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**An Introduction to Technetium in the  
Gaseous Diffusion Cascades**

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September 1996

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## LIST OF ACRONYMS

atm	atmosphere (i.e, pressure = 14.7 psia)
Ci	curie
F <sub>2</sub>	fluorine
(g)	gas
GDP	gaseous diffusion plant
gm	gram
HEU	highly enriched uranium
HF	hydrogen fluoride
kg	kilogram
(l)	liquid
MgF <sub>2</sub>	magnesium fluoride
μg	microgram
mm Hg	millimeters of mercury (pressure)
MW	megawatt
ORGDP	Oak Ridge Gaseous Diffusion Plant
PGDP	Paducah Gaseous Diffusion Plant
PORTS	Portsmouth Gaseous Diffusion Plant
ppm	parts per million
psia	pounds per square inch, absolute
RU	reprocessed uranium
(s)	solid
<sup>99</sup> Tc / Tc	technetium, isotope 99
TcO <sub>3</sub> F	pertechnetyl fluoride
TcO <sub>4</sub> F	technetium peroxyfluoride
TcO <sub>2</sub> F <sub>3</sub>	technetium dioxotrifluoride
TcF <sub>6</sub>	technetium hexafluoride
HTcO <sub>4</sub>	pertechnetic acid
TcO <sub>2</sub>	technetium dioxide
Tc <sub>2</sub> O <sub>7</sub>	technetium heptoxide
<sup>235</sup> U	uranium, isotope 235
<sup>238</sup> U	uranium, isotope 238
UF <sub>6</sub>	uranium hexafluoride
UF <sub>5</sub>	uranium pentafluoride
UF <sub>4</sub>	uranium tetrafluoride
UO <sub>2</sub> F <sub>2</sub>	uranyl fluoride
UO <sub>3</sub>	uranium oxide

## SUMMARY

The radioisotope technetium-99 ( $^{99}\text{Tc}$ ) was introduced into the gaseous diffusion plants (GDP) as a contaminant in uranium that had been reprocessed from spent nuclear reactor fuel.  $^{99}\text{Tc}$  is a product of the nuclear fission of uranium-235 ( $^{235}\text{U}$ ). The significantly higher emitted radioactivity of  $^{99}\text{Tc}$  generates concern in the enrichment complex and warrants increased attention (1) to the control of all site emissions, (2) to worker exposures and contamination control when process equipment requires disassembly and decontamination, and (3) to product purity when the enriched uranium hexafluoride ( $\text{UF}_6$ ) product is marketed to the private sector.

A total of 101,268 metric tons of RU (~96% of the total) was fed at the Paducah Gaseous Diffusion Plant (PGDP) between FY1953 and FY1976. An additional 5600 metric tons of RU from the government reactors were fed at the Oak Ridge Gaseous Diffusion Plant (ORGDP), plus an approximate 500 tons of foreign reactor returns. Only a small amount of RU was fed directly at the Portsmouth Gaseous Diffusion Plant (PORTS). The slightly enriched PGDP product was then fed to either the ORGDP or PORTS cascades for final enrichment. Bailey estimated in 1988 that of the 606 kg of Tc received at PGDP from RU, 121 kg was subsequently re-fed to ORGDP and 85 kg re-fed to PORTS.

By early estimates, approximately 85% of the total Tc received at Paducah was presumed fed to the cascade. Estimates from various sources have speculated that as much as 65-70% of the Tc introduced at PGDP was initially adsorbed (or condensed) on the cascade equipment as a non-volatile species. Cascade observations provided evidence that unfortunately the Tc did not remain non-volatile. Operational analysis at the GDP's have indicated an increasing concentration of  $^{99}\text{Tc}$  as the process gas moves toward the isotopic top of the plant. The remaining 30-35% of the  $^{99}\text{Tc}$  fed was believed to have been distributed to the  $\text{UF}_6$  product, chemically trapped, vented to the atmosphere, or discharged with aqueous decontamination streams. Sufficient operational evidence exists that clearly indicate the continued presence of unacceptable levels of  $^{99}\text{Tc}$  in the operating GDP's at both Paducah and Portsmouth.



The typical trace levels of  $^{99}\text{Tc}$  compounds in the operating GDP's is below the minimum detectable limit for any of the process gas analyzers. Consequently, it can not be definitively stated which Tc compounds are present in the operational cascades. The only gas phase Tc compound that has been reported to have been detected in the cascade gas stream is the pertechnetyl fluoride,  $\text{TcO}_3\text{F}$ , which was detected in the purge cascade during treatments to unplug barrier. The compounds of Tc that should be considered as potential cascade vapor phase compounds would consist of  $\text{TcO}_3\text{F}$ ,  $\text{HTcO}_4$ ,  $\text{TcOF}_4$  and  $\text{TcF}_6$ . The oxides  $\text{Tc}_2\text{O}_7$  and  $\text{TcO}_2$  could also possibly exist as condensed species, along with the liquid or solid pertechnetic acid,  $\text{HTcO}_4$ , and the oxyfluoride  $\text{TcO}_2\text{F}_3$ .

The unique nature of gas-phase technetium compounds has limited most of the investigation of these compounds to the gaseous diffusion complex. The trace levels of Tc present in the cascade, and the absence of regulatory concern at these levels of Tc for most of the operational lifetime, has inhibited the support for quantitative technetium chemistry by knowledgeable scientists. Consequently, significant gaps exist in the knowledge of the physical and chemical properties of the technetium compounds of interest to the gaseous diffusion cascade.

The reaction of the technetium fluorides with the uranium fluorides has been widely speculated but not quantitatively confirmed. The absence of thermodynamic or kinetic data for the volatile Tc fluorides currently prevents the calculation of confident estimates of technetium reactions with cascade uranium compounds. The hydrolysis of the cascade relevant technetium fluorides with water is well documented in the open literature, and the existence of these compounds in the cascade is supported by anecdotal reports, but not quantitative analysis. The undetermined chemical affinities of the various potential Tc compounds to the other cascade chemical compounds and cascade materials of construction prevent a definitive answer to questions of adsorption, absorption, and condensation of the Tc. The mechanism for any of the potential non-volatile species has not been defined, other than condensation. It is well known that technetium will collect in areas of low relative temperature, due to the lower vapor pressure of the technetium compounds. Chemical property data sufficient to support more than speculation about the chemical stability of these potential Tc species in the operating cascades does not currently exist in the available literature. The acquisition of kinetic and thermodynamic information

on relevant cascade Tc compounds would permit predictions of cascade Tc behavior.

