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**Compilation of the Melting Points  
Of the Metal Oxides**



**U.S. DEPARTMENT OF COMMERCE  
NATIONAL BUREAU OF STANDARDS**

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# Compilation of the Melting Points of the Metal Oxides

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A compilation has been made of the melting points of 70 metal oxides published prior to January 1963. Both the original melting point and the equivalent value based on the International Practical Temperature Scale of 1948 are presented. Included in the survey is information on pertinent experimental details such as the method of temperature measurement, purity, furnace type, and environmental conditions.

The melting points of the metal oxides are perhaps one of the most studied but least known of the physical properties. Seldom have two investigators completely agreed upon a single melting point value for a specific oxide. Many reported values were derived from excellent research and are quite reliable. Others are obviously in gross error. The result of this general disagreement has been that over the years a myriad of melting point values has been perpetuated through the literature. To complicate matters, necessary conversion from one temperature scale<sup>1</sup> to another has led to incorrect reporting of original data. Furthermore, various compilations of critical data have listed "best values" for the melting points of the metal oxides which often have been misinterpreted as unquestionably accurate and valid. At the present time our knowledge of the true melting points of the oxides can be regarded as in no less than a state of utter confusion.

The National Bureau of Standards has recently initiated a program of reevaluation of the melting points of the metal oxides. The program tentatively includes the acceptance of previously published values after experimental verification as well as accurate determinations of ill-defined or completely unknown melting points. The ultimate goal of the work is to have national and perhaps international recognition of the melting points of several specific oxides for use as calibration standards.

The initial effort in this program has been placed in making a comprehensive literature survey of previously published melting points of the various oxides. The results of this compilation of data are presented in table 1. The information given in table 1 was taken from the original published papers whenever possible. General surveys [1, 2, 3, 4]<sup>2</sup> are referred to only when the original source could not be established. Five primary sources [5, 6, 7, 8, 9] were used initially to obtain a bibliography of papers dealing with oxide melting point determinations. Reference to these papers in turn yielded another bibliography of reports which again supplied additional literature sources. The entire process was repeated until no new references were obtained. Through this cascading method of obtaining references it is believed that a majority of all

published data on oxide melting points up to January 1963 is included in the present work.

The oxides given in table 1 are arranged alphabetically according to the chemical symbol of the metal element. Excluded from consideration are oxides which are not solids at room temperature. Only oxides having stoichiometric metal to oxygen ratios are listed. Identifying symbols such as mineral names, crystal symmetry, etc., are given only to those oxides which are known to have polymorphs that melt metastably. In instances where an oxide has more than one reported melting point, individual values are given in the order of increasing magnitude according to the value based on the International Practical Temperature Scale of 1948 (Text Revision of 1960) [10].

Several criteria must be carefully considered in obtaining accurate melting point data. These factors in a general way are reflected by the column headings of table 1. Because of the wide divergence of melting point values, it is believed worthwhile, and in order, to discuss briefly each of these criteria more than normally would seem warranted in a compilation of data.

**Purity.** The ultimate accuracy with which any melting point can be determined is directly associated with the purity of the material. It would seem that this fact need not be emphasized; however, reports are frequently found in the literature in which purity was not considered or even mentioned. Terms such as "chemically pure," "pure," and "reagent grade" are generally meaningless unless specifically defined. Reasonable estimates of the percentages and types of impurities must be known along with their possible effect upon the melting point. Sufficient attention also must be given to the effects of contamination of the sample by its container during actual experimentation.

**Furnace type.** For obvious reasons it is particularly important in determining melting points to insure against thermal gradients throughout the sample as well as between the sample and adjacent areas where actual temperature measurements are made. Uniformity of temperature can be completely assured only by the attainment of blackbody conditions, which, of course, can never be entirely realized in practice. Excellent papers by Gouffé [11] and by De Voss [12] give methods for estimating the effective spectral emittance of several types of enclosures and cavities.

Resistance-type furnaces such as the wire-wound quench or tube furnace can be constructed

<sup>1</sup> Temperature scale refers not to the symbols denoting temperature (°K, °C, and °F) but to an actual scale with defining fixed points and formulas for interpolating between such points.

<sup>2</sup> Figures in brackets indicate the literature references at the end of this Monograph.

so as to have small gradients between the specimen and thermoelement. However, because of the close proximity of the specimen and thermoelement the deleterious effect of a gradient is usually negligible.

Whenever radiation pyrometry is employed for temperature measurements, near-blackbody conditions are considered to be a necessity. The induction furnace is generally one of the most suitable for approximating blackbody conditions. Temperature uniformity, with corresponding high emittance, is easily obtained through the use of inductively heated crucibles and cylinders. Another furnace frequently used for high-temperature application is the strip furnace. It usually consists of a narrow, short strip of refractory metal sometimes necked down at its center to provide a high-resistance area. Power is supplied to the strip through water-cooled leads. To facilitate temperature measurements, the strip is generally made into the shape of a V (or U). Specimens are generally placed at the apex of the V or at the midportion of a flat strip. Temperature measurements are made on areas immediately adjacent to the specimen.

The strip furnace provides an easy effective means of attaining high temperatures; however, it is not particularly conducive to accurate melting point determinations. Inherent with this furnace are the inevitable extreme thermal gradients throughout the specimen and area of measurement. Tungsten strip lamps used for calibration of optical pyrometers, although not strictly comparable, have thermal gradients along the filament which may amount to as much as 5 °C per millimeter [13]. Placement of a specimen on the strip will, in all probability, seriously alter the thermal gradient of the bare strip. The magnitude of this change will vary, depending upon such factors as the location, volume, thermal conductivity, and emissivity of the specimen.

It is a generally false assumption to consider that a V-shaped metal strip approximates a blackbody enclosure. As long as there is any deviation from isothermal conditions about the enclosure, blackbody conditions cannot be realized. However, brightness temperatures can approach true temperatures if the emittance of the strip is near unity. The use of a substance having a high natural emittance (i.e., graphite) as strip material is not always feasible.

*Temperature measurement.* The theory, use, and calibration of various temperature-measuring devices have been well documented [13, 14, 15, 16]. The optical pyrometer or thermocouple is only one part of an entire temperature-measuring system. Often, little attention has been given to the effect of auxiliary equipment or associated apparatus on temperature measurements. An inadequate reference junction of a thermocouple circuit can cause errors as serious as those produced by thermal gradients in a furnace. The effect of stray or induced voltage from furnace windings

or other power sources can be very harmful and yet go undetected.

The position of windows or enclosures relative to the target area is extremely important in making brightness temperature measurements with an optical pyrometer. Gross errors are produced by the reflection of radiation by the windows, enclosures, or associated pieces into the field of view of the pyrometer. Indiscriminate use of published spectral emissivity values to convert brightness temperatures to true temperatures should be avoided. Emissivity refers to a property of an opaque material whose surface has been polished optically flat. Unless the target area of the pyrometer conforms to these conditions, errors will result. The emittance of any enclosure supposedly built as a blackbody is extremely dependent upon the nature of internal reflections. Even an enclosure in which the geometric design is suitable<sup>3</sup> can have low emittance if the internal reflections are predominantly specular rather than diffuse [12, 13].

*Method.* It is not uncommon to observe that oxides appear to melt over a range of temperatures with a marked hysteresis between the apparent melting and freezing points. The magnitude of the hysteresis may vary from negligible quantities to many degrees. Whatever the reason for the hysteresis, whether it be partial dissociation, influence of impurities, or other causes, the melting point must be defined with respect to experimental conditions. The melting point of an oxide can be specified as that temperature at which solid and liquid are in equilibrium for a given confining pressure and for given partial pressures of environmental gases.

Obvious difficulties are inherent in establishing the melting temperature. Experimentally it can be taken as the temperature at which the last solid disappears on heating at a sufficiently slow rate to insure temperature uniformity. Conversely, the freezing point is the temperature at which crystallization first begins. Dynamic methods such as differential thermal analysis and heating or cooling curves have the advantage of the utilization of heat effects but the distinct disadvantage of being rate-dependent. Cooling curve methods have been successfully applied in the determination of the freezing points of palladium, platinum, rhodium, and iridium (secondary reference points on the International Temperature Scale of 1948 [10]). Pronounced supercooling and superheating tendencies of many oxides make this method somewhat less applicable, although it may be the best available.

The method of visual observation of a specimen during heating is generally unacceptable as an accurate means of establishing the melting point. If the specimen is visible, it is quite probable that thermal gradients are present, thereby preventing

<sup>3</sup> A large ratio of internal area of cavity to area of opening is required. The ratio is sometimes expressed in terms of depth of cavity and diameter of opening.

accurate temperature measurements. The attainment of blackbody conditions renders the visual method ineffectual, inasmuch as the specimen would be indistinguishable from adjacent areas.

The static method in which a specimen is heated and cooled prior to examination has proved satisfactory in many instances. Temperature measurement is not as much a problem as with dynamic methods. Auxiliary equipment such as the microscope and X-ray diffraction can be readily applied to aid in determining the degree of melting.

*Calibration points.* The determination of a melting point can be only thought of as being obtained through the use of an integral system of various types of equipment and procedures. The mere use of previously calibrated thermocouples and pyrometers does not insure accurate measurement of temperature. It is highly desirable to calibrate the entire system against the known melting points of several standards. Insofar as possible, the overall characteristics of the standards should conform to those of the test material. Because of the present lack of recognized standards, calibration materials should generally be limited to those defining the International Practical Temperature Scale [10]. It is as undesirable to use substances having ill-defined melting points as it is to use standards which have different characteristics than those of the test material.

*Environment.* The effect of environmental gas as well as the confining pressure on the melting points of oxides has been generally disregarded or considered noninfluential. Extensive studies have been made only on those oxides (i.e., iron and manganese oxides) which are pronouncedly affected. It is conceivable that most oxides are distinctly influenced by various gas partial pressures more than previously realized and partial dissociation always occurs to some degree. If dissociation exists, an oxide can be expected to melt incongruently.

*Melting point and temperature scale.* Sosman [17], in his paper "Temperature Scales and Silicate Research," has extensively discussed the development of temperature scales, especially during the period 1907 to 1948. In essence, a single scale had not been recognized until about 1910 to 1914, when the Geophysical Temperature Scale [18] was established. In 1927, the first internationally recognized scale [19] was adopted. A major revision of the 1927 scale was made in 1948 [20], with subsequent text revision in 1960 [10]. Some of the principal fixed and secondary reference points of the three scales are given in table 2. The Geophysical Scale is, for consistency of data, still being used by a few investigators. The 1948 scale has entirely superseded the 1927 scale as being a closer approach to the thermodynamic scale. Numerical differences between the three scales generally amount to less than one degree at temperatures below the gold point. At higher temperatures, the variance becomes much greater. The Geo-

TABLE 2. *Some fixed and secondary points of the geophysical and international temperature scales*

Material	Prop- erty <sup>a</sup>	Int. 1927 [19]	Int. 1948 [10]	Geo- physical [17]
Naphthalene	B.P.	217. 96	218. 0	217. 95
Tin	M.P.	231. 85	231. 91	231. 9
Benzophenone	R.P.	305. 9	305. 9	305. 9
Cadmium	M.P.	320. 9	321. 03	320. 9
Lead	M.P.	327. 3	327. 3	
Zinc	M.P.	419. 45	419. 505	419. 4
Sulfur	B.P.	444. 6	444. 6	444. 5 <sup>b</sup>
Antimony	M.P.	630. 5	630. 5	630. 0
Silver	M.P.	960. 5	960. 8	960. 2
Gold	M.P.	1063	1063	1062. 6
Copper	M.P.	1083	1083	1082. 8
Diopside	M.P.			1391. 5
Nickel	M.P.		1453	1452. 4
Palladium	M.P.	1555	1552	1549. 5
Platinum	M.P.		1769	1755
Rhodium	M.P.		1960	
Iridium	M.P.		2443	
Tungsten	M.P.	3400	3380	

<sup>a</sup> The symbols B.P. and M.P. roughly signify boiling point and melting point, respectively. For exact designation of the property under consideration, reference should be made to the appropriate publication.

physical Scale has an upper limit at the melting point of platinum, defined as 1755 °C. Approximate differences between the Geophysical and 1948 scales at various temperatures can be obtained from appropriate plots of the data given in table 2. The 1927 scale above the gold point is based on the Wien radiation formula,<sup>4</sup> which is not strictly valid at extremely high temperatures. The 1948 scale utilizes the Planck equation<sup>4</sup> for temperatures above the gold point. The equation is applicable to indefinitely high temperatures.

