

*Ion Exchange and Dehydration Experimental
Studies of Clinoptilolite: Implications to
Zeolite Dating*

Giday WoldeGabriel

MASTER

DISTRIBUTION OF THIS DOCUMENT IS UNLIMITED

HH

Los Alamos
NATIONAL LABORATORY

Los Alamos, New Mexico 87545

DISCLAIMER

This report was prepared as an account of work sponsored by an agency of the United States Government. Neither the United States Government nor any agency thereof, nor any of their employees, make any warranty, express or implied, or assumes any legal liability or responsibility for the accuracy, completeness, or usefulness of any information, apparatus, product, or process disclosed, or represents that its use would not infringe privately owned rights. Reference herein to any specific commercial product, process, or service by trade name, trademark, manufacturer, or otherwise does not necessarily constitute or imply its endorsement, recommendation, or favoring by the United States Government or any agency thereof. The views and opinions of authors expressed herein do not necessarily state or reflect those of the United States Government or any agency thereof.

DISCLAIMER

Portions of this document may be illegible in electronic image products. Images are produced from the best available original document.

ION EXCHANGE AND DEHYDRATION EXPERIMENTAL STUDIES OF CLINOPTILOLITE: IMPLICATIONS TO ZEOLITE DATING

by

Giday WoldeGabriel

ABSTRACT

Variable effects were noted on the argon (Ar) and potassium (K) contents of clinoptilolite fractions used in ion-exchange and dehydration experiments. The K contents of clinoptilolite fractions were differently affected during cation exchange with Ca-, Cs-, K-, and Na-chloride solutions. Ar was generally less affected during these experiments, except for a Na-clinoptilolite fraction exchanged for five days. Loss of Ar during organic heavy-liquid treatment and cleaning using acetone and deionized water does occur, as indicated by comparing the amounts of radiogenic Ar of treated and untreated fractions. Moreover, a regular decrease in radiogenic Ar contents was noted in clinoptilolite fractions during dehydration experiments at different temperatures for 16 hours. Comparable losses do not occur from saturated samples that were heated in 100 °C for more than five months.

Water appears to play a vital role in stabilizing the clinoptilolite framework structure and in the retention of Ar. The radiogenic Ar depletion pattern noted in clinoptilolite fractions dehydrated in unsaturated environment at different temperatures is similar to variations in the amount of radiogenic Ar observed in clinoptilolite samples from the unsaturated zone of an altered tuff. These results can be used to evaluate the extent of zeolitic water (and hence Ar) retention in unsaturated geologic settings.

I. INTRODUCTION

Zeolites are open-framework aluminosilicate minerals, generally formed at low temperatures, that are utilized widely in various environmental research applications for their adsorptive, molecular sieve, ion-exchange, and catalytic properties (Mumpton, 1978). The origin

and distribution of zeolites are generally well established; however, the timing of diagenesis is often poorly constrained.

Published information on zeolite dating is very sparse considering the widespread occurrences and importance of these industrial minerals. The presence of large cavities and channels in zeolites that allow cations, water molecules, and certain gases to move readily in and out of the crystal structure (e.g., Vaughan, 1978; Feng and Savin, 1993) implies the existence of poor Ar and K retention and this may explain the limited research in zeolite dating by the K/Ar and similar methods. However, zeolite diagenesis in the early Miocene John Day Formation of Oregon was indirectly determined by dating the coexisting K-bearing silicate minerals of authigenic feldspar and celadonite (Hay, 1963). Limited studies on K/Ar, Rb/Sr, and fission track dating of phillipsite, clinoptilolite, heulandite, stilbite, and chabazite separated from deep-sea and continental sediments, and from vesicles in basaltic lavas have also been described (Bernat et al., 1970; Koul et al., 1981; Clauer, 1982; Gundogdu et al., 1989).

The utility of alkali zeolites (e.g., phillipsite, clinoptilolite, and mordenite) from low-temperature, open-hydrologic alteration as potential dateable minerals was evaluated using the K/Ar method as part of the Yucca Mountain Site Characterization Project, which is evaluating Yucca Mountain, Nevada, as a potential high-level radioactive waste repository site (WoldeGabriel et al., 1993). The presence of zeolites was a major reason for choosing the site because these minerals provide favorable physical and chemical properties as important barriers to radionuclide migration. Thus, it is judicious to determine the time of diagenesis of zeolite and other coexisting minerals to understand the geologic history of the site. Authigenic minerals such as illite, illite/smectite, K-feldspar, etc., are generally dated to constrain burial and thermal alteration histories of rocks in various geological settings (e.g., Hoffman et al., 1976; WoldeGabriel and Goff, 1989; Elliot et al., 1991; Bish and Aronson, 1993). Thus, dating of K-bearing zeolites and other coexisting dateable minerals from Yucca Mountain could provide the diagenetic ages of these minerals to understand the geologic history of the site.