Various methods and operational schemes have been employed to control Tc in the GDP's, with varying success. The control of Tc in the enrichment complex involves methods necessary to maintain low Tc levels at (1)  $UF_6$  product withdrawal stations, (2) air emission vent streams, and (3) aqueous waste streams, and preferably (4) control in the isotopic and purge gas streams. Ideally, Tc control would involve the isolation of the Tc in one location, and the subsequent recovery of all the cascade material. This report presents an introduction to some of the methods employed to control the gas-phase technetium compounds in gaseous diffusion cascade operations.

## 1. INTRODUCTION

The radioisotope technetium-99 ( $^{99}\text{Tc}$ ) was introduced into the gaseous diffusion plants (GDP) as a contaminant in uranium that had been reprocessed from spent nuclear reactor fuel.<sup>1</sup>  $^{99}\text{Tc}$  is a product of the nuclear fission of uranium-235 ( $^{235}\text{U}$ ) with a thermal fission yield of ~6%.<sup>2</sup> Calculations indicate that a uranium reactor operating at 1000 MW would produce 25 gm of  $^{99}\text{Tc}$  per 1 kg of plutonium.<sup>3</sup>  $^{99}\text{Tc}$  is an emitter of weak beta activity ( $E_{\text{max}} = 0.292 \text{ MeV}$ ) with a shorter half-life (210,000 yr) than the uranium isotopes, and consequently a higher specific activity (0.0169 Curies (Ci)/gm) than uranium (see Tables 1 and 2).<sup>3,4,5</sup> For example, 100 gm of 5% enriched U with 5 ppm  $^{234}\text{U}$  impurity

**Table 1. Radioactivity of  $^{99}\text{Tc}$  and U isotopes**

	$^{99}\text{Tc}$	$^{234}\text{U}$	$^{235}\text{U}$	$^{238}\text{U}$
decay mode	beta	alpha	alpha	alpha
half-life (years)	210,000	246,000	7.04 E8	4.51 E9
specific activity (Ci/gm)	0.0169	.00621	2.16 E-6	3.33 E-7
activity in 100 gm (Ci)	1.69	0.621	2.16 E-4	3.33 E-5

**Table 2. Properties of  $^{99}\text{Tc}$  and  $^{235}\text{U}$**

Property	Uranium-235	Technetium-99
Atomic Number	92	43
Isotope	235	99
Atomic Weight	235.04 4	98.906
Density (g/cm <sup>3</sup> )	18.95	11.487
Melting point (°C)	1132	2440
Boiling point (°C)	3818	4900
Valence States		
stable compounds	+6, +5, +4, +3	+7, +4, +6
complexes		+7 to -1

would contain  $7 \times 10^{-5}$  Ci (70 microCi), while 100 gm of  $^{99}\text{Tc}$  would contain 1.69 Ci. It is this *significantly higher emitted radioactivity of  $^{99}\text{Tc}$  that generates the concern in the enrichment complex* and warrants increased attention (1) to the control of all site emissions, (2) to worker exposures and contamination control when process equipment requires disassembly and decontamination, and (3) to product purity when the enriched uranium hexafluoride ( $\text{UF}_6$ ) product is marketed to the private sector.

## 2. BACKGROUND

The government reactors at Hanford and Savannah River accounted for 96% of the total reprocessed uranium (RU) that was used as  $\text{UF}_6$  feed at the three GDP's. The remaining 4% was foreign commercial reactor fuel, primarily French. *A total of 101,268 metric tons of RU (~96% of the total) was fed at PGDP between FY1953 and FY1976.*<sup>6,7</sup> An additional 5600 metric tons of RU from the government reactors were fed at ORGDP, plus an approximate 500 tons of foreign reactor returns.<sup>6</sup> Only a small amount of RU was fed directly at PORTS.<sup>6,7,8,9</sup>

It has been estimated that *greater than 99% of the  $^{99}\text{Tc}$  introduced into the GDP's was from the government's military reactors.*<sup>6</sup> The chemical characteristics of  $^{99}\text{Tc}$  through the solvent extraction process used to separate the spent uranium and plutonium allowed a large fraction of the Tc ( $\sim 7 \text{ ppm} \pm 30\%$  on a total U basis for the depleted reactor tails<sup>7</sup>) to track with the uranium through the spent fuel reprocessing, particularly in the lower acidity PUREX process. The Paducah Gaseous Diffusion Plant received the oxide,  $\text{UO}_3$ , and produced the  $\text{UF}_{6(g)}$  that was then fed into the enrichment cascades. The operating conditions of the  $\text{UF}_6$  conversion process were also favorable to the conversion of a volatile Tc compound. An estimated 95% of the Tc in the  $\text{UO}_3$  was converted with the  $\text{UF}_6$ , with about 90% of the Tc in the  $\text{UF}_6$  feed vaporized into the cascade. By early estimates, *approximately 85% of the total Tc received at Paducah was presumed fed to the cascade.*<sup>7</sup> Significant quantities of the  $^{99}\text{Tc}$  were determined to have migrated to the top of the PGDP cascade and tended to accumulate in the same cell in which the  $\text{UF}_6$  product was withdrawn.<sup>9</sup> The slightly enriched PGDP product was then fed to either the ORGDP or PORTS cascades for final enrichment. *Bailey has estimated that of the 606 kg of Tc received at PGDP from RU, 121 kg was subsequently re-fed to ORGDP and 85 kg re-fed to PORTS.*<sup>9</sup>