Conversion from the 1927 scale to the 1948 scale, above the gold point, can be accomplished through the use of the following equation as derived by Corruccini [21].<sup>5</sup>

<sup>4</sup> (Wien)

$$\log \frac{J_t}{J_{A_u}} = \frac{C_2}{\lambda} \left[ \frac{1}{1336} - \frac{1}{(t+273)} \right] \quad (1)$$

If  $\lambda(t+273)$  is less than 0.3 cm deg the resulting error is less than 1 °C.

(Planck)

$$\frac{J_t}{J_{A_u}} = \frac{\exp \left[ \frac{C_2}{\lambda(t_{A_u} + T_0)} \right] - 1}{\exp \left[ \frac{C_2}{\lambda(t+T_0)} \right] - 1} \quad (2)$$

where

$J_{A_u}$  and  $J_t$ =radian energies per unit wavelength interval emitted per unit time by unit area of a blackbody at temperature  $t$  and at the gold point  $t_{A_u}$ .

$\lambda$ =wavelength (approximately  $0.65 \times 10^{-4}$  cm for most optical pyrometers)

$C_2$ =second radiation constant=  
1.432 cm deg (1927) or  
1.438 cm deg (1948)

$T_0$ =273.15 deg.

<sup>5</sup> The equation is derived by equating the right side of eq (1) with the logarithm of the right side of eq (2), expressing in exponential form and substituting appropriate values for  $T_0$ ,  $C_2$ , and  $t_{A_u}$  (1063 °C).

$$\frac{\exp\left[\frac{1.438}{1336.15\lambda}\right]-1}{\exp\left[\frac{1.438}{(T+273.15)\lambda}\right]-1} = \exp\left[\frac{1.432}{\lambda}\left(\frac{1}{1336} - \frac{1}{T+273}\right)\right]$$

where

$\lambda$ =wavelength (approximately  $0.65 \times 10^{-4}$  cm for most optical pyrometers)  
 $T$ =a temperature above the gold point, °C (Int. 1948)  
 $T'$ =a temperature above the gold point, °C (Int. 1927).

Approximate conversion from other temperature scales in use prior to 1927 can be accomplished in a similar manner. To ease calculations only Wien's relationship is utilized. The resulting equation necessary for the conversion is:

$$\frac{1.438}{\lambda} \left[ \frac{1}{1336} - \frac{1}{(T+273)} \right] = \frac{C_2'}{\lambda'} \left[ \frac{1}{(t_{Au}' + 273)} - \frac{1}{(T' + 273)} \right]$$

where

$C_2'$ =second radiation constant originally employed  
 $T$ =a temperature above the gold point, °C (Int. 1948)  
 $T'$ =a temperature above the gold point, °C (original scale)  
 $t_{Au}'$ =gold point originally employed (generally 1063 °C)  
 $\lambda$  and  $\lambda'$ =wavelength ( $\lambda$  and  $\lambda'$  can be assumed to be equal).

In the present work, all original melting points have been converted, whenever possible, to values based on the 1948 temperature scale by methods previously described. Unfortunately, some published papers fail to state the temperature scale utilized. Conversion to the 1948 scale was impossible in these instances unless the scale could be unambiguously assumed from information such as the date of publication or listed calibration temperatures and radiation constant.

*Comments.* Other than for the obvious purpose of the compilation of critical data, the present survey has another important function, in that it vividly illustrates the variance of melting point values found in the literature. No attempt will be made for the present to select one value over another as being absolutely correct. It is quite difficult to ascertain completely the validity of published data without additional experimental work. However, it is believed appropriate to make pertinent comments about the less obvious shortcomings inherent in any published work and to designate the preferred values with an asterisk.

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TABLE I

Oxide	Reference	Purity <sup>2</sup>	Furnace type	Temperature measurement	Method
<b>Ag<sub>2</sub>O</b>	Kracek [3].	Not stated.	Not stated.	Not stated.	Not stated.
	Shun-ichiro Iijima [22].	Not stated.	Not stated.	Not stated.	Weight loss measurements.
<b>Al<sub>2</sub>O<sub>3</sub></b>	Tiede and Birnbrauer [23].	Not stated.	Arc.	Optical pyrometer sighted on specimen.	Observation of specimen during heating.
	Geller and Yavorsky [24].	Three samples: (a) Soda free, 0.05% carbon; (b) 0.01% impurities; (c) <0.01% each of seven metals.	Resistance; ThO <sub>2</sub> -CeO <sub>2</sub> heating elements.	Optical pyrometer sighted on specimen.	Observation of pyramid-shaped specimen during heating. Rounded corners indicated melting.
	Ruff and Lauschke [25].	Commercially pure.	Resistance; graphite tube heating element.	Optical pyrometer sighted on specimen.	Observation of specimen during heating.
<b>Al<sub>2</sub>O<sub>3</sub></b>	Weigel and Kaysser [26].	Two samples: (a) Not stated (b) Alkali free.	Resistance; graphite tube heating element.	Optical pyrometer sighted on specimen.	Observation of specimen during heating.
	Mark [27].	Not stated.	Resistance; graphite tube heating element.	Optical pyrometer sighted on specimen.	Observation of specimen during heating.
	Diamond and Schneider [28].	About 99.9%.	Solar.	Optical pyrometer sighted on specimen.	Observation of bar-shaped specimen during cooling.
	Geller and Bunting [29].	99.98%.	Resistance.	Optical pyrometer sighted on specimen.	Observation of specimen during heating.
	Lambertson and Gunzel [30].	99.5%.	Resistance; W heating element.	Optical pyrometer sighted on specimen. Specimen not visible.	Examination of specimen after heating.
	Ruff [31].	Pure.	Resistance; graphite tube heating element.	Optical pyrometer sighted on specimen.	Observation of specimen during heating.
	Bunting [32].	99.95%.	Induction; Ir alloy susceptor, button-shaped.	Optical pyrometer sighted on small cavity in susceptor adjacent to specimen cavity.	Examination of specimen after heating.
	McNally, Peters, and Ribbe [33].	Not stated.	Induction; graphite tube susceptor.	Optical pyrometer sighted on specimen.	Observation of specimen suspended in furnace during heating.
	Ruff and Goecke [34].	Not stated.	Resistance; graphite tube heating element.	Optical pyrometer sighted on specimen.	Observation of cone-shaped specimen during heating. Fusion of cone tip indicated melting.
	Lang, Knudsen, Fillmore, and Roth [35].	99.9%.	Resistance; graphite tube heating element.	Optical pyrometer sighted on specimen.	Observation of pyramid-shaped specimen during heating. Deformation of pyramid tip indicated melting.
<b>Al<sub>2</sub>O<sub>3</sub></b>	v. Wartenberg, Linde, and Junq [36].	Pure.	Flame; oil-oxygen.	Optical pyrometer sighted on specimen.	Observation of specimen suspended in furnace during heating.

See footnotes at end of table.

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Calibration Points		Environment	Original Temp. Scale <sup>5</sup>	Melting Points		Comments	Oxide
Materials	Temp.			Original	Int. 1948 <sup>4</sup>		
Not stated	°C	Not stated	Not stated	191 dissociates	°C	Survey; not original data.	Ag <sub>2</sub> O
Not stated		Not stated	(Int. 1927)	230 dissociates	230		
Not stated		Vacuum	Not stated	1890		<sup>6</sup> ; partial dissociation probably occurred.	Al <sub>2</sub> O <sub>3</sub>
60Pt-40Rh	1940	(Air)	Int. 1927	2000 to 2030	1994 to 2024	( <sup>6</sup> )	
80Pt-20Rh	1905						
90Pt-10Rh	1845						
100Pt	1773.5						
Au	1062.4	(a) Air at 7.5 mm Hg (b) Air at 7.7 mm Hg	C <sub>2</sub> =1.437 cm deg	(a) 2005 (b) 2008	2005 2008	<sup>6</sup> ; partial dissociation probably occurred; reaction between specimen and graphite support probable.	
CaF <sub>2</sub>	1398						
Ag	960.5	(a1) Nitrogen at 1 atm	C <sub>2</sub> =1.44 cm deg	(a1) 2007±4 (a2) 2010	2009±4 2012	<sup>6</sup> ; reaction between specimen and charcoal support probable.	
Au	1063	(a2) Air		(b1) 2001	2003		
Pd	1557	(b1) Reducing		(b2) 2005 to 2010	2007 to 2012		
Pt	1770	(b2) Air					
Rh	1970						
Ir	2415						
Not stated		Neutral	(Int. 1948)	2020	2020	<sup>6</sup> ; reflection errors possible.	
None		Air	Int. 1948	2025	2025	<sup>6</sup> ; emissivity stated by authors to be about unity.	
Pt	Not stated	Air	Int. 1927	2035 ±10	2029 ±10	<sup>6</sup> ; reflection errors possible	
90Pt-10Rh	Not stated						
Pt	(1769)	Helium	(Int. 1948)	2034 ±16	2034 ±16		
Pt	1755	Not stated	C <sub>2</sub> =1.46 cm deg	2010 ±10	2035 ±10	<sup>6</sup> ; reaction between specimen and graphite support probable.	
Not stated		(Air)	Int. 1927	2045 ±25	2038 ±25	Temperature uniformity between specimen and target area of pyrometer not verified.	
Al <sub>2</sub> O <sub>3</sub>	2044	Air, argon, nitrogen	Int. 1948	2043 ±10	2043 ±10	( <sup>6</sup> )	
Au	1071	Nitrogen	(C <sub>2</sub> =1.48 cm deg)	2020	2044	<sup>6</sup> ; reaction between specimen and graphite support probable.	
Pt	1757						
Al <sub>2</sub> O <sub>3</sub>	2035	Argon	Int. 1948	2049	2049	( <sup>6</sup> )	
BeO	2510						
Pt	Not stated	Air	(Int. 1927)	2055	2049	( <sup>6</sup> )	
Mo	Not stated						

TABLE 11

Oxide	Reference	Purity <sup>2</sup>	Furnace type	Temperature measurement	Method
<b>Al<sub>2</sub>O<sub>3</sub>—Con.</b>	Kanolt [37]-----	Fe—0.001% Cu—none SiO <sub>2</sub> —0.001% Chloride—0.005% Sulfate—0.001%	Resistance; graphite tube heating element.	Optical pyrometer sighted on blackbody cavity immersed in specimen.	Thermal analysis, heating curves.
<b>As<sub>2</sub>O<sub>3</sub> Mono-clinic type.</b>	Smits and Beljaar [38]-----	Not stated-----	NaNO <sub>3</sub> -NaNO <sub>3</sub> bath-----	Pt resistance thermometer-----	Intersection of vapor pressure curve.
	Welch and Duschak [39].	C. P. grade-----	Not stated-----	Copper-constantan thermocouple.	Thermal analysis, heating curves; also examination of specimen after heating.
	Rushton and Daniels [40].	Not stated-----	Resistance; iron pot heating element.	Copper-constantan thermocouple.	Intersection of vapor pressure curves.
<b>B<sub>2</sub>O<sub>3</sub></b>	Kracek, Morey, and Merwin [41].	100 ± 0.1%-----	Resistance; quench type, wire-wound ceramic tube heating element.	Thermocouple-----	Examination with microscope of quenched specimen after heating. Last trace of crystals indicated melting point.
	McCulloch [42]-----	99.6%-----	Not stated-----	Not stated-----	Observation of specimen during heating.
<b>BaO</b>	Schumacher [43]-----	Not stated-----	Resistance; boat-shaped W strip heating element.	Optical pyrometer sighted on W strip adjacent to specimen. Spectral emissivity equals 0.431.	Observation of specimen during heating.
<b>BeO</b>	Tiede and Birnbrauer [23].	Not stated-----	Arc-----	Optical pyrometer sighted on specimen.	Observation of specimen during heating.
	Ruff and Lauschke [25].	Pure-----	Resistance; graphite tube heating element.	Optical pyrometer sighted on specimen.	Observation of specimen during heating.
	Lang, Knudsen, Filmore, and Roth [35].	99.9+%-----	Resistance; graphite tube heating element.	Optical pyrometer sighted on specimen.	Observation of pyramid shaped specimen during heating. Deformation of pyramid tip indicated melting.
	v. Wartenberg, Reusch, and Saran [44].	Commercially pure-----	Flame; gas-oxygen-----	Optical pyrometer-----	Observation of specimen during heating. Rounded corners of suspended specimen indicated melting.
	v. Wartenberg and Werth [45].	Commercially pure-----	(Flame); "ZrO <sub>2</sub> oven"-----	Optical pyrometer-----	Not stated.
	Ol'ShanskiI [46]-----	Not stated-----	Resistance; quench type, graphite tube heating element.	Optical pyrometer sighted on hole in side of heating element.	Examination of quenched specimen after heating. Formation of round balls indicated melting.
<b>Bi<sub>2</sub>O<sub>3</sub></b>	Ruff [31]-----	Pure-----	Resistance; carbon tube heating element.	Optical pyrometer sighted on specimen.	Observation of specimen during heating.
	Belladen [47]-----	Not stated-----	Resistance-----	Pt-PtRh thermocouple immersed in specimen.	Thermal analysis, cooling curves.
	Gattow and Schroder [131]	100±0.1% -----	Not stated-----	Pt-PtRh thermocouple immersed in specimen.	Thermal analysis, cooling curves.