Information on temperature and pressure conditions during zeolite diagenesis can often be inferred from authigenic silicate mineral assemblages, burial depth, and vitrinite reflectance data (Iijima, 1978; Bell, 1986; Turner and Fishman, 1991). Temperature is important in controlling

zeolite stability, whereas pressure appears to have a minimal effect as indicated by the occurrences of particular zeolites throughout a broad range of burial depths (Boles, 1977). Because the dominant authigenic minerals (smectite, clinoptilolite, opal-CT, etc.) at Yucca Mountain are characteristic of low-temperature diagenetic environments with upper stability limits of about 100 °C (Ijima, 1978; Bish, 1989), accurate ages on the K-bearing zeolites and other coexisting dateable minerals could also provide information on post-diagenetic processes and the past and future thermal history of the site.

Clinoptilolite fractions separated from outcrop and drill core samples of altered tuffs from Yucca Mountain and dated by the K/Ar method yielded variable isotopic ages (WoldeGabriel et al., 1993). Some of the clinoptilolite isotopic ages (9-12 Ma) from the saturated zone are generally similar to the eruption ages of the tuffs; however, those from the unsaturated zone have apparent ages that are much younger (2-5 Ma). Although the dated clinoptilolite fractions have the same mineralogy in both the saturated and the unsaturated zones, the factors responsible for the isotopic age variations are not fully understood. However, water appears to play a significant role in zeolite stability (Boles, 1972; Bish, 1988), and the difference in isotopic ages may be related to loss of water and then argon from the structural cavities of clinoptilolite in the unsaturated environment. Water molecules in the extraframework sites in clinoptilolite may in fact inhibit or slow the migration of extraframework Ar.

Because the K-bearing clinoptilolite fractions from outcrop and the unsaturated zone yielded younger isotopic ages, it is essential to evaluate the effects of cation exchange and variable saturation on the K⁺ and Ar contents of these minerals. Thus, the objective of this study was to evaluate experimentally the effects of post-crystallization ion-exchange and dehydration processes on K⁺ and Ar contents of clinoptilolite. Significant effects might be anticipated because zeolites are characterized by an open-framework structure with a large ion-exchange capacity.

II. EXPERIMENTAL METHODS

The clinoptilolite-rich bulk sample used in the ion-exchange and dehydration experiments for the present study is from the late Miocene Sucker Creek Formation of Sheaville, eastern Oregon, that formed in an open hydrologic system from percolating meteoric fluids (Sheppard et

al., 1983). The sample was chosen for its purity and is unlike the clinoptilolite-rich samples from Yucca Mountain that coexist with other diagenetic minerals. The time of zeolite formation in the Sucker Creek Formation is unknown; however, it is suggested that the alteration of silicic tuffs to zeolites in an open system occurs during early diagenesis (Hay, 1978). Both K- and Ca-clinoptilolites are present in the altered tuff units and it is reported that these zeolites were affected by post-diagenetic low-temperature (75-100 °C) hydrothermal alteration that resulted in cation exchange in the zeolites and the formation of quartz from opal-CT (Altaner and Grim, 1990). The clinoptilolite sample (27073) was obtained from Minerals Research, P. O. Box 591, Clarkson, New York, 14430 USA.

In most cases, zeolites were separated from coexisting authigenic silicate minerals by sedimentation in deionized water. Additional purification was carried out using organic heavy liquids. Clinoptilolite was thoroughly mixed with S-tetrabromoethane to evaluate the effect of the heavy liquid on the framework structure of the mineral during clinoptilolite purification. The heavy-liquid treated aliquot was thoroughly cleaned using acetone in an ultrasonic bath followed by centrifugation before it was cleaned with deionized water.

For the ion-exchange experiments, 1 M chloride solutions of Ca^{2+} , K^+ , Na^+ , and Cs^+ were prepared and added to several splits of bulk (i.e., not enriched by separation) clinoptilolite aliquots (5-50 m). The mixture was placed in an oven at 50 °C and periodically shaken during the 24-hr ion-exchange experiment. New batches of chloride solutions were added every 24 hours after the exchanged solution was removed by centrifugation. The experiments were conducted in 3 and 5 steps for total times of 72 and 120 hours, respectively. A clinoptilolite bulk powder and the solutions obtained by centrifugation after each of the ion-exchange experiments were analyzed for cations (e.g., Ca^{2+} , Cs^+ , K^+ , and Na^+) and major-element concentrations. Samples from the ion-exchange treatment were thoroughly cleaned using deionized water in an ultrasonic bath followed by centrifugation. The bulk powder was analyzed by X-ray fluorescence spectrometry and the exchange solutions were measured by atomic absorption spectrophotometry, combined with gravimetric determinations, along with bulk rock and solution standards.

