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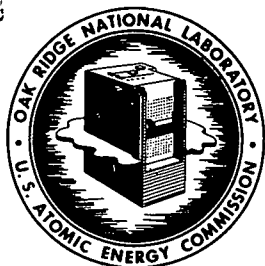
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SUBJECT: Darex Flowsheet (Batch Operations)Classification cancelled (or changed to **UNCLASSIFIED**)

TO: F. L. Culler

by authority of Str. M. Sect. R1.

FROM: W. E. Clark

by J. C. Ridenour TIE, date 12-11-58**AEC RESEARCH AND DEVELOPMENT REPORT****LEGAL NOTICE**

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The completed Darex flowsheet is transmitted herewith. Major contributors to its preparation include J. R. Flanary, J. H. Goode, and A. H. Kibbey of Chemical Development Section B and F. G. Kitts of Unit Operations.

It is recognized that this represents an interim flowsheet since continuous operation will doubtless prove to be advantageous. Continuous operation of the dissolver appears to be particularly desirable because of the relative ease with which the dissolution rate can be controlled as compared with a batch dissolver.

No provision has been made in this flowsheet for the absorption and/or rectification of the off gases from the dissolver and the stripper. The chemistry of this process is presently under investigation. It appears that this step also will lend itself to continuous operation and recycle in a system similar to that outlined by Guthrie (ORNL-CF-56-8-167).

Walter E. Clark
W. E. Clark

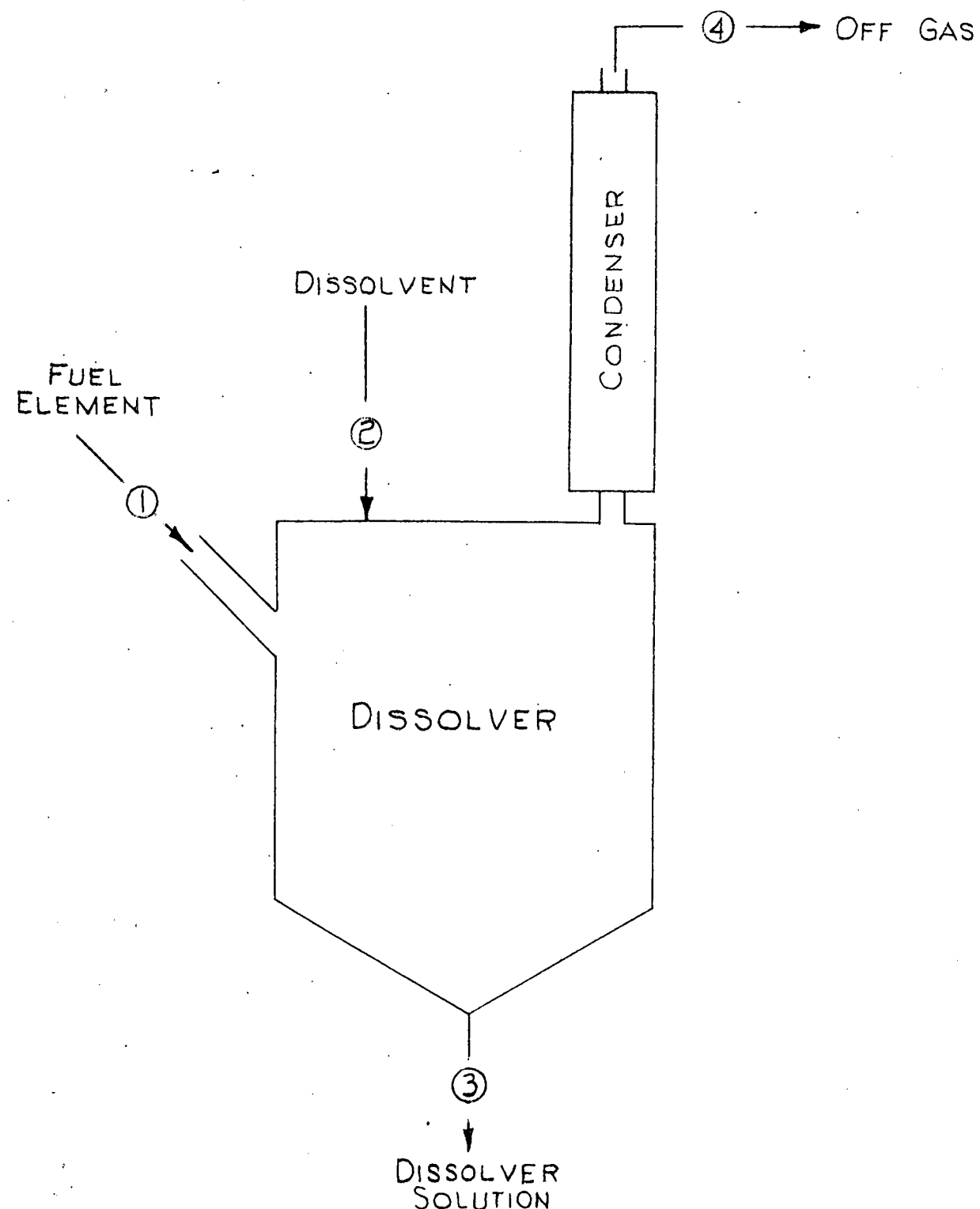
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164

