a function of the local surface environment, subject to the total bond order of A for both M-A and A-A interactions being conserved to unity [30,31].

The value of $Q_{A,n}^{(2)}$ can be expressed as

$$Q_{A,n}^{(2)} = \sum_{i=1}^n \sum_{l=1}^l 0.5 D_{AA} (2\delta_{il} - \delta_{il}^2) \quad (9)$$

where D_{AA} is the A-A bond dissociation energy, and δ_{il} is the bond order for the A-A interaction between the A atom coordinated with metal i and the l th nearest-neighbor A atom also coordinated with metal atom i . The summation over l in eq. 9 is to account for all nearest-neighbor A atoms. The occurrence of A-A interactions weakens the bond order associated with the M-A interactions and, as

a consequence, the bond order for each component of an M_n -A bond is given by

$$x_{i,n} = \frac{1}{n} - \sum_{l=1}^L \delta_{il} \quad (10)$$

The value of $Q_{A,n}(1)$ when more than one adsorbate is bonded to an individual metal atom is given by [30,31]

$$Q_{A,n}^{(1)} = \sum_{i=1}^n \left[\frac{Q_{0A}}{m_i} \left(2 - \frac{1}{m_i} \right) (2x_{i,n} - x_{i,n}^2) \right] \quad (11)$$

where m_i is the number of adsorbates bonded to the i th metal atom, and $x_{i,n}$ is given by eq. 10.

Up to this point, we have treated the energetics of a single type of species adsorbed on a surface. We now extend the model to include situations where two types of adsorbates are present on a surface. When coadsorption of species A and B occurs, interactions through the metal between M, A, and B are possible as well as direct A-A, B-B, and A-B interactions. The total heat of adsorption for species A can again be partitioned (see eq. 8) between the metal-adsorbate interactions $Q_{A,n}(1)$ and the direct adsorbate-adsorbate interactions $Q_{A,n}(2)$. A similar expression can be written for the total heat of adsorption of species B.

When A and B are coadsorbed on the surface, the calculation of the direct A-A, B-B, and A-B interactions can be determined from eq. 9 in the same manner as already described for A-A interactions. The values of D_{XY} and δ appearing in eq. 9 are then the values for the respective X-Y interactions.

The values of $Q_{A,n}(1)$ and $Q_{B,n}(1)$ can be calculated for the case

of coadsorption in the following way. The total metal-adsorbate energy, Q_{M-A-B} , for the coadsorption of species A and B with a metal atom M is given by [31]:

$$Q_{M-A-B} = N_A \frac{Q_{A,n}}{n_A} (2x_A - x_A^2) + N_B \frac{Q_{B,n}}{n_B} (2x_B - x_B^2) \quad (12)$$

where N_A and N_B represent the number of A and B species, respectively, bonded to the metal atom, and $Q_{A,n}$ and $Q_{B,n}$ represent the heats of adsorption for isolated A and B species in their respective n-fold coordination (see eqs. 6 and 7). Equation 12 is minimized with respect to x_A and x_B under the bond order constraint for the metal atom that

$$N_A x_A + N_B x_B = 1 \quad (13)$$

The resulting expressions for x_A and x_B are:

$$x_A = \frac{-N_A \frac{Q_{B,n}}{n_B} + \frac{N_A Q_{B,n}}{N_B n_B} + N_A \frac{Q_{A,n}}{n_A}}{\frac{N_A^2 Q_{B,n}}{N_B n_B} + N_A \frac{Q_{A,n}}{n_A}} \quad (14)$$

$$x_B = \frac{1 - N_A x_A}{N_B} \quad (15)$$

The magnitude of the total through-metal interactions for A and B, $Q_{A,n}^{(1)}$ and $Q_{B,n}^{(1)}$, are then obtained from the expressions:

$$Q_{A,n}^{(1)} = \frac{Q_{A,n}}{n_A} \sum_i (2x_{Ai} - x_{Ai}^2) ; Q_{B,n}^{(1)} = \frac{Q_{B,n}}{n_B} \sum_j (2x_{Bj} - x_{Bj}^2) \quad (16)$$

The bond orders x_{Ai} and x_{Bj} are calculated from eqs. 14 and 15, respectively. The summations over i and j in eq. 16 include all of the metal atoms to which an adsorbate is bonded. It should be noted

that because of the structure of eqs. 14 and 15, the bond orders x_A and x_B can become negative, depending on the magnitudes of $Q_{A,n}$, $Q_{B,n}$, n_A , n_B , N_A , and N_B . A negative value of x_A or x_B simply implies that adsorption under these conditions would not take place.

The expression for $Q_{A,n}^*$ derived above (see eq. 8) can be used to determine the activation energy for desorption. For the desorption of a single atom or adsorbed molecule, A, the activation energy is given by the heat of adsorption, $E_{d,A} = Q_{A,n}^*$. Implicit in this is that adsorption is not an activated process. For homonuclear associative desorption, the activation energy $E_{d,AA'}$ is given by [31]

$$E_{d,AA'} = \frac{Q_{A,n}^* Q_{A',n}^*}{Q_{A,n}^* + Q_{A',n}^*} \quad (17)$$

where $Q_{A,n}^*$ and $Q_{A',n}^*$ represent the heats of adsorption of A and A', respectively. The two recombining atoms are designated as A and A' to denote that the local environments of the two atoms may differ.

The energy and probability formulations described above were incorporated into a Monte Carlo algorithm for simulating the temperature-programmed desorption of adsorbates. The metal surface was represented by a 100 by 100 array of numbered sites, and periodic boundary conditions were used to eliminate edge effects. Adsorbate atoms or molecules were placed on the metal surface lattice to achieve a desired initial coverage, θ_0 . The temperature was initialized at T_0 and taken to be constant at this value for the time interval Δt . A surface site (or pair of adjacent surface sites for associative desorption) was then selected in a random fashion. If the site (or both sites for associative

desorption) were occupied, the probability of desorption was calculated using eq. 1. The local activation energy for desorption was calculated using eq. 8 (or eq. 17 for associative desorption). The calculated desorption probability was then compared with a random number, R , $0 < R < 1$. If $R < P_i(1)$, the adsorbate (or pair of adsorbates for associative desorption) was removed from the lattice, and N_d in eq. 2 was incremented by one (two for associative desorption). Conversely, if $R \geq P_i(1)$, the site (or pair of sites for associative desorption) remained occupied. The preceding steps were repeated $N_{s\theta}$ times ($0.5 N_{s\theta}^2$ times for associative desorption) [30]. After completing the sampling of the surface, the rate of desorption for this time interval was calculated using eq. 2. When coadsorbates existed on the surface, each species was sampled alternately.

A redistribution of the remaining adsorbates was carried out next to account for the effects of surface diffusion. A surface site and an adjacent site were chosen at random. If the surface site were occupied and the nearest-neighbor site were vacant, a probability of diffusion was calculated using eq. 3. Q_i was taken as the heat of adsorption in the initial site, and Q_j was taken as the heat of adsorption in the final site. The calculated value of $P_{ij}(2)$ was then compared with a random number, R , $0 < R < 1$. If $R < P_{ij}(2)$, the adsorbate was moved from the initial site, i , to the adjacent site, j . Conversely, if $R \geq P_{ij}(2)$, movement of the adsorbate was not allowed. A sufficient number of surface visitations was allowed to attain an equilibrium distribution of adsorbates on the surface. When coadsorbates existed on the surface, each was allowed to

diffuse alternately.

The desorption and diffusion calculations described above constitute a Monte Carlo Step (MCS). As noted above, the time interval associated with an MCS is Δt , and over this interval the temperature is constant. At the end of an MCS, the temperature was increased by the increment $\beta\Delta t$, where β is the desired heating rate. A new MCS was then carried out at the next temperature. This process was repeated until a temperature was reached for which the surface was depleted of adsorbate. A plot of the desorption rate as a function of temperature for each species then yielded a TPD spectrum for each species. In addition, the average activation energy of desorption could also be determined as a function of the total surface coverage. For both the TPD spectra and the average activation energy profiles, best-fit curves were drawn by eye through the data points. All of the calculations described above were carried out on an IBM 3090 computer. Typical run times were 2-10 CPU minutes per simulation. Random numbers were generated by the IMSL linear congruential number generator GGUBFS [33].

3.0 RESULTS AND DISCUSSION

3.1 Associative desorption in the presence of a fixed coadsorbate

The associative desorption of A atoms in the absence of a coadsorbate was considered first. The A atoms were positioned in two-fold bridge sites on a bcc(100) surface at an initial coverage of two and were free to diffuse. Both nearest- and next-nearest-neighbor A atoms were considered as possible partners for associative desorption as A_2 . The heat of adsorption for individual A

atoms was calculated taking into account both M-A and A-A interactions. The parameter values used for the simulation of the TPD spectra are listed in table 1.

Figure 1a shows the TPD spectrum for the desorption of A_2 . Three peaks are observed at 285, 381, and 459 K. The dependence of the activation energy on the coverage of A, shown in fig. 1b, exhibits a step-wise decrease as a function of increasing coverage. The flat portions of the activation energy versus coverage curve reflect local surface configurations with relatively constant energetics whereas the steep portions, which occur after changes in coverage due to the evolution of the peaks, correspond to rearrangement and stabilization of the adsorbates by surface diffusion. A map of the surface for $\theta_A = 1.0$, illustrated in fig. 2, shows no long range order. Short range order, however, is observed with two A atoms being bonded on the average to every metal atom so as to constitute nearest neighbors. This type of bonding configuration results from the trade-off between the repulsive effects of more than one A atom being bonded to individual metal atoms and the attractive effects of the direct A-A interactions.

Simulations were conducted next for the associative desorption of A in the presence of an immobile coadsorbate B. The B atoms were positioned in hollow sites at a coverage of $\theta_B = 0.25$ in a $p(2 \times 2)$ structure. In calculating the heat of adsorption of A atoms, M-A, M-B, and A-A interactions were taken into account, but no A-B interactions were included. The binding energy of B was chosen to be sufficiently high so that the rate of B desorption could be neglected.

Table 1 Model parameters for the TPD simulations of associative desorption in the presence of a fixed coadsorbate.^{a)}

Species	Coordination	$Q_{i,n}$ (kcal/mol)	k_d^* (s ⁻¹) ^{b)}	$\theta_{i,0}$ ^{c)}	Direct interactions	D_{XY} (kcal/mol) ^{d)}	Fig.
A	Bridge	63	1×10^{16}	2.0	A-A	109.5	1
A	Bridge	63	1×10^{16}	2.0	A-A	109.5	3
B	Hollow	140	1×10^{16}	$0.25 p(2x2)$	-	-	
A	Bridge	63	1×10^{16}	2.0	A-A	109.5	3
B	Hollow	190	1×10^{16}	$0.25 p(2x2)$	-	-	
A	Bridge	63	1×10^{16}	2.0	A-A	109.5	5
B	Hollow	140	1×10^{16}	$0.50 c(2x2)$	-	-	
A	Bridge	63	1×10^{16}	2.0	A-A	109.5	5
B	Hollow	190	1×10^{16}	$0.50 c(2x2)$	-	-	
A	Bridge	63	1×10^{16}	2.0	A-A,A-B	109.5,100	6
B	Hollow	140	1×10^{16}	$0.25 p(2x2)$	A-B	100.0	
A	Bridge	63	1×10^{16}	2.0	A-A,A-B	109.5,100	6
B	Hollow	190	1×10^{16}	$0.25 p(2x2)$	A-B	100.0	
A	Bridge	63	1×10^{16}	2.0	A-A,A-B	109.5,100	7
B	Hollow	140	1×10^{16}	$0.50 c(2x2)$	A-B	100.0	
A	Bridge	63	1×10^{16}	2.0	A-A,A-B	109.5,100	7
B	Hollow	190	1×10^{16}	$0.50 c(2x2)$	A-B	100.0	

a) Simulations were conducted for a bcc(100) surface with a heating rate $\beta=74$ K/s. Each simulation was the average of 1 run with $\Delta t = 0.41$

b) For second order desorption, k_d^* represents the product of the second order preexponential factor times the concentration of adsorption sites.

c) Initial coverage and structure of adsorbate overlayer.

d) The value of D_{XY} is for the respective direct interactions listed in the preceding column.

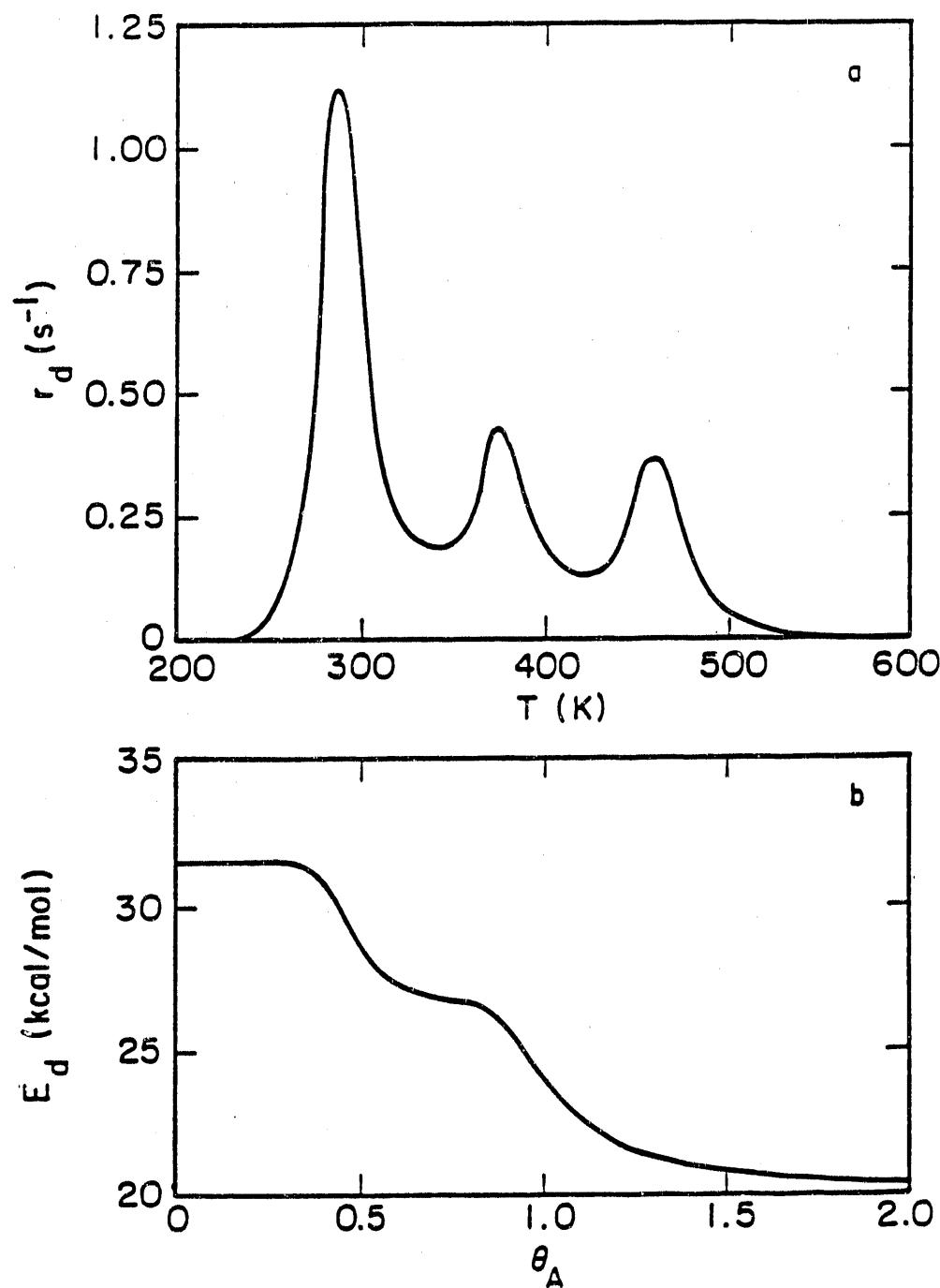


Fig. 1 a. Simulation of A_2 desorption. See table 1 for model parameters. b. Variation of the average E_d with θ_A .

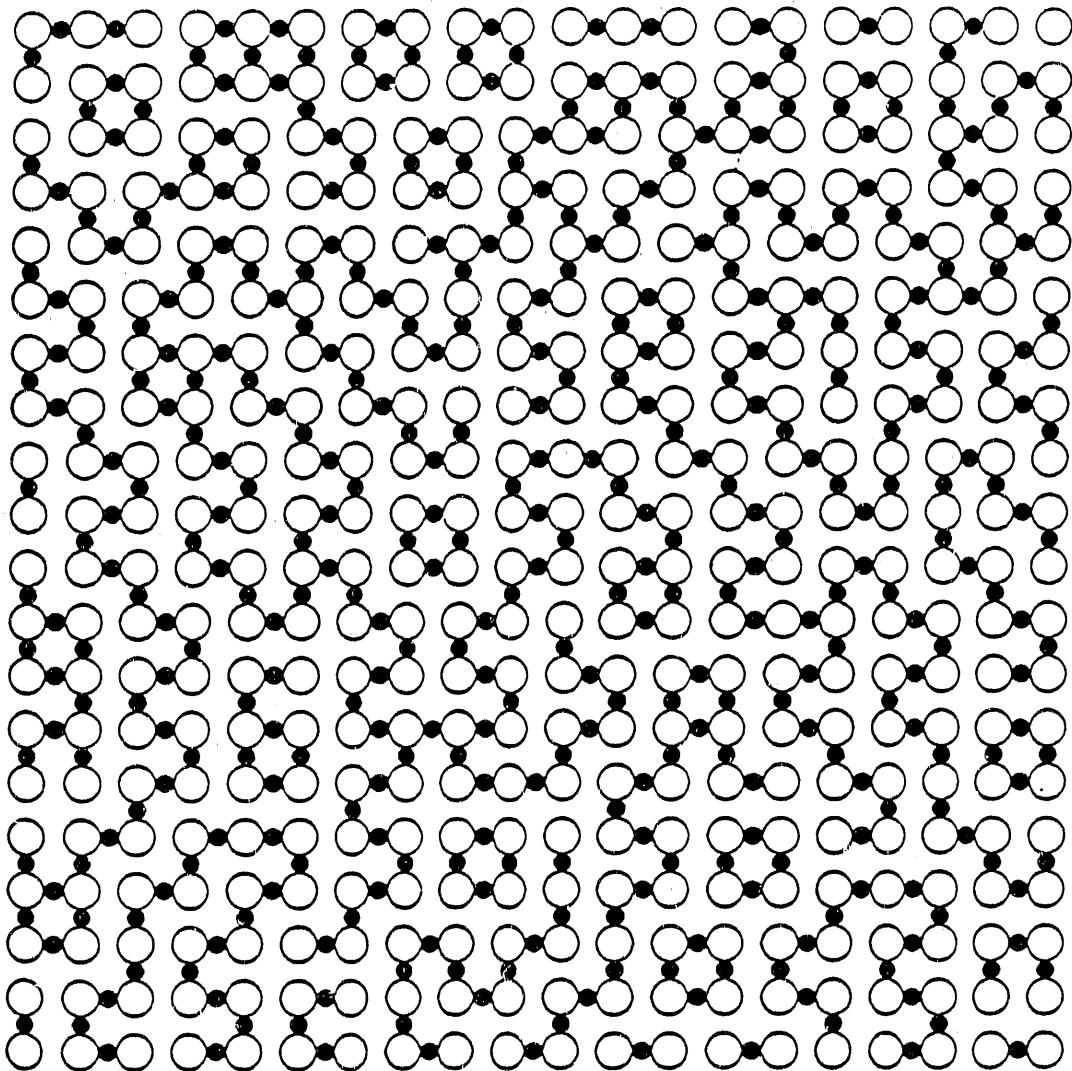


Fig. 2 Surface map for A(●) at $\theta_A=1.0$ for the simulation in fig. 1a. The open circles are metal atoms.

TPD spectra for two different values of the initial binding energy of B are displayed in fig. 3a. For $Q_{B,n} = 140$ kcal/mol, three peaks for A_2 are observed. Relative to the spectrum shown in fig. 1a, however, the peaks in fig. 3a are shifted to lower temperatures and exhibit less peak separation. For $Q_{B,n} = 190$ kcal/mol, three peaks are also observed, and the entire spectrum is shifted to even lower temperatures relative to the spectrum in fig. 1a. The activation energy profiles for both values of $Q_{B,n}$, displayed in fig. 3b, show less variation with θ_A than is seen in the case of pure A (see fig. 1b). As $Q_{B,n}$ increases, the activation energy for desorption of A_2 decreases. This trend is a direct result of the reduction in the heat of adsorption of A due to the decrease in x_A . Figure 4 illustrates the surface map for $\theta_A = 1.0$ when $Q_{B,n} = 140$ kcal/mol. The local structure seen here is very similar to that in fig. 2 and, in fact, shows little influence due to the presence of the B atoms. This is a direct consequence of the absence of A-B interactions and the equivalence of all the metal atoms (i.e., each metal atom is bonded to one B atom).

Figure 5a shows TPD spectra for two cases in which the coverage of B is increased to 0.5 in a $c(2\times 2)$ overlayer. For $Q_{B,n} = 140$ kcal/mol, one peak is seen at 255 K with a high-temperature shoulder. For $Q_{B,n} = 190$ kcal/mol, only one peak is observed at 200 K. The corresponding activation energy profiles for A_2 versus θ_A are given in fig. 5b. Compared to the results for $\theta_B = 0.25$, the values of the activation energy are lower and exhibit less variation in magnitude with increasing θ_A .

A final set of simulations was conducted to illustrate the

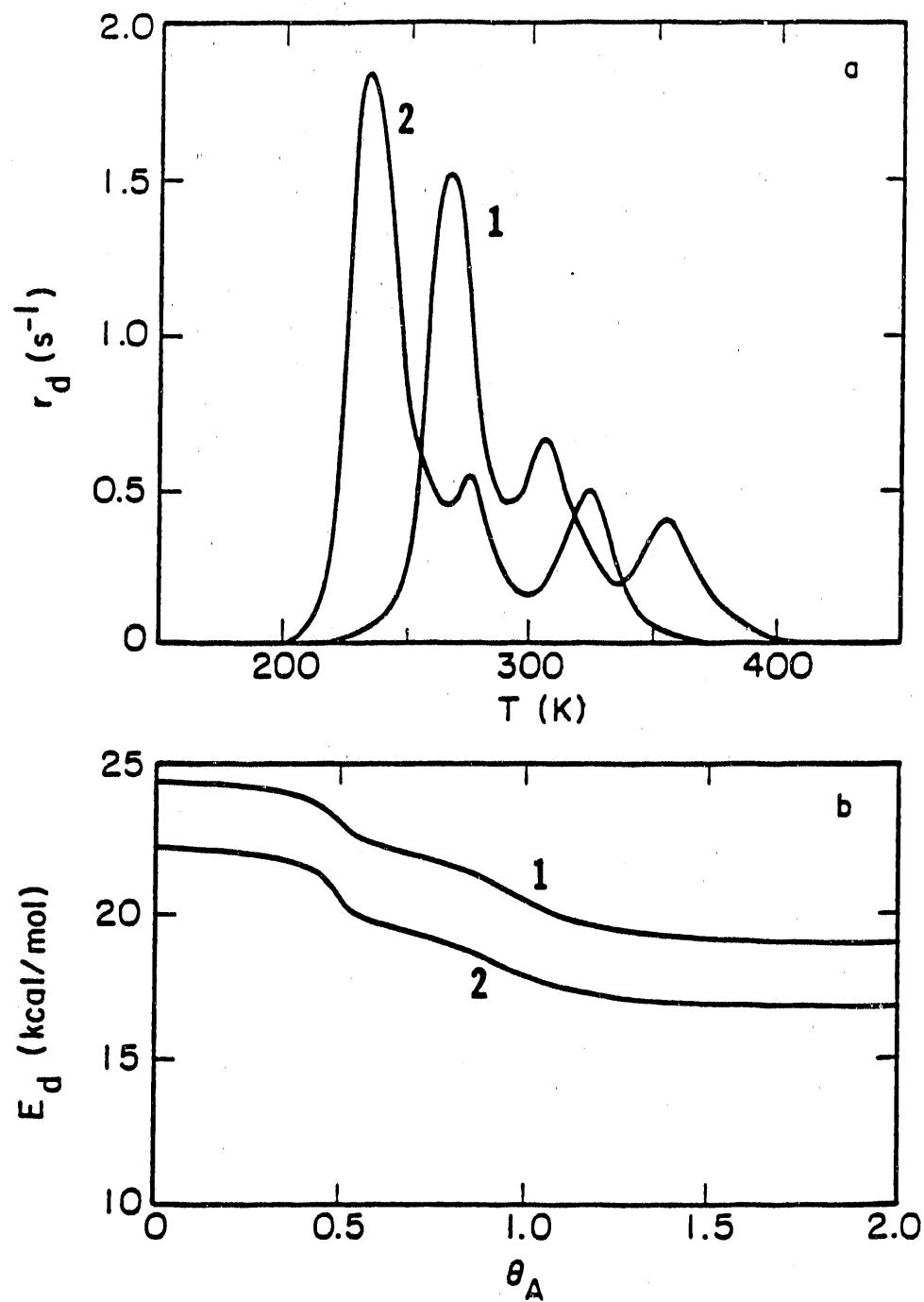


Fig. 3 a. Simulation of A_2 desorption with $\theta_B=0.25$. Curve 1 is for $Q_{B,n}=140$ kcal/mol and curve 2 is for $Q_{B,n}=190$ kcal/mol. See table 1 for model parameters. b. Variation of the average E_d with θ_A .