An accurate material balance for the Tc in the cascades has not been published, and would be difficult to adequately compile with current knowledge. Estimates from various sources have speculated that *as much as 65-70% of the Tc introduced at PGDP was initially adsorbed (or condensed) on the cascade equipment.* Saraceno estimated  $\sim 80\%$  of the Tc fed at PORTS remained in the cascade in 1978.<sup>8</sup> Operational observations

indicate that all the  $^{99}\text{Tc}$  did not remain adsorbed, but rather tended to migrate over time depending on the cascade operations. Analyzed barrier that have been removed from cascade service have indicated an *increasing concentration of  $^{99}\text{Tc}$  as the gas moves toward the isotopic top of the plant.*<sup>8,10</sup> A considerable fraction of the adsorbed  $^{99}\text{Tc}$  was thought to have been removed with the aqueous decontamination processes during the two major cascade improvement programs at PGDP. The remaining 30-35% of the  $^{99}\text{Tc}$  fed was believed to have been distributed to the  $\text{UF}_6$  product, chemically trapped, or vented to the atmosphere. However, *sufficient evidence exists that clearly indicate the continued presence of high levels of  $^{99}\text{Tc}$  in the operating GDP's at both Paducah and Portsmouth.*

### 3. POTENTIAL CASCADE TECHNETIUM COMPOUNDS

The detection of vapor phase Tc compounds has proven to be a difficult task. The dominant method of technetium detection in the gaseous diffusion complex is liquid scintillation counting of an aqueous solution to determine the beta activity of the sample. However, this method is an elemental analysis and does not provide information on the specific compound. *The typical trace levels of  $^{99}\text{Tc}$  in the operating GDP's is below the minimum detectable limit for any of the cascade gas analyzers.* The occasional laboratory evaluations have provided some usable spectroscopic evidence, but the focus of these efforts has usually been to quickly resolve a cascade operational concern. Sufficient resources have not been provided to adequately define the gas phase detection capabilities of cascade technetium compounds. Consequently, *it can not be definitively stated which Tc compounds are present in the operational cascades.*

The initial researchers at PGDP expected that they were dealing with the technetium hexafluoride,  $\text{TcF}_6$ . The Tc was found to quantitatively accompany the U in the first two steps of the  $\text{UF}_6$  conversion process at PGDP.<sup>1</sup> The fluorination of the resultant  $\text{UF}_4$  to  $\text{UF}_6$  gas was presumed to fluorinate the Tc to  $\text{TcF}_6$ . The excellent work performed to define the separation of Tc from  $\text{UF}_6$  with magnesium fluoride ( $\text{MgF}_2$ ) chemical sorbent traps assumed the compound to be  $\text{TcF}_6$ , but stated "the compound form of the volatile technetium fluoride has not been established".<sup>11</sup> The preparation and isolation of  $\text{TcF}_6$  was not reported until January of 1961 by Selig et al. at the Argonne National Laboratory.<sup>12</sup> A vapor phase infrared spectrum has been reported. The reported vapor pressure of  $\text{TcF}_6$  (~ 1 atm at 55°C) is slightly less than that of  $\text{UF}_6$  at cascade operating conditions.<sup>13</sup>

*The only gas phase Tc compound that has been reported to have been detected in the operating cascade gas stream is the pertechnetyl fluoride,  $\text{TcO}_3\text{F}$  (Figure 1).<sup>14,15</sup> The vapor pressure reported for  $\text{TcO}_3\text{F}$  (~ 1 atm at 100°C) is less than the hexafluorides.<sup>16</sup> The recovery of a barrier plug (i.e., permeability reduction) at K-402.9 (ORGDP purge cascade) in 1979 that was assumed to be caused by a condensed Tc compound volatilized detectable quantities of  $\text{TcO}_3\text{F}$  mixed with  $\text{UF}_6$  and  $\text{HF}$ <sup>14,15</sup>. The Tc identification was performed with an on-line dispersive infrared spectrometer. The level of Tc volatilized*

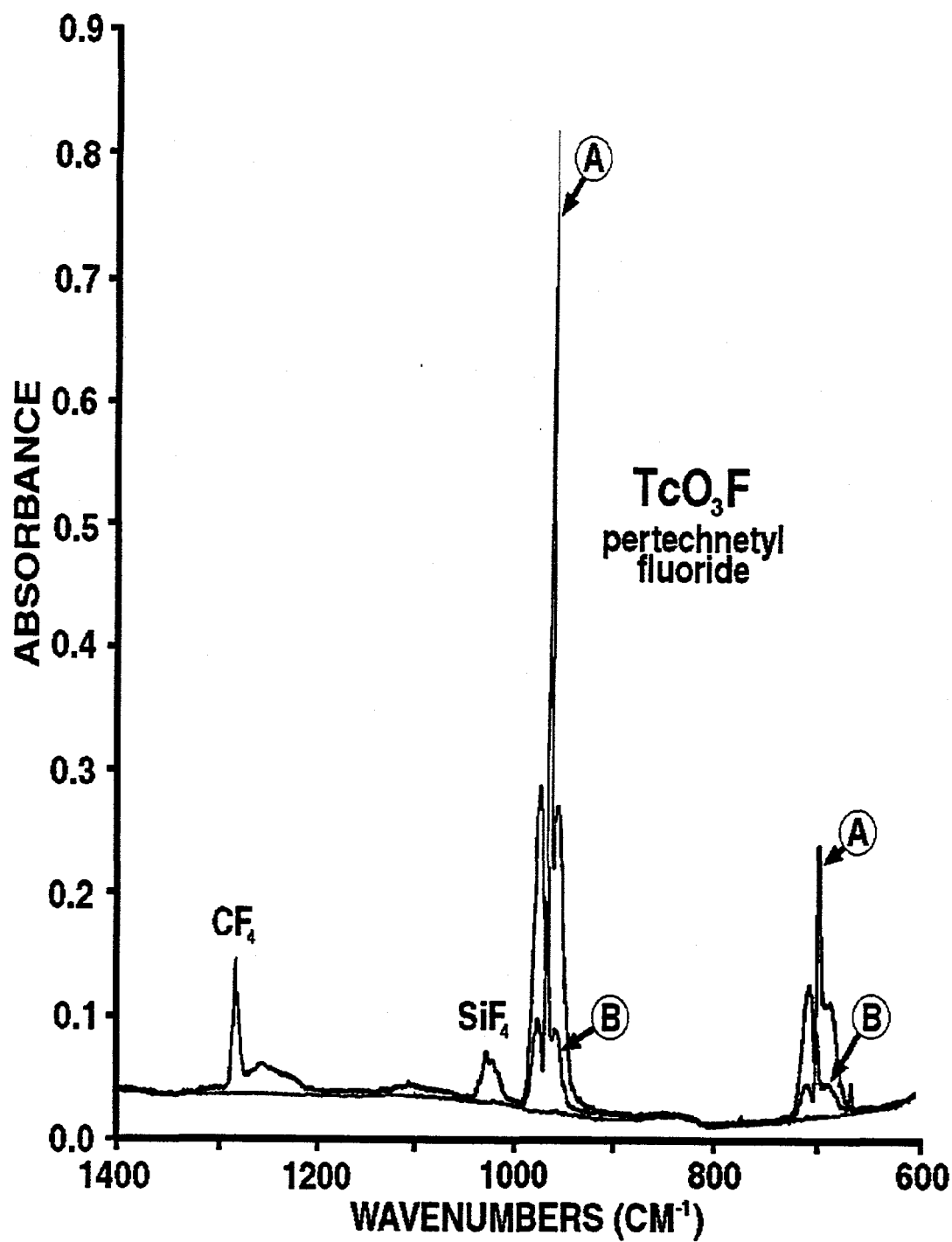
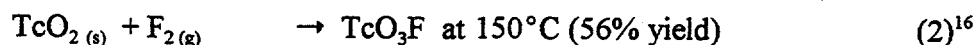
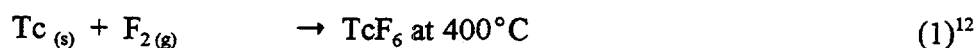