See footnotes at end of table.

—Continued

Calibration Points		Environment	Original Temp. Scale <sup>3</sup>	Melting Points		Comments	Oxide
Materials	Temp.			Original	Int. 1948 <sup>4</sup>		
Sb.....	630..... °C	(a) Vacuum at 2 mm Hg. (b) Hydrogen.....	.....	2050..... °C	2072 *..... °C	Partial dissociation probably occurred at reduced pressure.	Al <sub>2</sub> O <sub>3</sub> — Con.
Cu-Ag eutectic.....	779.....	.....	.....	.....	.....	.....	.....
Ag.....	960.5.....	.....	.....	.....	.....	.....	.....
Cu.....	1083.....	.....	.....	.....	.....	.....	.....
Diopside.....	1391.....	.....	.....	.....	.....	.....	.....
Pt.....	1755.....	.....	.....	.....	.....	.....	.....
Not stated.....	.....	66.1 mm Hg.....	(Int. 1927).....	312.3..... Triple point.	312.3 *.....	.....	As <sub>2</sub> O <sub>3</sub> Mono- clinic type.
Water.....	(100 °C).....	Not stated.....	Not stated.....	313.....	.....	Pressure at melting not stated; melting point should be design- ated as triple point.	.....
Naphthalene.....	Not stated.....	.....	.....	.....	.....	.....	.....
Benzophe- none.....	Not stated.....	.....	.....	.....	.....	.....	.....
Not stated.....	.....	Not stated.....	Not stated.....	315.....	.....	Pressure at melting not stated; melting point should be design- ated as triple point.	.....
Not stated.....	.....	(Air).....	(Geophysical).....	450±2.....	450±2*.....	.....	B <sub>2</sub> O <sub>3</sub>
Not stated.....	Not stated.....	Not stated.....	Not stated.....	460 to 470.....	.....	Reaction between specimen and SiO <sub>2</sub> container probable.	.....
Not stated.....	.....	Hydrogen at 0.2 atmospheres.....	C <sub>2</sub> =1.433 cm deg.....	1923.....	1918.....	*; temperature uniformity be- tween specimen and target area of pyrometer not verified; reflection errors probable.	BaO
Not stated.....	.....	Vacuum.....	Not stated.....	2400 ± 100.....	.....	(*).....	BeO
Au.....	1062.4.....	Air at 15 mm Hg.....	C <sub>2</sub> =1.437 cm deg.....	2410.....	2410.....	*; reaction between specimen and graphite support probable.	.....
CaF <sub>2</sub> .....	1398.....	.....	.....	.....	.....	.....	.....
Al <sub>2</sub> O <sub>3</sub> .....	2035.....	Argon.....	Int. 1948.....	2452.....	2452.....	(*).....	.....
BeO.....	2510.....	.....	.....	.....	.....	.....	.....
Not stated.....	.....	Oxidizing.....	(Int. 1927).....	2520±30.....	2508±30.....	(*).....	.....
Not stated.....	.....	Oxidizing.....	(Int. 1927).....	2570.....	2557.....	.....	.....
Not stated.....	.....	Nitrogen.....	(Int. 1948).....	2570±30.....	2570±30.....	Specimen dropped through heat- ed furnace; thermal lag be- tween specimen and measured temperature probable.	.....
Pt.....	1755.....	Nitrogen at 4 to 10 mm Hg.....	C <sub>2</sub> =1.46 cm deg.....	2525.....	2573.....	Reaction between specimen and graphite support probable.	.....
Not stated.....	.....	Not stated.....	Not stated.....	817.....	.....	.....	Bi <sub>2</sub> O <sub>3</sub>
Not stated.....	.....	(Air).....	(Int. 1948).....	824±2.....	824±2.....	.....	.....

TABLE 1

Oxide	Reference	Purity <sup>2</sup>	Furnace type	Temperature measurement	Method
<b>Bi<sub>2</sub>O<sub>3</sub>—Con.</b>	Guertler [48]	Not stated.....	Not stated.....	Thermocouple immersed in specimen.	Thermal analysis, cooling curves.
	Levin and McDaniel [49].	Si and Fe<0.01%..... Al and Pb<0.001%..... Ag, Ca, Cr, Mg, and Mn <0.0001%.....	Resistance; quench type, Pt alloy wire-wound ceramic tube heating element.	Pt-90Pt10Rh thermocouple adjacent to specimen.	Examination with microscope of quenched specimen after heating.
<b>CaO</b>	Schumacher [43]	C.P. grade.....	Resistance; boat-shaped W strip heating element.	Optical pyrometer sighted on W strip adjacent to specimen; spectral emissivity equals 0.418.	Observation of specimen during heating.
	Kanolt [37]	Na and K—nil..... Mg—trace..... H <sub>2</sub> S metals—nil..... Cl—0.004%..... Sulfate—nil.....	Resistance; graphite tube heating element.	Optical pyrometer sighted on blackbody cavity immersed in specimen. Also, specimen acted as blackbody enclosure.	Thermal analysis, heating curves.
<b>Cr<sub>2</sub>O<sub>3</sub></b>	O' Shanskill [46]	Not stated.....	Resistance; quench type, graphite tube heating element.	Optical pyrometer sighted on hole in side of heating element.	Examination of quenched specimen after heating. Formation of round balls indicated melting.
	Doman, Barr, McNally, and Alper [50].	Not stated.....	Not stated.....	Not stated.....	Not stated.....
<b>CdO</b>	Roth [51]	Si—0.001 to 0.01%..... Al, B, Ca, Cu, Fe, Mg and Pb—0.0001 to 0.001%..... Ag—<0.0001%.....	Resistance; quench type, Pt alloy wire-wound ceramic tube heating element.	Pt-90Pt10Rh thermocouple adjacent to specimen.	Examination of quenched specimen after heating.
<b>CeO<sub>2</sub></b>	Ruff [31]	Not stated.....	Resistance; carbon tube heating element.	Optical pyrometer sighted on specimen.	Observation of specimen during heating.
	v. Wartenberg and Gurr [52].	Not stated.....	Flame; acetylene-O <sub>2</sub> .....	Not stated.....	Observation of specimen during heating.
<b>CoO</b>	Trombe [2]	Not stated.....	Not stated.....	Not stated.....	Not stated.....
	v. Wartenberg, Reusch, and Saran [44].	Ni Free.....	Flame; oil.....	Optical pyrometer sighted on specimen.	Observation of specimen suspended in furnace during heat. Rounded corners indicated melting.
	v. Wartenberg and Prophet [53].	99.9%.....	Flame; oxyhydrogen or gas-air.	Optical pyrometer.....	Not stated.....
<b>Cr<sub>2</sub>O<sub>3</sub></b>	v. Wartenberg and Gurr [52].	Pure.....	Not stated.....	Optical pyrometer.....	Not stated.....
	Ruff [31]	Not stated.....	Resistance; carbon tube heating element.	Optical pyrometer sighted on specimen.	Observation of specimen during heating.
	Kanolt [37]	Iron—0.01%..... Chromate—trace..... Sulphate—0.005%.....	Resistance; graphite tube heating element.	Optical pyrometer sighted on blackbody cavity immersed in specimen.	Thermal analysis, heating curves.

See footnotes at end of table.

—Continued

Calibration Points		Environment	Original Temp. Scale <sup>3</sup>	Melting Points		Comments	Oxide
Materials	Temp.			Original	Int. 1948 <sup>4</sup>		
Not stated	°C	Not stated	Not stated	°C	°C		<b>Bi<sub>2</sub>O<sub>3</sub>— Con.</b>
NaCl	800.4	Air	Int. 1948	825±3	825±3*		
Au	1063						
Not stated		Hydrogen at 0.2 atm	C <sub>2</sub> =1.433 cm deg.	2576	2565	<sup>5</sup> ; temperature uniformity be- tween specimen and target area of pyrometer not verified; reflection errors probable.	<b>CaO</b>
Sb	630	Hydrogen	( <sup>6</sup> )	2572	2614*		
Cu—Ag Eutectic.	779						
Ag	960.5						
Cu	1083						
Diopside	1391						
Pt	1755						
Not stated		Nitrogen	(Int. 1948)	2620	2620	Specimen dropped through heated furnace; thermal lag between specimen and mea- sured temperature probable.	
Not stated		Not stated	(Int. 1948)	2630	2630	Melting point taken from ab- stract.	
Not stated		Not stated	Int. 1948	>1500	>1500	Author noted that CdO sub- limes rapidly.	<b>CdO</b>
Pt	1755	Not stated	C <sub>2</sub> =1.46 cm deg.	1950	1973	Reaction between specimen and ZrO <sub>2</sub> support probable.	<b>CeO<sub>2</sub></b>
Not stated		Air	(Int. 1927)	>2600	>2600		
Not stated		Not stated	Not stated	2800		Survey; not original data.	
Not stated		Oxidizing	(Int. 1927)	1800±20	1795±20	( <sup>6</sup> )	<b>CoO</b>
Not stated		Air	(Int. 1927)	1810	1805	Oxidation state of cobalt not investigated.	
Not stated		Air	(Int. 1927)	1935	1929	Authors stated this melting point to be in error; see pre- ceding reference [53] for cor- rected value.	
Pt	1755	Nitrogen at (a) 30 mm Hg (b) 1 atmosphere.	C <sub>2</sub> =1.46 cm deg.	(a) 1830 to 2080 (b) 1960	1849 to 2107 1983	<sup>6</sup> ; reaction between specimen and graphite support proba- ble. Author concluded re- duction of specimen occurred.	<b>Cr<sub>2</sub>O<sub>3</sub></b>
Sb	630	Vacuum	( <sup>6</sup> )	1990	2011	Oxidation state of chromium not investigated. Reduction is probable.	
Cu—Ag Eutectic.	779						
Ag	960.5						
Cu	1083						
Diopside	1391						
Pt	1755						