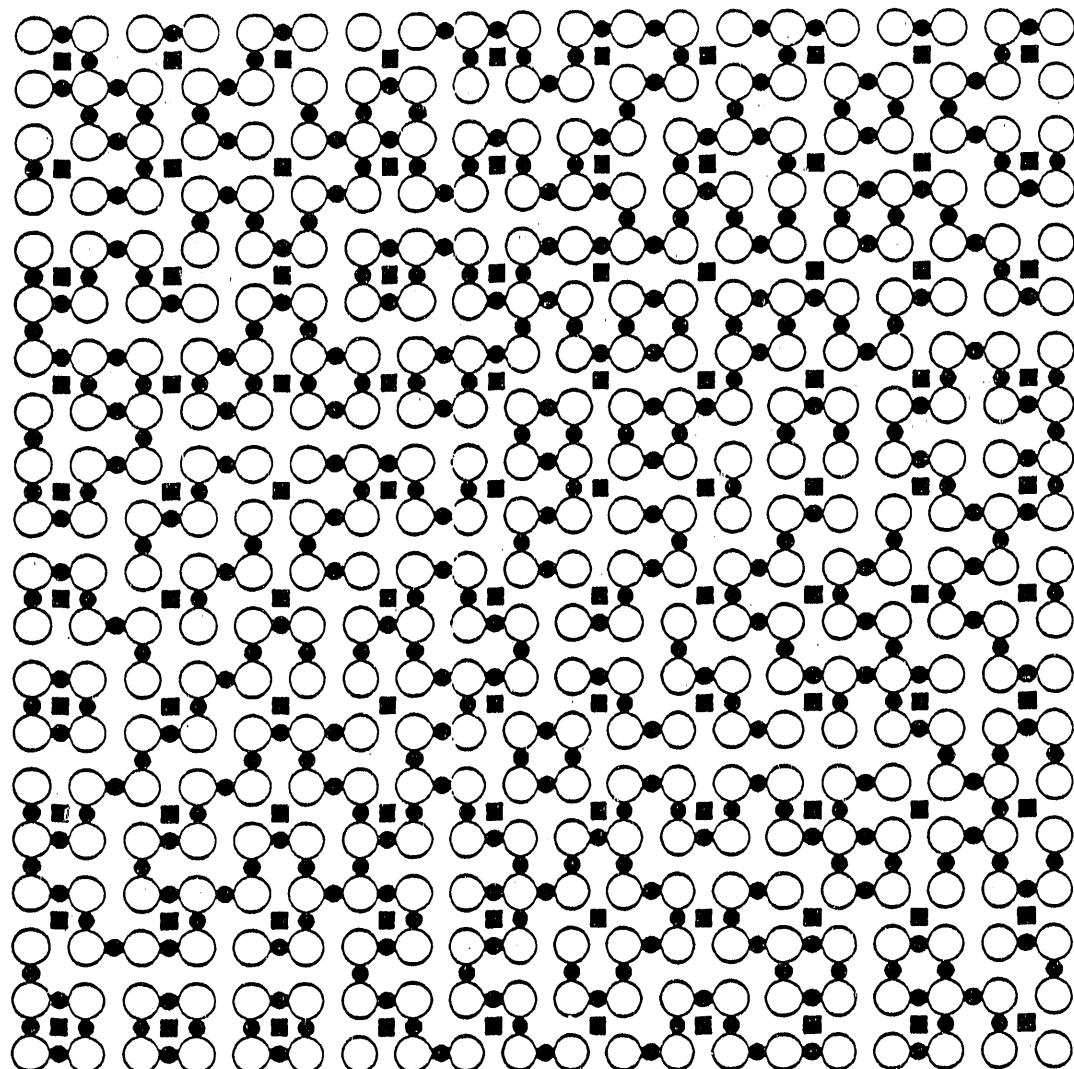


Fig. 4 Surface map for A(●) and B(■) at $\theta_A=1.0$ and $\theta_B=0.25$ for the simulations in fig. 3a. The open circles are metal atoms.

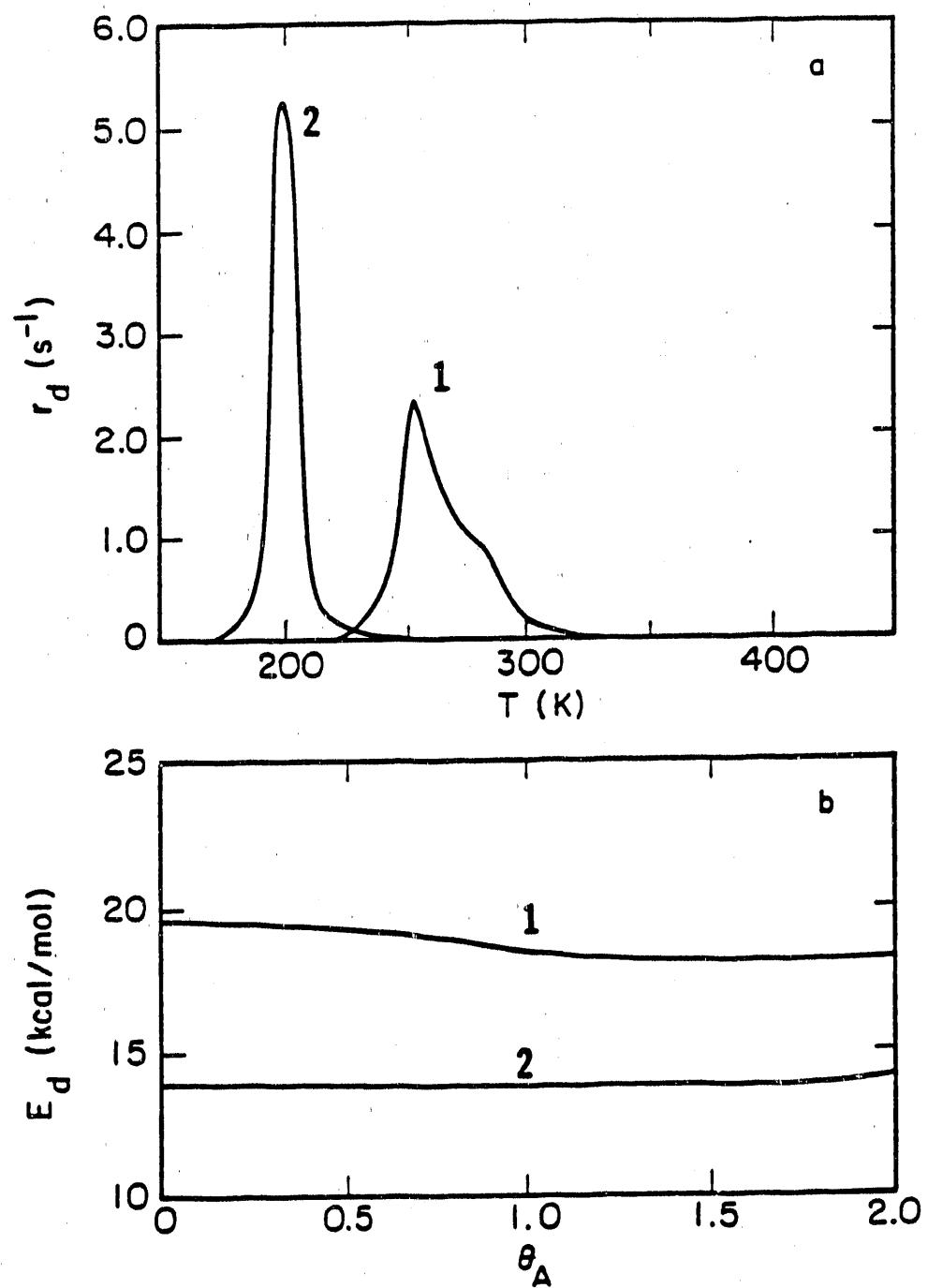


Fig. 5 a. Simulation of A_2 desorption with $\theta_B=0.50$. Curve 1 is for $Q_{B,n}=140$ kcal/mol and curve 2 is for $Q_{B,n}=190$ kcal/mol. See table 1 for model parameters. b. Variation of the average E_d with θ_A .

effects of attractive A-B interactions (see table 1). TPD spectra for A_2 desorption are shown in figs. 6a and 7a. For $\theta_B = 0.25$, the spectra shown in fig. 6a display three well resolved peaks. The peak positions are shifted relative to the peaks seen in fig. 3a for the case in which A-B interactions are neglected. For $\theta_B = 0.50$, the spectra shown in fig. 7a display very little difference in shape from those seen in fig. 5a but are shifted to higher temperatures. The activation energy profiles, displayed in figs. 6b and 7b, are similar in shape to the curves in figs. 3b and 5b but exhibit an approximately 2-3 kcal/mol increase in energy for a given coverage which is a consequence of the attractive A-B interactions. The surface map for the case of $\theta_A = 1.0$, $\theta_B = 0.25$, and $Q_{B,n} = 140$ kcal/mol is shown in fig. 8. Comparison of figs. 4 and 8 demonstrates that the effect of the A-B interactions is to produce a more ordered distribution of A atoms than in the absence of these interactions. As seen in fig. 8, most of the A atoms form a square array around each B atom.

The variations in the spectral features discussed above are all attributable to changes in the coverage and binding strength of species B. Increasing θ_B and increasing $Q_{B,n}$ both weaken the through-metal bonding of A which in turn lowers the activation energy barrier for the associative desorption of A_2 . This trend is reflected in all of the simulations presented above which show a progressively stronger down-scale shift in spectrum location with increasing coverage and increasing binding energy of species B. The presence of B atoms can also cause the loss of specific spectral features; for example, in the case of $\theta_B = 0.50$, $Q_{B,n} = 190$ kcal/mol, the three peaks for A_2 observed in the absence of B merge into one

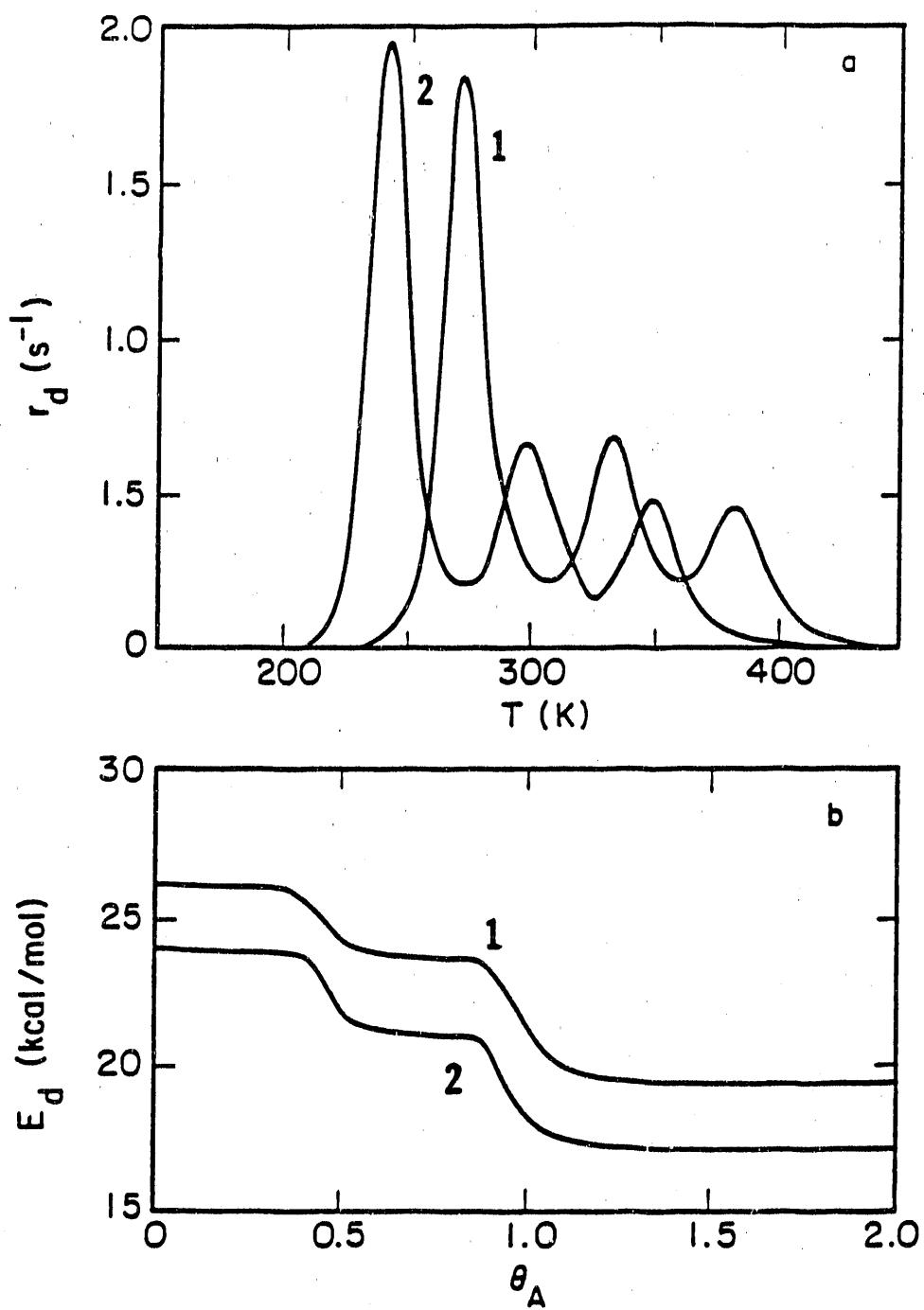


Fig. 6 a. Simulation of A_2 desorption with $\theta_B=0.25$ when attractive A-B interactions occur. Curve 1 is for $Q_{B,n}=140$ kcal/mol and curve 2 is for $Q_{B,n}=190$ kcal/mol. See table 1 for model parameters. b. Variation of the average E_d with θ_A .

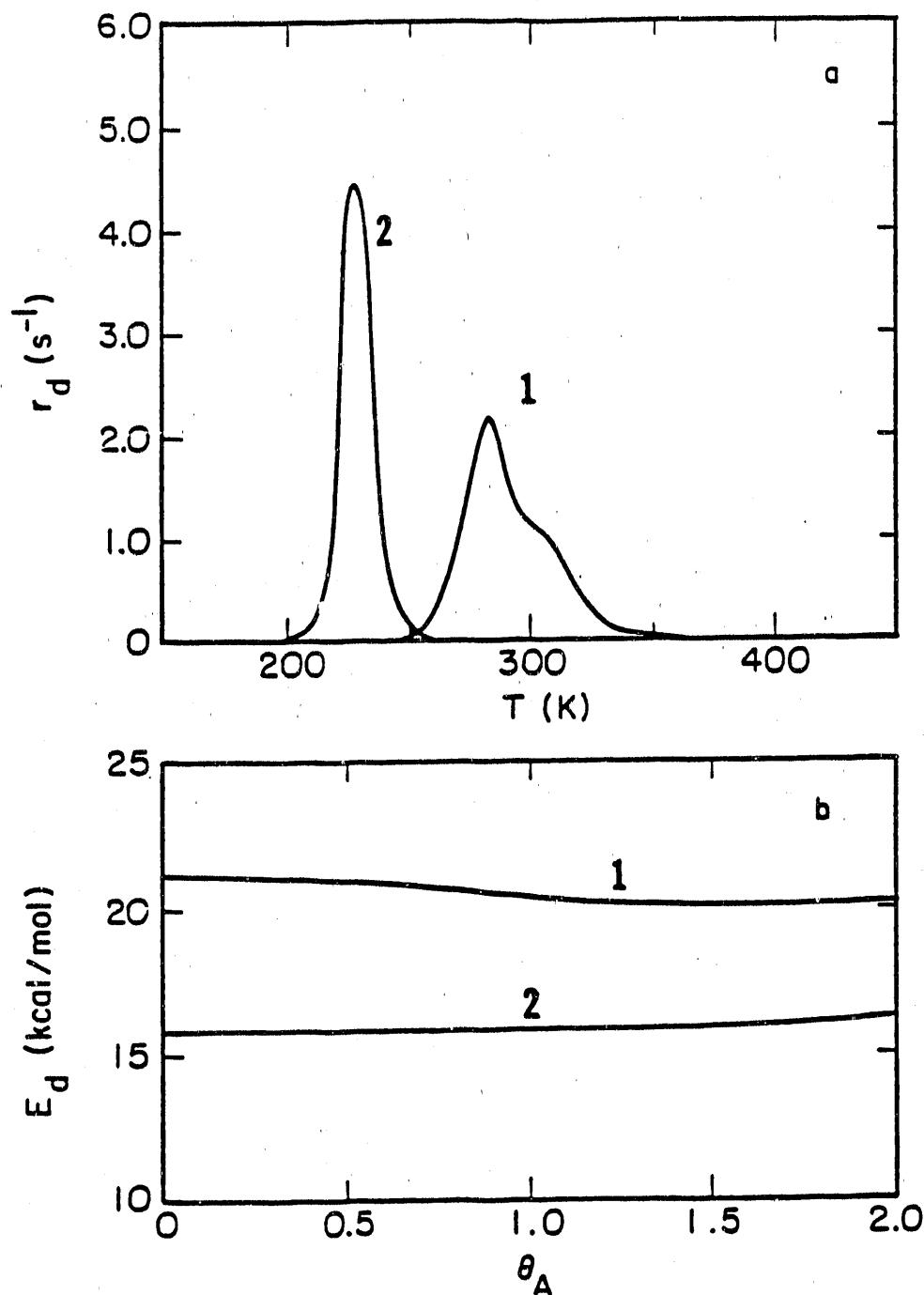


Fig. 7 a. Simulation of A_2 desorption with $\theta_B=0.50$ when attractive A-B interactions occur. Curve 1 is for $Q_{B,n}=140$ kcal/mol and curve 2 is for $Q_{B,n}=190$ kcal/mol. See table 1 for model parameters. b. Variation of the average E_d with θ_A .

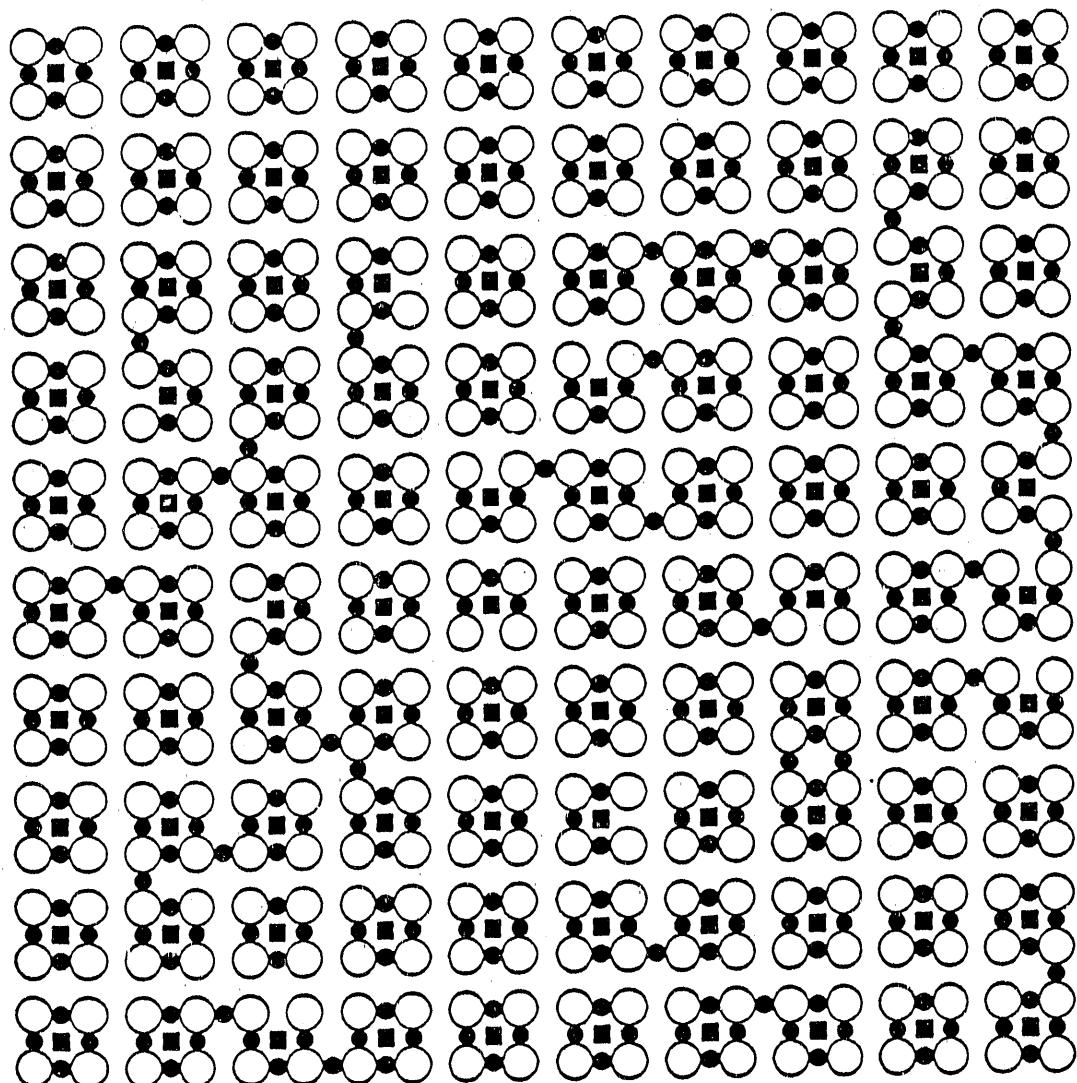


Fig. 8 Surface map for A(●) and B(■) at $\theta_A=1.0$ and $\theta_B=0.25$ for the simulation in fig. 6a. The open circles are metal atoms.

peak. When attractive A-B interactions are included, the binding energy increases and with it the activation energy for desorption. This increase, however, is insufficient to completely offset the weakening from the through-metal interactions and a trade-off in spectrum shape and location occurs. Attractive interactions are also seen to strongly affect the structure of the local adsorbate overlayer.