Aliquots of about 2 g of clinoptilolite from the same bulk sample used in the ion-exchange experiment were heated to 50, 100, 150, and 200 °C in air for 16 hours. The samples were kept in

a desiccator until they were analyzed for K and Ar. Additional heating experiments were conducted on other aliquots from the same sample held in deionized water (2 to 1 water/sample ratio) for a week, a month, three months, and about six months at 100 °C. All samples from the experiments were X-rayed using an automated Siemens D-500 diffractometer with Cu K radiation, 0.02° 2 θ steps, and a counting time of 1 s per step for all mounts (2-36° 2 θ).

The K₂O contents were determined using an atomic absorption spectrophotometer on 70 to 100 mg of sample, and aliquots of about 120 to 250 mg of the same sample were used for the radiogenic Ar measurements using an MS10 mass spectrometer on line with bulb-pipetted ³⁸Ar. The K/Ar system was prebaked while samples were cooled by running water. Standard samples (e.g., BCR-1, G-2, and LP-6) analyzed with the unknown samples for K₂O and Ar varied by less than 2.5 from the published values (Odin et al., 1982; Govindaraju, 1989).

III. RESULTS

Heavy-liquid treated and ion-exchanged clinoptilolite fractions showed no apparent changes in their XRD patterns compared with the untreated sample. This is also true for samples heated under unsaturated and saturated conditions at variable temperatures for 16 hours and more than five months, respectively (Fig. 1).

Results of chemical analysis of the untreated clinoptilolite fraction and a standard sample and published results of the standard (NIM-G) are given in Table 1, and the 5.34 wt % K₂O of the clinoptilolite is used as a reference to evaluate the extent of potassium enrichment and depletion during the ion-exchange experiments. The chemical compositions of the solutions obtained by centrifugations after each of the 24-hr Ca-, Cs- K-, and Na-chloride solution exchanges are given in Table 2. The results indicate that the alkali and alkaline earth cations were replaced stepwise, with most replacement occurring during the first phase of the ion-exchange experiment and exchange decreasing substantially during subsequent steps (Fig. 2).

Results of Ca²⁺ analysis from the Cs-, K-, and Na-exchanged solutions indicate that similar amounts of Ca²⁺ were replaced by these cations (Table 2 and Fig. 2a). Clinoptilolite shows a greater affinity for Cs⁺, with Na⁺ replacing slightly more K⁺ than Ca²⁺ (Fig. 2b). More Na⁺ was replaced by Cs⁺ and K⁺ during the first step of the exchange experiment compared with Ca²⁺