TABLE 11

Oxide	Reference	Purity <sup>a</sup>	Furnace type	Temperature measurement	Method
<b>Cr<sub>2</sub>O<sub>3</sub>—Con.</b>	Wilde and Rees [54]-----	Not stated-----	Resistance; graphite tube heating element.	Photoelectric pyrometer sighted on specimen.	Thermal analysis, heating curves. Also, observation of specimen during heating.
	Bunting [55]-----	99.9%-----	Induction; Ir alloy susceptor, button-shaped.	Optical pyrometer sighted on small cavity adjacent to specimen cavity.	Examination of specimen after heating.
	Bunting [32]-----	C. P. grade-----	Induction; Ir susceptor, button-shaped.	Optical pyrometer sighted on small cavity adjacent to specimen cavity.	Examination of specimen after heating.
	v. Wartenberg and Reusch [56].	Analytical grade-----	(Flame); "Fletscher oven".	Optical pyrometer sighted on specimen.	Observation of specimen suspended in furnace during heating.
	McNally, Peters, and Ribbe [33].	Co <sub>3</sub> O <sub>4</sub> , SiO <sub>2</sub> , and V <sub>2</sub> O <sub>5</sub> —0.01-0.1% Al <sub>2</sub> O <sub>3</sub> , MgO, and Fe <sub>2</sub> O <sub>3</sub> —0.005-0.05% NiO—0.001-0.01% MnO and TiO <sub>2</sub> —1-10 ppm.	Induction; graphite tube susceptor.	Optical pyrometer sighted on specimen.	Observation of specimen suspended in furnace during heating.
	v. Wartenberg and Eckhardt [57].	Not stated-----	Flame; acetylene-O <sub>2</sub> -----	Optical pyrometer sighted on specimen.	Observation of specimen suspended in furnace during heating.
<b>Cs<sub>2</sub>O</b>	Rengade [58]-----	Purified-----	Flame-heated Al block-----	Thermocouple-----	Thermal analysis, cooling curves. Also, observation of specimen during heating.
<b>Cu<sub>2</sub>O</b>	Ruer and Nakamoto [59].	Not stated-----	Resistance; carbon tube heating element.	Pt-PtRh thermocouple immersed in specimen.	Examination of specimen after heating.
	v. Wartenberg, Reusch, and Saran [44].	Pure-----	Flame; gas-oxygen-----	Optical pyrometer-----	Observation of specimen suspended in furnace during heating.
	Roberts and Smyth [60].	0.0083% impurities-----	Not stated-----	Pt-PtRh thermocouple-----	Intersection of pressure-temperature-composition curves.
<b>Dy<sub>2</sub>O<sub>3</sub></b>	Wisnyl and Pijanowski [61].	Not stated-----	Resistance; V-shaped W strip heating element.	Optical pyrometer sighted on specimen.	Observation of specimen during heating. Apparent disappearance of specimen indicated melting.
<b>Er<sub>2</sub>O<sub>3</sub></b>					
<b>Eu<sub>2</sub>O<sub>3</sub></b>	Wisnyl and Pijanowski [61].	Not stated-----	Resistance; V-shaped W strip heating element.	Optical pyrometer sighted on specimen.	Observation of specimen during heating. Apparent disappearance of specimen indicated melting.
	Schneider [62]-----	Ba, Cu, Er, and Si—<0.01%; Cr, Cu, Fe, Mg, and Ni—<0.001%.	Induction; crucible-shaped Ir susceptor.	Optical pyrometer sighted on small hole in crucible lid.	Examination of specimen after heating. Flat button indicated melting.

See footnotes at end of table.

—Continued

Calibration Points		Environment	Original Temp. Scale <sup>4</sup>	Melting Points		Comments	Oxide
Materials	Temp.			Original	Int. 1948 <sup>4</sup>		
Al <sub>2</sub> O <sub>3</sub>	2040 °C	Air	(Int. 1927)	2060±25 °C	2053±25 °C	<sup>5</sup> ; reduction of specimen probable.	Cr <sub>2</sub> O <sub>3</sub> —Con.
Al <sub>2</sub> O <sub>3</sub>	2040	(Air)	(Int. 1927)	2140±25	2133±25	Temperature uniformity between specimen and target area of pyrometer not verified. Bunting in later publication [32] stated 2140 °C value is low because of reduction during calcine.	
Not stated		(Air)	(Int. 1927)	2275±25	2266±25	Temperature uniformity between specimen and target area of pyrometer not verified.	
Pt	(1773)	Air	(Int. 1927)	2275±25	2266±25	( <sup>6</sup> )	
Al <sub>2</sub> O <sub>3</sub>	2044	(a) Nitrogen (b) Air	Int. 1948	(a) 2315 (b) 2330	2315 2330	<sup>5</sup> ; reduction of specimen is probable.	
Not stated		Air	(Int. 1927)	2435±10	2424±10	<sup>5</sup> ; reduction of specimen is probable.	
Not stated		Nitrogen	Not stated	490±10			Cs <sub>2</sub> O
Cu	1084	Nitrogen	Not stated	1222	1222		Cu <sub>2</sub> O
Ni	1451						
Fe	1528						
Not stated		Oxidizing	(Int. 1927)	1230±20	1229±20	( <sup>6</sup> )	
Quartz trans- formation	Not stated	0.6 mm Hg Pressure	Geophysical	1235	1236*		
Au	1062.6						
Al <sub>2</sub> O <sub>3</sub>	2040	Either helium, hydro- gen, or vacuum.	Int. 1948	2340±10	2340±10	<sup>6</sup> ; author's stated emissivity of approximately 0.9 to 0.95 is probably overestimated. Reflection error probable.	Dy <sub>2</sub> O <sub>3</sub>
Al <sub>2</sub> O <sub>3</sub>	2040	Either helium, hydro- gen or vacuum.	Int. 1948	2050±30	2050±30	No melting point data located in the literature.	Er <sub>2</sub> O <sub>3</sub>
Al <sub>2</sub> O <sub>3</sub>	2040		Int. 1948			<sup>6</sup> ; author's stated emissivity of approximately 0.9 to 0.95 is probably overestimated. Reflection errors and reduction in vacuum or hydrogen probable.	Eu <sub>2</sub> O <sub>3</sub>
Au	1063	Air	Int. 1948	2240±10	2240±10*		
Pt	1769						
Rh	1960						

TABLE 11

Oxide	Reference	Purity <sup>1</sup>	Furnace type	Temperature measurement	Method
<b>FeO</b>	Chipman and Marshall [63].	Fe and C—<0.005%; S <0.004%; Si<0.008%; Mn<0.003%; Cu<0.004%; Cu<0.004%; P< 0.002%; Ni<0.031%.	Resistance; SiC tube heating element.	Pt-90Pt10Rh thermocouple adjacent to specimen.	Weight loss data; specimen suspended from balance arm. Sudden weight loss indicated melting.
	Hay, Howat, and White [64].	Not stated-----	Resistance; Mo wire-wound ceramic tube heating element.	Mo-W thermocouple adjacent to the specimen.	Differential thermal analysis.
	Darken and Gurry [65].	Fe and C—0.012%; Mn<0.01%; P—0.004%; S—0.003%; Cu—0.045%; Si—0.007%; Ni—0.02%; Pb—0.04%.	Resistance; quench type, Pt wire-wound ceramic tube heating element.	Pt-PtRh thermocouple; temperature measured before and after each determination.	Examination of specimen after heating.
	Bowen and Schairer [66]	Not stated-----	Resistance; quench type, Pt wire-wound ceramic tube heating element.	Pt-90Pt10Rh thermocouple; temperature measured before and after each determination.	Examination with microscope of quenched specimen after heating.
<b>Fe<sub>2</sub>O<sub>3</sub></b>	Moruzzi and Shafer [67]	High purity-----	Resistance; quench type, Pt wire-wound ceramic tube heating element.	Thermocouple adjacent to specimen.	Observation of rod-shaped specimen during heating. Rounded corners indicated melting.
	Darken and Gurry [65]	Not stated-----	Resistance; quench type, Pt wire-wound ceramic tube heating element.	Pt-PtRh thermocouple; temperature measured before and after each determination.	Examination of specimen after heating. Formation of droplets indicated melting.
	Greig, Posnjak, Merwin, and Sosman [68]	Fe <sub>2</sub> O <sub>3</sub> —99.7%; FeO—0.03%; SiO <sub>2</sub> —0.01%; Na <sub>2</sub> O—0.04%; TiO <sub>2</sub> —0.01%; MnO—0.01%; P <sub>2</sub> O <sub>5</sub> —0.005%	Resistance; quench type, Pt wire-wound ceramic tube heating element.	Pt-90Pt10Rh thermocouple adjacent to specimen.	Examination with microscope of quenched specimen after heating.
	v. Wartenberg and Eckhardt [57]	Commercially pure-----	Flame; acetylene-oxygen.	Optical pyrometer sighted on specimen.	Observation of specimen suspended in furnace during heating.
<b>Ga<sub>2</sub>O<sub>3</sub></b>	Hill, Roy, and Osborn [69].	Not stated-----	Resistance, U-shaped 60 Pt-40Rh strip heating element.	Optical pyrometer-----	Examination with microscope of specimen after heating.
	v. Wartenberg and Reusch [56].	Not stated-----	(Flame) "Fletscher Oven".	Optical pyrometer sighted on specimen.	Observation of specimen suspended in furnace during heating.
	Schneider and Waring [70].	99.9%-----	(a) Induction; crucible-shaped Ir susceptor.  (b) Resistance; quench type, Pt alloy wire-wound ceramic tube heating element.	(a) Optical pyrometer sighted on small hole in crucible lid.  (b) 95Pt5Rh-80Pt20Rh thermocouple adjacent to specimen.	(a) and (b) Examination of specimen after heating.
<b>Gd<sub>2</sub>O<sub>3</sub></b>	Wisnyi and Pijanowski [61].	Not stated-----	Resistance; V-shaped W strip heating element.	Optical pyrometer sighted on specimen.	Observation of specimen during heating. Apparent disappearance of specimen indicated melting.

See footnotes at end of table.

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Calibration Points		Environment	Original Temp. Scale <sup>3</sup>	Melting Points		Comments	Oxide
Materials	Temp.			Original	Int. 1948 <sup>4</sup>		
Cu.....	°C Not stated.....	Slightly oxidizing.....	(Int. 1927).....	1369.....	1368.....		FeO
Ni.....	Not stated.....						
Pd.....	1555						
Not stated.....		Nitrogen at one atm.....	(Int. 1927).....	1370.....	1368.....		
Not stated.....		Nitrogen at one atm.....	(Int. 1927).....	1371±1.....	1369±1*.....		
Li <sub>2</sub> SiO <sub>3</sub> .....	1201.....	Nitrogen—slightly oxidizing.	Geophysical.....	1380±5.....	1382±5.....		
Diopsid.....	1391.5						
Pd.....	1549.5						
Not stated.....		Air.....	(Int. 1948).....	1591.....	1591.....		Fe <sub>3</sub> O <sub>4</sub>
Not stated.....		(a) Oxygen at 1 atm..... (b) Air at 1 atm..... (c) Oxygen at 0.0575- atm.....	(Int. 1927).....	(a) 1583±2..... (b) 1594±2..... (c) 1597±2.....	1580±2*..... 1591±2*..... 1594±2*.....		
Au.....	1062.6.....	Small oxygen pres- sure	Geophysical.....	1591±5.....	1594±5*.....		
Diopsid.....	1391.5						
Pd.....	1549.5						
Pt.....	1755						
Not stated.....		Air.....	(Int. 1927).....	1650.....	1647.....	( <sup>5</sup> ).....	
15% CaO..... 85% SiO <sub>2</sub>	1710.....	Not stated.....	Int. 1948.....	1725±15.....	1725±15.....	<sup>6</sup> ; temperature uniformity be- tween specimen and target area of pyrometer not verified.	Ga <sub>2</sub> O <sub>3</sub>
Not stated.....		Air.....	(Int. 1927).....	1740±25.....	1736±25.....	( <sup>6</sup> ).....	
(a) Au..... Pd..... Pt..... Rh.....	1063..... 1552 1769 1960	(a) and (b) Air.....	Int. 1948.....	(a) and (b) 1795±15.	1795±15*.....		
(b) Au..... Pd..... Pt.....	1063 1552 1769						
Al <sub>2</sub> O <sub>3</sub> .....	2040.....	Either helium, hy- drogen, or vacuum.	Int. 1948.....	2330±20.....	2330±20.....	<sup>6</sup> ; authors stated emissivity of approximately 0.9 to 0.95 is probably overestimated. Re- flection errors probable.	Gd <sub>2</sub> O <sub>3</sub>