The heat of adsorption of A used in the above simulations is representative for atomic hydrogen adsorbed on early-transition metals, and the range of the heats of adsorption of B used in the simulations span the range of values for atomic nitrogen, carbon, and oxygen adsorbed on early-transition metals. The TPD spectra predicted with these values of the heats of adsorption as inputs to the model are qualitatively consistent with experimental observation. For example, hydrogen adsorbed on a clean Mo(100) surface exhibits a TPD spectrum similar to that presented in fig. 1a [34,35]. When hydrogen is coadsorbed in the presence of one-half monolayer of nitrogen, the integrated intensity of the hydrogen spectrum is attenuated, and all of the spectral features are shifted slightly to lower temperatures and are more poorly resolved [36]. A second illustration of the effects of coadsorbates is given in ref. [13] which reports on the adsorption of hydrogen on an Fe(100) surface in the presence of coadsorbed carbon and oxygen. In the absence of a coadsorbate, the maximum hydrogen coverage is 0.50 of a monolayer. The TPD spectrum for this case consists of a single TPD peak similar to the high-temperature peak observed in fig. 1a. Coadsorption of either carbon or oxygen results in an attenuation in

the hydrogen signal intensity and a shift of the TPD peak to lower temperature.

While the model presented here cannot reproduce the details of the experimental spectra, the qualitative features of such spectra are observed. We also note that the structure of eqs. 14 and 15 indicate that for strongly bound coadsorbates, a decrease in the adsorption capacity should occur with increasing strength of adsorption of the coadsorbate. This effect was not observed in the simulations presented here because the coadsorbate coverages and adsorption strength were insufficiently high. Finally, we note that the loss or gain of specific spectral features may not necessarily arise from blocking of a particular binding site, but rather, from the energetic interactions between coadsorbed species.

3.2 *Associative desorption and molecular desorption*

We consider next cases in which two mobile species, labeled for convenience as A and B, are coadsorbed on an fcc(100) surface. The energetics are chosen such that both species are capable of desorbing. Atomic species A associatively desorbs to form A_2 , and species B desorbs molecularly. A relevant example would be the coadsorption of atomic hydrogen and molecular carbon monoxide.

We first examine the desorption of pure A in which the A atoms were assigned to hollow sites at an initial coverage of unity, and both nearest- and next-nearest neighbors were considered as desorption partners. The heat of adsorption of an isolated A atom was chosen as 55 kcal/mol. Heats of adsorption as a function of coverage for individual A atoms were calculated taking into account

both M-A and A-A interactions. Table 2 lists the parameters used. As seen in fig. 9a, the TPD spectrum for A₂ exhibits a predominant peak at 290 K and a smaller peak at 355 K superimposed on a broad, high-temperature tail. The activation energy profile, displayed in fig. 9b, is constant at 26 kcal/mol up to a coverage of 0.25 and then decreases steadily until it reaches a value of 20 kcal/mol. A second TPD simulation was conducted for $Q_{A,n} = 52$ kcal/mol and is also shown in fig. 9a. It is similar in shape to the previous spectrum except that the high-temperature peak is less pronounced. In addition, the entire spectrum is shifted by 10 K to lower temperature. The activation energy profile for this case, shown in fig. 9b, is similar in shape to the profile when $Q_{A,n} = 55$ kcal/mol but is shifted slightly to lower energies as a consequence of the initially lower binding energy of the A atoms.

The desorption of molecular species B from the clean surface was considered for two cases of B coordination with the surface, namely, on-top and bridge bonding. The B species were allowed to interact with each other only through the metal atoms. TPD spectra calculated for both on-top and bridge bonding for initial coverages of one-half monolayer in a c(2×2) overlayer are shown in Fig. 9a. Each of these spectra exhibits one peak at 485 K. For bridge-bonded B, however, a low-temperature shoulder is also observed. The activation energy profiles, seen in fig. 9b, are initially flat for both species, except that as the coverage approaches 0.35, the activation energy profile for bridge-bonded B decreases with increasing coverage.

Three cases were considered for the desorption of A and B

Table 2 Model parameters for the TPD simulations of associative desorption and molecular desorption.a)

Species	Coordination	$Q_{i,n}$ (kcal/mol)	k_d^o (s^{-1}) ^{b)}	$\theta_{i,0}$ c)	Direct interactions	D_{xy} (kcal/mol) ^{d)}	Fig.
A	Hollow	55	1×10^{15}	1.0	A-A	109.5	9
	Hollow	52	1×10^{15}	1.0	A-A	109.5	9
B	Bridge	33	1×10^{15}	0.5 c(2x2)	-	-	9
	Top	33	1×10^{15}	0.5 c(2x2)	-	-	9
<i>Case I</i>							
A	Hollow	55	1×10^{15}	1.0	A-A	109.5	10
	Bridge	33	1×10^{15}	0.5 c(2x2)	-	-	10
<i>Case II</i>							
A	Hollow	52	1×10^{15}	1.0	A-A	109.5	12
	Bridge	33	1×10^{15}	0.5 c(2x2)	-	-	12
<i>Case III</i>							
A	Hollow	55	1×10^{15}	1.0	A-A	109.5	13
	Top	33	1×10^{15}	0.5 c(2x2)	-	-	13

a) Simulations were conducted for a fcc(100) surface with a heating rate $\beta=25$ K/s. Each simulation was the average of 1 run with $\Delta t = 0.12$.

b) For second order desorption, k_d^o represents the product of the second order preexponential factor times the concentration of adsorption sites.

c) Initial coverage and structure of adsorbate overlay.

d) The value of D_{xy} is for the respective direct interactions listed in the preceding column.

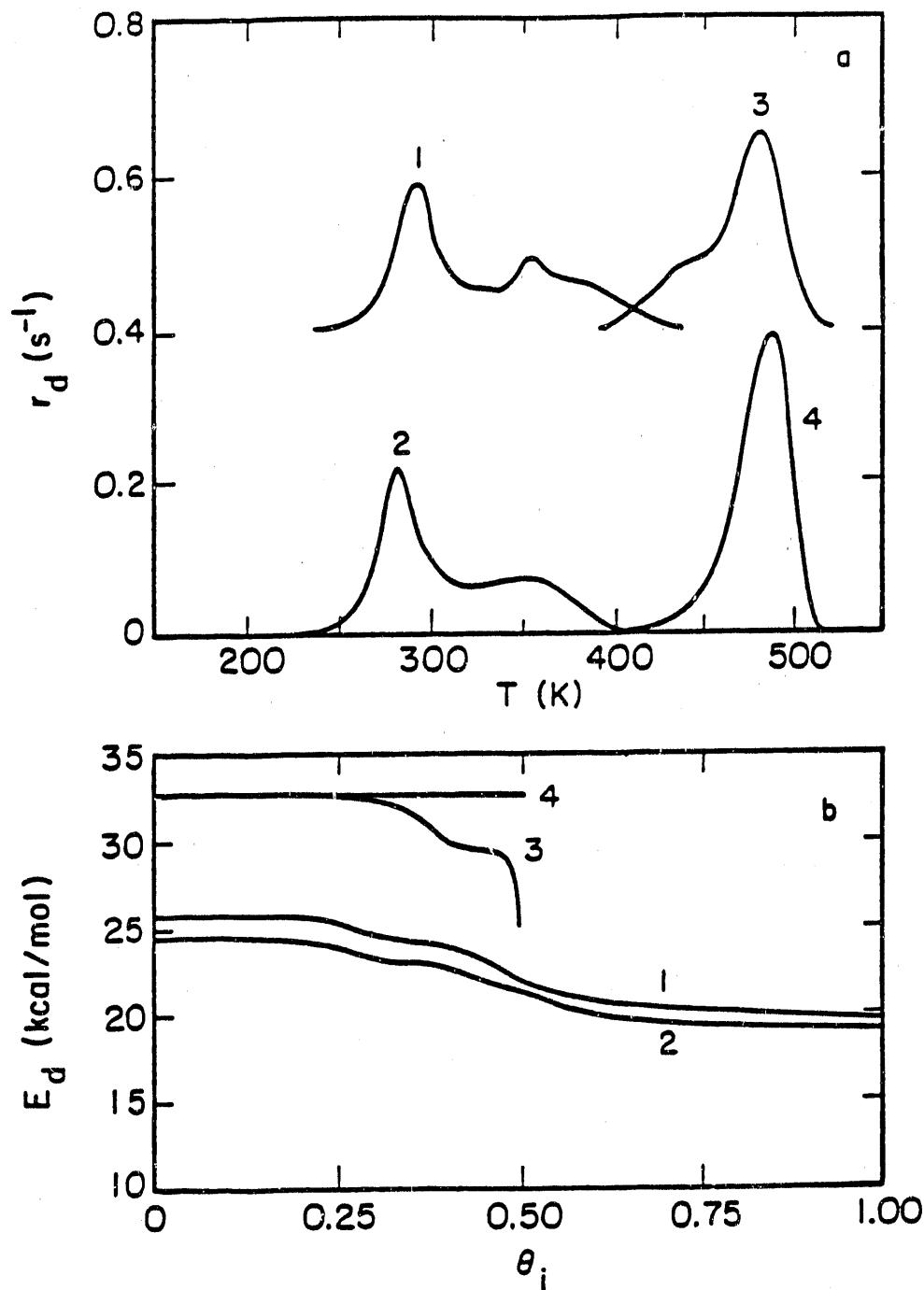


Fig. 9 a. Simulation of pure component desorption for A₂ and B. Curve 1 is for $Q_{A,n}=55$ kcal/mol and curve 2 is for $Q_{A,n}=52$ kcal/mol. Curve 3 is for bridge-bonded species B with $Q_{B,n}=33$ kcal/mol. Curve 4 is for on-top bonded species B with $Q_{B,n}=33$ kcal/mol. See table 2 for model parameters. b. Variation of the average E_d with coverage.

when both were present on the surface. In case I, species A occupied hollow positions, and species B occupied bridge positions. The initial coverage of A was unity and that of B was one-half in a $c(2\times 2)$ overlayer. The binding energies of the isolated species A and B were 55 and 33 kcal/mol, respectively. The binding energy of species A as a function of its local environment was calculated taking into account A-A, M-A, and M-B interactions, and the binding energy of species B as a function of its local environment was calculated by taking account M-B and M-A interactions. For both species, no direct A-B interactions were included.

The TPD spectra for case I are shown in fig. 10a. The spectra for both A and B exhibit new features which were not present in the spectra when each species was adsorbed separately (see fig. 9a). The spectrum of species A exhibits a new peak at 240 K in addition to a high-temperature peak at 300 K accompanied by a high-temperature tail. The peak at 300 K and the high-temperature tail occur in the same temperature region as the features observed during the desorption of pure A but with different intensities. The spectrum for B exhibits a low-temperature peak at 230 K, a small peak at 355 K, and a more intense peak at 480 K. The peaks at 230 K and 355 K are new, while the peak at 480 K appears in the same region as the spectrum for pure B.

The activation energy profiles, shown in fig. 10b, differ significantly from the profiles found in the case of single component desorption (see fig. 9b). The profile for species A is shifted to lower energies by about 1-4 kcal/mol for most of the coverage range between 0.10 to 1.0. The curve for species B shows very strong

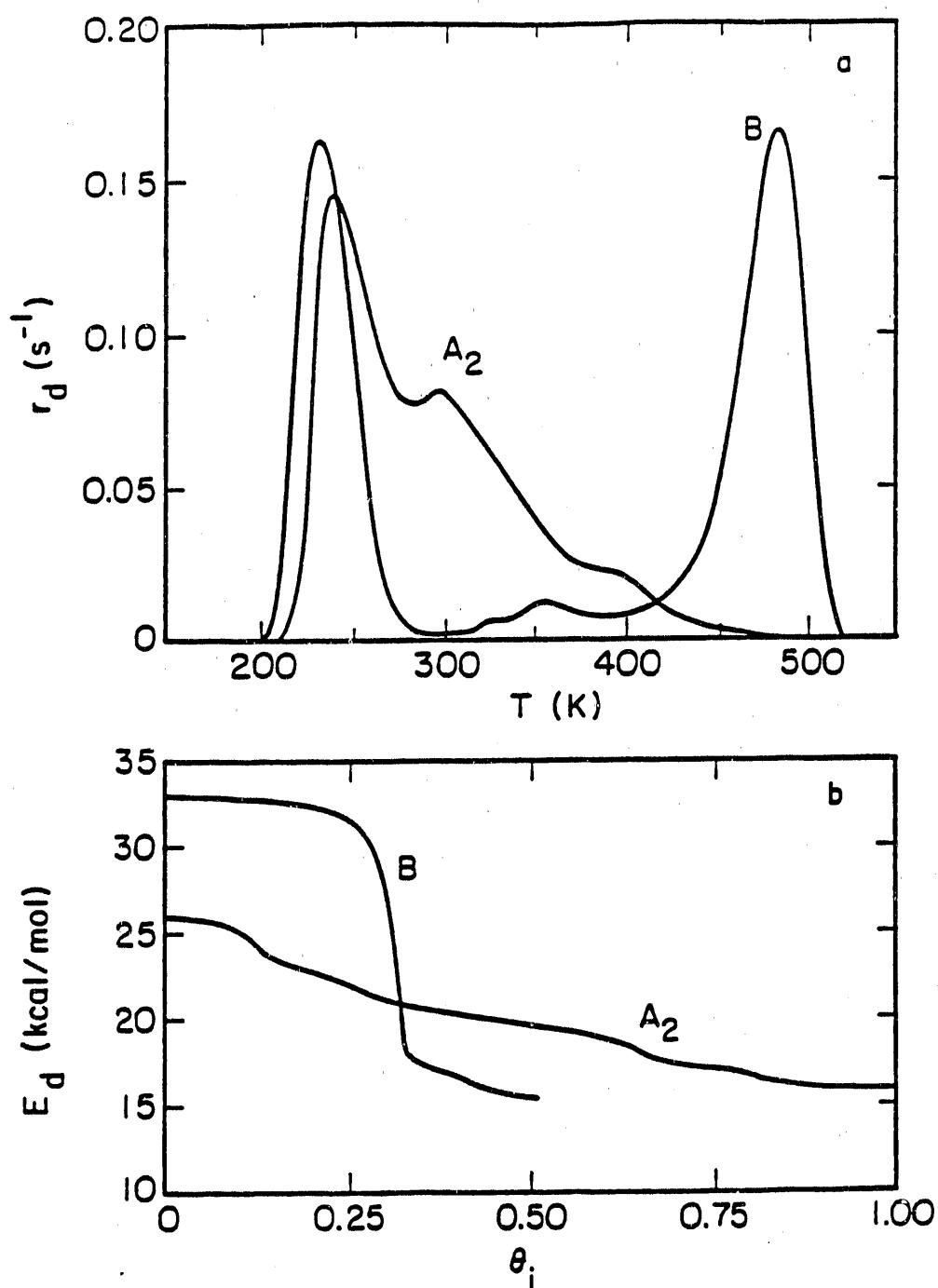


Fig. 10 a. Simulation of A₂ and B codesorption for case I. See table 2 for model parameters. b. Variation of the average E_d with coverage.

coverage-dependent behavior, exhibiting a sharp decrease at $\theta_B = 0.30$. It is significant to note that the activation energy for the desorption of B at the start of the TPD (i.e. at $\theta_B = 0.50$) is lower than the corresponding value for the desorption of A₂ ($\theta_A = 1.0$). A surface map for $\theta_A = 0.51$ and $\theta_B = 0.32$, shown in fig. 11, displays segregation of the the two species into regions rich in A and B. Furthermore, species B, which was initially ordered in a c(2×2) overlayer, is now seen to be densely packed in the B-rich regions.

In case II, the binding energy of an isolated species A was decreased from 55 to 52 kcal/mol. The binding energy and coordination of B remained the same as for case I. The calculated TPD spectra are shown in fig. 12a. Species A desorbs predominantly at low temperature with a peak at 230 K and a broad, high-temperature tail. Species B exhibits peaks at 370 and 480 K, and no peak at 230 K. The activation energy profiles as a function of coverage are shown in fig. 12b. The curve for A, as compared with the curve for A in case I in fig. 10b, is lower for a given coverage and decreases more rapidly to its value at high coverage. The profile for species B, however, is much closer in shape and location to that observed for the desorption of pure B (see fig. 9b). It is interesting to note that reducing the value of $Q_{A,n}$ from 55 to 52 kcal/mol almost totally eliminates the influence of adsorbed A on the desorption of B. The large change in the appearance of the desorption spectra for a relatively small change in the initial binding energy of species A demonstrates the sensitivity of desorption kinetics to the relative binding energies of the coadsorbed species.

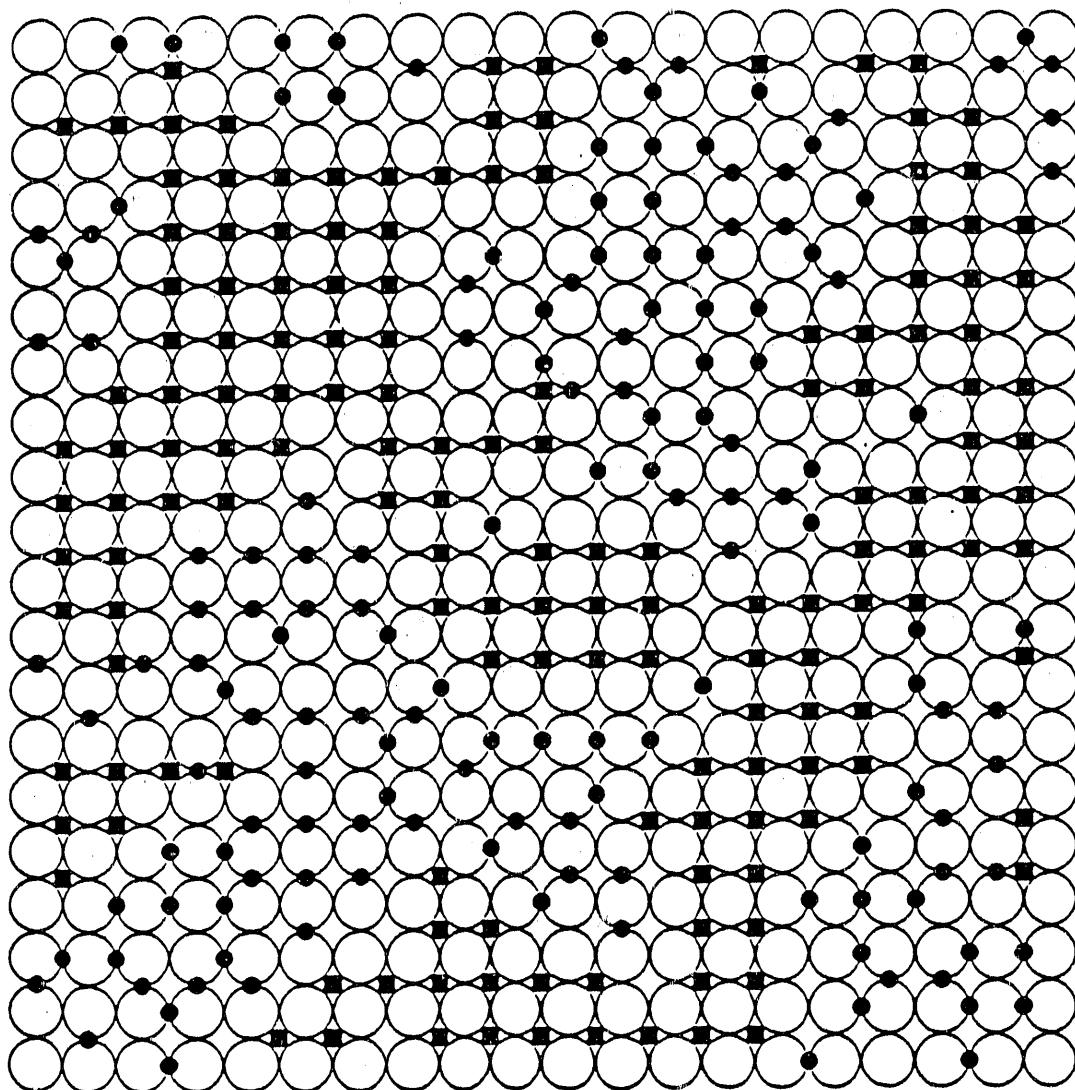


Fig. 11 Surface map for A(■) and B(●) at $\theta_A=0.51$ and $\theta_B=0.32$ for the simulation in fig. 10a. Open circles are metal atoms.

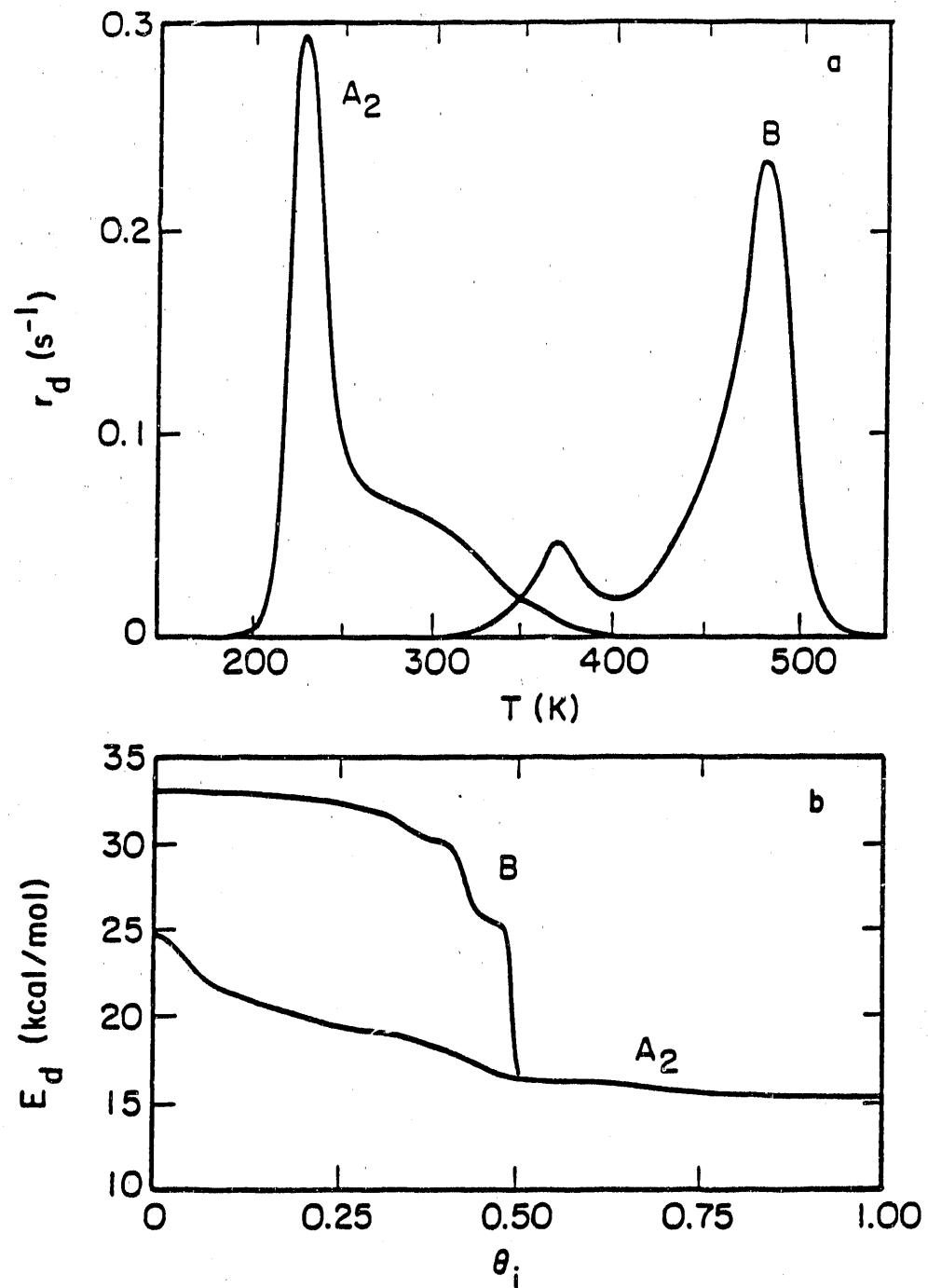


Fig. 12 a. Simulation of A_2 and B codesorption for case II. See table 2 for model parameters. b. Variation of the average E_d with coverage.