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FLOW SHEET:



* REACTIONS:

1. $\text{HNO}_3 + 3\text{HCl} \rightarrow \text{NOCl} + \text{Cl}_2 + 2\text{H}_2\text{O}; \Delta H = 42.2 \text{ K-cal/mole } \text{Cl}_2 \text{ FORMED}$
2. $\text{M} + n\text{NOCl} \rightarrow \text{MCl}_n + n\text{NO}; \Delta H (\text{K-cal/mole REACTED}): \text{Fe} = 95.7, \text{Cr} = 133.5, \text{Ni} = 35.8, \text{Mn} = 73.5, \text{UO}_2 = 47.5$
3. $\text{UO}_2 + 4\text{HNO}_3 \rightarrow \text{UO}_2(\text{NO}_3)_2 + 2\text{NO}_2 + 2\text{H}_2\text{O}; \Delta H = 17.9 \text{ K-cal/mole } \text{UO}_2 \text{ DISSOLVED}$
4. $x\text{M} + (2x+2)\text{HNO}_3 \rightarrow x\text{M}(\text{NO}_3)_n + \frac{(2x+2)}{4}\text{NO} + (x+1)\text{H}_2\text{O}; \Delta H (\text{K-cal/mole DISSOLVED}): \text{Fe} = 217.3; \text{Cr} = 154.5; \text{Ni} = 199.5; \text{Mn} = 100.0; \text{UO}_2 = 29.3$
5. $\text{M} + 2n\text{HNO}_3 \rightarrow \text{M}(\text{NO}_3)_n + n\text{NO}_2 + n\text{H}_2\text{O}; \Delta H (\text{K-cal/mole DISSOLVED}): \text{Fe} = 87.0; \text{Cr} = 137.4; \text{Ni} = 55.1; \text{Mn} = 88.6; \text{UO}_2 = 17.9$
6. $\text{Fe}(\text{NO}_3)_2 + 2\text{HNO}_3 \rightarrow \text{Fe}(\text{NO}_3)_3 + \text{NO}_2 + \text{H}_2\text{O}; \Delta H = 14.8 \text{ K-cal/mole } \text{Fe}^{+2} \rightarrow \text{Fe}^{+3}$
(NOTE: $\text{M} = \text{Fe}; \text{Ni}; \text{Cr}; n = 2, 3; x = 1 \text{ or } 3$)

PROCEDURE:

1. CHARGE FUEL ELEMENTS TO DISSOLVER.
2. ADD THE MIXED ACIDS SLOWLY WHILE STILL WARM ($\sim 50^\circ\text{C}$) FROM PREPARATION, UNTIL TEMPERATURE RISES TO BOILING ($\sim 110^\circ\text{C}$). FINAL ADDITIONS OF ACID MUST BE MADE CAUTIOUSLY TO AVOID UNCONTROLLABLE GAS EVOLUTION.
3. RETURN VOLATILIZED HCl , HNO_3 AND EXCESS WATER TO THE DISSOLVER BY USING CONDENSER ON OFF-GAS STREAM