Fig. 1. Gas Phase Infrared Spectrum of  $\text{TcO}_3\text{F}$  obtained at Portsmouth. (A) 12.7 mm Hg (torr) pressure, (B) 3.4 mm Hg pressure in a  $10\text{ cm}^{-1}$  gas cell with silver chloride ( $\text{AgCl}$ ) optical windows. Lower trace is background at vacuum. Trace impurities of  $\text{CF}_4$  and  $\text{SiF}_4$  were also detected.



with circulating hot air treatments was reported to range 1—11 mole % across the stages.<sup>14</sup> The  $\text{TcO}_3\text{F}$  was cold trapped and removed from the cascade gas stream. Analysis of the collected material yielded inconsistent results. Information from PGDP and PORTS have provided additional spectroscopic evidence of the existence of  $\text{TcO}_3\text{F}$  in areas of the operating cascades. This report will assume that  $\text{TcO}_3\text{F}$  is the dominant gas phase species in the operating cascade.

The laboratory preparation of the relevant volatile technetium fluorides can be summarized as:



Selig reported a yield of ~90% for reaction (1) when excess  $\text{F}_{2(g)}$  was used.

Reaction (3) was reported to have near quantitative yield ( $\approx 100\%$ ) with excess fluorine.

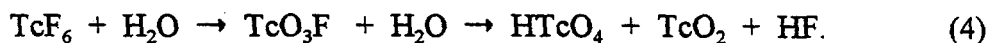
The preparation of  $\text{TcO}_3\text{F}$  is also possible from the reaction of anhydrous HF with both the pertechnetate anion, as  $\text{NH}_4^+\text{TcO}_4^-$ ,<sup>17</sup> and the heptoxide,  $\text{Tc}_2\text{O}_7$ .<sup>18</sup> The addition of an excess of the hexafluoride of xenon,  $\text{XeF}_6$ , to the latter solution yielded the technetium dioxotrifluoride ( $\text{TcO}_2\text{F}_3$ ).<sup>18</sup>

Sufficient kinetic or thermodynamic data does not exist that will permit accurate assessment of the fluorination of Tc at conditions other than those reported. The fluorination of Tc with other cascade fluorinating agents, such as chlorine trifluoride ( $\text{ClF}_3$ ), has only recently been reported.<sup>19</sup>

*The reaction of the technetium fluorides with the uranium fluorides has been widely speculated but not quantitatively confirmed.*<sup>10,14,15,20</sup> The absence of thermodynamic or kinetic data for the volatile Tc fluorides currently prevents the calculation of confident estimates of technetium reactions with cascade uranium compounds. The acquisition of this kinetic and thermodynamic information would permit predictions of cascade Tc behavior. Although experimental evaluations to determine the

chemical reactivity of technetium and uranium compounds would be straightforward to conduct, no laboratory work has been commissioned to date.

*The reaction of the cascade relevant technetium fluorides with water is well documented in the literature.*<sup>2,3,4,5,10,12,15,16</sup> The hydrolysis is fairly rapid and will continue to a final product mixture of HTcO<sub>4</sub>, TcO<sub>2</sub>, and HF if allowed to go to completion. The hydrolysis pathway of the technetium fluorides can be simply represented as



The pertechnetetic acid (HTcO<sub>4</sub>) has been described in the literature as a dark, red fuming liquid by several investigators. A significant number of anecdotal reports of "pouring Tc" from cascade instrument lines exist. Observations of a fuming, viscous brownish-red material with high beta activity suggests the presence of this acid, or perhaps a mixture of it, in low(er) temperature copper lines. HTcO<sub>4</sub> has a relatively low vapor pressure (61 torr at 100°C) at temperatures typical to the cascade,<sup>21</sup> and could also easily migrate as a gas phase compound. A vapor pressure curve for TcO<sub>2</sub> has not been reported, but Peacock states that it is only slightly volatile at 900°C.<sup>5</sup> The dioxide, TcO<sub>2</sub>, or its hydrate, TcO<sub>2</sub>·xH<sub>2</sub>O, would not be expected to be volatile at cascade conditions.

The subsequent dissociation and disproportionation of the acid to a mixture of the dioxide and the heptoxide in atmosphere is also well documented, and may account for conflicting analytical reports of recovered Tc deposits. The vapor pressure for the heptoxide, Tc<sub>2</sub>O<sub>7(s)</sub>, was reported by Smith<sup>21</sup> to be 0.1 mm Hg at 100°C. The oxy tetrafluoride, TcOF<sub>4(s)</sub>, has a reported vapor pressure with volatility slightly less than the pertechnetetic acid (HTcO<sub>4</sub>) at temperatures of interest.<sup>5</sup> The technetium dioxotrifluoride (TcO<sub>2</sub>F<sub>3</sub>) has been prepared and observed in mass spectrometry studies but a vapor pressure curve has not been reported.<sup>10,15,18,22</sup> The melting point was reported to be 200°C and it is known to be volatile at high temperature under high vacuum conditions.