In case III, the binding energy for species A was the same as that in case I (see Table 2). Species B was positioned on on-top sites in a $c(2\times 2)$ overlayer at an initial coverage of 0.5. The TPD spectra for this set of conditions are shown in fig. 13a. Species A desorbs with a peak at 225 K and exhibits a high-temperature tail. Compared to the spectrum for A_2 desorption in the absence of B (see fig. 9a), the peak of the principal feature for A_2 is shifted to lower temperatures by 65 K. The spectrum for species B is relatively unchanged from that seen in fig. 9a. The activation energy profiles are shown in fig. 13b. The profile for species A is lower than that calculated for the desorption of pure A. The activation energy profile for species B, however, is quite similar to that for the desorption of pure B. This is not surprising, since the desorption of B occurs only after the surface has been substantially depleted of A. The surface map for case III is shown in fig. 14. As in fig. 11, segregation of A and B is observed. Since species B can only be affected by species A via through-metal interactions, only those B species located at the periphery of B-rich regions are affected energetically by the presence of A.

To summarize, the cases considered above show two general types of behavior: in case I, the spectra of both adsorbates are strongly influenced by interactions between the coadsorbed species and new features are observed; in cases II and III, associative desorption of the atomic adsorbates is influenced by the molecular species, but the molecular species are only mildly affected by the atomic adsorbates. It is also noted that the appearance or loss of spectral features is determined by the nature and strength of the

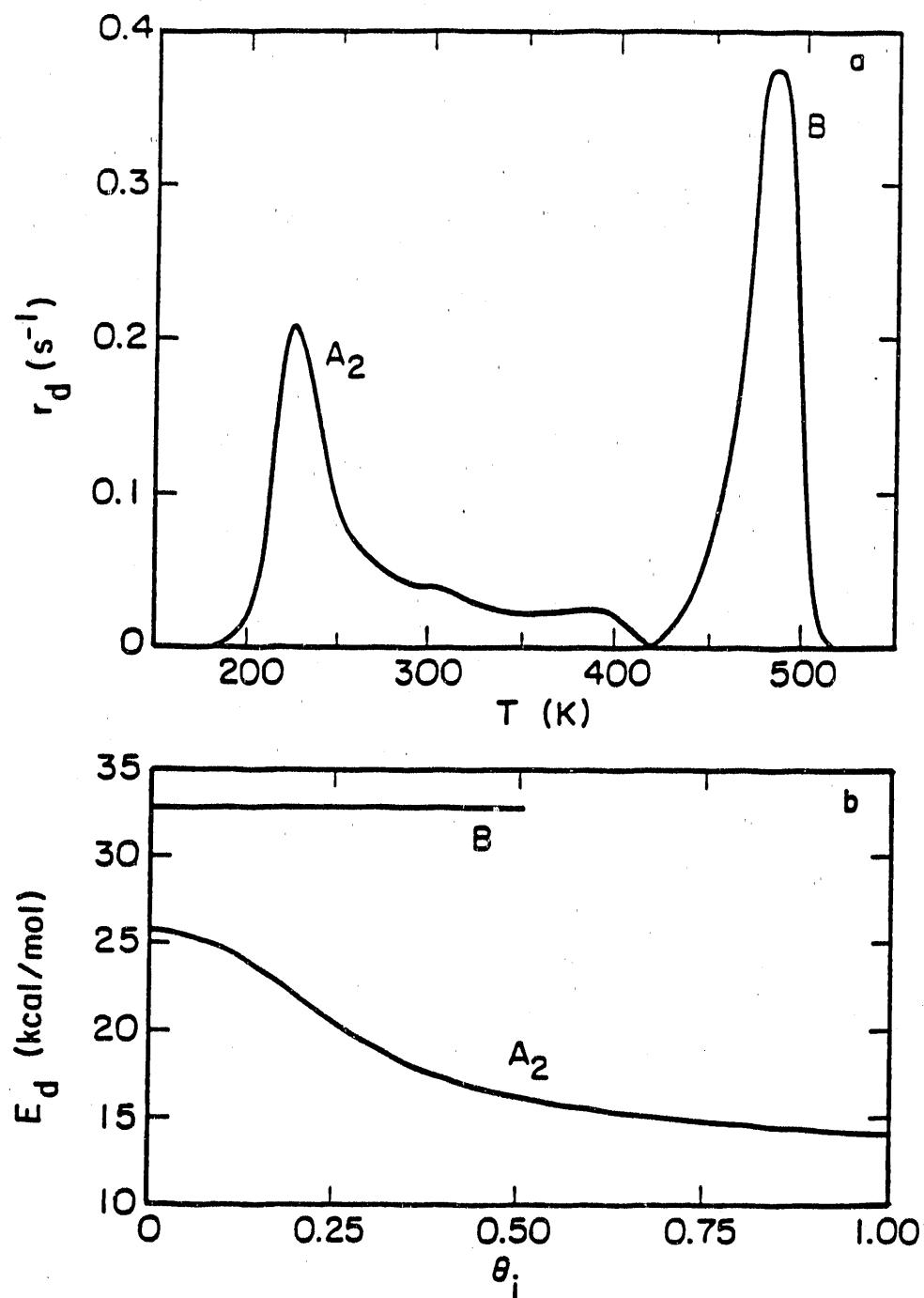


Fig. 13 a. Simulation of A₂ and B codesorption for case III. See table 2 for model parameters. b. Variation of the average E_d with coverage.

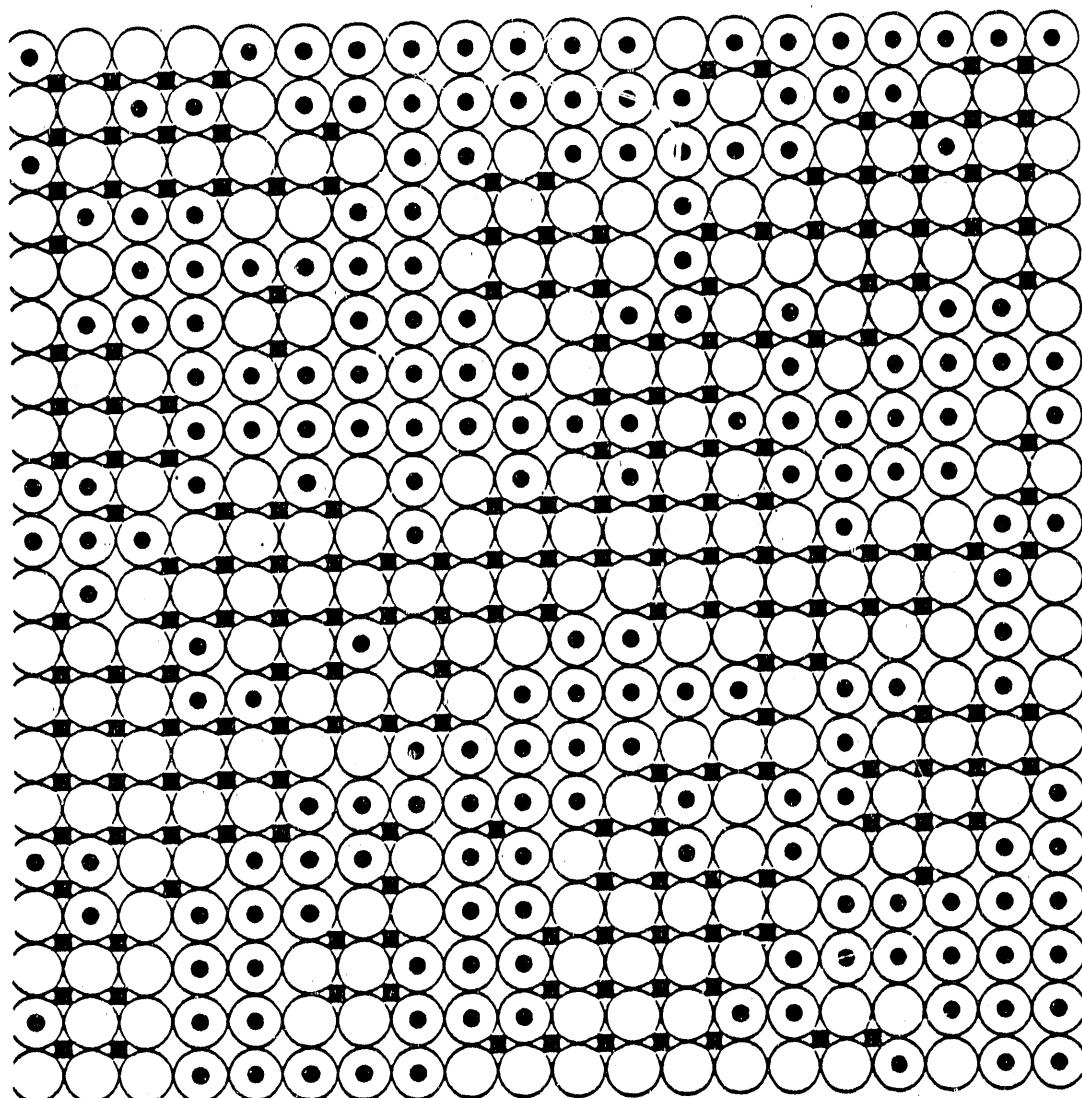


Fig. 14 Surface map for A(■) and B(●) at $\theta_A=0.52$ and $\theta_B=0.50$ for the simulation in fig. 13a. The open circles are metal atoms.

interactions between species and not by the loss or creation of new binding states.

The heats of adsorption for A and B used in the above simulations are representative for atomic hydrogen and molecular carbon monoxide, respectively, adsorbed on Group VIII transition metals. The TPD spectra predicted with these values of the heats of adsorption as inputs to the model are qualitatively consistent with experimental observation. For example, the spectra for hydrogen and carbon monoxide adsorbed alone on a Ni(100) face each exhibit a single peak [16-18], as shown in fig. 15. Figure 15 also shows that when carbon monoxide and hydrogen are coadsorbed, low-temperature peaks appear for both species. Comparison of figs. 10a and 15 shows that the spectra simulated for the codesorption of A and B under the conditions of case I resemble qualitatively the spectra observed experimentally. This suggests that the simultaneous evolution of two species at temperatures far below those required to desorb the pure components may simply be a manifestation of the interaction between the coadsorbed species. It is also interesting to note that the segregation of the adsorbates A and B is driven by the attractive interaction between A atoms. When this interaction is strong enough to counter the repulsive effects arising from multiple A atoms being bonded to individual metal atoms, segregation into regions rich in A is energetically favorable. As a consequence of this driving force for A segregation, the B species populate the regions deficient in A. In this connection, it is noted that in the studies of the codesorption of hydrogen and carbon monoxide from Ni(100), it was hypothesized that at low

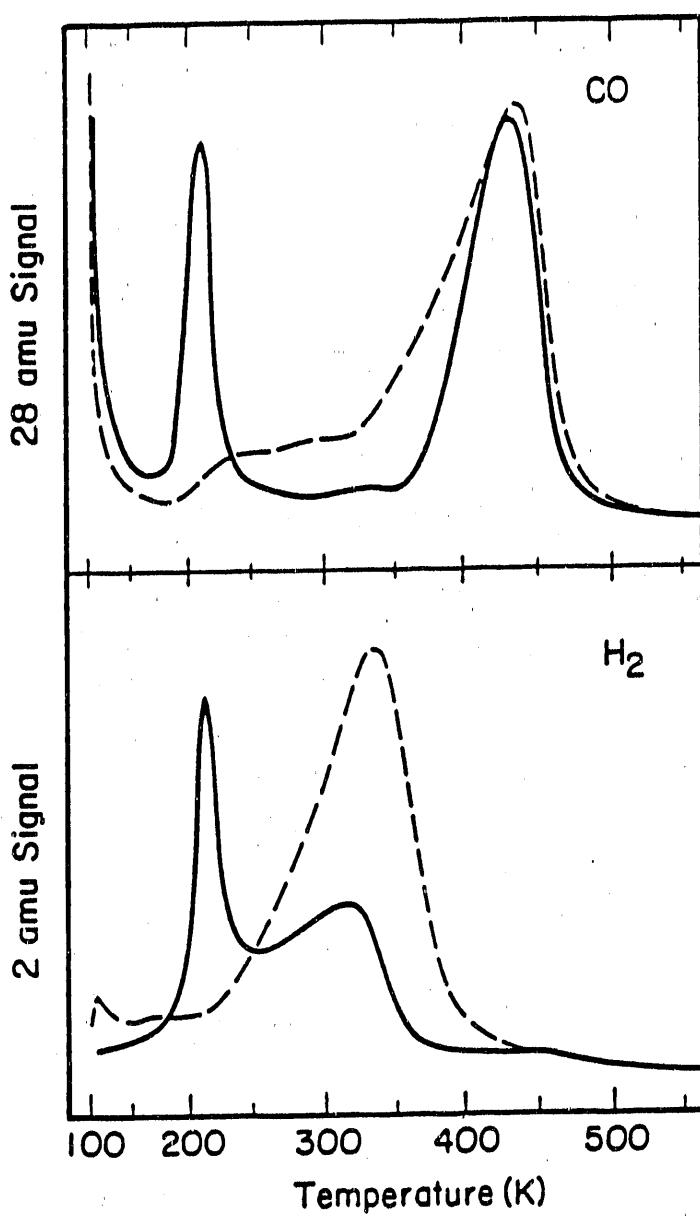


Fig. 15 Experimentally observed TPD spectra for H_2 and CO desorption from Ni(100) [18]. The dashed lines are for desorption of each species adsorbed separately, and the solid lines are for desorption of both species coadsorbed.

temperature, where $\theta_H = 1.0$ and $\theta_{CO} = 0.5$, adsorbed H and CO are intimately mixed on the surface, but as the temperature is raised and sites are vacated, the hydrogen and carbon monoxide segregate, giving rise to desorption features similar to those observed when each species is adsorbed separately [17,18].

On Rh(100), the spectra for carbon monoxide and deuterium adsorbed separately exhibit one and two peaks, respectively [19]. When coadsorbed, however, the carbon monoxide is only weakly affected by the presence of the deuterium, for a variety of deuterium and carbon monoxide coverages. The deuterium spectra in the presence of coadsorbed carbon monoxide, on the other hand, are shifted to lower temperatures, and new features appear. Cases II and III presented above display this kind of behavior.

4.0 CONCLUSIONS

The present study demonstrates that the BOC-MP method can be used successfully to represent the energetics of coadsorption of two species on well-defined metal surfaces. Monte Carlo simulations of the temperature-programmed desorption of coadsorbed species demonstrates that the number and location of the desorption peaks and the activation energy versus coverage profiles are sensitive to the relative coverages and binding energies for each species. In addition, the combined effects of the through-metal interactions and the direct adsorbate-adsorbate interactions are seen to have important consequences on the activation energies for desorption and on the distribution of adsorbates on the surface. The appearance, disappearance, and shifting of spectral features are

shown to arise from the nature of the energetic interactions between the coadsorbed species with the metal and with each other, and not from the appearance, disappearance, or modification of distinct binding sites.

The presence of strongly bound, immobile coadsorbates was found to lower the activation energy for the associative desorption of atomically adsorbed species. The magnitude of this effect increases with increasing coverage and binding strength of the coadsorbate. When molecular and atomic adsorbates of comparable activation energies for desorption are coadsorbed, either both adsorbates are strongly affected or only the atomic species is strongly affected. This dependence is sensitive to the coordination and relative binding energies of the two species. For certain circumstances, the activation energy for the desorption of the molecular adsorbate may become smaller than that for associative desorption of the atomic species. It is also found that when both adsorbates are mobile, surface segregation of the two coadsorbed species can occur, such segregation being driven by the attractive interactions between the species undergoing associative desorption.

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NOMENCLATURE

A,B,AB Adsorbates

a Morse potential constant (\AA)
 D_{XY} Bond dissociation energy for X-Y bond (kcal/mol)
 $E_{d,i}$ Activation energy for desorption from the i th environment (kcal/mol)
 $E_{d,A}$ Activation energy for the non-associative desorption of A (kcal/mol)
 $E_{d,AA}$ Activation energy for associative desorption of A and A' (kcal/mol)
 $k_{d,i}^*$ Preexponential factor for desorption from the i th environment (s^{-1})
 m_i Number of adsorbates bonded to the i th metal atom
 N_A Number of adsorbates A
 N_d Number of adsorbate atoms or molecules desorbing (mol)
 N_s Number of surface metal atoms (mol)
 n Number of metal atoms bonded to an adsorbate
 n_A Number of metal atoms bonded to A
 $P_{i(1)}$ Probability of desorption from the i th environment
 $P_{ij(2)}$ Probability of diffusion from site i to site j
 $Q(x)$ Bond energy (kcal/mol)
 Q_b Equilibrium bond energy (kcal/mol)
 Q_{oA} Heat of adsorption of A in the on-top position (kcal/mol)
 Q_i, Q_j Heats of adsorption in sites i and j (kcal/mol)
 $Q_{A,n}^*$ Total heat of adsorption for A (kcal/mol)
 $Q_{A,n}$ Heat of adsorption of A as a function of coordination (kcal/mol)
 $Q_{A,n(1)}$ Heat of adsorption of A due to through-metal interactions (kcal/mol)
 $Q_{A,n(2)}$ Heat of adsorption of A due to adsorbate-adsorbate interactions (kcal/mol)
 R Gas constant (kcal/mol K) or random number
 r_d Turnover frequency (s^{-1})
 r Bond length (\AA)
 r_0 Equilibrium bond length (\AA)
 T Temperature (K)
 Δt Time interval (s)
 x Bond order

$x_{i,n}$ Bond order of an adsorbate bonded to n metal atoms
 x_A Bond order of a metal atom bonded to an adsorbate A
 α Desorption order
 β Heating rate (K/s)
 δ Direct adsorbate-adsorbate bond order
 θ Coverage
 $\theta_{A,0}$ Initial coverage of A

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Chapter V

**Monte Carlo Simulations of the Effect of Pressure on
Isothermal and Temperature-Programmed Desorption
Kinetics****ABSTRACT**

A Monte Carlo simulation technique is presented for describing the adsorption, surface diffusion, and desorption kinetics of molecules from metal surfaces. Lateral interactions between adsorbed molecules are taken into account using the Bond-Order-Conservation-Morse-Potential method. The rate of desorption observed in the presence of a gas-phase species is higher than that observed in a vacuum. The increase in the apparent rate coefficient for desorption with increasing pressure can be ascribed to the effects of repulsive lateral interactions on the activation energy for desorption. The simulated kinetics are in good agreement with the experimentally-observed kinetics for the isothermal desorption of CO from polycrystalline Pd and for the temperature-programmed desorption of CO from Ni(100).

1.0 INTRODUCTION

A number of studies [1-13] have shown that the rate of gas desorption from a metal surface is enhanced when desorption occurs in the presence of a finite pressure of the desorbing gas. Based on the results of isothermal studies of CO desorption from polycrystalline Pd, Yamada et al. [5,6] have proposed that the rate coefficient of CO desorption depends explicitly on the pressure of CO. This behavior was attributed to an unspecified chemical interaction in the adsorbed layer. The effects of adsorbate fluxes have also been observed in temperature-programmed desorption studies. Yates and Goodman [4] have noted that when CO desorbs from Ni(100) in the presence of gas-phase CO, the positions of the desorption peaks observed in vacuum shift to lower temperatures and new, low-temperature features appear. The authors ascribed these effects to the influence of lateral interactions on the activation energy for desorption.

The possibility of collision-induced desorption has also been considered. Yamada et al. [5] found that when CO desorption was carried out in a background of Ar the rate of desorption was identical to that observed in vacuum. Similar results were obtained by Yates and Goodman [4], who found that the presence of N₂ had no effect on the TPD spectrum of CO. More recently, Ceyer and coworkers [14-17] have observed that the rate of CO desorption from a Ni(111) surface can be enhanced by a flux of mono-energetic Ar atoms but only when the translational kinetic energy of the Ar atoms exceeds 60 kcal/mol. Taken together, these observations

suggest that for gas temperatures normally encountered in catalysis (300-1300 K), collision-induced desorption is not significant.

Here, we present a theoretical analysis of the influence of molecular adsorption on the desorption of adsorbed species. Our approach is to use a Monte Carlo simulation technique which accounts for the elementary surface processes of adsorption, surface diffusion, and desorption. Also included in the model is the influence of lateral interactions between adsorbed species. We compare the results of the simulations with the experimental data for the isothermal desorption of CO from polycrystalline Pd [5,6] and the temperature-programmed desorption of CO from Ni(100) [4]. The model calculations demonstrate that the apparent pressure-dependence of the desorption kinetics is a direct consequence of lateral interactions between adsorbed molecules.

2.0 THEORY

Adsorbate molecules are assumed to participate in three elementary processes: adsorption, diffusion, and desorption. While on the surface of the metal, molecules are assumed to occupy well-defined sites on a single-crystal lattice. The coordination of the adsorbate is defined by the composition of the metal and the adsorbate. Since the local occupancy of adsorption sites varies with position on the surface and adsorbate-adsorbate interactions are known to affect the heat of adsorption, the dynamics of adsorption and desorption are taken to be site-specific. In view of the inhomogeneous distribution of adsorbates, Monte Carlo techniques

are used to represent the effects of adsorption, diffusion, and desorption on the net rate of adsorption or desorption of adsorbates.

The probability of adsorption on a vacant site i in the time interval Δt can be defined as (see Appendix for the derivation of eq. 1):

$$P_i^a = S_0 P a_s / (2\pi m k_b T_g)^{1/2} \Delta t = S_0 F \Delta t \quad (1)$$

where S_0 is the sticking coefficient at zero coverage, P is the pressure, a_s is the area per site, m is the mass of an adsorbate, k_b is Boltzmann's constant, and T_g is the temperature of the gas phase.

The factor F on the far right-hand side of eq. 1 is the flux of adsorbates per site. The total rate of adsorption, expressed as a turnover frequency based on the number of surface metal atoms, N_s , is given by

$$r_a = \sum_i N_{a,i} / (N_s \Delta t) \\ = N_a / (N_s \Delta t) \quad (2)$$

The probability of desorption from site i in the time interval Δt can be represented as

$$P_i^d = v_i \exp[-E_{d,i} / k_b T] \Delta t \quad (3)$$

where v_i is the frequency factor for desorption and $E_{d,i}$ is the activation energy for desorption. The total rate of desorption from all local environments is simply

$$r_d = \sum_i N_{d,i} / (N_s \Delta t) \\ = N_d / (N_s \Delta t) \quad (4)$$

Since the activation energy for surface diffusion is 10-15% of

that for desorption, whereas the frequency factors for the two processes are comparable, surface diffusion is expected to occur much more rapidly than desorption. As a consequence, it is assumed that the spatial distribution of adsorbate molecules will never be far from that corresponding to equilibrium. Perturbations to the equilibrium distribution caused by the adsorption or desorption of molecules are eliminated by repositioning the adsorbed molecules in accordance with Kawasaki statistics. The probability that an adsorbate on site i moves to an adjacent site j is given by [18]:

$$P_{ij} = \frac{\exp [-(Q_i - Q_j) / k_b T]}{1 + \exp [-(Q_i - Q_j) / k_b T]} \quad (5)$$

where Q_i and Q_j are the heats of adsorption for an adsorbate at sites i and j , respectively. Equation 5 weights diffusional jumps according to the magnitude of $(Q_i - Q_j)$. It should be noted that if $Q_i = Q_j$, then $P_{ij}(2) = 0.5$, and the two sites have equal probability of occupancy.

To complete the descriptions of the probabilities of desorption and diffusion given by eqs. 3 and 5, the values of Q_i and $E_{d,i}$ appearing therein must be determined. Our approach is to use the Bond-Order-Conservation-Morse-Potential (BOC-MP) method developed by Shustorovich [19,20] because it allows us to account for the effects of local site occupancy without the introduction of arbitrarily assigned energy parameters. Since the BOC-MP method has been discussed extensively elsewhere [19-22], only those relationships required for this study are summarized here.