- * (1) HEAT OF REACTION PER BATCH DISSOLUTION OF ONE FUEL ELEMENT ≈ 88.6 MOLES OF S.S. + $\text{UO}_2 = 19,684.8 \text{ K-cal}$. (ASSUME 20,000 K-cal/APPR ELEMENT A GOOD APPROXIMATION FOR MAXIMUM ΔH VALUE.) THIS VALUE IS BASED ON USE OF EQUATIONS 1, 2, 3, 4, AND 6, ABOVE, SINCE EQUATION 5 IS BELIEVED TO BE A MINOR REACTION TAKING PLACE ONLY AT THE BEGINNING OF THE DISSOLUTION

DAREX PROCESS FLOWSHEET

I. BATCH DISSOLUTION OF APPR. FUEL

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PER: *M.E. Clark*
ORNL-LR-DWG
24310

CHEMICAL TECHNOLOGY DIVISION

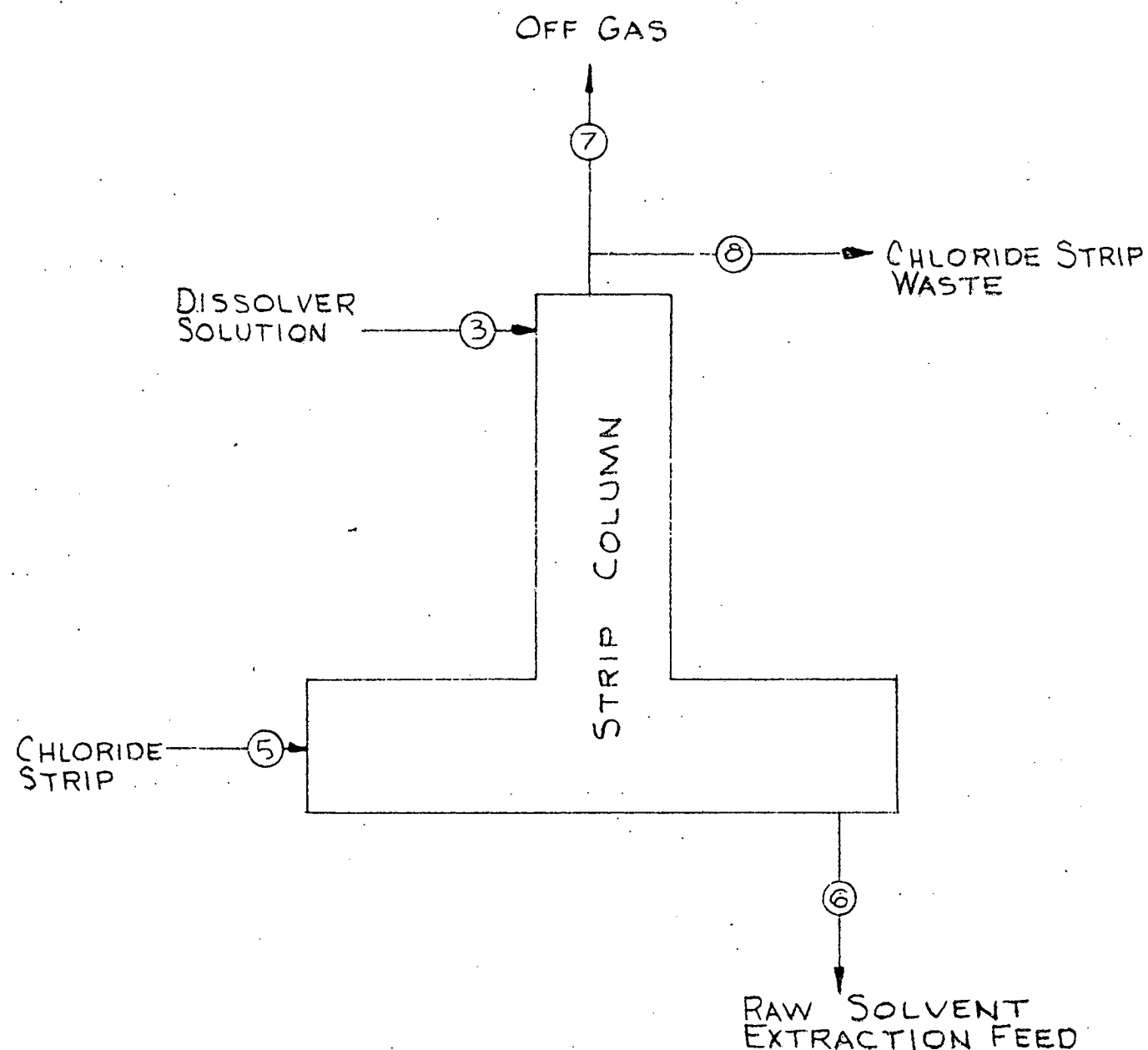
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DESIGNED <i>M.E. Clark</i>	DATE 7-11-57	APPROVED <i>M.E. Clark</i>	DATE 9/11/57