TABLE 11

Oxide	Reference	Purity <sup>2</sup>	Furnace type	Temperature measurement	Method
<b>Gd<sub>2</sub>O<sub>3</sub>—Con.</b>	Curtis and Johnson [71].	Y—0.2%; La<0.05%; Ce—0.5%; Pr<0.1%; Nd—0.5%; Sm—0.5%; Eu—0.5%; Dy—0.5%; Tb—2.0%; Ho<0.5%; Er<0.05%; Yb<0.05%; Lu—0.05%.	Not stated.....	Optical pyrometer sighted on specimen.	Observation of rod shaped specimen during heating. Definite fluidity of specimen indicated melting.
<b>GeO<sub>2</sub> Quartz type.</b>	Schwarz, Schenk, and Giese [72]	Not stated.....	Resistance; quench type, Pt wire-wound ceramic tube heating element.	Thermocouple.....	Examination with microscope of quenched specimen after heating. Disappearance of last crystal indicated melting.
	Laubengayer and Morton [73].	<0.01% impurities.....	Resistance; quench type, Pt wire-wound ceramic tube heating element.	Pt-PtRh thermocouple adjacent specimen.	Examination with microscope of quenched specimen after heating. Disappearance of last crystal indicated melting.
<b>HfO<sub>2</sub></b>	Clausing [74].....	ZrO <sub>2</sub> <1%.....	Resistance; bar-shaped W strip heating element.	Optical pyrometer sighted on small cavity adjacent to specimen cavity.	Observation of specimen during heating. Flowing specimen indicated melting.
	Mark [27].....	Fe and Si—0.5-1%; Zr—5%.	Resistance; graphite or metal strip heating element.	Optical pyrometer sighted on specimen.	Observation of specimen during heating.
	Henning [75].....	"Pure theoretical".....	Resistance; W tube heating element.	Optical pyrometer sighted on specimen.	Observation of specimen during heating.
	Curtis, Doney, and Johnson [76].	Zr-2 ppm; Ti-1000 ppm; Al-600 ppm; Fe-100 ppm; Si—100 ppm.	Flame; oxyacetylene.....	Optical pyrometer.....	Not stated.....
<b>Ho<sub>2</sub>O<sub>3</sub></b>					
<b>In<sub>2</sub>O<sub>3</sub></b>	Schneider [62].....	Al, Ca, Cu, Fe, Mg, Ni, P, and Si, each <0.01%; Er <0.001%; Ag and Mn <0.00001%	Induction; crucible-shaped Ir susceptor.	Optical pyrometer sighted on small hole in crucible lid.	Examination of specimen after heating. Flat button indicated melting.
<b>IrO<sub>2</sub></b>	Cordfunke and Meyer [77].	Not stated.....	Resistance; "Sillbar".....	Pt-90Pt10Rh thermocouple.....	Vapor pressure data.....
<b>K<sub>2</sub>O</b>					
<b>La<sub>2</sub>O<sub>3</sub></b>	Ruff [31].....	Contained didymium impurities.	Resistance; carbon tube heating element.	Optical pyrometer sighted on specimen.	Observation of specimen during heating.
	Lambertson and Gunzel [30].	99%.....	Resistance; W heating element.	Optical pyrometer sighted on specimen. Specimen not visible.	Examination of specimen after heating.
	v. Wartenberg and Reusch [56].	100%.....	(Flame) "Fletscher oven".....	Optical pyrometer sighted on specimen.	Observation of specimen during heating.
<b>Li<sub>2</sub>O</b>	Chemistry and Physics Handbook [4].	Not stated.....	Not stated.....	Not stated.....	Not stated.....

See footnotes at end of table.

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Calibration Points		Environment	Original Temp. Scale <sup>3</sup>	Melting Points		Comments	Oxide
Materials	Temp.			Original	Int. 1948 <sup>4</sup>		
Not stated	°C	Air	(Int. 1948)	2350±50	2350±50	<sup>5</sup> ; authors stated an attempt was made to correct for nonblackbody conditions.	Gd <sub>2</sub> O <sub>3</sub> —Con.
Na <sub>2</sub> SO <sub>4</sub>	884.7	Air	(Geophysical)	1115±3	1115±3*		GeO; Quartz type.
K <sub>2</sub> SO <sub>4</sub>	1069.1						
Li <sub>2</sub> SiO <sub>3</sub>	1201.8						
Not stated		(Air)	(Int. 1927)	1116±4	1116±4*		
Not stated		Hydrogen	(Int. 1927)	2774±25	2758±25	<sup>5</sup> ; temperature uniformity between specimen and target area of pyrometer not verified.	HfO <sub>2</sub>
Not stated		Neutral	(Int. 1948)	2770	2770	<sup>5</sup> ; reflection errors probable.	
Not stated		Nitrogen or hydrogen	Not stated	2812±25		<sup>5</sup> ; melting point determined for impure HfO <sub>2</sub> sample and extrapolated to 100% es	
Not stated		Not stated	(Int. 1948)	2900±25	2900±25		
						No melting point data located in the literature.	Ho <sub>2</sub> O <sub>3</sub>
Au	1063	Air	Int. 1948	1910±10	1910±10		In <sub>2</sub> O <sub>3</sub>
Pt	1769						
Rh	1960						
Not stated		Oxygen at 1 atm	Int. 1948	1100 dissociates	1100		IrO <sub>2</sub>
						No melting point data located in the literature. Because of the reactivity of K <sub>2</sub> O, it is unlikely that a melting point can be determined.	K <sub>2</sub> O
Pt	1755	Not stated	C <sub>2</sub> =1.46 cm deg	1840	1859	<sup>5</sup> ; reaction between specimen and ZrO <sub>2</sub> support probable.	La <sub>2</sub> O <sub>3</sub>
Pt	(1769)	Helium	(Int. 1948)	2210±20	2210±20		
Pt	(1773)	Air	(Int. 1927)	2315	2307	( <sup>5</sup> )	
Not stated		Not stated	Not stated	>1700		Survey, not original data. Because of the reactivity of Li <sub>2</sub> O it is unlikely that a melting point can be determined.	Li <sub>2</sub> O

TABLE 11

Oxide	Reference	Purity <sup>2</sup>	Furnace type	Temperature measurement	Method
<b>Lu<sub>2</sub>O<sub>3</sub></b>					
<b>MgO</b>	Ruff [31].	Contained: H <sub>2</sub> CO <sub>3</sub> , H <sub>2</sub> O, Fe <sub>2</sub> O <sub>3</sub> , Si(OH) <sub>4</sub> and NaCl.	Resistance; carbon tube heating element.	Optical pyrometer sighted on specimen.	Observation of specimen during heating.
	Kelley [1].	Not stated.	Not stated.	Not stated.	Not stated.
	McNally, Peters, and Ribbe [33].	Co <sub>3</sub> O <sub>4</sub> and WO <sub>3</sub> —0.01-0.1%; SiO <sub>2</sub> and Fe <sub>2</sub> O <sub>3</sub> —0.005-0.05%; Al <sub>2</sub> O <sub>3</sub> and Cr <sub>2</sub> O <sub>3</sub> —0.001-0.01%; B <sub>2</sub> O <sub>3</sub> , V <sub>2</sub> O <sub>5</sub> , TiO <sub>2</sub> , and CuO—5-50 ppm; MnO and Ag <sub>2</sub> O—1-10 ppm.	Induction; graphite tube susceptor.	Optical pyrometer sighted on small hollow graphite cylinder adjacent to specimen.	Observation of specimen during heating.
	Kanolt [37].	CaO—none. Al <sub>2</sub> O <sub>3</sub> —0.0002%. Fe—0.0005%. CO <sub>2</sub> —0.130%. Cl—0.275%. SO <sub>3</sub> —0.001%. HNO <sub>3</sub> —none.	Resistance; graphite tube heating element.	Optical pyrometer sighted on blackbody cavity immersed in specimen.	Thermal analysis; heating curves.
<b>MnO</b>	Tiede and Birnbrauer [23].	Not stated.	Arc.	Optical pyrometer sighted on specimen.	Observation of specimen during heating.
	White, Howat and Hay [78].	Not stated.	Resistance; Mo wire-wound ceramic tube heating element.	Mo-W thermocouple adjacent to the specimen.	Differential thermal analysis.
<b>Mn<sub>2</sub>O<sub>4</sub></b>	v. Wartenberg and Prophet [53].	Not stated.	Flame; oxyhydrogen or gas-air.	Optical pyrometer.	Not stated.
	Van Hook and Keith [79].	Not stated.	Resistance; quench type, Pt wire-wound ceramic tube heating element.	Pt-90Pt10Rh thermocouple adjacent to specimen.	Examination with microscope of quenched specimen after heating.
	Ranganathan, Mackean and Muan [80].	Reagent grade.	Resistance; Mo wire-wound ceramic tube heating element.	Pt-90Pt10Rh thermocouple adjacent to specimen.	Examination with X-rays and microscope of quenched specimen after heating.
	v. Wartenberg, Reusch and Saran [44].	Commercially pure.	Flame; oil-oxygen.	Optical pyrometer.	Observation of specimen suspended in furnace during heating.
	v. Wartenberg and Gurr [52].	Pure.	Not stated.	Optical pyrometer.	Not stated.
<b>MoO<sub>3</sub></b>	Carnelley [81].	Not stated.	Flame; "Bunsen lamp".	Mercury thermometer.	Not stated.
	Groschuff [82].	Not stated.	Not stated.	Not stated.	Thermal analysis; heating and cooling curves.
	v. Jaeger and Germs [83].	Purified.	Resistance; nichrome wire heating element.	Thermocouple immersed in specimen.	Thermal analysis, heating curves.
	Hoermann [84].	Not stated.	Resistance; Pt wire heating element.	Pt-PtRh thermocouple immersed in specimen.	Observation of specimen during heating.

See footnotes at end of table.

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Calibration Points		Environment	Original Temp. Scale <sup>3</sup>	Melting Points		Comments	Oxide
Materials	Temp.			Original	Int. 1948 <sup>4</sup>		
	°C			°C	°C		
Pt.....	1755.....	(a) Nitrogen at 10-30 mm Hg. (b) Nitrogen at 1 atm (c) Nitrogen at 10-30 mm Hg. (d) Nitrogen at 1 atm (e) Nitrogen at 10-30 mm Hg.	C <sub>2</sub> =1.46 cm deg.	(a) 2120..... (b) 2250-2280 (c) 2450 (d) 2500 (e) 2550	2150..... 2285-2316 2494 2546 2599	No melting point data located in the literature. <sup>5</sup> ; various melting points obtained using different shaped specimens and graphite supports. Reaction with graphite and partial reduction of MgO occurred.	Lu <sub>2</sub> O <sub>3</sub>
Not stated.....	Not stated.....	Not stated.....	Not stated.....	2642.....		Survey; not original data.	MgO
Al <sub>2</sub> O <sub>3</sub> .....	2044.....	Nitrogen.....	Int. 1948.....	2825±20.....	2825±20.....	Temperature uniformity between specimen and target area of pyrometer not verified.	
Sb.....	630.....	Carbon monoxide and nitrogen at atmospheric pressure.	(t).....	2800.....	2852*.....		
Cu-Ag eutectic.....	779.....						
Ag.....	960.5.....						
Cu.....	1083.....						
Diopside.....	1391.....						
Pt.....	1755.....						
Not stated.....	Not stated.....	Vacuum.....	Not stated.....	1650.....		(t).....	MnO
Fe transformation.....	1400.....	Not stated.....	(Int. 1927).....	1785.....	1781.....		
Fe.....	1530.....						
Not stated.....	Not stated.....	Air.....	(Int. 1927).....	1560.....	1557.....		Mn <sub>3</sub> O <sub>4</sub>
CaSiO <sub>3</sub> .....	1544.....	Air.....	(Geophysical).....	1562.....	1564*.....		
CaSiO <sub>3</sub> .....	1546.....	Air.....	Int. 1948.....	1567±5.....	1567±5*.....		
90SiO <sub>2</sub> :10CaO.....	1707.....						
Pt.....	1769.....						
Not stated.....	Not stated.....	Air.....	(Int. 1927).....	1590±20.....	1587±20.....	(t).....	
Not stated.....	Not stated.....	Air.....	(Int. 1927).....	1705.....	1701.....		
Not stated.....	Not stated.....	Not stated.....	Not stated.....	759.....			MoO <sub>3</sub>
Not stated.....	Not stated.....	Air.....	Not stated.....	791.....			
Not stated.....	Not stated.....	Oxidizing.....	Not stated.....	795.....			
Not stated.....	Not stated.....	Not stated.....	(Int. 1927).....	795.....			