The heat of adsorption for a molecular adsorbate AB coordinated via atom A to n metal atoms can be approximated by

$$Q_{AB,n} = \frac{Q_{oA}^2}{D_{AB} + Q_{oA}/n} \quad (6)$$

where D_{AB} is the gas-phase dissociation energy for the A-B bond and Q_{oA} is the heat of adsorption of A in the on-top position.

Equation 6 is appropriate for isolated adsorbate molecules. When the adsorbate coverage increases, situations arise in which more than one adsorbate is coordinated with a single metal atom. In this case, the heat of adsorption is given by

$$Q_{AB,n} = \sum_{i=1}^n \frac{Q_{AB,n}}{nm_i} (2-1/m_i) \quad (7)$$

where m_i is the number of adsorbates bonded to the i th metal atom.

The expression for $Q_{AB,n}^*$ can be used to determine the activation energy for desorption. For the non-associative desorption of a molecule, the activation energy is given by the heat of adsorption, $E_d = Q_{AB,n}^*$. Implicit in this relationship is that adsorption is non-activated.

The energy and probability formulations described above were incorporated into a Monte Carlo algorithm for simulating isothermal and temperature-programmed kinetic experiments. The metal surface was represented by a 100 by 100 array of numbered sites, and periodic boundary conditions were used to eliminate edge effects. When an initial coverage was required to begin a simulation, molecules were placed randomly on the lattice to achieve an initial coverage, θ^* .

Each simulation was divided into a sequence of Monte Carlo steps (MCS) and each MCS was subdivided into three segments

corresponding to the adsorption, diffusion, and desorption of molecules. The time interval for each MCS, Δt , was chosen sufficiently small so that the change in surface coverage during each segment of an MCS was less than ten percent of the saturation coverage. Specifying Δt in this manner is necessary for eqs. 1 and 3 to be valid representations of the probabilities of adsorption and desorption, respectively, (see Appendix).

For the adsorption segment of the first MCS, the temperature was initialized at T_0 and taken to be constant at this value for the time interval Δt . During this interval, molecules were adsorbed on the surface in the following manner. A surface site was selected in a random fashion. If the site were unoccupied, the probability of adsorption was calculated using eq. 1. The calculated adsorption probability was then compared with a random number, R ($0 < R < 1$). If $R < P_{1a}$, the adsorbate was placed on the lattice, and N_a in eq. 2 was incremented by one. Conversely, if $R \geq P_{1a}$, the site remained unoccupied. The number of unoccupied sites visited during an MCS was $[S(\theta)/S_0]N_s$, where $S(\theta)$ is the sticking coefficient. The form of $S(\theta)$ used for specific simulations is discussed in the next section. After completing the sampling of the surface, the rate of adsorption for this segment of the MCS was calculated using eq. 2 and the surface coverage was updated to account for the adsorption of N_a molecules.

During the second segment of the first MCS, a redistribution of the adsorbates was carried out to account for the effects of surface diffusion. A surface site and an adjacent site were chosen at random. If the surface site were occupied and the nearest-neighbor

site were vacant, a probability of diffusion was calculated using eq. 5. Q_i was taken as the heat of adsorption in the initial site, and Q_j was taken as the heat of adsorption in the final site. The calculated value of P_{ij} was then compared with a random number, R ($0 < R < 1$). If $R < P_{ij}$, the adsorbate was moved from the initial site, i , to the adjacent site, j . Conversely, if $R \geq P_{ij}$, movement of the adsorbate was not allowed. A sufficient number of surface visitations was allowed to attain an equilibrium distribution of adsorbates on the surface. The attainment of equilibrium was defined by a constant average heat of adsorption.

During the final segment of the first MCS, the desorption of adsorbates was treated in the following manner. A surface site was selected in a random fashion. If the site were occupied, the probability of desorption was calculated using eq. 3. The local activation energy for desorption was calculated using eq. 7. The value of P_{id} was then compared with a random number, R ($0 < R < 1$). If $R < P_{id}$, the adsorbate was removed from the lattice, and N_d in eq. 4 was incremented by one. Conversely, if $R \geq P_{id}$, the site remained occupied. The number of occupied sites visited during an MCS was θN_s . After completing the sampling of the surface, the rate of desorption for this MCS was calculated using eq. 4 and the surface coverage was updated to account for the desorption of N_d molecules.

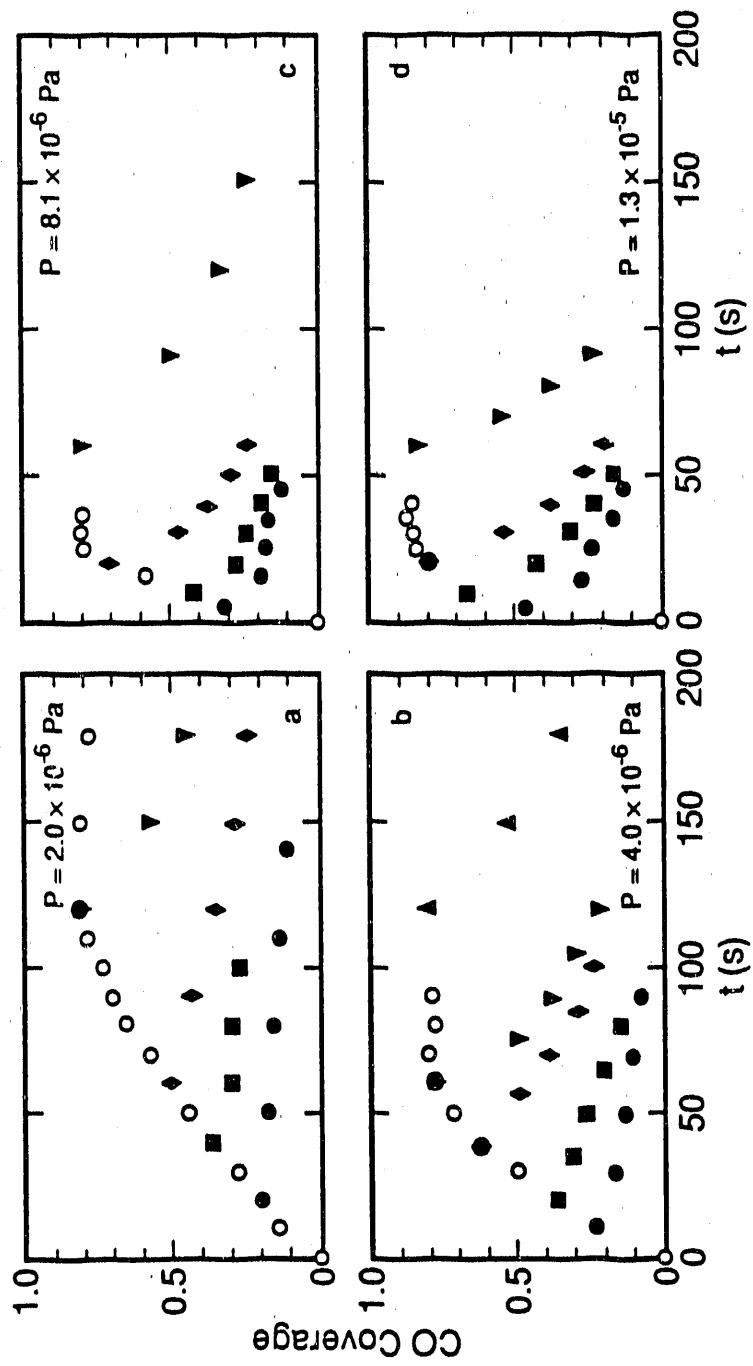
The adsorption, diffusion, and desorption segments described above constitute a Monte Carlo Step (MCS). As noted above, the time interval associated with an MCS is Δt . For simulations in which the surface temperature was maintained constant, a new MCS was begun by incrementing the time by Δt . For temperature-programmed

desorption simulations, the temperature was increased by the increment $\beta\Delta t$, where β is the heating rate. A new MCS was then carried out next. This process was repeated for either a specific length of time or until a desired temperature was reached. The results of the simulations were obtained as plots of coverage versus time or as the rate of desorption versus temperature. All of the calculations described above were carried out on an IBM 3090 computer. Random numbers were generated by the IMSL linear congruential number generator GGUBFS [23].

3.0 RESULTS AND DISCUSSION

3.1 Isothermal desorption of CO from Pd

The desorption of CO from polycrystalline Pd has been investigated by Yamada et al. [5,6] with the following type of experiment. At $t = 0$, C^{18}O at a pressure P is adsorbed onto a clean Pd surface held at a surface temperature T_s . At $t = t_1$, the gas-phase is quickly switched from C^{18}O to C^{16}O while maintaining a constant pressure. At $t = t_1 + t_2$, the surface is flashed to high temperature. The resulting TPD spectra resulting are integrated to determine the coverages of C^{18}O and C^{16}O , $\theta(\text{C}^{18}\text{O})$ and $\theta(\text{C}^{16}\text{O})$, and the total coverage, $\theta_T = \theta(\text{C}^{18}\text{O}) + \theta(\text{C}^{16}\text{O})$. Repetition of this procedure for a series of t_1 and t_2 generates plots of coverage versus time. Figures 1a-d show examples of the data obtained by Yamada et al. [6] with this approach. As seen in the figures, the total coverage initially rises with time until it reaches its equilibrium value for a given pressure. As the pressure is increased, higher equilibrium coverages



Variation in CO coverage versus time on polycrystalline Pd at $T_s = 380$ K (taken from Yamada et al. [6]). The open circles represent θ_T and the filled symbols represent the decay of $\theta(C^{18}O)$ for $t > t_1$. a) The isotope switch was made at a) $t_1 = 20$ s (●), 40 s (■), 60 s (◆), and 120 s (▼); b) at $t_1 = 10$ s (●), 20 s (■), 40 s (◆), 60 s (▼), and 120 s (▲); c) at $t_1 = 5$ s (●), 10 s (■), 20 s (◆), and 60 s (▼); d) at $t_1 = 5$ s (●), 10 s (■), 20 s (◆), and 60 s (▼).

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are achieved in shorter periods of time. Also shown in figs. 1a-d are the curves of the decay of $\theta(C^{18}O)$ for $t > t_1$. The rates of $\theta(C^{18}O)$ decay are seen to depend on the pressure and the coverage at which the isotope switch is made.

Yamada et al. [5] obtained apparent rate coefficients for the desorption of $C^{18}O$ by fitting the decay curves of $\theta(C^{18}O)$ to single exponential functions of the form $\theta(C^{18}O) = \theta^*(C^{18}O) \exp[-k_d(t-t_1)]$, where $\theta^*(C^{18}O)$ is the value $\theta(C^{18}O)$ at $t = t_1$. The initial rate of $C^{18}O$ desorption at the time of the isotope switch $t = t_1$ (i.e., the product $\theta^*(C^{18}O)k_d$) was found to be approximately linear with coverage. The rate coefficients determined from the slopes of the initial rate versus coverage curves were then plotted versus P . As seen in fig. 2, the rate constants k_dP exhibit an apparent linear dependence on the pressure. Based on this result, Yamada et al. [5] postulated that the rate of desorption is given by the expression

$$r_d = k'_d \theta + k''_d P^\alpha \theta = k_d^P \theta \quad (8)$$

where k_d' is the desorption rate coefficient when $P = 0$, k_d'' is the desorption rate coefficient which accounts for the effects of pressure when $P \neq 0$, and α is the power law exponent describing the dependence on pressure. It should be noted that while Yamada et al. [5] assumed k_d' to be coverage independent, their plots of r_d versus θ for the case when $P = 0$ are parabolic in shape, indicating that k_d' increases with increasing coverage.

In the simulations of the experiments of Yamada et al. [5,6], CO molecules were assumed to occupy both types of bridge sites on a Pd(100) surface [26,27]. A maximum of two CO molecules was

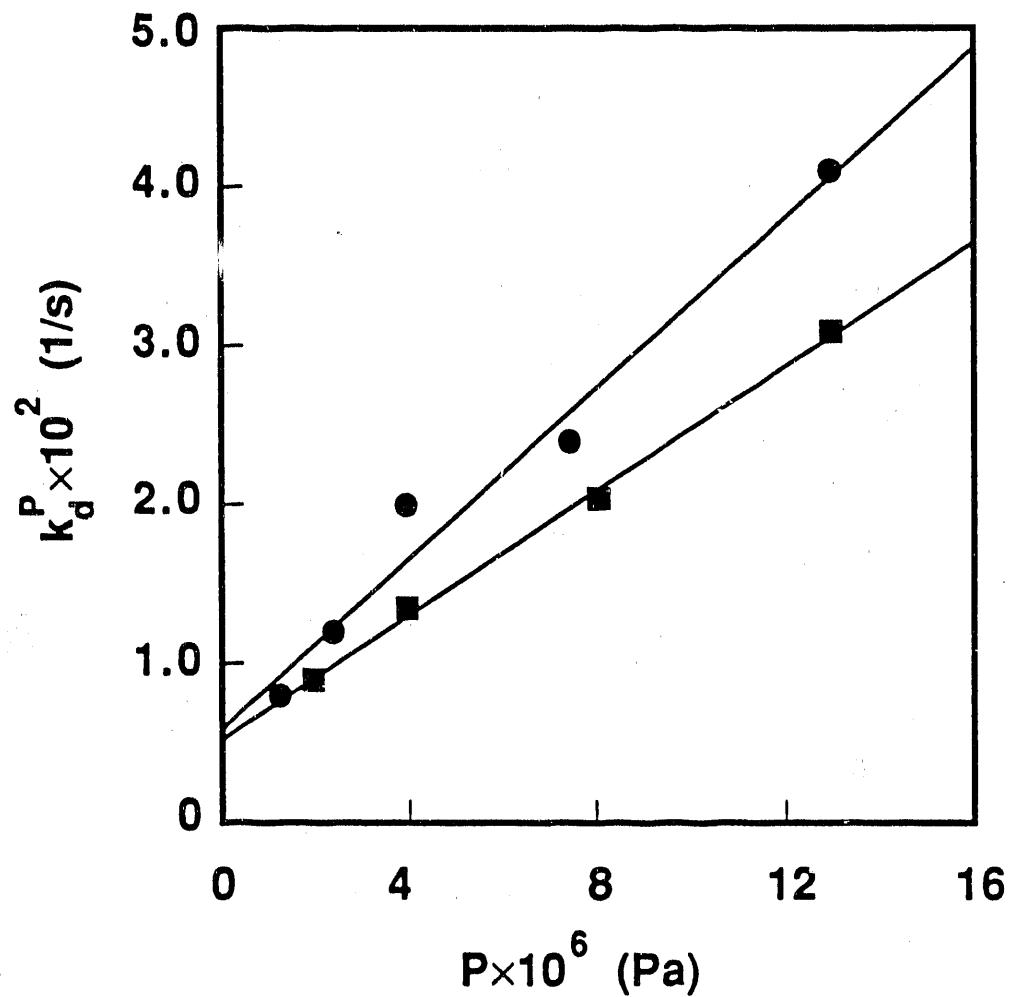


Fig. 2 Variation of $k_d P$ versus P . The points designated by (●) are from the analysis of experimental data presented by Yamada et al. [5] and the points designated by (■) are the values obtained from the Monte Carlo simulations.

allowed to be bonded to each metal atom, and because of steric constraints, the two molecules had to be located at 180° from each other relative to the shared metal atom. Exclusion of two CO molecules located at 90° relative to the shared metal atom is justified by the fact that in such a case the CO-CO distance would be 1.94 Å, whereas the hard-sphere diameter for CO is approximately 2.7 Å [24]. The saturation coverage of CO was taken to be 6.4×10^{-14} molecules cm^{-2} , the value reported by Yamada et al. [5]. This coverage corresponds to one CO molecule per two metal atoms for a Pd(100) surface.

The sticking coefficient for adsorption and its dependence on coverage were taken from the work of Yamada et al. [5]. In this study, the authors concluded that the function $S(\theta)$ could be described by a precursor model [25]. The specific form of $S(\theta)$ used for the simulations was

$$S(\theta) = S_0 \left(1 + \frac{\theta T}{1 - \theta T} K \right)^{-1} \quad (9)$$

where S_0 is the sticking coefficient at zero coverage and $K = k_d^{**}/(k_d^* + k_a^*)$. The rate constant k_d^{**} is the desorption rate constant from extrinsic precursor states, and k_d^* and k_a^* are the rate constants for desorption and adsorption, respectively, from intrinsic precursor states. The values of S_0 and K are 0.88 and 0.05, respectively.

To calculate desorption probabilities, the value of v in eq. 3 was taken as 10^{16} s^{-1} . This value fall within the range reported for CO on Pd surfaces [26,28]. The activation energy for desorption was calculated from eq. 7. The value of $Q_{\text{CO},n}$ in this equation was taken

as 31.5 kcal/mol, a value that lies within in the range of 30-40 kcal/mol reported in the literature for CO on Pd surfaces [24,26,28,29].

The results of the simulations are shown in figs. 3a-d. The equilibrium coverage is seen to increase with increasing CO pressure whereas the time required to reach the equilibrium coverage decreases. Comparison of figs. 3a-d with figs. 1a-d indicates that the general trends in the simulations resemble those observed in the experimental data. Exact agreement between the simulations and the experiments was not sought, since the simulations were carried out assuming a Pd(100) surface, whereas the experimental data were obtained on polycrystalline Pd. Nevertheless, the level of agreement between the experimental and simulated data is sufficiently high to suggest that the simulations capture the essential features of the experiments.

The curves of CO coverage versus time presented in fig. 3a-d were analyzed according to the method used by Yamada et al. [5]. As shown in fig. 2, treating the data in this fashion leads to a linear pressure dependence of the apparent rate coefficient for desorption. Both the slope and intercept of the dependence of k_dP on pressure determined from the simulations agrees rather well with the values reported by Yamada et al. [5].

Before addressing the origin of the apparent pressure dependence of k_dP , it should be noted that the isotope switch from C^{18}O to C^{16}O is made under two sets of conditions: before the total coverage is at its equilibrium value and after the total coverage has reached its equilibrium value. When the switch is made before the

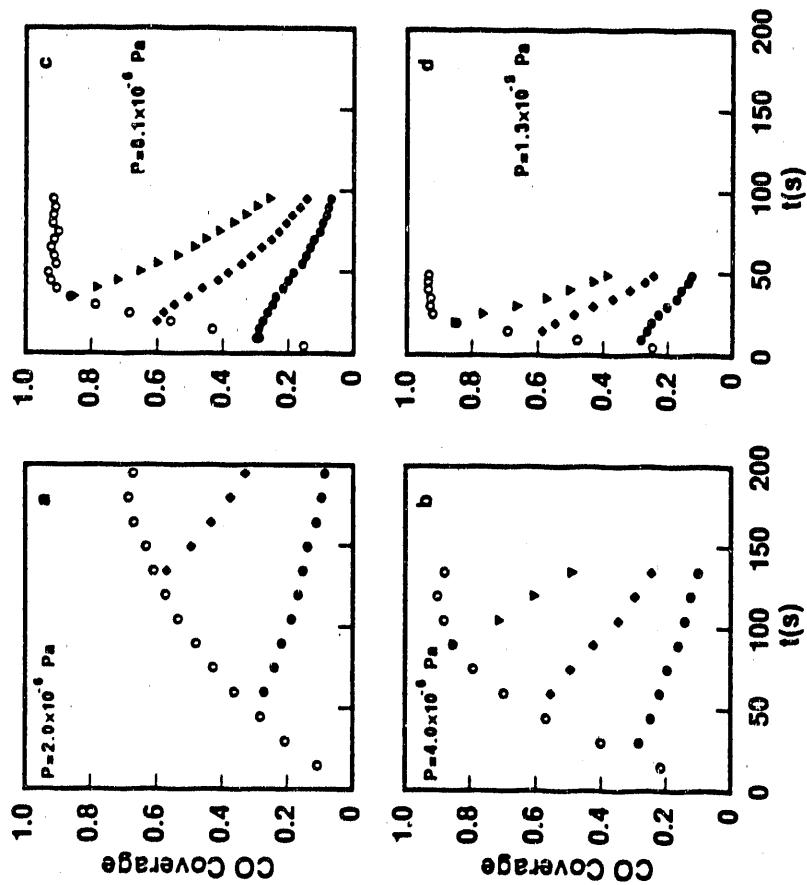


Fig. 3

Simulated variation in CO coverage versus time on Pd(100) at $T_s = 380$ K. The open circles represent θ_1 and the filled symbols represent the decay of $\theta(C^{18}O)$ for $t > t_1$.

a) The isotope switch was made at $\theta(C^{18}O) = 0.30$ (●) and 0.60 (◆). b-d) The isotope switch was made at $\theta(C^{18}O) = 0.30$ (●), 0.60 (◆), and 0.85 (▼).

equilibrium coverage is reached, the desorption of C^{18}O occurs while the total coverage is increasing. The rate at which the total coverage increases is proportional to the gas pressure. When the switch is made after the equilibrium coverage is reached, the desorption of C^{18}O occurs at a constant total coverage which is proportional to the pressure. Thus, in both cases, the pressure determines the total coverage trends.

Figure 4 shows that the average activation energy of the desorbing C^{18}O molecules decreases from 31.5 kcal/mol at $\theta_T = 0.35$ to 27.5 kcal/mol at $\theta_T = 0.90$. As a consequence, changes in θ_T due to the effects of pressure will influence the activation energy for desorption. It follows, therefore, that the pressure dependence of k_dP observed in fig. 2 can be ascribed to the effects of lateral interactions on the activation energy for CO desorption, which, in turn, can be related back to changes in θ_T with pressure.