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FLWSHEET:



REACTION:



(THIS APPEARS TO BE THE DOMINANT REACTION AT NO_3/Cl -mole RATIO OF APPROXIMATELY 17/1)

PROCEDURE:

COLUMN OPERATION CONDITIONS:

- (1) V/L MOLE RATIO = 1
- (2) NUMBER OF STAGES = 3
- (3) TEMPERATURE:
 - (a) TOP = 110°C
 - (b) BOTTOM = 123°C

CHLORIDE CONTENT OF STREAM 6 IS SPECIFIED AS $\leq 30\text{ ppm}$

DAREX PROCESS FLOWSHEET
II CHLORIDE REMOVAL FROM
DISSOLVER SOLUTION

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PER: *W.E. Clark*

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24382

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DESIGNED <i>G.H. Kelly</i>	DATE 9-11-57	APPROVED <i>W.E. Clark</i>	DATE 9/11/57

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				U		Fe		Ni		Cr		Mn		Si	P	B	H ⁺		NO ₃ ⁻		Cl ⁻		H ₂ O		NO		NO ₂		Cl ₂	
STREAM NO	VOL L	MASS Kg	STD 3/4	M	Kg	M	Kg	M	Kg	M	Kg	M	Kg	Kg	Kg	Kg	M	Kg	M	Kg	M	Kg	g MOLES	Kg	g MOLES	Kg	g MOLES	Kg	g MOLES	Kg
STREAMS IN																														
3	61	84.9	1.39	0.026	.377	1.00	3.42	0.11	.397	0.28	.893	0.03	.099	.099	.050	.002	1.50	.092	3.95	14.9	1.74	3.77	3380	60.8						
5	118	167	1.42														15.2	1.81	15.2	111			3000	54.0						