*The compounds of Tc that should be considered as potential cascade vapor phase compounds would consist of TcO<sub>3</sub>F, HTcO<sub>4</sub>, TcOF<sub>4</sub> and TcF<sub>6</sub>.* The oxides Tc<sub>2</sub>O<sub>7</sub> and TcO<sub>2</sub>·xH<sub>2</sub>O could also possibly exist as condensed species, along with the liquid or solid pertechnetetic acid, HTcO<sub>4</sub>, and the oxyfluoride TcO<sub>2</sub>F<sub>3</sub>. Figure 2 shows the reported vapor

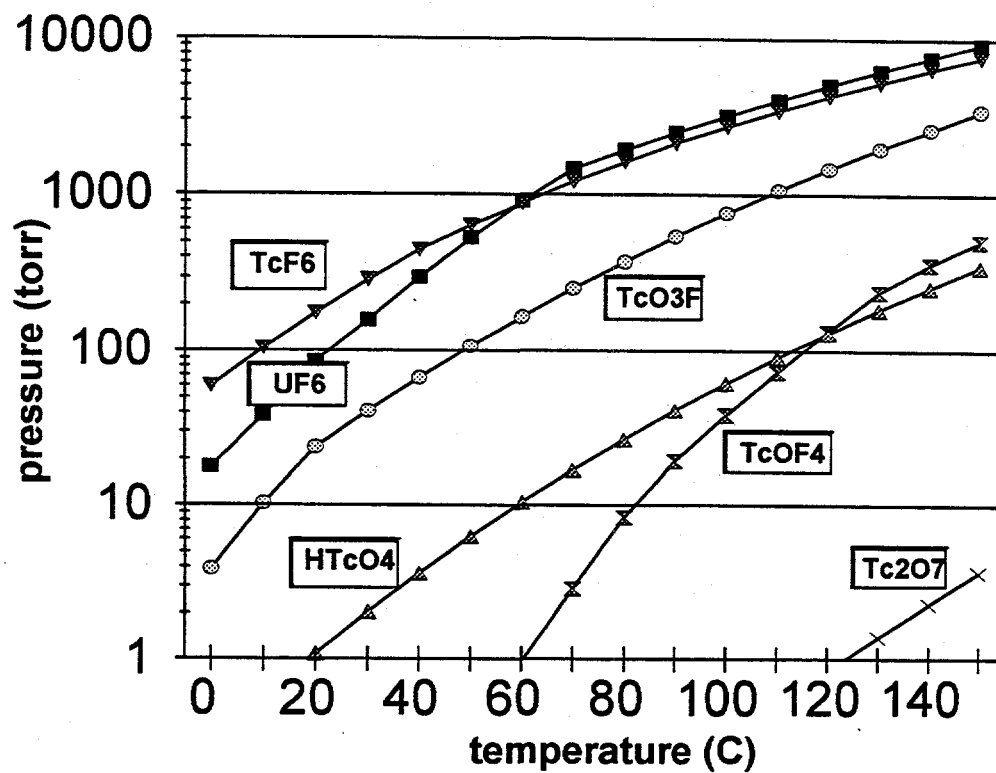


Fig. 2. Vapor Pressure of Potential Cascade Technetium Compounds and  $\text{UF}_6$ .

pressure of the potential Tc compounds compared to  $\text{UF}_6$ . The published chemical property data for these compounds is compiled in Table 3. *Chemical property data sufficient to support more than speculation about the chemical stability of these potential Tc species in the operating cascades does not currently exist in the available literature.*

Table 3. Property data for possible cascade Technetium compounds and Uranium Hexafluoride.

	TcO <sub>3</sub> F	TcF <sub>6</sub>	TcO <sub>2</sub> F <sub>3</sub>	TcOF <sub>4</sub>	HTcO <sub>4</sub>	TcO <sub>2</sub>	Tc <sub>2</sub> O <sub>7</sub>	UF <sub>6</sub>
molecular weight	166	213	188	191	164	131	310	352
melting point, C	18.3	37.4	200	133			119.5	64 <sup>a</sup>
boiling point, C	~100	55.3		165			310.6	56.5 <sup>b</sup>
vapor pressure								
@ 50°C	107	635		0.3	6			526
@ 100°C	766	2739		38	61			3200
(mm Hg)							0.1	
valence	+7	+6	+7	+6	+7	+4	+7	+6

<sup>a</sup>UF<sub>6</sub> has a triple point pressure in excess of 1 atmosphere and cannot be melted at atmospheric pressure. The triple point, measured at 1137 mm. Hg, is presented instead.

<sup>b</sup>UF<sub>6</sub> has a triple point pressure in excess of 1 atmosphere and cannot be melted at atmospheric pressure. Consequently, a normal boiling point cannot be obtained. The sublimation point, the temperature at which the solid vapor pressure obtains 760 mm Hg is typically presented.

#### 4. TECHNETIUM IN THE OPERATING CASCADE

*The volatile technetium compounds are intermediate in weight to  $UF_6$  and nitrogen.* The fraction of technetium that remains volatile will move promptly up the diffusion cascade. Significant evidence exists from all three GDP's that significant levels of *Tc tends to concentrate in the top of the isotopic and purge cascade stages.* A "bubble", or pocket can then form between the  $UF_6$  and the "lights" (gases of low molecular weight). The ability of cascade operations to maintain a sufficient cascade length with proper temperature and pressure profiles appears to be critical to the control of the Tc in regards to compliance with air emission regulations and current  $UF_6$  product specification.

The undetermined chemical affinities of the various potential Tc compounds with the other cascade chemical compounds and cascade materials of construction prevent a definitive answer to questions of adsorption, absorption, and condensation of the Tc. Basic knowledge of Tc, U, and fluorine chemistry can fuel speculation, but until focused evaluations of these chemical systems are performed these assumptions must be considered as only that. Experimental investigations of relevant cascade reactions could be rather simply accomplished with existing laboratory systems.

It is well known that a volatile Tc compound, presumably  $TcO_3F$ , moves through the diffusion cascade. It is also known that volatile Tc species are either sorbed, condensed, complexed, or reacted with the material surfaces in the cascade, or other chemical compounds (i.e.,  $UF_6$ ,  $UF_5$ ,  $UF_4$ ,  $UO_2F_2$ , or moisture). Any, or all, of these methods can result in non-volatile Tc species at given conditions. The exposed surfaces of the cascade are fluoride surfaces, from the initial fluorination film treatments and the continued fluoride gas exposure. Some of the Tc compounds are more attracted to certain fluoride surfaces than others, and various available fluoride surfaces exist in the cascade that could sorb trace levels of Tc at the process conditions.