It is interesting to compare the results of the present model with the continuum model presented earlier by Zhdanov [30]. In Zhdanov's model, the rates of adsorption and desorption of each isotopically labeled form of CO are given by

$$r_a = F_j S(\theta_T) \quad (10)$$

and

$$r_d = v \exp[-E_d(\theta_T) / k_b T] \theta_j \quad (11)$$

where F_j is the flux of isotope j to the surface and θ_j is the coverage of the surface by isotope j . The function $S(\theta_T)$ is given by eq. 9 with $S_0 = 0.90$ and $K = 0.10$. The coverage dependence of the activation energy for desorption is given by

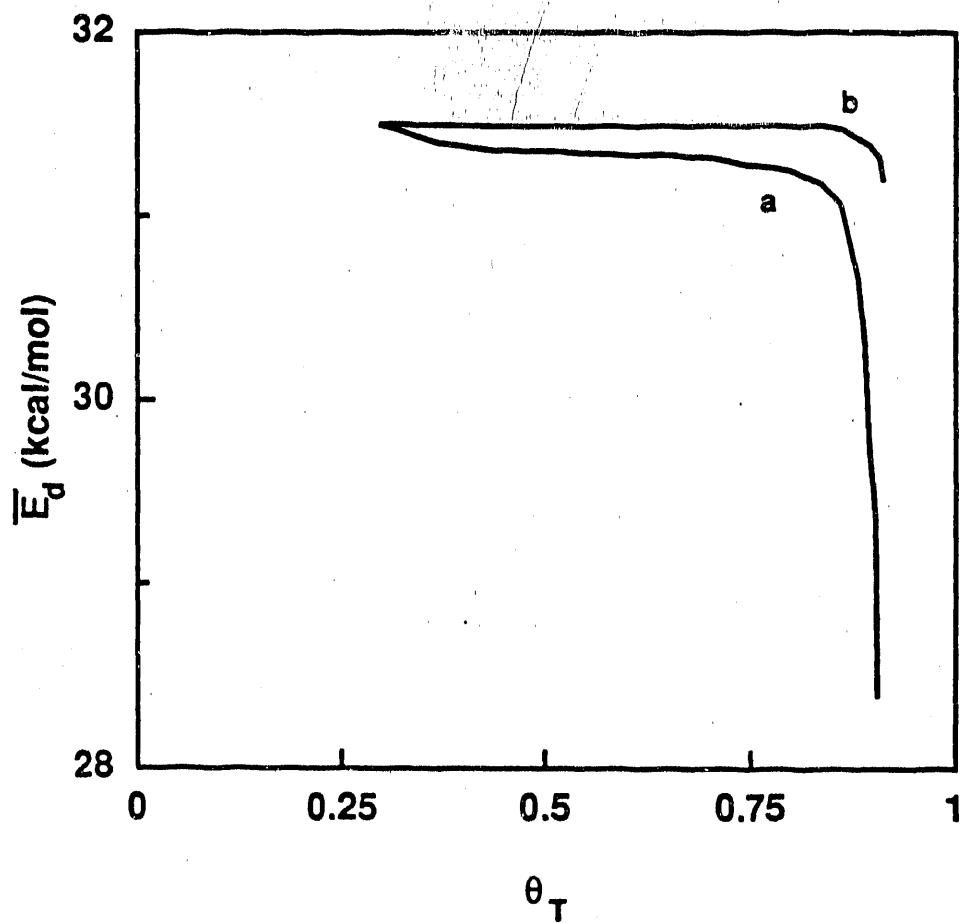


Fig. 4 Variation in the average value of E_d as a function of θ_T . Curve a is for the $C^{18}O$ molecules which desorb and curve b is for all of the $C^{18}O$ molecules on the surface. The results shown are for simulation presented in fig. 3c for the case where the isotope switch is made at $\theta(C^{18}O) = 0.30$.

$$E_d(\theta_T) = E_d^0 - \phi \theta_T^2 \quad (12)$$

where E_d^0 (= 39 kcal/mol) is the activation energy at zero coverage and ϕ (= 12 kcal/mol) is the coefficient describing the magnitude of the coverage dependence. The value of v in Zhdanov's model is 10^{16} s⁻¹. The variation in θ_j with time is then dictated by

$$\frac{d\theta_j}{dt} = r_a - r_d \quad (13)$$

Equations 10-13 were solved numerically to obtain coverage versus time profiles. As seen in fig. 5, the decay curves of $\theta(C^{18}O)$ determined from Zhdanov's model exhibit a flat induction period before decreasing. This pattern is inconsistent with that observed in the experiments of Yamada et al. [6] (see fig. 1) and in the Monte Carlo simulations reported here (see fig. 3).

To determine whether the observed induction period is a consequence of the manner in which the curves of $\theta(C^{18}O)$ versus time are calculated or the manner in which the activation energy dependence on coverage is represented, Monte Carlo simulations were conducted in which the activation energy for desorption appearing in eq. 3 was calculated from eq. 12. The resulting coverage versus time profiles are shown in fig. 5. The Monte Carlo simulations are seen to be in close agreement with those obtained by numerical solution of eqs. 10-13, and both methods of simulation produce an induction period in the decay curves of $\theta(C^{18}O)$. [The deviations of the Monte Carlo results from the numerical solution are attributable to the effects of finite step sizes and the approximation made in deriving the probabilities of adsorption and

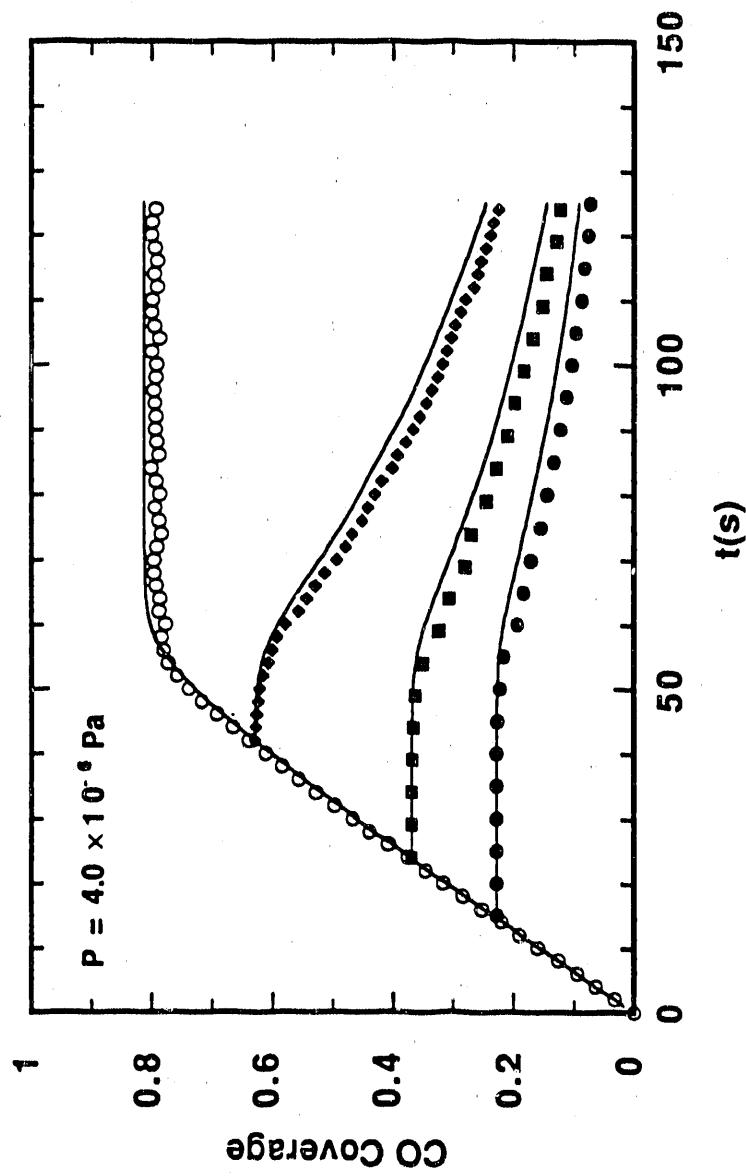


Fig. 5 Variation in CO coverage versus time at $T_s = 380$ K based on Zhdanov's model [30]. The solid curves represent the numerical solutions of eqs. 10-13 and the symbols represent the results of Monte Carlo simulation using eq. 12 to describe $E_d(\theta)$.

desorption (see Appendix).] It is, therefore, apparent that the appearance of an induction period can be traced to the use of eq. 12 for describing the effects of adsorbate coverage on the activation energy for desorption.

In contrast to Zhdanov's model, which makes use of the mean field (Bragg-Williams) approximation, the model reported here accounts explicitly for the local coverage dependence of E_d , through eq. 7. An important consequence of this feature is that the average activation energy of desorbing CO molecules is lower than the average activation energy for desorption calculated for all molecules present on the surface at a given coverage. This point is well illustrated in fig. 4, where curve a represents the average activation energy of desorbing CO molecules and curve b represents the activation energy for desorption averaged over all adsorbed CO molecules. Figure 4 also shows that the coverage dependence of curve a is much stronger than that for curve b. As discussed in ref. [21], the difference between curves a and b can be ascribed to the fact that desorption occurs preferentially from sites having a local coverage higher than the average. It is exactly this aspect of the present model that results in the absence of an induction period in the simulations of the decay in CO coverage versus time shown in fig. 3.

Finally, we note that the shape of curve b in fig. 4 closely resembles that determined from the analysis of equilibrium isobars and isosteres for CO on Pd(100) surfaces [24,26]. This indicates that the BOC-MP representation of the dependence of E_d on θ provides a physically correct description.

3.2 Temperature-programmed desorption of CO from Ni(100)

Yates and Goodman [4] have examined the desorption of C¹⁸O from Ni(100) in the presence and in the absence of a flux of C¹⁶O. The TPD spectra for C¹⁸O desorbing in the absence of a C¹⁶O flux are shown in fig. 6a as a function of gas exposure. The high-temperature β states are seen to fill first followed by the low-temperature α states. The principal feature in the spectrum is the β_2 peak at 425 K.

The effect of desorbing a saturated surface layer of C¹⁸O into a flux of C¹⁶O is seen in fig. 6b. The presence of the C¹⁶O flux is seen to shift the TPD spectra of C¹⁸O to lower temperatures relative to the spectrum observed in the absence of the beam flux. In addition, the integrated intensity of the α_1 and α_2 features is larger than that observed in the absence of the C¹⁶O flux.

Simulations of the experiments of Yates and Goodman [4] were conducted in the following manner. CO molecules were allowed to occupy both types of bridge sites on a Ni(100) surface with a maximum of 2 molecules per metal atom. As described in the previous section, the bonding of a sorbates 90° to each other on the shared metal atom was not allowed. It should be mentioned that HREELS studies of CO adsorption on Ni(100) have shown that CO occupies both bridge and on-top sites and that the relative occupancy of the two types of sites is dependent on the temperature and coverage [31]. This work demonstrated that at high coverage, bridge-bonding of CO predominates. For this reason, CO molecules were assigned to bridge sites at all coverages. The coverage in

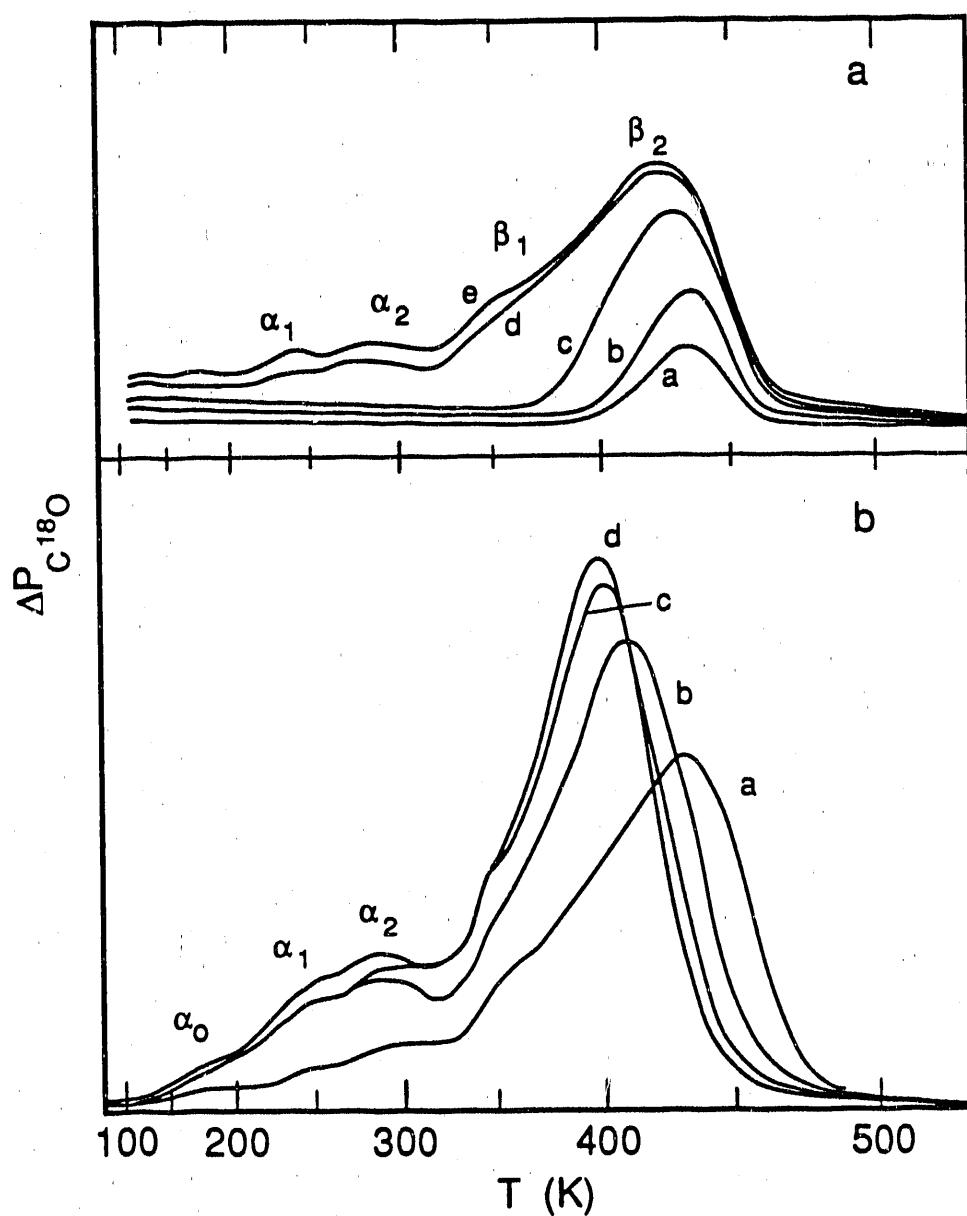


Fig. 6 TPD spectra of $C^{18}O$ desorption from $Ni(100)$ reported by Yates and Goodman [4]. a) Desorption in the absence of gas-phase $C^{16}O$. For curve a, the $C^{18}O$ exposure is 0.4 L; for curve b, 0.8 L; for curve c, 1.5 L; for curve d, 3.0 L; and for curve e, 8.0 L. b) Desorption of $C^{18}O$ (4.0 L dose) in the presence of gas-phase $C^{16}O$. For curve a, the $C^{16}O$ flux is $0 \text{ cm}^2 \text{ s}^{-1}$; for curve b, $4.1 \times 10^{14} \text{ cm}^2 \text{ s}^{-1}$; for curve c, $8.3 \times 10^{14} \text{ cm}^2 \text{ s}^{-1}$; and for curve d, $14.5 \times 10^{14} \text{ cm}^2 \text{ s}^{-1}$.

these simulations was defined as the number of CO molecules divided by the number of surface metal atoms.

Equation 9 was used to represent the sticking coefficient for CO adsorption. Based on molecular beam studies of CO adsorption on Ni(100) carried out by D'Evelyn et al. [32], the values of S_0 and K were taken to be 0.9 and 0.25, respectively.

The desorption probabilities of CO were calculated with the parameters $Q_{CO,n} = 30$ kcal/mol and $v = 10^{15}$ s⁻¹. These values are in the range of values reported in the literature for CO on Ni(100) [33-36].

The simulation of C¹⁸O desorption in the absence of a C¹⁶O flux is displayed in fig. 7a. The principal feature of the spectrum is a peak at 430 K which is in good agreement with the location of the β_2 feature in the experimentally obtained spectrum shown in fig. 6a. At coverages greater than 0.50, low-temperature features in the 300-400 K range are present.

Simulations of the desorption of C¹⁸O in the presence of a flux of C¹⁶O are shown in fig. 7b. The peak at 430 K which is observed in the absence of the C¹⁶O flux is seen to decrease in intensity as the flux is increased. In addition, the C¹⁶O flux is seen to enhance the desorption of C¹⁸O at low temperature relative to the amount of C¹⁸O observed at high temperature. The effect of the C¹⁶O flux on the C¹⁸O TPD spectra is qualitatively very similar to what is observed in the experimentally recorded spectra presented in fig. 6b.

The observation of enhanced desorption of C¹⁸O in the presence of a C¹⁶O flux can be understood in the following way. The presence of gas-phase C¹⁶O during the experiment leads to the adsorption of

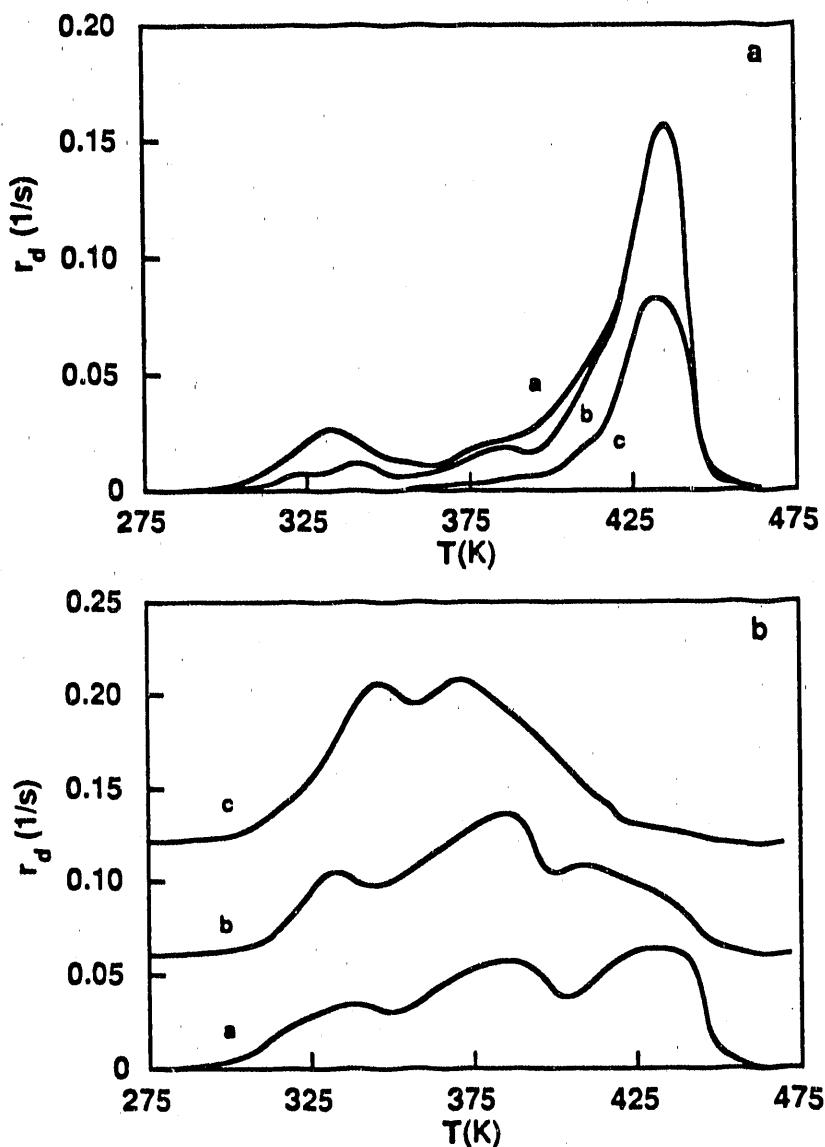


Fig. 7 Simulated TPD spectra of C^{18}O desorption from $\text{Ni}(100)$.
 a) Desorption in the absence of gas-phase C^{16}O . For curve a, $\theta^\circ(\text{C}^{18}\text{O}) = 0.60$; for curve b, 0.55; and for curve c, 0.25; b) Desorption of C^{18}O ($\theta^\circ(\text{C}^{18}\text{O}) = 0.60$) in the presence of gas-phase C^{16}O . For curve a, the C^{16}O flux is $4.1 \times 10^{14} \text{ cm}^2 \text{ s}^{-1}$; for curve b, $7.3 \times 10^{14} \text{ cm}^2 \text{ s}^{-1}$; and for curve c, $14.5 \times 10^{14} \text{ cm}^2 \text{ s}^{-1}$.

$C^{16}O$ and therefore to a larger value of θ_T at a given temperature. This conclusion is supported by fig. 8a which shows the variation in θ_T as a function of temperature for different fluxes of $C^{16}O$. At low temperatures, the surface is saturated with CO independent of the magnitude of the $C^{16}O$ flux. At higher temperatures, the surface is no longer saturated and the total coverage is seen to increase with the intensity of the $C^{16}O$ flux.

The influence of coverage on the activation energy is illustrated in figs. 8a and b. At low temperature, where the surface is saturated, both the coverage and activation energy are constant. In the temperature range between 350-450 K, the total coverage increases with the flux, whereas the activation energy varies inversely with the coverage. The lowering of the activation energy with increased coverage is thus responsible for the enhanced rate of desorption at lower temperatures. This explanation is consistent with that offered by Yates and Goodman [4] to explain their experimental data.

4.0 CONCLUSIONS

A Monte Carlo simulation technique has been developed to represent the effects of adsorbate pressure on the rate of adsorbate desorption. The model accounts for the adsorption, surface diffusion, and desorption of adsorbates and for the effects of lateral interactions between adsorbed species. The influence of adsorbate coverage on the activation energy for desorption is described using the BOC-MP approach. Simulations of both isothermal and

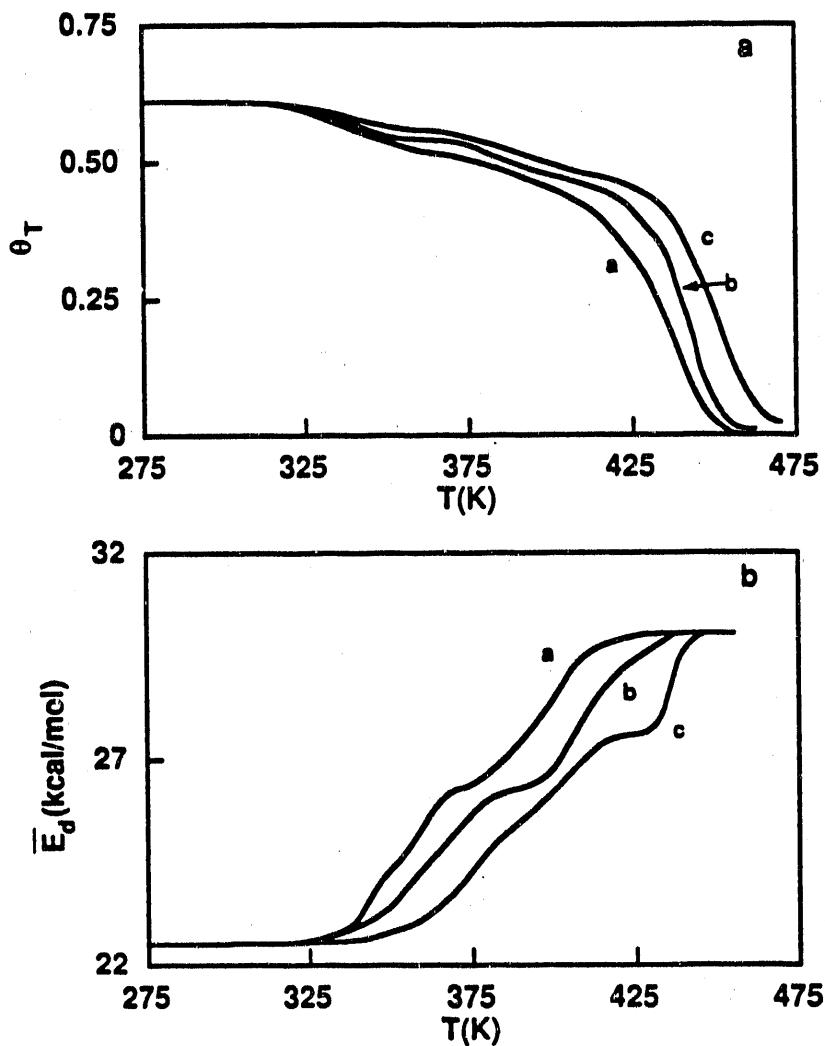


Fig. 8 a) Variation in θ_T versus T . b) Variation in the average \bar{E}_d for the desorbing CO molecules versus T . In both panels, curve a corresponds to a C^{16}O flux of $0 \text{ cm}^2 \text{ s}^{-1}$; curve b, $4.1 \times 10^{14} \text{ cm}^2 \text{ s}^{-1}$; and curve c, $14.5 \times 10^{14} \text{ cm}^2 \text{ s}^{-1}$.

temperature-programmed desorption demonstrate that the rate of desorption observed in the presence of adsorbing gas-phase species is enhanced relative to that observed in the absence of gas-phase species. This behavior can be ascribed to repulsive lateral interactions between adsorbates which lower the activation energy for desorption. The Monte Carlo simulations also demonstrate that for a given adsorbate coverage, the activation energy for desorption of desorbing molecules can be significantly lower than that for all molecules present on the surface. This difference is a direct consequence of the nonuniform distribution of adsorbate molecules on the surface.

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APPENDIX

The rationale for using the expressions for the probability of adsorption and desorption given by eqs. 1 and 3, respectively, is presented here. We show that these expressions can be derived from continuum rate expressions. The specific case considered is when the probabilities are normalized and all the environments on the surface are equivalent.