DRAWN	DATE	CHECKED	DATE
JW DAY	8-29-57		
DESIGNED	DATE	APPROVED	DATE
W. B. K. B. B.	9-11-57	W. B. K. B. B.	9-11-57

DAREX PROCESS FLOWSHEET
II CHLORIDE REMOVAL FROM
DISSOLVER SOLUTION

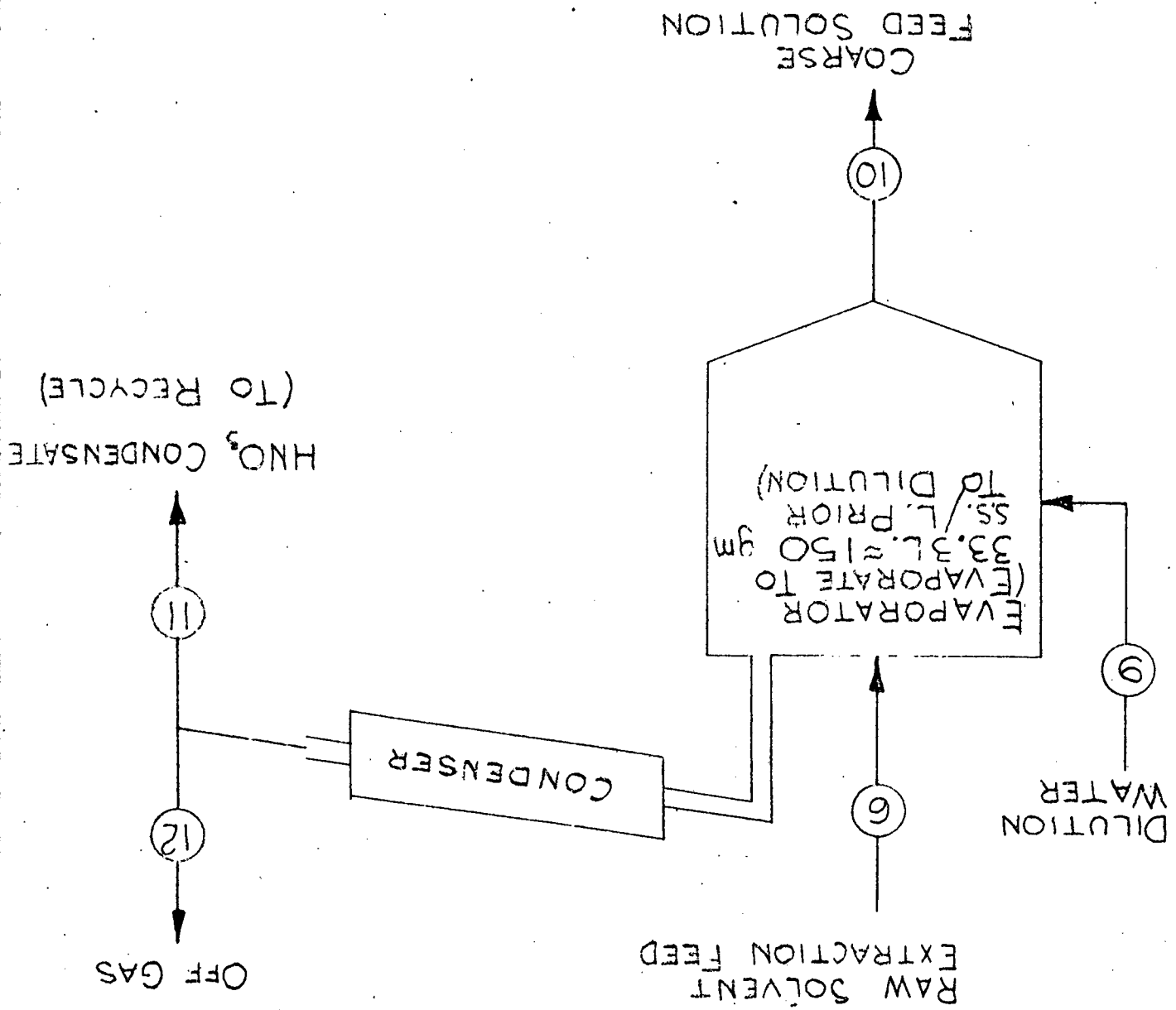
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FLOWSHEET



DAREX PROCESS FLOWSHEET III FEED ADJUSTMENT TANK (FAT)-NITRIC ACID REMOVAL

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2484
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REACTION