For example, it is widely known that a volatile cascade Tc *can be selectively removed from the  $UF_6$  gas stream with  $MgF_2$  chemical traps at  $\sim 140^\circ F$ , with only trace  $UF_6$  removed.* In comparison, sodium fluoride (NaF) will quantitatively trap  $UF_6$  with (typically) only trace levels of the cascade Tc removed at  $200^\circ F$ . (However, some investigators have reported the removal of gas phase Tc compounds by NaF.) By

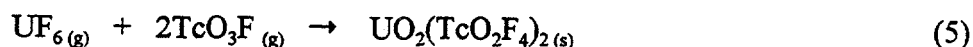
maintaining the higher temperature of the NaF trap, one can also reduce the ability of the trap to complex HF and preserve the trap longevity while maintaining the purity of recovered UF<sub>6</sub>. None of the other major fluoride gases in the cascade are readily sorbed to either of these trapping agents at these conditions. The relationship between chemical sorbents and gases is a complex mixture of thermodynamic relationships and kinetic potential. To understand the sorbent capabilities to a given compound, these properties must be known and the system's behaviour calculated at the application's operating conditions. To completely understand these relationships with technetium compounds would require some applied experimental evaluation at these conditions. Good work has been performed trapping gas phase Tc on selected sorbents, but neither the Tc compound or the sorbing mechanism is currently known.<sup>1,8,11</sup>

*The mechanism for any of the potential alternatives to non-volatility has not been defined*, other than condensation. Trace amounts of Tc are continually moving, and continually "sitting down" in the cascade. The continued presence of Tc in Paducah product feed adds additional Tc yearly to the Portsmouth cascade inventory. Recent observations at Portsmouth have also shown increased levels of Tc believed due to the cleanup fluorination treatments of the HEU Shutdown Program. Occasional elevated levels of Tc are observed that are attributed to the increase in ambient temperature and/or humidity. A focused attempt to develop correlaries for these elevated levels needs to be pursued to attempt adequate explanation of the conditions responsible. Perhaps with additional theoretical and experimental investigations the "plating out of Tc" in the cascade could be defined in such a manner as to allow control of the movement of Tc during the various cascade operations, with the subsequent removal of all the Tc from the enrichment complex prior to decommissioning.

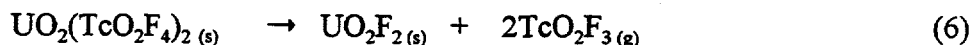
The volatility of the suspected cascade Tc fluorides is known to be less than that of UF<sub>6</sub>. The process gas lines in the cascade are typically sufficiently hot to maintain UF<sub>6</sub> as a gas at the required cascade process pressures. However there can be areas of auxilliary piping that may not be consistently maintained at temperatures sufficiently high to prevent the simple condensation of a fraction of the available TcO<sub>3</sub>F, and UF<sub>6</sub>, at these system pressures. *It is well known that technetium will collect in areas of low relative temperature*. Additionally, the presence of any system moisture will readily hydrolyze the

technetium fluorides to the pertechnetic acid,  $\text{HTcO}_4$  and the solid dioxide,  $\text{TcO}_2$  (see equation 4). As previously mentioned,  $\text{HTcO}_4$  is a viscous liquid with a much lower vapor pressure than the fluorides. The vapor pressure of a uranium-technetium-oxygen-fluoride complex or liquid mixture would also (probably) be lower than the vapor pressure of either single constituent.

The existence of such complexes or mixtures in the cascade has been postulated by Farrar and Barber.<sup>10,15</sup> Farrar speculated from operational observations in the ORGDP purge cascade that a uranyl per(fluoro)-technetate with a low vapor pressure could exist, that readily dissociated when heated. The complex could form via



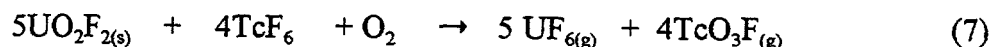
Gentle heating of a thin film of this "pertechnetate" would dissociate back to the originating gas phase precursors. In contrast, when a bulk material was strongly heated, the complex would instead decompose to yield the associated oxyfluorides



Analytical results for some of the materials recovered during Farrar's work at the ORGDP purge cascade exist, but were not published.<sup>10</sup> Consequently, the accuracy of the results are unknown without explanation and validation by the experimenters involved. It is noteworthy here to recall the report of excess  $\text{XeF}_6$  added to a solution of  $\text{TcO}_3\text{F}$  in  $\text{HF}$  yielding  $\text{TcO}_2\text{F}_3$ .<sup>18</sup> It seems reasonable to suggest that an excess of  $\text{UF}_6$  could similarly react with a deposit of  $\text{HF}$  and  $\text{TcO}_3\text{F}$  in the cascade to yield  $\text{TcO}_2\text{F}_3$  and  $\text{UO}_2\text{F}_{2(s)}$ .

Barber presented as a plausible alternative to Farrar's suggested complex the existence of a maximum boiling azeotrope in the  $\text{UF}_6$ - $\text{TcO}_3\text{F}$  phase system.<sup>15</sup> It was suggested that the presence of variable compositions of Tc bearing liquids could be more readily explained by this postulate. Barber also reported that the hexafluoride,  $\text{TcF}_6$  is converted into  $\text{TcO}_3\text{F}$  by reaction with  $\text{UO}_2\text{F}_2$  with the accompanying release of  $\text{UF}_{6(g)}$  and a minor unidentified non-volatile compound. He suggested that the reaction may proceed

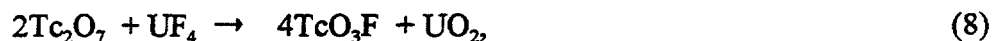
as shown by



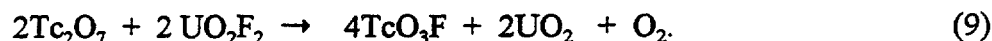
but noted that the primary importance to cascade chemistry was the short life span of the  $\text{TcF}_6$  in the presence of  $\text{UO}_2\text{F}_2$ . Unfortunately, Barber did not provide a reference to this observation. At this time it is not quantitatively known if  $\text{TcF}_6$  is actually introduced into the cascade.