TABLE 1

Oxide	Reference	Purity <sup>a</sup>	Furnace type	Temperature measurement	Method
<b>MoO<sub>3</sub>—Con.</b>	Rieck [85]	Purified	Flame; gas-air	Pt-PtRh thermocouple immersed in specimen.	(Thermal analysis)
	Cosgrove and Snyder [86]	Trace—heavy and alkali metals.	Resistance; nichrome wire-wound ceramic tube heating element.	Pt-PtRh thermocouple	Thermal analysis, cooling curves.
<b>Na<sub>2</sub>O</b>					
<b>Nb<sub>2</sub>O<sub>5</sub></b>	Brauer [87]	TiO <sub>2</sub> —0.28%	"Tammann oven"	Optical pyrometer sighted on specimen.	Observation of specimen during heating. Flowing specimen indicated melting.
	Shafer and Roy [88]	Ta <sub>2</sub> O <sub>5</sub> <0.2%	Resistance; Pt wire-wound ceramic tube heating element.	Thermocouple	Not stated
	Ibrahim, Bright, and Rowland [89]	99+%	Resistance; Pt wire-wound tube heating element.	Pt-90Pt10Rh thermocouple	Observation of cone shaped specimen during heating. Complete slumping indicated melting.
	Roth and Waring [90]	Si <0.01%; Ca and Mg <0.001%; As, Cu, and Ta perhaps present.	Resistance; quench type Pt alloy wire-wound ceramic tube heating element.	Pt-90Pt10Rh thermocouple adjacent to the specimen.	Examination of quenched specimen after heating.
	Reisman and Holtzberg [91]	Ta <sub>2</sub> O <sub>5</sub> <0.2%	Resistance; Pt wire-wound ceramic tube heating element.	Pt-90Pt10Rh thermocouple immersed in specimen.	Differential thermal analysis.
	Roth and Waring [92]	99.7+%	Resistance; quench type, Pt alloy wire-wound ceramic tube heating element.	Pt-90Pt10Rh thermocouple adjacent to specimen.	Examination of quenched specimen after heating.
	Holtzberg, Reisman, Berry, and Berkenblit [93]	Ta<0.2%	Resistance; kanthal wire heating element.	Pt-90Pt10Rh thermocouple immersed in specimen.	Differential thermal analysis.
	Diamond and Schneider [28]	99.7+%	Solar	Optical pyrometer sighted on specimen.	Observation of bar-shaped specimen during cooling.
	Roth and Coughanour [94]	99.7+%	Resistance; Pt alloy heating element.	Pt-90Pt10Rh thermocouple	Examination of pyramid-shaped specimen after heating.
	Orr [95]	Si<0.03%; Mg<0.05%; Ti<0.01%	Not stated	Thermocouple	Not stated
	Ruff [31]	Ta <sub>2</sub> O <sub>5</sub> —1-2%	Resistance; carbon tube heating element.	Optical pyrometer sighted on specimen.	Observation of specimen during heating.
<b>Nd<sub>2</sub>O<sub>3</sub></b>	Lambertson and Ganzel [30]	99%	Resistance; W heating element.	Optical pyrometer sighted on specimen. Specimen not visible.	Examination of specimen after heating.
<b>NiO</b>	Merica and Wartenberg [96]	Not stated	Not stated	Optical pyrometer	Not stated
	v. Wartenberg and Prophet [53]	99.9%	Flame; oxyhydrogen or gas-air	Optical pyrometer	Not stated
	v. Wartenberg and Gurr [52]	Pure	Not stated	Optical pyrometer	Not stated

See footnotes at end of table.

—Continued

Calibration Points		Environment	Original Temp. Scale <sup>3</sup>	Melting Points		Comments	Oxide
Materials	Temp.			Original	Int. 1948 <sup>4</sup>		
NaCl.....	°C Not stated.....	Air.....	(Int. 1927).....	795±2.....	795±2.....		MoO <sub>3</sub> Con.
Sn, Pb, Al, Cu, Ag, Sl. Steam.....	Not stated.....	Nitrogen at 1 atmo.....	Int. 1948.....	795.36.....	795.36*.....		
Not stated.....						No melting point data located in the literature. Because of the reactivity of Na <sub>2</sub> O, it is unlikely that a melting point can be determined.	Na <sub>2</sub> O
Not stated.....		Oxygen.....	(Int. 1927).....	1460±5.....	1458±5.....	(?).....	Nb <sub>2</sub> O <sub>5</sub>
Not Stated.....		Not Stated.....	(Int. 1948).....	1465±5.....	1465±5.....		
Au..... Pd..... Pseudowol- lastonite.....	Not stated..... Not stated..... Not stated.....	Air.....	Int. 1948.....	1479.....	1479.....	Authors stated that recorded temperatures may be as much as 20 °C low.	
Au.....	1063.....	Air.....	Int. 1948.....	1485±5.....	1485±5*.....		
K <sub>2</sub> CO <sub>3</sub> ..... NaCl.....	Not stated..... Not stated.....	Oxygen plus air.....	(Int. 1948).....	1486±5.....	1486±5*.....		
Au..... Barium disili- cate.....	1063..... 1420.....	Air.....	Int. 1948.....	1487.....	1487.....		
KNbO <sub>3</sub> ..... Transforma- tions.....	215..... 425.....	(Air).....	(Int. 1948).....	1491±2.....	1491±2*.....		
None.....		Air.....	Int. 1948.....	1496.....	1496.....	<sup>5</sup> ; emissivity stated by authors to be about unity.	
Not stated.....		Air.....	Int. 1948.....	1500 ± 10.....	1500 ± 10.....		
Au.....	(1063).....	Not stated.....	(Int. 1948).....	1512.....	1512.....		
Pt.....	1755.....	Not stated.....	C <sub>2</sub> -1.46 cm deg.....	1520.....	1530.....	<sup>5</sup> ; reaction between specimen and Mgo support probable.	
Pt.....	(1769).....	Helium.....	(Int. 1948).....	2272 ± 20.....	2272 ± 20.....		Nd <sub>2</sub> O <sub>3</sub>
Not stated.....		(a) Vacuum (b) Air.....	Not stated.....	(a) 1552 (b) 1660.....			NiO
Not stated.....		Air.....	(Int. 1927).....	1990.....	1984.....	(?).....	
Not stated.....		Air.....	(Int. 1927).....	2090.....	2083.....	<sup>5</sup> ; authors stated this melting point to be in error. See preceding reference [53] for corrected value.	

TABLE 11

Oxide	Reference	Purity <sup>2</sup>	Furnace type	Temperature measurement	Method
<b>OsO<sub>4</sub></b> yellow-type	v. Wartenberg [97]-----	Not stated-----	Resistance-----	Thermocouple adjacent to specimen.	Observation of specimen during heating. Also from intersection of vapor pressure curves.
	Chemistry and Physics Handbook [4].	Not stated-----	Not stated-----	Not stated-----	Not stated-----
	Kelley [1]-----	Not stated-----	Not stated-----	Not stated-----	Not stated-----
<b>P<sub>2</sub>O<sub>5</sub></b>	Hoeflaake and Scheffer [98].	Purified-----	Resistance; "oven"-----	Pt-PtRh thermocouple-----	Observation of specimen during heating. Also from vapor pressure data.
<b>PbO</b> yellow type.	Balladen [47]-----	Not stated-----	Resistance-----	Pt-PtRh thermocouple immersed in specimen.	Thermal analysis, cooling curves.
	Hilpert and Weiller [99].	Not stated-----	Resistance-----	Pt-PtRh thermocouple immersed in specimen.	Thermal analysis-----
	v. Jaeger and Germs [83].	(a) 99.2%----- (b) Not stated (c) Not stated	Resistance; nichrome heating element.	Thermocouple immersed in specimen.	Thermal analysis-----
	Schenck and Rassbach [100].	Pure-----	Resistance-----	Pt-PtRh thermocouple-----	Thermal analysis, cooling curves.
	Geller, Creamer, and Bunting [101].	0.02% total impurities-----	Not stated-----	Pt-PtRh thermocouple-----	Differential thermal analysis-----
	Krakau [102].-----	0.1-0.2% impurities-----	Not stated-----	Pt-PtRh thermocouple-----	Thermal analysis, cooling curves.
	Kroll [103].-----	Not stated-----	Resistance; Pt heating element.	Thermocouple immersed in specimen.	Thermal analysis, cooling curves; also observation of specimen.
	Cooper, Shaw, and Loomis [104].	Pure-----	Resistance-----	Pt-PtRh thermocouple-----	(Thermal analysis.)-----
<b>PdO<sub>2</sub></b>					
<b>Pr<sub>6</sub>O<sub>11</sub></b>					
<b>PtO<sub>2</sub></b>	Chemistry and Physics Handbook [4].	Not stated-----	Not stated-----	Not stated-----	Not stated-----
<b>Rb<sub>2</sub>O</b>					
<b>Re<sub>2</sub>O<sub>7</sub></b>	Kelley [1]-----	Not stated-----	Not stated-----	Not stated-----	Not stated-----
<b>Rh<sub>2</sub>O<sub>3</sub></b>					

See footnotes at end of table.

—Continued

Calibration Points		Environment	Original Temp. Scale <sup>3</sup>	Melting Points		Comments	Oxide
Materials	Temp.			Original	Int. 1948 <sup>4</sup>		
Not stated	°C	About 11 mm Hg	Not stated	40.1	°C	Crystal form not indicated. The reported melting point is probably determined from mixture of both the white and yellow types.	OsO <sub>4</sub> yellow type
Not stated		Not stated	Not stated	41		Survey, not original data	
Not stated		Not stated	Not stated	56		Survey, not original data	
Naphthaline	218	4600 mm Hg	Not stated	569	569		P <sub>2</sub> O <sub>5</sub>
Benzo- phenone.	306			Triple point.			
Sulphur	444.5						
Sb	630						
Not stated		Not stated	Not stated	870			PbO yellow type.
Not stated		Air	Not stated	876			
Not stated		Oxidizing	Not stated	(a) 877 (b) 879 (c) 879			
Not stated		Air	Not stated	879			
K <sub>2</sub> SO <sub>4</sub> Trans- formation.	583±1	Air	Int. 1927	886	886*		
KCl	770.3						
K <sub>2</sub> SO <sub>4</sub>	1069.1						
Not stated		Air	(Int. 1927)	886	886		
Not stated		Air	Not stated	888			
Not stated		(Air)	Not stated	888			PdO <sub>2</sub>
						No melting point data located in the literature. The oxide probably dissociates to the metal before melting.	
						No melting point data located in the literature.	Pr <sub>2</sub> O <sub>11</sub>
Not stated		Not stated	Not stated	450		Survey; not original data. The oxide probably dissociates to the metal before melting.	PtO <sub>2</sub>
						No melting point data located in the literature.	Rb <sub>2</sub> O
Not stated		Not stated	Not stated	296		Survey, not original data	Re <sub>2</sub> O <sub>7</sub>
						No melting point data located in the literature. The oxide probably dissociates to the metal before melting.	Rh <sub>2</sub> O <sub>3</sub>