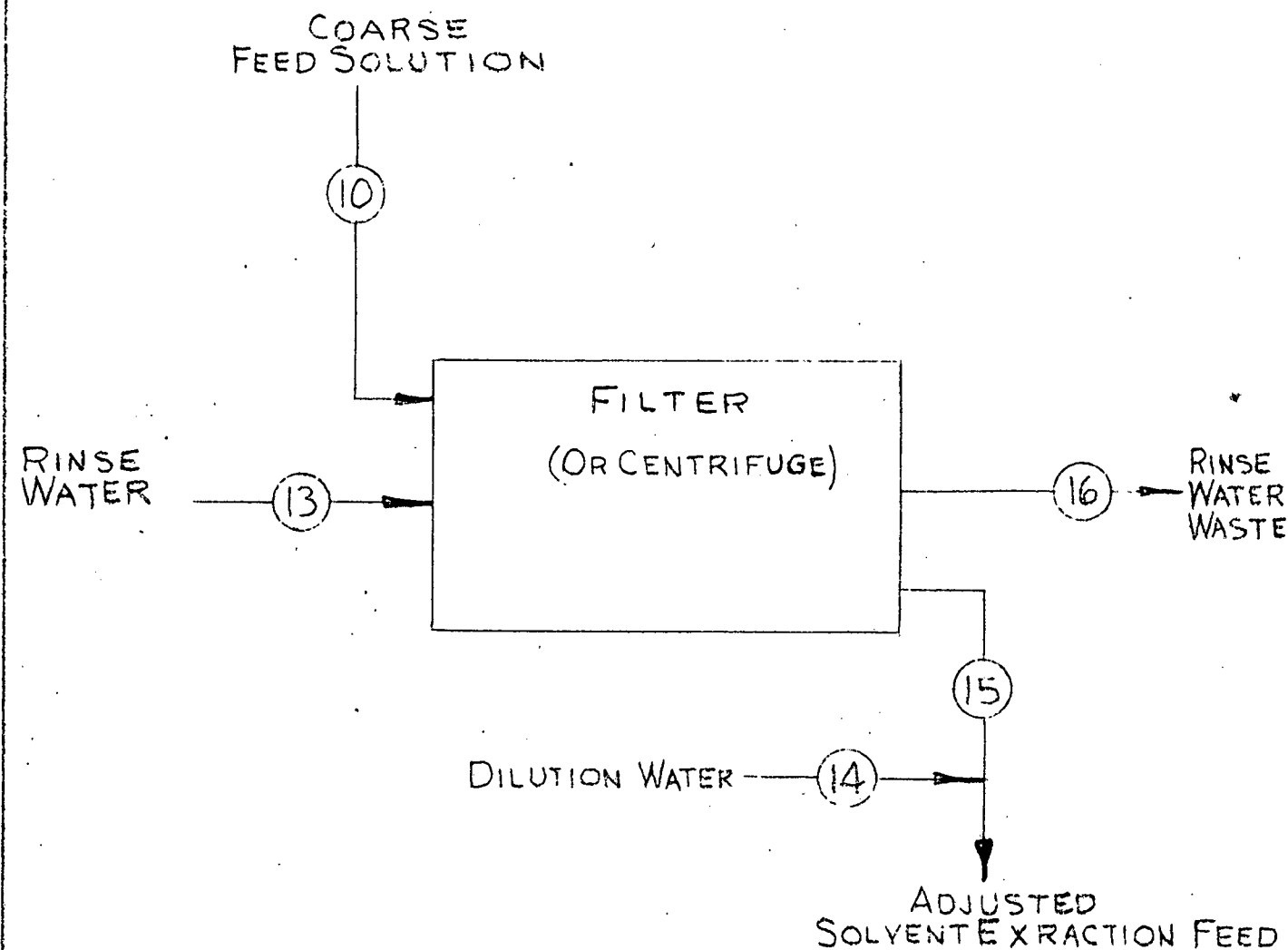
PROCEDURE:

1. EVAPORATE RAW SOLVENT EXTRACTION FEED TO REMOVE EXCESS HNO₃. AFTER EVAPORATION THE SOLUTION IS ~4M IN FREE ACID. THE AZEOTROPE (15.2M) DISTILLS OFF IMMEDIATELY UNDER THESE CONDITIONS SINCE BOILING POINT OF 118°C-125°C IS HIGHER THAN THAT OF THE HNO₃ AZEOTROPE (BP=120.5°C)
2. DILUTE BACK TO ORIGINAL GIL VOLUME WITH WATER

DRAWN	J.W. DAY	DATE	9-4-57
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DESIGNED		DATE	
APPROVED	M.E. Clark	DATE	7-11-57

FLOWSHEET

REACTION



PROCEDURE:

1. COARSE FEED SOLUTION IS FILTERED (OR CENTRIFUGED)
2. WATER FOR DILUTION IS ADDED AS 2 WASHES OF 7L EACH FOR THE FILTER CAKE.
3. RINSE WATER IS ADDED AS A BACKWASH TO REMOVE SILICA FROM FILTER (OR CENTRIFUGE), PROBABLY IN 2 ALIQUOTS. SINCE THE VOLUME REQUIRED FOR SILICA REMOVAL IS LARGELY DEPENDENT UPON EQUIPMENT DESIGN, THE RINSE VOLUME WAS ARBITRARILY CHOSEN AS EQUAL TO THE FEED VOLUME.

DAREX PROCESS FLOWSHEET

IV FEED CLARIFICATION -
SI REMOVAL

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24386

PER: *W. E. [Signature]*

DRAWN JW DAY	DATE 10-5-57	CHECKED	DATE
DESIGN <i>G. H. [Signature]</i>	DATE 7-11-57	APPROVED <i>W. E. [Signature]</i>	DATE 9/11/57

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