An early mass spectrometry study of Tc at ORNL yielded some interesting observations.<sup>20</sup> An ammonium pertechnetate salt was converted to pertechnetic acid and analyzed to attempt determination of the upper limit of the isotopic ratio. The most dominant ion observed was the  $\text{Tc}_2\text{O}_7^+$ , and as temperature increased the  $\text{TcO}_4^+$  intensity decreased. The authors postulated the thermal decomposition of  $\text{HTcO}_4$ , but it is interesting to note that  $2\text{HTcO}_4$  is formulaically equivalent to  $\text{Tc}_2\text{O}_7 \cdot \text{H}_2\text{O}$ . Most interesting to the enrichment complex are the results obtained when  $\text{UF}_4$  was added to this system to provide a reference mass cracking pattern.

*Following the addition of  $\text{UF}_4$  to the  $\text{HTcO}_4$ , no ditechneium ions (i.e.,  $\text{Tc}_2$  compounds) were detected. Instead the most abundant peak was  $\text{TcO}_3\text{F}^+$ , with an intensity 20X the next most abundant ions  $\text{TcOF}^+$  and  $\text{TcO}_2\text{F}^+$ . Also, the  $\text{UF}_4$  was not detected, but instead the mass pattern revealed  $\text{UF}_5^+$ . The authors speculated that the  $\text{UF}_5$  was a result of reaction between  $\text{UF}_4$  and some assumed  $\text{UO}_2\text{F}_2$  impurity. It seems more probable that the  $\text{UF}_5^+$  was resultant from the presence of  $\text{UF}_6$  gas generated in the vessel. The authors proposed the following reactions at  $160^\circ\text{C}$*



and



*The Tc trapping affinity of both  $\text{UO}_2\text{F}_2$  and  $\text{UF}_4$  was demonstrated by Golliher during the developmental evaluation of  $\text{MgF}_2$  traps.<sup>1</sup> Golliher reported 28% Tc recovery*



from  $\text{UF}_6$  at 200°F and atmospheric pressure for  $\text{UO}_2\text{F}_2$  powder, and 17% sorbance in  $\text{UF}_4$ . Neither the Tc compound nor the concentration, or the U surface area was given.  $\text{UF}_5$  was not tested but presumably could sorb Tc to a similar extent. Uranium deposits of various mass exist in all cascade areas, if only as a relatively thin surface "film". Trace levels of Tc, which is all that exists in most areas of the operational cascade, could be rather readily sorbed within these uranium deposits.

The analysis of an ORGDP deposit with elevated beta radioactivity, indicative of  $^{99}\text{Tc}$ , revealed an X-ray Diffraction (XRD) pattern of amorphous  $\text{UO}_2\text{F}_2$  (that is, the deposit had absorbed water and a definitive crystal pattern was not observed).<sup>23</sup> The presence of the Tc was analytically detected by additional elemental analyses. Again, determination of the Tc compound was unsuccessful.

The ability to determine the Tc absorption and reaction potentials, the probabilities of reactions with cascade surfaces and compounds, and the conditions favorable to formation of these cascade Tc compounds will require theoretical evaluation and experimental investigation. Experimental systems fabricated for previous efforts currently exist to study these reactions, and quantitative evaluations could be simply performed with instrumental methods currently on-hand.



## 5. TECHNETIUM CONTROL IN THE CASCADE

Various methods and operational schemes have been employed to control Tc in the GDP's. Some of these have been reported in classified reports, but most have only been passed along through technical discussions. An excellent review of techniques current at PORTS in 1981 was presented by Saraceno.<sup>8</sup> Only minor improvements have been instituted since that time. The control of Tc in the enrichment complex would include methods necessary to maintain low Tc levels at (1) UF<sub>6</sub> product withdrawal stations, (2) air emission vent streams, and (3) aqueous waste streams, and preferably (4) control in the isotopic and purge gas streams. Ideally, Tc control would involve the isolation of the Tc in one location, and the subsequent recovery of all the cascade material.

*The unique nature of gas phase Tc compounds in the DOE GDP's has limited the evaluation of the technetium fluorides in the scientific community.* Very little information on the gas phase species relevant to the cascades has been compiled. The more common problem of Tc in aqueous waste streams has been much better characterized worldwide. Several internal investigations have focused on Tc control of the varied aqueous waste streams at the operating GDP's. A recent overview of the treatment of Tc contaminated process wastewaters and the remediation of groundwater originating from uranium enrichment activities at the GDP's was compiled by Bostick.<sup>24</sup> This report will consider only some of the gas phase control methods.

An increased movement of Tc in the enrichment cascades has been reported to be influenced by the increase in temperature, and/or the reduction in system pressure. Consequently, *areas of lower temperature and/or higher pressure will tend to accumulate the Tc.* This reported observation has been employed at the GDP's to attempt to control the Tc level by reducing the temperature of given cells at the top of the cascade. These cells would tend to retain a higher fraction of the gas-phase Tc. The cell could then be taken "off-stream" and an inert gas such as dry air introduced, for "hot-air treatments".<sup>14</sup> The cell temperature would be increased to ~ 225—250°F by the circulation of an inert "light" gas with the process compressors. The diffusion of the hot, light gas would unplug the barrier by volatilizing the Tc. A cryogenic trap could then be valved in to the flow loop to remove the volatilized Tc from the inert gas stream. The temperature range reported to

maintain Tc in the gas phase is 200—250°F, depending on the gas stream pressure. Cell plugging has been reported to occur at 140—150°F.