TABLE 11

Oxide	Reference	Purity <sup>2</sup>	Furnace type	Temperature measurement	Method
<b>RuO<sub>4</sub></b>					
<b>Sb<sub>2</sub>O<sub>3</sub> or- thorom- bic type</b>	Hincke [105]	Not stated	Not stated	Not stated	Calculated from vapor pres- sure data.
<b>Sc<sub>2</sub>O<sub>3</sub></b>	Schneider and War- ing [70].	99.9%	Induction; crucible shaped Ir suscepter.	Optical pyrometer sighted on small hole in crucible lid.	Examination of specimen after heating.
<b>SeO<sub>2</sub></b>	Chemistry and Phys- ics Handbook [4].	Not stated	Not stated	Not stated	Not stated
<b>SiO<sub>2</sub>, cristo- balite type</b>	Wietzel [106]	Not stated	Resistance; Ir heating ele- ment.	Optical pyrometer sighted on Pt blacked with Fe <sub>2</sub> O <sub>3</sub> .	Examination of specimen after heating. First glass forma- tion indicated melting.
	Endell and Rieke [107]	99.9%	Resistance; Ir heating ele- ment.	Ir-IrRh thermocouple adja- cent to specimen.	Examination of specimen after heating.
	White, Howat, and Hay [78]	Not stated	Resistance; Mo wire- wound ceramic tube heating element.	Mo-W thermocouple	Observation of specimen during heating.
	Ferguson and Merwin [108]	Not stated	Resistance; quench type, Pt alloy wire-wound ceramic tube heating element.	Pt-90Pt10Rh thermocouple. Temperature measured be- fore and after each determina- tion.	Examination with microscope of quenched specimen after heating. Last trace of crys- tals indicated melting.
	Greig [109]	Very pure	Resistance; quench type, Pt alloy wire-wound ceramic tube heating element.	Pt-90Pt10Rh thermocouple adjacent to the specimen.	Examination with microscope of quenched specimen after heating. Last trace of crys- tals indicated melting.
	Zhirnova [110]	Not stated	Flame; acetylene-O <sub>2</sub>	Optical pyrometer sighted on specimen.	Observation of cone shaped specimen during heating. Slumping of the cone indi- cated melting.
	Ruff and Lauschke [25]	Not stated	Resistance; graphite tube heating element.	Optical pyrometer sighted on specimen.	Observation of specimen during heating.
<b>Sm<sub>2</sub>O<sub>3</sub></b>	Wisnyi and Pijanowski [61]	Not stated	Resistance; V-shaped W strip heating element.	Optical pyrometer sighted on specimen.	Observation of specimen during heating. Apparent disappearance of specimen indicated melting.
	Curtis and Johnson [71]	Y<0.05%; La<0.05%; Pr<0.2%; Nd<0.05%; Ho<0.02%; Er<0.05%; Eu<0.4%; Gd<0.03%; Tb<0.01%; Dy<0.02%; Yb<0.05%.	Not stated	Optical pyrometer sighted on specimen.	Observation of rod shaped specimen during heating. Definite fluidity of speci- men indicated melting.
<b>SnO<sub>2</sub></b>	Ruff [31]	Not stated	Resistance, carbon tube heating element.	Optical pyrometer sighted on specimen.	Observation of specimen dur- ing heating.
	Barczak and Insley [111].	Not stated	Not stated	Not stated	Differential thermal analysis. Also examination of speci- men after heating.

See footnotes at end of table.

—Continued

Calibration Points		Environment	Original Temp. Scale <sup>a</sup>	Melting Points		Comments	Oxide
Materials	Temp.			Original	Int. 1948 <sup>b</sup>		
	°C			°C	°C		
Not stated		8.5 mm Hg	(Int. 1927)	655	655	No melting point data located in the literature.	$\text{RuO}_4$
Au	1063	Air	Int. 1948	<2405	<2405		$\text{Sc}_2\text{O}_3$
Pd	1552						$\text{Sb}_2\text{O}_3$ orthorhombic type
Pt	1769						
Rh	1960						
Not stated		Not stated	Not stated	240-250		Survey, not original data. The oxide is extremely volatile.	$\text{SeO}_3$
Pd	1556	Air	( $C_2=1.437$ cm deg).	1696±10	1691±10		$\text{SiO}_3$ cristobalite type
Au	1063	Nitrogen	(Geophysical)	1685	1692		
Pd	1549						
Pt	1755						
Fe transformation.	1400	Not stated	(Int. 1927)	1705	1701		
Fe	1530						
Pd	1549.5	Air	Geophysical	1710±10	1720±10		
Ice	0	Air	Geophysical	1713±5	1723±5 <sup>c</sup>		
Au	1062.6						
Pd	1549.5						
Pt	1755						
Pt	1755	Air	( $C_2=1.457$ cm deg).	1715±20	1728±20	( <sup>d</sup> )	
$\text{Al}_2\text{O}_3$	2050						
CaO	2570						
Au	1062.4	17.0 mm Hg	( $C_2=1.437$ cm deg.)	1850	1850	<sup>e</sup> ; reaction between specimen and graphite support probable.	
$\text{CaF}_2$	1398						
$\text{Al}_2\text{O}_3$	2040	Either helium, hydrogen, or vacuum	Int. 1948	2300±50	2300±50	<sup>e</sup> ; authors' stated emissivity of approximately 0.9 to 0.95 is probably overestimated. Reflection errors probable.	$\text{Sm}_2\text{O}_3$
Not given		Air	(Int. 1948)	2350±50	2350±50	<sup>e</sup> ; authors' stated an attempt was made to correct for non-blackbody conditions.	
Pt	1755	Not stated	$C_2=1.46$ cm deg	(a) 1385 (b) 1625	1391 1637	<sup>e</sup> reaction between specimen and (a) $\text{MgO}$ or (b) $\text{ZrO}_2$ supports probable.	$\text{SnO}_3$
Not stated		Not stated	(Int. 1948)	1630±5	1630±5	Possible reaction of $\text{SnO}_3$ with Pt container probable.	

TABLE 14

Oxide	Reference	Purity <sup>2</sup>	Furnace type	Temperature measurement	Method
<b>SrO</b>	Schumacher [43].....	C. P. grade.....	Resistance; boat-shaped W strip heating element.	Optical pyrometer sighted on W strip adjacent to specimen; spectral emissivity equals 0.431.	Observation of specimen during heating.
<b>Ta<sub>2</sub>O<sub>5</sub> alpha type</b>	Reisman, Holtzberg, Berkenblit, and Berry [112].	TiO <sub>2</sub> <0.001%.....	Resistance, U-shaped Rh or Pt alloy strip heating element.	Optical pyrometer sighted on specimen.	Observation of specimen during heating.
	Ruff [31].....	Nb <sub>2</sub> O <sub>5</sub> —7.1%.....	Resistance; carbon tube heating element.	Optical pyrometer sighted on specimen.	Observation of specimen during heating.
<b>Tb<sub>4</sub>O<sub>7</sub></b>	.....	.....	.....	.....	.....
<b>TeO<sub>3</sub></b>	Kracek [3].....	Not stated.....	Not stated.....	Not stated.....	Not stated.....
<b>ThO<sub>2</sub></b>	Tiede and Birnbrauer [23].	Commercially pure.....	Arc.....	Optical pyrometer sighted on specimen.	Observation of specimen during heating.
	Ruff [31].....	Pure.....	Resistance; carbon tube heating element.	Optical pyrometer sighted on specimen.	Observation of specimen during heating.
	Trombe [2].....	Not stated.....	Not stated.....	Not stated.....	Not stated.....
	Ruff, Ebert, and Woitinek [113].	"Pure".....	Flame; oxyacetylene.....	Optical pyrometer.....	Not stated.....
<b>TiO<sub>2</sub></b>	Lambertson and Gunzel [30].	99.7%.....	Resistance; W heating element.	Optical pyrometer sighted on specimen. Specimen not visible.	Examination of specimen after heating.
	Statton [114].....	Reagent grade.....	Resistance; Mo strip heating element.	Optical pyrometer sighted on Mo strip adjacent to specimen. Corrected for spectral emissivity of Mo.	Observation of specimen during heating.
	v. Wartenberg and Prophet [53].	Not stated.....	Flame; oxyhydrogen or gas-air.	Optical pyrometer.....	Not stated.....
	Sigurdson and Cole [115].	Not stated.....	Resistance; Pt strip heating element.	Optical pyrometer sighted on Pt strip adjacent to specimen. Corrected for spectral emissivity of Pt.	Observation of specimen during heating.
	Rase and Roy [116]....	Not stated.....	Resistance; U-shaped Pt alloy heating element.	Optical pyrometer sighted on Pt alloy strip.	Observation of specimen during heating.
	Coughanour and De Prossé [117].	99.9%.....	Resistance; ThO <sub>2</sub> heating element.	Optical pyrometer sighted on specimen.	Observation of pyramid-shaped specimen during heating.
	St. Pierre [118].....	Si—0.04%; Mg—0.02%; Ca—0.01%.	Induction; carbon tube susceptor.	Optical pyrometer sighted on specimen.	Observation of specimen during heating.
Diamond and Schneider [28].	About 99.9%.....	Solar.....	.....	Optical pyrometer sighted on specimen.	Observation of bar-shaped specimen during cooling.

See footnotes at end of table.

Calibration Points		Environment	Original Temp. Scale <sup>3</sup>	Melting Points		Comments	Oxide
Materials	Temp.			Original	Int. 1948 <sup>4</sup>		
Not stated	°C	Hydrogen at 0.2 atm	$C_2=1.433$ cm deg.	2430	2420	<sup>5</sup> ; temperature uniformity between specimen and target area of pyrometer not verified; reflection errors probable.	SrO
Diopside 15% CaO: 85% SiO <sub>2</sub>	1392 1712	(Air)	Int. 1948	1872±10	1872±10	( <sup>6</sup> )	
Pt	1755	Nitrogen at reduced pressure.	$C_2=1.46$ cm deg	1875	1895	<sup>5</sup> ; reaction between specimen and ZrO <sub>2</sub> support probable.	Tb <sub>2</sub> O <sub>7</sub>
						No melting point data located in the literature.	
Not stated	Not stated	Not stated	732.6			Survey; not original data. This oxide is very volatile and has a tendency to sublime rather than to melt.	TeO <sub>2</sub>
Not stated	Vacuum	Not stated	2000			( <sup>6</sup> )	ThO <sub>3</sub>
Pt	1755	Nitrogen at reduced pressure.	$C_2=1.46$ cm deg	2425 2440 2470	2468 2483 2515	<sup>5</sup> ; reaction between specimen and ceramic supports probable.	
Not stated	Not stated	Not stated	3000			Survey, not original data.	
Not stated	Not stated	(Int. 1927)	3050±50	3030±50		Melting point extrapolated from liquidus curves of the ZrO <sub>2</sub> -ThO <sub>3</sub> system.	
Pt	(1769)	Helium	(Int. 1948)	3220±50	3220±50*		
Pd	1555	Vacuum	(Int. 1927)	1720	1716	<sup>5</sup> ; temperature uniformity between specimen and target area of pyrometer not verified.	TiO <sub>2</sub>
Pt	1774					Reduction of TiO <sub>2</sub> very probable.	
Not stated	Air	(Int. 1927)	1825	1820		<sup>5</sup> ; partial reduction of TiO <sub>2</sub> probably occurred.	
Au	1063	Oxidizing	(Int. 1948)	1825	1825	( <sup>6</sup> )	
Na <sub>2</sub> Ti <sub>3</sub> O <sub>7</sub>	1128						
BaF <sub>2</sub>	1280						
CaMgSi <sub>2</sub> O <sub>6</sub>	1391						
MgTi <sub>2</sub> O <sub>5</sub>	1660						
Diopside Pseudo- wollas- tonite. 15% CaO: 85% SiO <sub>2</sub> .	1391.5 1544 1698	(Air)	(Geophysical)	1830		( <sup>6</sup> )	
Pt	1769	Air	Int. 1948	1839±10	1839±10	( <sup>6</sup> )	
Not stated	Air	(Int. 1948)	1840±10	1840±10		( <sup>6</sup> )	
None	Air	Int. 1948	1840	1840		<sup>5</sup> ; emissivity stated by authors to be about unity.	