The rate of adsorption is given by

$$\frac{d\theta}{dt} = FS(\theta) = k_a PS(\theta) \quad (A1)$$

where k_a is the rate constant for adsorption and the other symbols have the same meaning as discussed earlier in this work. Equation A1 is a probability density function. Using standard relationships from probability theory, the probability of adsorption during a time interval Δt is defined as:

$$P' = \Delta\theta = \int_{t_1}^{t_2} k_a PS(\theta) dt = k_a PS(\theta) \Delta t \quad (A2)$$

The number of adsorbates, N_a , which adsorb in the time interval Δt is the product of P' and G , the sampling frequency of sites (i.e., the number of sites to which eq. A2 is applied). For a process which only depends only a single site, $G = N_s$, and N_a is given as

$$N_a = P'G = k_a PS(\theta) \Delta t N_s \quad (A3)$$

In order to incorporate eq. A3 into a Monte Carlo algorithm, it is convenient to factor eq. A3 into two terms, P^a and G^a which are given by:

$$P^a = S_o k_a P \Delta t \quad (A4)$$

$$G^a = [S(\theta)/S_0]N_s \quad (A5)$$

Equation A4 is identical to eq. 1 and eq. A5 is the expression used to determine the sampling frequency during the adsorption segment of an MCS.

Similar arguments apply to the process of desorption. The rate of non-associative desorption is given by

$$- \frac{d\theta}{dt} = k_d \theta \quad (A6)$$

Equation A6 can be used to define the probability of desorption:

$$P'' = |\Delta\theta| = \int_{t_1}^{t_2} k_d \theta dt = k_d \theta \Delta t \quad (A7)$$

The number of adsorbates which desorb in the time interval Δt is

$$N_d = P''G = k_d \theta \Delta t N_s \quad (A8)$$

Equation A8 can be factored into two terms, P^d and G^d , which are given by:

$$P^d = k_d \Delta t \quad (A9)$$

$$G^d = \theta N_s \quad (A10)$$

Equation A9 is identical to eq. 3 and eq. A10 is the expression used to determine the sampling frequency during the desorption segment of an MCS.

The derivations presented above establish the connection between the more familiar continuum rate expressions (eqs. A1 and A6) and the Monte Carlo probability expressions (eqs. A4 and A9).

The accuracy of the approximation $dt \approx \Delta t$ in eqs. A2 and A7 depends on the value of the step size Δt . As a rule of thumb, Δt was chosen so that the maximum number of adsorbates which adsorb or desorb

in a segment of an MCS was less than ten percent of the saturation coverage. As seen in fig. 5, specifying Δt in this fashion yields results which are in reasonable agreement with the numerical solution of the differential equations. The accuracy of the Monte Carlo approach has also been discussed in ref. [21] for the case of first- and second-order desorption kinetics.

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Appendix

FORTRAN Programs for Simulating Temperature-Programmed and Isothermal Desorption Kinetics

A description of the FORTRAN programs and subroutines used to perform the calculations described in this work is presented here. MONCAR70 is a main calling program used to simulate temperature-programmed desorption of two adsorbates coadsorbed on a surface. MONCAR15 is a main calling program used to simulate isothermal desorption in the presence of an adsorbing species. The listing of these two programs is provided at the end of the Appendix.

Description of MONCAR70

MONCAR70 is a calling program, written in FORTRAN, for simulating TPD spectra of one or two species, named A for adsorbate and C for coadsorbate, adsorbed on either an fcc or bcc (100) surface. The adsorbates A and C can occupy on-top, bridge, or hollow sites. The adsorbates must, however, be restricted to one type of site, and they both cannot have the same type of coordination (i.e., both cannot be bridged bonded). All parameters which are needed to perform a simulation are contained in MONCAR70. Variable initialization, overall program flow, and data output are all controlled in this program.

The surface array used in the simulations is a 100 by 100 logical array of numbered sites and all sites (on-top, bridge, and hollow) are numbered sequentially with the indices (I,J). If a site is occupied, it has a value "TRUE"; if unoccupied, it has a value of "FALSE". Site numbering begins in the upper left corner of the array with the on-top site numbered (1,1). A schematic of the site numbering is shown in fig. 1. Examples of on-top locations are (1,1), (3,5), and (99,99); examples of bridge locations are (1,2), (2,99), and (99,100); and examples of hollow locations are (2,2),(2,4), and (98,98). The location of an (I,J) pair of coordinates can be determined by the following set of rules applied to the array indices (I,J):

For on-top sites, the sum I+J is even and I is odd;

For bridge sites, the sum I+J is odd;

For hollow sites, the sum I+J is even and I is even.

In order to eliminate edge effects, periodic boundary conditions are used. In addition, the following corner or near-corner sites are never allowed to be occupied: (1,1),(1,2), (2,1),(2,2), (1,99), (2,99), (1,100), (2,100), (99,1), (99,2), (100,1), (100,2), (99,99), (99,100), (100,99), and (100,100). These 16 sites, which represent a small fraction of the total number of sites, are kept unoccupied because it is computationally expensive to account for them in simulations.

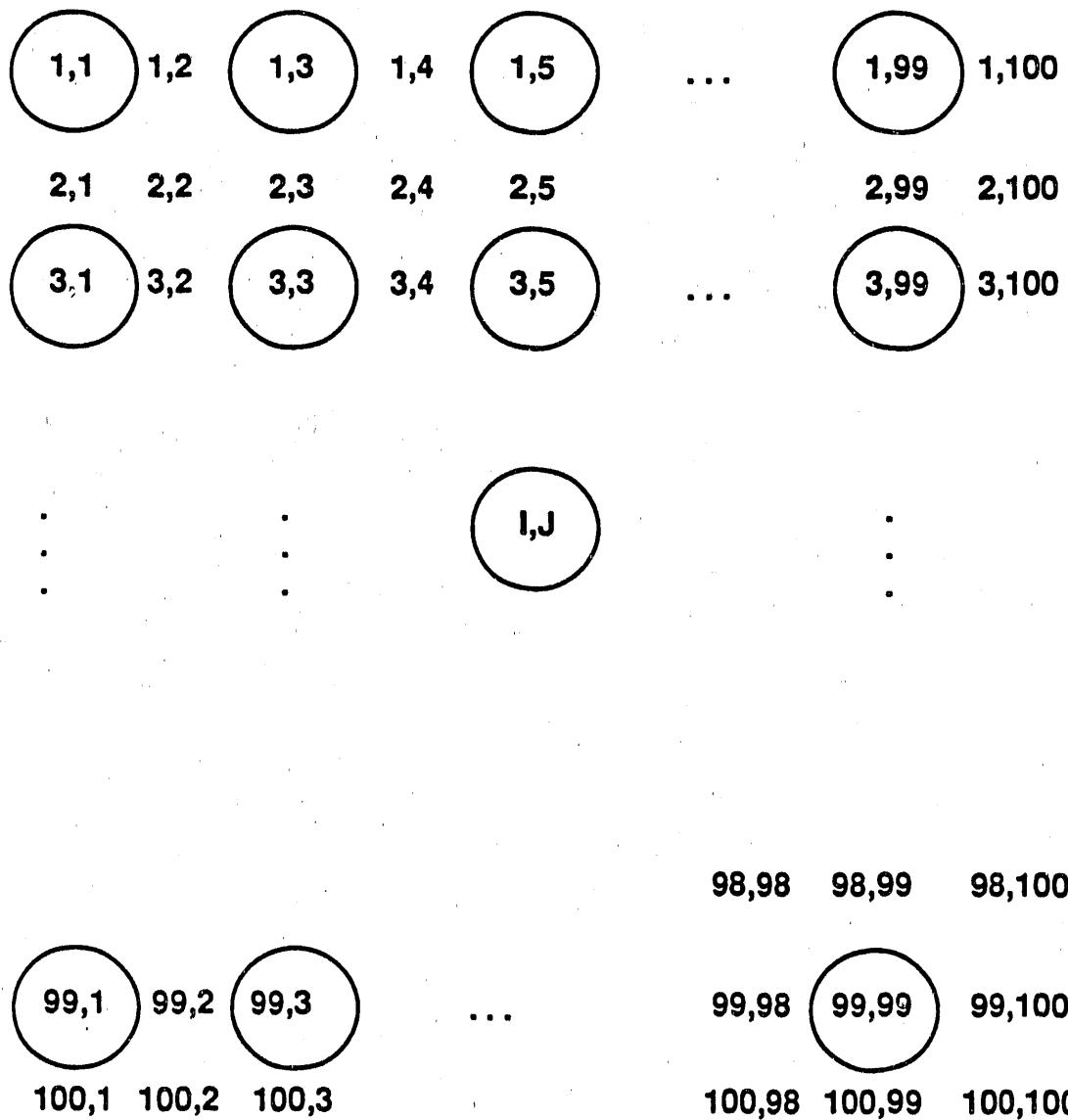


Fig. 1 Surface map showing the numbering scheme (I,J) of the on-top, bridge, and hollow sites. The open circles are the metal atoms.

To perform a simulation, MONCAR70 calls a number of subroutines. The overall hierarchy of MONCAR70 is as follows:

- 1) Variables are defined and initialized.
- 2) The initial coverages are placed on the surface arrays by calling subroutines INTTOP, INTBRI, and INTHOL.
- 3) The temperature for the Monte Carlo step (MCS) is initialized. The desorption segment of the MCS is conducted by calling subroutines DESORT, DESORB, and DESORH.
- 4) The diffusion segment of the MCS is conducted by calling subroutines DIFTOP, DIFBRI, and DIFHOL.
- 5) The output variables for the MCS are written to a data file. The temperature is increased and steps 3,4, and 5 are repeated a specified number of times.
- 6) The simulation ends.

Each subroutine called by MONCAR70 is highly modular in nature and performs a specific function. The subroutines and their arguments are listed below along, with a brief description. In general, names of subroutines, arguments, or variables having to do with on-top sites or species have "T" or "TOP" in their name. In a similar fashion, names which apply to bridge sites or species have "B" or "BRI" in their name and names which apply to hollow sites or species have "H" or "HOL".

Subroutines(Arguments)

INTTOP(COVTI,LOCT)

Initializes the coverage in the on-top sites. COVTI is the value of the coverage to be initialized and LOCT indicates whether the coverage is to be random or ordered.

INTBRI(COVBI,LOCB)

Initializes the coverage in the bridge sites. COVBI is the value of the coverage to be initialized and LOCB indicates whether the coverage is to be random or ordered.

INTHOL(COVHI,LOCH)

Initializes the coverage in the hollow sites. COVHI is the value of the coverage to be initialized and LOCH indicates

whether the coverage is to be random or ordered.

DESORT(SAMTOP,COUNTT,NSVIST,BET,AET,BETD,AETD)

Treats the non-associative desorption of adsorbates from on-top sites. SAMTOP is the sampling frequency of occupied on-top sites during an MCS. COUNTT is a running count of the number of on-top species which desorb in an MCS. NSVIST is a running count of the number of successful visits to on-top sites (i.e., occupied sites) during an MCS. BET is a running sum of the binding energies for all on-top species visited during an MCS. BETD is a running sum of the binding energies for all on-top species which desorb during MCS. AET is a running sum of activation energies of desorption for all on-top species visited during an MCS. AETD is a running sum of the activation energies of desorption for the on-top species desorbing during an MCS.

DESORB(SAMBRI,COUNTB,NSVISB,BEB,AEB,BEBD,AEBD)

Treats the non-associative and associative desorption of adsorbates from bridge sites. The arguments are similar to those described for DESORT.

DESORH(SAMHOL,COUNTH,NSVISH,BEH,AEH,BEHD,AEHD)

Treats the non-associative and associative desorption of adsorbates from hollow sites. The arguments are similar to those described for DESORT.

DIFTOP(NJMPT)

Treats jumps of adsorbates between on-top sites. NJMPT is a running count of the number of jumps between on-top sites during an MCS.

DIFBRI(NJMPB)

Treats jumps of adsorbates between bridge sites. NJMPB is a running count of the number of jumps between bridge sites during an MCS.

DIFHOL(NJMPC)

Treats jumps of adsorbates between hollow sites. NJMPC is a

running count of the number of jumps between hollow sites during an MCS.

QONTOP(I,J,QT)

Calculates the binding energy QT for an on-top site (I,J) with the BOC-MP method.

QBRIDG(I,J,QB)

Calculates the binding energy QB for a bridge site (I,J) with the BOC-MP method.

QHOLLO(I,J,QH)

Calculates the binding energy QH for an fcc(100) hollow site (I,J) with the BOC-MP method.

QBCCHO(I,J,QH)

Calculates the binding energy QH for a bcc(100) hollow site with the BOC-MP method.

LOCTT(I,J,ITT,JTT)

Given an on-top (I,J) location, returns arrays ITT and JTT of the four nearest-neighbor on-top locations.

LOCTB(I,J,ITB,JTB)

Given an on-top (I,J) location, returns arrays ITB and JTB of the four nearest-neighbor bridge locations.

LOCTH(I,J,ITH,JTH)

Given an on-top (I,J) location, returns arrays ITH and JTH of the two nearest-neighbor hollow locations.

LOCBT(I,J,IBT,JBT)

Given a bridge (I,J) location, returns arrays IBT and JBT of the four nearest-neighbor on-top locations.

LOCBB(I,J,IBB,JBB)

Given a bridge (I,J) location, returns arrays IBB and JBB of the four nearest-neighbor bridge locations.

LOCBH(I,J,IBH,JBH)

Given a bridge (I,J) location, returns arrays IBH and JBH of the two nearest-neighbor hollow locations.

LOCBBN(I,J,IBBN,JBBN)

Given a bridge (I,J) location, returns arrays IBBN and JBBN of the four next-nearest-neighbor bridge locations.

LOCHT(I,J,IHT,JHT)

Given a hollow (I,J) location, returns arrays IHT and JHT of the four nearest-neighbor on-top locations.

LOCHB(I,J,IHB,JHB)

Given a hollow (I,J) location, returns arrays IHB and JHB of the four nearest-neighbor bridge locations.

LOCHH(I,J,IHH,JHH)

Given a hollow (I,J) location, returns arrays IHH and JHH of the four nearest-neighbor hollow locations.

LOCHHN(I,J,IHHN,JHHN)

Given a hollow (I,J) location, returns arrays IHHN and JHHN of the four next-nearest-neighbor hollow locations.

SURMAP

Creates a TELAGRAF plotting file of the positions of the surface atoms and the adsorbates.

A listing of the important variables and arrays in MONCAR70 and their meanings are given next. Names which have already been mentioned above as subroutine arguments are not described again. The convention for using "T" or "TOP" to relate to on-top, "B" or "BRI" to relate to bridge, and "H" or "HOL" to relate to hollow, still applies.

Name	Meaning
T(100,100)	Surface array for the occupancy of on-top sites.
B(100,100)	Surface array for the occupancy of bridge sites.
H(100,100)	Surface array for the occupancy of hollow sites.

DSEED	Seed variable for the random number generator.
R	Gas constant.
SITES	Number of surface metal atoms or sites.
TEMP	Temperature of the surface (K).
TMIN	Initial temperature.
BETA	Heating rate (K/s).
DELT	Time interval (s).
KD0	Preexponential factor for desorption (1/s).
DELKD0	The product of DELT and KD0.
TOP	Denotes if on-top sites sites are used (1) or not (0).
BRI	Denotes if bridge sites sites are used (1) or not (0).
HOL	Denotes if hollow sites sites are used (1) or not (0).
BCC	Denotes if the hollow site is a bcc(100) hollow site (1) or a fcc(100) hollow site (0).
MODT	Denotes the mechanism of desorption for species in on-top sites: (1) for non-associative and (2)for associative desorption.
MODB	Denotes the mechanism of desorption for species in bridge sites: (1) for non-associative and (2)for associative desorption.
MODH	Denotes the mechanism of desorption for species in hollow sites: (1) for non-associative and (2)for associative desorption.
Q0T	Heat of adsorption in the on-top position for the on-top species.
Q0B	Heat of adsorption in the on-top position for the bridge species.
Q0H	Heat of adsorption in the on-top position for the hollow species.
TT	Denotes whether direct on-top to on-top adsorbate interactions occur (1) or not (0).
TB	Denotes whether direct on-top to bridge adsorbate interactions occur (1) or not (0).
TH	Denotes whether direct on-top to hollow adsorbate interactions occur (1) or not (0).
BB	Denotes whether direct bridge to bridge adsorbate interactions occur (1) or not (0).

BH	Denotes whether direct bridge to hollow adsorbate interactions occur (1) or not (0).
HH	Denotes whether direct hollow to hollow adsorbate interactions occur (1) or not (0).
BOTT	Bond order for on-top to on-top direct adsorbate interactions.
BOTB	Bond order for on-top to bridge direct adsorbate interactions.
BOTH	Bond order for on-top to hollow direct adsorbate interactions.
BOBB	Bond order for bridge to bridge direct adsorbate interactions.
BOBH	Bond order for bridge to hollow direct adsorbate interactions.
BOHH	Bond order for hollow to hollow direct adsorbate interactions.
DETT	Dissociation energy for on-top to on-top direct adsorbate interactions.
DETB	Dissociation energy for on-top to bridge direct adsorbate interactions.
DETH	Dissociation energy for on-top to hollow direct adsorbate interactions.
DEBB	Dissociation energy for bridge to bridge direct adsorbate interactions.
DEBH	Dissociation energy for bridge to hollow direct adsorbate interactions.
DEHH	Dissociation energy for hollow to hollow direct adsorbate interactions.
SURPLO	Denotes whether surface maps are (1) or are not (0) to be made.
IM	Switch variable to limit the printing of the coordinates of metal atoms to once.
IP(1)	Switch variable to limit the printing of adsorbate coordinates for a specified coverage to once.
COVDES(1)	Array containing the coverage at which a surface map is to be made.
GCOUNT	Global (total) running count of the number of species which have desorbed from on-top sites.
GCOUNB	Global (total) running count of the number of

GCOUNH	species which have desorbed from bridge sites. Global (total) running count of the number of species which have desorbed from hollow sites.
COVT	Running value of the coverage in the on-top sites.
COVB	Running value of the coverage in the bridge sites.
COVH	Running value of the coverage in the hollow sites.
RATET	Rate of desorption for species in on-top sites for an MCS.
RATEB	Rate of desorption for species in bridge sites for an MCS.
RATEH	Rate of desorption for species in hollow sites for an MCS.

Description of MONCAR15

MONCAR15 is a calling program, written in FORTRAN, for simulating isothermal adsorption, diffusion, and desorption of one or two species, named A (for adsorbate) and C (for coadsorbate), adsorbed on an fcc(100) surface. Species A and C have identical energy parameters (they are isotopes) and they both occupy bridge sites. All parameters which are needed to perform a simulation are contained in MONCAR15. Variable initialization, overall program flow, and data output are all controlled in this program. MONCAR15, and the subroutines called by it, are very similar in structure to MONCAR70 and the subroutines described earlier. The surface array (see fig. 1) is identical to that described earlier.

In order to perform a simulation, MONCAR15 calls a number of subroutines. The overall hierarchy of MONCAR15 is as follows:

- 1) Variables are defined and initialized.
- 2) The initial coverages are placed on the surface arrays by calling subroutine INTBTW.
- 3) The adsorption segment of a Monte Carlo step (MCS) is conducted by calling subroutine ADSBRI.
- 4) The diffusion segment of an MCS is conducted by calling subroutine DIFBTW.
- 5) The desorption segment of an MCS is conducted by calling subroutine DESBTW.
- 6) The output variables for an MCS are written to a data file. The time is incremented and steps 3, 4, 5, and 6 are repeated a specified number of times.
- 6) The simulation ends.

The subroutines mentioned above and their arguments are listed below along with a brief description. Names of subroutines, arguments, or variables having to do with bridge sites or species have "B" or "BRI" in their name, and, in addition, an "A" or "C" is added to the names to differentiate between species A or C.

Subroutines(Arguments)

INTBTW(COVBAI,LOCBA,COVBCI,LOCBC)

Initializes the coverage in the bridge sites. COVBAI and COVBCI are the values of the coverages of A and C to be initialized, respectively. LOCBA and LOCBC indicates whether the coverages of A and C are to be random or ordered.

ADSBRI(SAMFRB,CTADC)

Treats the non-dissociative adsorption of species C. SAMFRB is the sampling frequency of unoccupied bridge sites. CTADC is a running count of the number of C species adsorbed.

DIFBTW(NJMP2)

Treats jumps of adsorbates between bridge sites. NJMP2 is a running count of the number of jumps between bridge sites during an MCS.

DESBTW(SAMFDA,SAMFDC,COUNTA,COUNTC,AEAD,AECD,AEA,AEC,
NSVISA,NSVISC)

Treats the non-associative desorption of adsorbates A and C from bridge sites. SAMFDA and SAMFDC are the sampling frequencies of occupied bridge sites for A and C during an MCS. COUNTA and COUNTC are running counts of the number of bridge A and C species which desorb in a MCS. AEAD and AEAC are running sums of the activation energies for all bridge A and C species visited during an MCS. AEA and AEC are running sums of activation energies of desorption for all bridge A and C species visited during an MCS. NSVISA and NSVISC is a running count of the number of successful visits to bridge A and C sites (i.e., occupied sites) during an MCS.

The calling programs MONCAR70 and MONCAR15 are listed on the next pages.