The presence of high freon (R-114) levels in the cascade gas stream can effect the control of Tc in the purge cascade. The Tc is lighter in weight than R-114 and the Tc "bubble" or front, moves up or down the purge cascade slightly ahead of the R-114 bubble. With significant R-114 inleakage the R-114 bubble can extend over several cells in both the side and the top purge. Consequently the heavier UF<sub>6</sub> front is pushed back down the cascade. The control of this UF<sub>6</sub> / "lites" front is very difficult with a short length cascade. If the high concentration Tc bubble coincides with the product withdrawal point the product cylinder will contain an elevated Tc level. Several operational alternatives exist to potentially allow more optimal control of the fronts, as reported by Faulkner.<sup>25</sup>

The use of chemical trapping agents has been previously mentioned. The development of MgF<sub>2</sub> pellets at the PGDP allowed the selective removal of the volatile Tc from the UF<sub>6</sub> gas stream. Trapping studies at the PGDP and PORTS have adequately defined the removal efficiency of MgF<sub>2</sub> traps for targeted gas-phase cascade applications. Unfortunately, the ability of the MgF<sub>2</sub> trap to remove Tc decreases at very low levels (ppm or ppb) and will allow a low level bleed-through (~ 1 ppm) at typical cascade gas velocities. *It is speculated that the existing gas phase MgF<sub>2</sub> trapping technology will not adequately remove Tc to meet the proposed new UF<sub>6</sub> ASTM product specification of 0.2 μg Tc per g <sup>235</sup>U.*

*Recent experiments at PORTS have successfully evaluated the low level Tc trapping efficiency of MgF<sub>2</sub> traps from the liquid phase UF<sub>6</sub>.*<sup>26</sup> The use of smaller MgF<sub>2</sub> pellets at typical product withdrawal station conditions reduced the Tc bleed through and resulted in a long term cumulative recovery of ~99% of the total <sup>99</sup>Tc fed. A small MgF<sub>2</sub> trap was successful in stripping 718 μg of Tc from 28 kg liquid UF<sub>6</sub>. The trap loading and life time has yet to be determined, but the early indications are very encouraging. If a long term feed stream trapping program had been maintained at the GDP's during the processing of RU, it is arguable that the Tc level today would not be a problem. With the addition of consistently maintained on-line MgF<sub>2</sub> traps at the product withdrawal stations, the current levels of Tc would certainly be below regulatory concern.

Additional solid chemical sorbents have been evaluated for the recovery of both Tc

and U at cascade air emission points. An excellent overview by Barber was mentioned previously.<sup>15</sup> Saraceno has performed tests with many alternatives at the PORTS purge cascade<sup>8,27</sup>, and numerous trapping studies have been performed at the PGDP and reported by Otey and colleagues. Adequately tailored to specific gas streams and applications (i.e., vent stacks), some of these sorbents perform exceptionally well. These materials can not be considered as alternatives to  $MgF_2$  - they are not selective for Tc - as they will trap most all fluoride compounds at typical process conditions. Solid sorbents that have been found effective at various conditions include 1) activated  $\gamma$ -alumina ( $2AlO(OH)$ ), 2) soda lime (a mixture of sodium hydroxide (NaOH) and calcium oxide (CaO) or calcium hydroxide ( $Ca(OH)_2$ ), 3) various limestone compositions (pure calcium carbonate ( $CaCO_3$ ) reported to be superior) and 4) again, sodium fluoride (NaF) by some investigators. Calcium sulfate ( $CaSO_4$ ), dolomite, sodium carbonate ( $NaCO_3$ ), and iron oxide have also been evaluated.

A liquid gas phase scrubber, such as the potassium hydroxide (KOH) unit employed at the ORGDP purge cascade can also reduce Tc emissions, as well as the U and  $F^-$  (fluoride ion). Personnel at K-25 attributed part of the success of low  $^{99}Tc$  site emissions in the early '80's to the KOH scrubber's additional 50% reduction of residual Tc entering the vent stack.

The cascade operations can affect control of the Tc level in many ways. Obviously, the power levels and feed rates can affect the Tc level at the top of the cascade. Cell surges due to high intermediate gas concentrations at the top purge, gas pressure and/or velocity fluctuations, and associated cell temperature excursions can "burp" the vent stacks. Unfortunately, good correlation of the cause of increased Tc levels with the observation is not always acquired. Frequently, the information casually interpreted will conflict what had been previously accepted as historical fact.

*A good historical review and compilation of the correlary cause and effects of known high Tc "bubbles", vent stack emissions, and product levels could provide valuable future operational information.*



## 6. RECOMMENDATIONS

The unique nature of gas-phase technetium compounds has limited most of the investigation of these compounds to the gaseous diffusion complex. The trace levels of Tc present in the cascade, and the absence of regulatory concern at these levels of Tc for most of the operational lifetime, has inhibited the support for quantitative technetium chemistry by knowledgeable enrichment scientists. Consequently, *significant gaps exist in the knowledge of the physical and chemical properties of the technetium compounds of interest to the gaseous diffusion cascade.*

The resolution of current and future cascade operational concerns due to the presence of technetium could be enhanced by attaining the most relevant information on the chemistry of technetium in the gaseous diffusion cascade. The recommended areas of investigation that would resolve questions currently being asked that cannot be confidently answered should focus on a few priority tasks:

- Determine the on-line spectroscopic capabilities of cascade gas analyzers for compounds,
- Determine the chemical reactivity of  $\text{TcO}_3\text{F}$  with prevalent cascade compounds at cascade temperatures and pressures,
  - uranium compounds
  - fluoride surfaces (chemical sorbents, cascade materials)
  - other cascade chemicals (e.g.  $\text{F}_2$ ,  $\text{ClF}_3$ ,  $\text{HF}$ , R-114, etc.)
- Determine what non-volatile Tc compounds can exist, at what conditions, and the mechanism(s) responsible,
- Determine chemical and thermodynamic properties for the identified priority compounds.

The acquisition of applied technetium cascade chemistry information would provide background knowledge sufficient to project solutions to enrichment complex problems regarding  $\text{UF}_6$  product specification, cascade compliance to ES&H issues, and would enhance the potential for removal of the technetium from the GDP's valuable

recyclable scrap metal prior to decommissioning.

An annotated bibliography of technetium fluorides and oxyfluorides prepared in 1974 has been discovered. The literature searched included progress and internal reports at all three sites that are not abstracted in the open literature. An updated bibliography was prepared in 1977. Both listings provide references to GDP observations and reports that should provide additional insight into cascade technetium chemistry. A thorough review of these documents and compilation of the information contained should be performed.

Additionally, a good historical review and compilation of the correlary cause and effects of known high Tc "bubbles", vent stack emissions, and  $UF_6$  product levels could provide valuable future operational information.



## 7. ACKNOWLEDGEMENTS

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