TABLE 11

Oxide	Reference	Purity <sup>2</sup>	Furnace type	Temperature measurement	Method
$\text{TiO}_2$ Con.	v. Wartenberg and Gurr [52].	Not stated.....	Not stated.....	Optical pyrometer.....	Not stated.....
	Lang, Fillmore, and Maxwell [119].	99.9%.....	Resistance; $\text{ThO}_2$ heating elements.	Optical pyrometer sighted on specimen.	Observation of pyramid-shaped specimen during heating.
	v. Wartenberg and Eckhardt [57].	Pure; $\text{SiO}_2$ free.....	Flame; acetylene- $\text{O}_2$ .....	Optical pyrometer sighted on specimen.	Observation of specimen suspended in furnace during heating.
	Brauer and Littke [120].	>99.8%.....	Solar.....	Radiation pyrometer sighted on specimen.	Observation of specimen during heating.
$\text{Ti}_2\text{O}_5$	Duncan [121].....	Not stated.....	Not stated.....	Pt-90Pt10Rh thermocouple immersed in specimen.	Thermal analysis, cooling curves.
$\text{Tm}_2\text{O}_5$	.....	.....	.....	.....	.....
$\text{UO}_3$	Ruff and Goecke [34]..	Not stated.....	Resistance; graphite tube heating element.	Optical pyrometer sighted on specimen.	Observation of cone-shaped specimen during heating.
	Wisnyl and Pijanowski [61].	Not stated.....	Resistance; V-shaped W strip heating element.	Optical pyrometer sighted on specimen.	Observation of specimen during heating. Apparent disappearance of specimen indicated melting.
	Ehert and Margrave [122].	Not stated.....	Induction; graphite susceptor.	Optical pyrometer sighted on specimen. Correction made for spectral emissivity.	Observation of specimen during heating.
	Lambertson and Gunzel [30].	Very pure.....	Resistance; W heating element.	Optical pyrometer sighted on specimen. Specimen not visible.	Examination of specimen during heating.
$\text{V}_2\text{O}_5$	Kracek [3].....	Not stated.....	Not stated.....	Not stated.....	Not stated.....
	Carnelley [81].....	Not stated.....	Flame; "bunsen lamp".....	Mercury thermometer.....	Not stated.....
	Cook [123].....	Purified.....	Not stated.....	Thermocouple.....	Heat content curves.....
	Illarionov, Ozeron, and Kil'disheva [124].	Not stated.....	Not stated.....	Pt-PtRh thermocouple.....	Differential thermal analysis.....
	Holtzberg, Reisman, Berry, and Berkenblit [125].	99.94%.....	Not stated.....	Pt-90Pt10Rh thermocouple immersed in specimen.	Thermal analysis.....
	McDaniel [126].....	99.5%.....	Resistance; Pt alloy wire-wound ceramic tube heating element.	Pt-90Pt10Rh thermocouple immersed in specimen.	Thermal analysis; electrical conductance versus temperature curves.
	Burdese [127].....	Not stated.....	Not stated.....	Not stated.....	Examination with microscope of specimen after heating.
	Chemistry and Physics Handbook [4].	Not stated.....	Not stated.....	Not stated.....	Not stated.....
$\text{WO}_3$	v. Jaeger and Germs [83].	Fe-0.05%.....	Resistance; nichrome wire heating element.	Thermocouple immersed in specimen.	Thermal analysis, heating curves.

See footnotes at end of table.

Calibration Points		Environment	Original Temp. Scale <sup>3</sup>	Melting Points		Comments	Oxide
Materials	Temp.			Original	Int. 1948 <sup>4</sup>		
Not stated	°C	Air	(Int. 1927)	1850	1845		TiO <sub>2</sub> — Con.
Not stated		Air	Int. 1948	1845	1845	(s)	
Not stated		Air	(Int. 1927)	1855	1850	(s)	
CaF <sub>2</sub>	Not stated	(a) Oxygen at 300 torr. Argon at 460 torr.	(Int. 1948)	(a) 1840±15	1840±15	(s)	
NiO	Not stated	(b) Oxygen at 500 torr. Argon at 260 torr.		(b) 1860±15	1860±15		
Cr <sub>2</sub> O <sub>3</sub>	Not stated	(c) Oxygen at 600 torr. Argon at 160 torr.		(c) 1870±15	1870±15		
Al <sub>2</sub> O <sub>3</sub>	Not stated	(d) Oxygen at 760 torr. (e) Oxygen at 1140 torr.		(d) 1870±15	1870±15		
Not stated		Oxygen at 1 atm	(Int. 1927)	717±5	717±5		Tl <sub>2</sub> O <sub>3</sub>
						No melting point data located in the literature.	Tm <sub>2</sub> O <sub>3</sub>
Au	1071	Nitrogen	(C <sub>2</sub> =1.48 cm deg.)	2176	2208	<sup>4</sup> ; reaction between specimen and uranium carbide support probable.	UO <sub>3</sub>
Pt	1757						
Al <sub>2</sub> O <sub>3</sub>	2040	Either helium, hydrogen, or vacuum.	Int. 1948	2760±30	2760±30	<sup>4</sup> ; author's stated emissivity of 0.9 to 0.95 probably overestimated. Reflection errors probable.	
Not stated		Vacuum	(Int. 1948)	2860±45	2860±45		
Pt	(1769)	Helium	(Int. 1948)	2878±22	2878±22*		
Not stated		Not stated	Not stated	656		Survey; not original data	V <sub>2</sub> O <sub>5</sub>
Not stated		Not stated	Not stated	658			
Au	(1063)	Not stated	(Int. 1927)	670	670		
Pd	(1555)						
Not stated		Not stated	(Int. 1948)	672	672		
NaCl	800.4	Air + Oxygen	Int. 1948	674±5	674±5*		
K <sub>2</sub> SO <sub>4</sub>	1069						
NaCl	800	Air	Int. 1948	675±3	675±3*		
Not stated		Carbon dioxide	(Int. 1948)	685±5	685±5		
Not stated		Not stated	Not stated	690		Survey, not original data	
Not stated		Oxidizing	Not stated	1473±1			WO <sub>3</sub>

TABLE 1

Oxide	Reference	Purity <sup>1</sup>	Furnace type	Temperature measurement	Method
WO <sub>3</sub> —Con.	Hoerman [84]	Not stated	Resistance; Pt wire heating element.	Pt-PtRh thermocouple immersed in specimen.	Observation of specimen during heating.
Y <sub>2</sub> O <sub>3</sub>	Ruff and Lausche [25]	Commercially pure	Resistance; graphite tube heating element.	Optical pyrometer sighted on specimen.	Observation of specimen during heating.
	Ruff [31]	Traces of impurities	Resistance; carbon tube heating element.	Optical pyrometer sighted on specimen.	Observation of specimen during heating.
Yb <sub>2</sub> O <sub>3</sub>					
ZnO	Bunting [128]	>99.9%	Induction, button-shaped Ir-Pt susceptor.	Optical pyrometer sighted on small cavity adjacent to specimen cavity.	Examination of specimen after heating.
ZrO <sub>2</sub>	Tiede and Birnbrauer [23]	Very pure	Arc	Optical pyrometer sighted on specimen.	Observation of specimen during heating.
	Ruff and Lauschke [25]	SiO <sub>2</sub> —0.95% Fe <sub>2</sub> O <sub>3</sub> —0.27%	Resistance; graphite tube heating element.	Optical pyrometer sighted on specimen.	Observation of specimen during heating.
	Ruff [31]	93%	Resistance; graphite tube heating element.	Optical pyrometer sighted on specimen.	Observation of specimen during heating.
	Clausing [74]	HfO <sub>2</sub> <1%	Resistance; bar-shaped W strip heating element.	Optical pyrometer sighted on small cavity adjacent to the specimen cavity.	Observation of specimen during heating. Flowing specimen indicated melting.
	Podszus [129]	Pure	Arc, carbon electrodes	Optical pyrometer sighted on specimen.	Observation of specimen during heating.
	Henning [75]	Pure	Resistance; W tube heating element.	Optical pyrometer sighted on specimen.	Observation of specimen during heating.
	Mark [27]	Not stated	Resistance; graphite or metal heating element.	Optical pyrometer sighted on specimen.	Observation of specimen during heating.
	Trombe [2]	Not stated	Not stated	Not stated	Not stated
	Lambertson and Gunzel [30]	HfO <sub>2</sub> —2.03% Others—0.03%	Resistance, W heating element.	Optical pyrometer sighted on specimen. Specimen not visible.	Examination of specimen after heating.
	Zhirnova [110]	Not stated	Flame; acetylene-O <sub>2</sub>	Optical pyrometer sighted on specimen.	Observation of cone shaped specimen during heating. Slumping of the cone indicated melting.
Curtis, Doney, and Johnson [76]	Hf—80 ppm Ti—60 ppm Al—150 ppm Fe—600 ppm Si—200 ppm	Flame; oxyacetylene		Optical pyrometer	Not stated

<sup>1</sup> All phrases, numbers, words, etc., enclosed in parentheses indicate that these items are not directly expressed in the quoted reference. However, based on information contained in the published report, the enclosed items can be reasonably assumed to be valid.

<sup>2</sup> Purities are listed as quoted in the reference. No attempt has been made to convert them to a common basis.

<sup>3</sup> Int. 1927—The International Temperature Scale (of 1927) [19].

Int. 1948—The International Practical Temperature Scale of 1948 [10].

Geophysical—Geophysical Temperature Scale [18].

$C_2$ —second radiation constant. This constant is required in the definition of the temperature scale. See text for more complete description of appropriate equations.

—Continued

Calibration Points		Environment	Original Temp. Scale <sup>3</sup>	Melting Points		Comments	Oxide
Materials	Temp.			Original	Int. 1948 <sup>4</sup>		
Not stated	°C	Not stated	(Int. 1927)	1473	°C	1471	
Au	1062.4	Air at 21.5 mm Hg	$C_2=1.437$ cm deg.	2410	2410		WO <sub>3</sub> —Con.
CaF <sub>2</sub>	1398						Y <sub>2</sub> O <sub>3</sub>
Pt	1755	Nitrogen at 15 mm Hg	$C_2=1.46$ cm deg	2415	2458	<sup>5</sup> ; reaction between specimen and graphite support probable.	
						No melting point data located in literature.	Yb <sub>2</sub> O <sub>3</sub>
Not stated		Air	(Int. 1927)	1975±25	1969±25	Temperature uniformity between specimen cavity and target area of pyrometer not verified.	ZnO
Not stated		Vacuum	Not stated	2430		( <sup>6</sup> )	ZrO <sub>2</sub>
Au	1062.4	(a) Hydrogen at 760 mm Hg.	$C_2=1.437$ cm deg.	(a) 2519	2519		
CaF <sub>2</sub>	1398	(b) Air at 8.22 mm Hg		(b) 2563±10	2563±10	<sup>5</sup> ; reaction between specimen and graphite support probable.	
Pt	1755	Nitrogen	$C_2=1.46$ cm deg.	2585	2636	( <sup>6</sup> )	
Not stated		Hydrogen	(Int. 1927)	2677±25	2663±5	( <sup>6</sup> )	
Pt	Not stated	Air	Not stated	(a) 2677 (b) 2727		( <sup>6</sup> )	
Not stated		Hydrogen and nitrogen	Not stated	2687		( <sup>6</sup> )	
Not stated		Neutral	(Int. 1948)	2690	2690	( <sup>6</sup> )	
Not stated		Not stated	Not stated	2700		Survey; not original data	
Pt	(1769)	Helium	(Int. 1948)	2710±15	2710±15		
Pt	1755	Air	( $C_2=1.457$ cm deg.)	2715	2765	( <sup>6</sup> )	
Al <sub>2</sub> O <sub>3</sub>	2050						
CaO	2570						
Not stated		Not stated	(Int. 1948)	2850±25	2850±25	( <sup>6</sup> )	

<sup>4</sup> (A) The melting points marked with an asterisk are considered to be the better values of those listed. The values so designated do not necessarily represent the true melting points, but merely highlight those values believed to be more reliable.

(B) See text for method of conversion to values based on International Practical Temperature Scale of 1948. No entry in this column indicates insufficient data in original paper to permit conversion of melting point to 1948 basis.

<sup>5</sup> A close enough approach to blackbody conditions to facilitate accurate temperature measurement probably has not been obtained. As a general guideline, it should be noted that if an object in an enclosure is distinguishable from its surroundings, blackbody conditions probably have not been realized.

<sup>6</sup> Kanolt did not specifically employ Wien's equation in the measurement of high temperatures. A complete discussion of the method of conversion to temperatures on the 1948 scale is given elsewhere [130].