Main: MONCAR70

C TWO TYPES OF ADSORBATES CAN BE ACCOMMODATED ON A FOUR-FOLD MON00010
 C ROTATIONAL SURFACE. THE THREE SURFACE ARRAYS ARE T(100,100), MON00020
 C B(100,100) AND H(100,100). SITE DIFFERENTIATION IS BY MON00030
 C ON-TOP SUM OF I+J IS EVEN AND I IS ODD; MON00040
 C BRIDGE SUM OF I+J IS ODD; MON00050
 C HOLLOW SUM OF I+J IS EVEN AND I IS EVEN. MON00060
 C WRITTEN 11/22/88 BY STEPHEN J. LOMBARDO. MON00070
 C*****MON00080

IMPLICIT DOUBLE PRECISION (A-H,O-Z) MON00090
 DOUBLE PRECISION DSEED, TEMP, R, QOT, QOB, QOH, BOTT, BOTB, BOTH, BOBB, MON00100
 1 BOBH, BOHH, DETT, DETB, DETH, DEBB, DEBH, DEHH, KDO, MON00110
 1 GCOUNT, GCOUNB, GCOUNH, COV(3), COVDES(3) MON00120
 INTEGER I, J, MODT, MODB, MODH, TT, TB, TH, BB, BH, HH, TOP, BRI, HOL, MON00130
 1 COUNTT, COUNTB, COUNTH, BCC, SURPLO, IP(3) MON00140
 LOGICAL T(100,100), B(100,100), H(100,100) MON00150
 COMMON T, B, H, DSEED, TEMP, R, QOT, QOB, QOH, BOTT, BOTB, BOTH, BOBB, BOBH, MON00160
 1 BOHH, DETT, DETB, DETH, DEBB, DEBH, DEHH, DELKDO, MON00170
 1 MODT, MODB, MODH, TT, TB, TH, BB, BH, HH, TOP, BRI, HOL, BCC MON00180
 C*****MON00190

C INITIALIZE VARIABLES MON00200
 DSEED=149753D0 MON00210
 R=1.987D-3 MON00220
 SITES=2500D0 MON00230
 TMIN=200D0 MON00240
 BETA=74D0 MON00250
 DELT=3D0/BETA MON00260
 KDO=1D16 MON00270
 DELKDO=DELT*KDO MON00280
 TEMP=TMIN-BETA*DELT MON00290
 TOP=0 MON00300
 BRI=1 MON00310
 HOL=0 MON00320
 BCC=1 MON00330
 MODT=1 MON00340
 MODB=2 MON00350
 MODH=1 MON00360
 QOT=33.45D0 MON00370
 QOB=42.0D0 MON00380
 QOH=57.0D0/1.75D0 MON00390
 TT=0 MON00400
 TB=0 MON00410
 TH=0 MON00420
 BB=1 MON00430
 BH=0 MON00440
 HH=0 MON00450
 BOTT=.01D0 MON00460
 BOTB=.01D0 MON00470
 BOTH=.01D0 MON00480
 BOBB=.038D0 MON00490
 BOBH=.01D0 MON00500
 BOHH=.03D0 MON00510
 DETT=100.D0 MON00520
 DETB=100.D0 MON00530
 DETH=100.D0 MON00540
 DEBB=109.5D0 MON00550
 DEBH=100.D0 MON00560
 DEHH=109.5D0 MON00570
 SURPLO=1 MON00580
 IM=1 MON00590
 IP(1)=1 MON00600
 COVDES(1)=1.00D0 MON00610
 IP(2)=1 MON00620
 COVDES(2)=0.5D0 MON00630
 IP(3)=0 MON00640
 COVDES(3)=0.5D0 MON00650
 GCOUNT=0 MON00660
 GCOUNB=0 MON00670
 GCOUNH=0 MON00680
 C*****MON00690

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C MON00700
C*****MON00710
C      INITIALIZE COVERAGE EITHER RANDOM(1) OR ORDERED(0) BY LOCI. MON00720
C      DATA COVTI, LOCT, COVBI, LOCB, COVHI, LOCH/0.5D0, 1, 2.0D0, 0, 1D0, 0/ MON00730
C      IF(TOP.EQ.1) THEN MON00740
C          CALL INTTOP(COVTI, LOCT) MON00750
C      ENDIF MON00760
C      IF(BRI.EQ.1) THEN MON00770
C          CALL INTBRI(COVBI, LOCB) MON00780
C      ENDIF MON00790
C      IF(HOL.EQ.1) THEN MON00800
C          CALL INTHOL(COVHI, LOCH) MON00810
C      ENDIF MON00820
C*****MON00830
C      MON00840
C*****MON00850
C      START OF SIMULATION MON00860
C      BEGIN MONTE CARLO STEP: DESORPTION SEGMENT MON00870
C      DO 100 NM=1, 133 MON00880
C          TEMP=TEMP+BETA*DELT MON00890
C          COVT=((COVTI*SITES)-GCOUNT)/SITES MON00900
C          COVB=((COVBI*SITES)-GCOUNB)/SITES MON00910
C          COVH=((COVHI*SITES)-GCOUNH)/SITES MON00920
C          COV(1)=COVB MON00930
C          COV(2)=COVB MON00940
C          COV(3)=COVB MON00950
C          IF(SURPLO.EQ.1) THEN MON00960
C              CALL SURMAP(COVT, COVB, COVH, COV, COVDES, IP, IM) MON00970
C          ENDIF MON00980
C          COUNTT=0 MON00990
C          COUNTB=0 MON01000
C          COUNTH=0 MON01010
C          NSVIST=0 MON01020
C          NSVISB=0 MON01030
C          NSVISH=0 MON01040
C          AET=0 MON01050
C          AEB=0 MON01060
C          AEH=0 MON01070
C          BET=0 MON01080
C          BEB=0 MON01090
C          BEH=0 MON01100
C          AETD=0 MON01110
C          AEBD=0 MON01120
C          AEHD=0 MON01130
C          BETD=0 MON01140
C          BEBD=0 MON01150
C          BEHD=0 MON01160
C          NJMPT=0 MON01170
C          NJMPB=0 MON01180
C          NJMPH=0 MON01190
C          IF(MODT.EQ.1) THEN MON01200
C              SAMTOP=SITES*COVT MON01210
C          ELSE MON01220
C              SAMTOP=SITES*(COVT**2)*.5D0 MON01230
C          ENDIF MON01240
C          IF(MODB.EQ.1) THEN MON01250
C              SAMBRI=SITES*COVB MON01260
C          ELSE MON01270
C              SAMBRI=SITES*(COVB**2)*.5D0 MON01280
C          ENDIF MON01290
C          IF(MODH.EQ.1) THEN MON01300
C              SAMHOL=SITES*COVH MON01310
C          ELSE MON01320
C              SAMHOL=SITES*(COVH**2)*.5D0 MON01330
C          ENDIF MON01340
C*****MON01350
C      DETERMINE DESORPTION OF SPECIES. MON01360
C      DO 55 NM=1, 5000 MON01370
C          CALL DESORT(SAMTOP, COUNTT, NSVIST, BET, AET, BETD, AETD) MON01380
C          CALL DESORB(SAMBRI, COUNTB, NSVISB, BEB, AEB, BEBD, AEBD) MON01390
C          CALL DESORH(SAMHOL, COUNTH, NSVISH, BEH, AEH, BEHD, AEHD) MON01400

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55      CONTINUE                               MON01410
C*****MON01420
C*****MON0143
C*****MON01440
C      COMPLETE MONTE CARLO STEP: DIFFUSION SEGMENT  MON01450
DO 80 KK=1,25000                               MON01460
C      CALL DIFTOP (NJMPT)                      MON01470
DO 78 LL=1,2                                     MON01480
78      CALL DIFBRI (NJMPB)                      MON01490
C      CALL DIFHOL (NJMPH)                      MON01500
80      CONTINUE                               MON01510
C*****MON01520
C*****MON01530
C*****MON01540
C      CALCULATE OUTPUT VARIABLES AND WRITE OUT RESULTS MON01550
GCOUNT=GCOUNT+COUNTT                         MON01560
GCOUNB=GCOUNB+COUNTB                         MON01570
GCOUNH=GCOUNH+COUNTH                         MON01580
RATET=COUNTT/(DELT*SITES*MODT)                MON01590
RATEB=COUNTB/(DELT*SITES*MODB)                MON01600
RATEH=COUNTH/(DELT*SITES*MODH)                MON01610
IF (NSVIST.NE.0) THEN
  AET=AET/(NSVIST)                           MON01620
  BET=BET/(NSVIST*MODT)                      MON01630
ENDIF                                         MON01640
IF (NSVISB.NE.0) THEN
  AEB=AEB/(NSVISB)                           MON01650
  BEB=BEB/(NSVISB*MODB)                      MON01660
ENDIF                                         MON01670
IF (NSVISH.NE.0) THEN
  AEH=AEH/(NSVISH)                           MON01680
  BEH=BEH/(NSVISH*MODH)                      MON01690
ENDIF                                         MON01700
IF (COUNTT.NE.0) THEN
  AETD=AETD/COUNTT*MODT                      MON01710
  BETD=BETD/COUNTT                           MON01720
ENDIF                                         MON01730
IF (COUNTB.NE.0) THEN
  AEBD=AEBD/COUNTB*MODB                      MON01740
  BEBD=BEBD/COUNTB                           MON01750
ENDIF                                         MON01760
IF (COUNTH.NE.0) THEN
  AEHD=AEHD/COUNTH*MODH                      MON01770
  BEHD=BEHD/COUNTH                           MON01780
ENDIF                                         MON01790
IF (TOP.EQ.1) THEN
  WRITE(7,90) TEMP, RATET, COVT, AET, AETD, BET, BETD, COUNTT, NJMPT,
1      NSVIST                                     MON01800
ENDIF                                         MON01810
IF (BRI.EQ.1) THEN
  WRITE(8,90) TEMP, RATEB, COVB, AEB, AEBD, BEB, BEBD, COUNTB, NJMPB,
1      NSVISB                                     MON01820
ENDIF                                         MON01830
IF (HOL.EQ.1) THEN
  WRITE(9,90) TEMP, RATEH, COVH, AEH, AEHD, BEH, BEHD, COUNTH, NJMPH,
1      NSVISH                                     MON01840
ENDIF                                         MON01850
90      FORMAT(1X,F6.1,1X,6(1PE10.3,1X),3(I6,1X)) MON01860
100     CONTINUE                               MON01870
      STOP                                     MON01880
      END                                     MON01890
                                         MON01900
                                         MON01910
                                         MON01920
                                         MON01930
                                         MON01940
                                         MON01950
                                         MON01960
                                         MON01970
                                         MON01980
                                         MON01990
                                         MON02000
                                         MON02010

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Main: MONCAR15

C TWO TYPES OF ADSORBATES CAN BE ACCOMMODATED ON A FOUR-FOLD MON00010
 C ROTATIONAL SURFACE. THE THREE SURFACE ARRAYS ARE T(100,100), MON00020
 C B(100,100) AND H(100,100). SITE DIFFERENTIATION IS BY MON00030
 C ON-TOP SUM OF I+J IS EVEN AND I IS ODD; MON00040
 C BRIDGE SUM OF I+J IS ODD; MON00050
 C HOLLOW SUM OF I+J IS EVEN AND I IS EVEN. MON00060
 C THE TWO TYPES OF ADSORBATE BOTH OCCUPY SITES WITH THE SAME MON00070
 C COORDINATION. ONE SURFACE ARRAY IS FOR THE ABSOLUTE COVERAGE, THE MON00080
 C OTHER IS TO IDENTIFY THE SPECIES. THIS PROGRAM IS FOR PRESSURE MON00090
 C ASSISTED DESORPTION WITH TWO ISOTOPES. MON00100
 C WRITTEN 1/25/90 BY STEPHEN J. LOMBARDO. MON00110
 C*****MON00120
 C IMPLICIT DOUBLE PRECISION (A-H, O-Z) MON00130
 C DOUBLE PRECISION DSEED, TEMP, R, QOT, QOB, QOH, BOTT, BOTB, BOTH, BOBB, MON00140
 1 BOBH, BOHH, DETT, DETB, DETH, DEBB, DEBH, DEHH, KDO, MON00150
 1 GCOUNT, GCOUNB, GCOUNH, COV (3), COVDES (3), GCCTAD, MON00160
 1 NCA(90), NCC(90), NCB(90), N2(90), NVA(90), NVC(90), NVB(90), MON00170
 1 NCTC(90), AA(90), AAD(90), AC(90), ACD(90), AB(90), ABD(90), MON00180
 1 CA(90), CC(90), CB(90) MON00190
 C INTEGER I, J, MODT, MODB, MODH, TT, TB, TH, BB, BH, HH, TOP, BRI, HOL, MON00200
 1 COUNTA, COUNTC, COUNTB, BCC, SURPLO, IP(3), CTADC MON00210
 1 LOGICAL T(100,100), B(100,100), H(100,100) MON00220
 COMMON T, B, H, DSEED, TEMP, R, QOT, QOB, QOH, BOTT, BOTB, BOTH, BOBB, BOBH, MON00230
 1 BOBH, DETT, DETB, DETH, DEBB, DEBH, DEHH, DELKDO, MON00240
 1 MODT, MODB, MODH, TT, TB, TH, BB, BH, HH, TOP, BRI, HOL, BCC MON00250
 C*****MON00260
 C INITIALIZE VARIABLES MON00270
 DSEED=1233449121D0 MON00280
 R=1.987D-3 MON00290
 SITES=1250D0 MON00300
 NAVE=1 MON00310
 NSTEP=63 MON00320
 NADS=1 MON00330
 TEMP=380.D0 MON00340
 BETA=1D0 MON00350
 DELT= 2.0D0/BETA MON00360
 DELTA=DELT MON00370
 KDO=1D16 MON00380
 DELKDO=DELT*KDO MON00390
 STCOEF=.9D0 MON00400
 PRESS=4.0D-11 MON00410
 C FLUX=STCOEF*PRESS*2.8129D8 !2D15 BRIDGE SITES/CM**2 MON00420
 FLUX=STCOEF*PRESS*4.4051D8 !FOR PD SURFACE ONLY! MON00430
 TOP=0 MON00440
 BRI=1 MON00450
 HOL=0 MON00460
 BCC=0 MON00470
 MODT=1 MON00480
 MODB=1 MON00490
 MODH=2 MON00500
 QOT=165.D0/1.80D0 MON00510
 QOB=31.5D0/1.5D0 MON00520
 QOH=180.D0/1.80D0 MON00530
 TT=0 MON00540
 TB=0 MON00550
 TH=0 MON00560
 BB=0 MON00570
 BH=0 MON00580
 HH=0 MON00590
 BOTT=.025D0 MON00600
 BOTB=.02D0 MON00610
 BOTH=.025D0 MON00620
 BOBB=.02D0 MON00630
 BOBH=.025D0 MON00640
 DETT=130.0D0 MON00650
 DETB=100.D0 MON00660
 DETH=256.0D0 MON00670
 DEBB=100.0D0 MON00680
 DEHH=100.0D0 MON00690

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DEBH=100.D0          MON00700
DEHH=130.0D0          MON00710
SURPLO=0              MON00720
IM=1                  MON00730
IP(1)=0               MON00740
COVDES(1)=1.00D0      MON00750
IP(2)=1               MON00760
COVDES(2)=0.5D0       MON00770
IP(3)=0               MON00780
COVDES(3)=0.5D0       MON00790
WRITE(7,*) DSEED      MON00800
WRITE(7,90) TEMP,PRESS,FLUX,STCOEF MON00810
WRITE(7,93) DELT,QOB,NAVE,NSTEP    MON00820
C*****MON00830
C*****MON00840
C*****MON00850
DO 200 NL=1,NAVE      MON00860
GCOUNA=0              MON00870
GCOUNC=0              MON00880
GCOUNB=0              MON00890
GCCTAD=0              MON00900
TIM=DELT              MON00910
C INITIALIZE COVERAGE EITHER RANDOM(1) OR ORDERED(0) BY LOCI. MON00920
IF(TOP.EQ.1) THEN      MON00930
  DATA COVTAI,LOCTA,COVTCI,LOCTC/0.1D0,1,1.0D0,1/ MON00940
  CALL INTTW(COVTAI,LOCTA,COVTCI,LOCTC) MON00950
ENDIF                 MON00960
IF(BRI.EQ.1) THEN      MON00970
  LOCBA=1              MON00980
  COVBAI=0.63D0         MON00990
C USE 1/2 OF COVBAI FOR INTBTW MON01000
  COVBAX=0.5D0*COVBAI MON01010
C DATA COVBAI,LOCBA,COVBCI,LOCBC/0.3000D0,1,0.0D0,1/ MON01020
  CALL INTBTW(COVBAX,LOCBA,COVBCI,LOCBC) MON01030
ENDIF                 MON01040
IF(HOL.EQ.1) THEN      MON01050
  DATA COVHAI,LOCHA,COVHCI,LOCHC/0.5D0,0,0.5D0,0/ MON01060
  CALL INTHTW(COVHAI,LOCHA,COVHCI,LOCHC) MON01070
ENDIF                 MON01080
C*****MON01090
C BEGIN MONTE CARLO STEP MON01100
DO 100 NM=1,NSTEP      MON01110
  TIM=TIM+DELT          MON01120
  COVA=((COVBAI*SITES)-GCOUNA)/SITES MON01130
  COVC=((COVBCI*SITES)-GCOUNC+GCCTAD)/SITES MON01140
  COVB=COVA+COVC        MON01150
  CTADC=0               MON01160
  COUNTA=0              MON01170
  COUNTC=0              MON01180
  COUNTC=0              MON01190
  NSVISA=0              MON01200
  NSVISC=0              MON01210
  NSVISB=0              MON01220
  AEA=0                  MON01230
  AEC=0                  MON01240
  AEAD=0                 MON01250
  AECD=0                 MON01260
  NJMP2=0                 MON01270
C*****MON01280
C ADSORPTION SEGMENT AND EQUILIBRATE BY DIFFUSION MON01290
IF(NADS.EQ.1) THEN      MON01300
  SAMFRB=SITES*(2D0-COVB)*FLUX*DELTA      MON01310
  SAMFRB=SITES*FLUX*DELTA*(1D0-COVB)/(1D0-COVB+COVB*0.10D0) MON01320
  DO 59 NN=1,10000          MON01330
    CALL ADSBRI(SAMFRB,CTADC) MON01340
59  CONTINUE             MON01350
  ENDIF                 MON01360
C DO 60 KK=1,200000        MON01370
C   DO 68 LL=1,2           MON01380
C     CALL DIFBTW(NJMP2)    MON01390
C8  CONTINUE             MON01400
60

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C*****MON01410
COVA=((COVBA1*SITES)-GCOUNA)/SITES      MON01420
COVC=((COVBC1*SITES)-GCOUNC+GCCTAD+CTADC)/SITES MON01430
COVB=COVA+COVC                           MON01440
COV(1)=COVB                            MON01450
COV(2)=COVB                            MON01460
COV(3)=COVB                            MON01470
IF(SURPLO.EQ.1) THEN                   MON01480
  CALL SURMAP(COVT,COVB,COVH,COV,COVDES,IP,IM) MON01490
ENDIF
IF(MODB.EQ.1) THEN                   MON01500
  SAMFDA=SITES*COVA                   MON01510
  SAMFDC=SITES*COVC                   MON01520
ELSE
  SAMBRI=SITES*(COVB**2)*.5D0        MON01530
ENDIF
C*****MON01540
C
C*****MON01550
C      DESORPTION SEGMENT             MON01560
DO 55 N=1,10000                         MON01570
  CALL DESBTW(SAMFDA,SAMFDC,COUNTA,COUNTC,AEAD,AECD,AEA,AEC, MON01580
  NSVISA,NSVISC,COVB)                  MON01590
1      CONTINUE                           MON01600
C*****MON01610
C
C*****MON01620
C      CALCULATE OUTPUT VARIABLES AND WRITE OUT RESULTS      MON01630
C      OUTPUT COVERAGES TO CORRESPOND TO BEGINNING OF THE TIME STEP MON01640
COVA=((COVBA1*SITES)-GCOUNA)/SITES      MON01650
COVC=((COVBC1*SITES)-GCOUNC+GCCTAD)/SITES MON01660
COVB=COVA+COVC                           MON01670
COUNTB=COUNTA+COUNTC                     MON01680
NSVISB=NSVISA+NSVISC                     MON01690
GCOUNA=GCOUNA+COUNTA                     MON01700
GCOUNC=GCOUNC+COUNTC                     MON01710
GCOUNB=GCOUNB+COUNTB                     MON01720
GCCTAD=GCCTAD+CTADC                     MON01730
RATEA=COUNTA/(DELT*SITES*MODB)          MON01740
RATEC=COUNTC/(DELT*SITES*MODB)          MON01750
RATEB=COUNTB/(DELT*SITES*MODB)          MON01760
RATADC=CTADC/(DELT*SITES*MODB)          MON01770
IF(COUNTB.NE.0) THEN                   MON01780
  AEBD=(AEAD+AECD)/COUNTB*MODB        MON01790
ENDIF
IF(NSVISB.NE.0) THEN                   MON01800
  AEB=(AEA+AEC)/(NSVISA+NSVISC)        MON01810
ENDIF
IF(NSVISA.NE.0) THEN                   MON01820
  AEA=AEA/(NSVISA)                     MON01830
ENDIF
IF(NSVISC.NE.0) THEN                   MON01840
  AEC=AEC/(NSVISC)                     MON01850
ENDIF
IF(COUNTA.NE.0) THEN                   MON01860
  AEAD=AEAD/COUNTA*MODB               MON01870
ENDIF
IF(COUNTC.NE.0) THEN                   MON01880
  AECD=AECD/COUNTC*MODB               MON01890
ENDIF
RATNET=RATADC-RATEB                   MON01900
C      WRITE(7,90) RATEA,COVA,AEA,AEAD,COUNTA,NSVISA      MON01910
C      WRITE(7,90) RATEC,COVC,AEC,AECD,COUNTC,NSVISC      MON01920
C      WRITE(7,90) RATEB,COVB,AEB,AEBD,COUNTB,NSVISB      MON01930
C      WRITE(7,91) RATNET,RATADC,TIM,CTADC,NJMP2          MON01940
C      WRITE(7,*) ' '
90      FORMAT(1X,4(1PE10.3,1X),3(16,1X))      MON01950
91      FORMAT(1X,3(1PE10.3,1X),3(16,1X))      MON01960
92      FORMAT(1X,5(1PE10.3,1X))                MON01970
93      FORMAT(1X,2(1PE10.3,1X),2(16,1X))      MON01980
C      DETERMINE WEIGHTED CONTRIBUTION TO AVERAGE IN A 1-D VECTOR MON01990
C

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N2 (NM) =NJMP2+N2 (NM)                               MON02120
CA (NM) =COVA*SITES+CA (NM)                         MON02130
AA (NM) =AEA*NSVISA+AA (NM)                         MON02140
AAD (NM) =AEAD*COUNTA+AAD (NM)                      MON02150
NCA (NM) =COUNTA+NCA (NM)                           MON02160
NVA (NM) =NSVISA+NVA (NM)                           MON02170
CC (NM) =COVC*SITES+CC (NM)                         MON02180
AC (NM) =AEC*NSVISC+AC (NM)                         MON02190
ACD (NM) =AECD*COUNTC+ACD (NM)                      MON02200
NCC (NM) =COUNTC+NCC (NM)                           MON02210
NVC (NM) =NSVISC+NVC (NM)                           MON02220
CB (NM) =COVB*SITES+CB (NM)                         MON02230
AB (NM) =AEB*NSVISB+AB (NM)                         MON02240
ABD (NM) =AEBD*COUNTB+ABD (NM)                      MON02250
NCB (NM) =COUNTB+NCB (NM)                           MON02260
NVB (NM) =NSVISB+NVB (NM)                           MON02270
NCTC (NM) =CTADC+NCTC (NM)                          MON02280
100  CONTINUE                                         MON02290
200  CONTINUE                                         MON02300
C   DETERMINE AVERAGES
      WRITE (7,*) 'AVERAGES'
      TIM=-DELT
      DO 300 NZ=1, NSTEP
      TIM=TIM+DELT
      DAVE=DBLE (NAVE)
      XRA=NCA (NZ) / (DELT*SITES*MODB*DAVE)
      XCA=CA (NZ) / (SITES*DAVE)
      IF (NVA (NZ) .EQ. 0.0) THEN
          XAA=0.0
      ELSE
          XAA=AA (NZ) / (NVA (NZ) )
      ENDIF
      IF (NCA (NZ) .EQ. 0.0) THEN
          XAAD=0.0
      ELSE
          XAAD=AAD (NZ) / (NCA (NZ) )
      ENDIF
      XNCA=NCA (NZ) / DAVE
      XRC=NCC (NZ) / (DELT*SITES*MODB*DAVE)
      XCC=CC (NZ) / (SITES*DAVE)
      IF (NVC (NZ) .EQ. 0.0) THEN
          XAC=0.0
      ELSE
          XAC=AC (NZ) / (NVC (NZ) )
      ENDIF
      IF (NCC (NZ) .EQ. 0.0) THEN
          XACD=0.0
      ELSE
          XACD=ACD (NZ) / (NCC (NZ) )
      ENDIF
      XNCC=NCC (NZ) / DAVE
      XRB=NCB (NZ) / (DELT*SITES*MODB*DAVE)
      XCB=CB (NZ) / (SITES*DAVE)
      IF (NVB (NZ) .EQ. 0.0) THEN
          XAB=0.0
      ELSE
          XAB=AB (NZ) / (NVB (NZ) )
      ENDIF
      IF (NCB (NZ) .EQ. 0.0) THEN
          XABD=0.0
      ELSE
          XABD=ABD (NZ) / (NCB (NZ) )
      ENDIF
      XNCB=NCB (NZ) / DAVE
      XNCTC=NCTC (NZ) / DAVE
      XRADC=NCTC (NZ) / (DELT*SITES*MODB*DAVE)
      XRN=XRADC-XRB
      XN2=N2 (NZ) / DAVE
      WRITE (7, 92) XRA, XCA, XAA, XAAD, XNCA
      WRITE (7, 92) XRC, XCC, XAC, XACD, XNCC
      WRITE (7, 92) XRB, XCB, XAB, XABD, XNCB

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300 WRITE(7,92) XRNET,XRADC,TIM,XNCTC,XN2
 WRITE(7,*)
 CONTINUE
 STOP
 END

MON02830
MON02840
MON02850
MON02860
MON02870

END

DATE FILMED

11/19/90

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