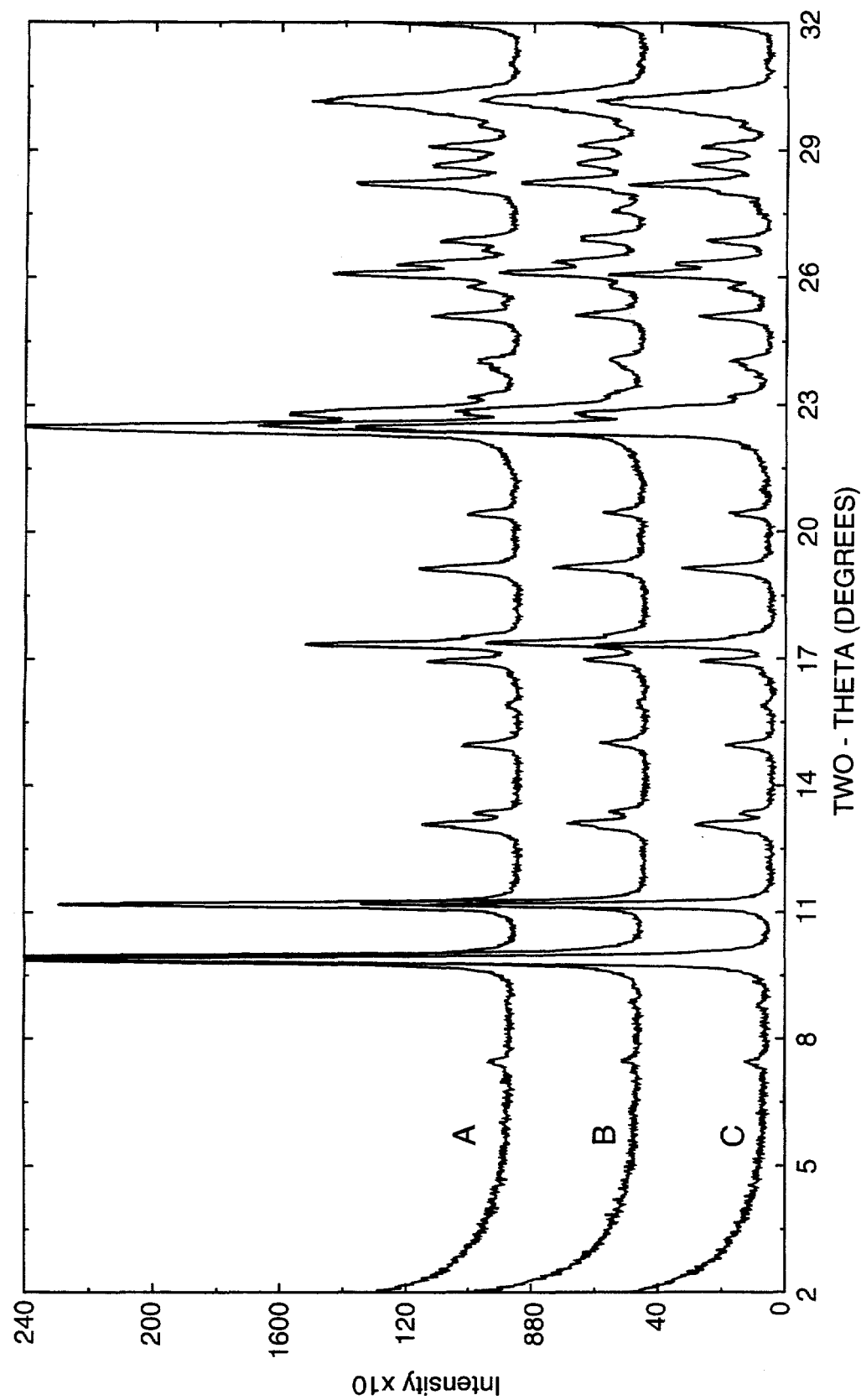


Figure 1. X-ray diffraction patterns of clinoptilolite fractions used in the dehydration and ion-exchange experiments. The patterns are for unsaturated heating at 200 °C (A), saturated heating at 100 °C for more than five months (B), and an untreated fraction (C).

Table 1. Major-element composition of clinoptilolite-rich bulk sample in weight percent used for ion-exchange and dehydration experiments. Clinoptilolite and standard sample NIM-G analyzed using X-ray fluorescence to check the analytical data against published information.

	27073/CPT#1	NIM-G	NIM-G†
SiO ₂	64.62	75.70	75.70
TiO ₂	0.07	0.059	0.09
Al ₂ O ₃	12.22	12.20	12.08
Fe ₂ O ₃	0.42	0.49	0.58
FeO	0.03	1.27	1.30
MnO	0.03	0.01	0.02
MgO	0.28	0.03	0.06
CaO	1.43	0.78	0.78
Na ₂ O	1.77	3.39	3.36
K ₂ O	5.34	5.02	4.99
H ₂ O ⁻	4.19	0.52	0.49
H ₂ O ⁺	8.40	0.0	0.01
P ₂ O ₅	0.01	0.01	0.0
BaO	0.23	0.01	
SrO	0.03	<0.004	
Total	99.03	99.41	99.46

†Published major element composition of standard NIM-G included for comparison (Govindaraju, 1989).

(Fig. 2c). These results are generally consistent with the cation affinity sequence for clinoptilolite reported by Ames (1961).

Radiogenic Ar contents of clinoptilolite fractions from the various experiments are given in Table 3. These values range from 1.1109×10^{-12} to 3.7402×10^{-12} moles and probably do not represent the amount of radiogenic Ar accumulated since the time of alteration from silicic glass to clinoptilolite because the clinoptilolite fractions were subjected to ion-exchange and heating experiments. The K₂O contents of the clinoptilolite fractions (1 to 9 wt %) were significantly modified during the ion-exchange experiments, whereas the radioagenic Ar contents show less drastic variations except for the Cs-clinoptilolite exchanged for five days (Fig. 3a). The Cs- and Na-clinoptilolite fractions exchanged for five days contain larger and smaller amounts of radiogenic Ar, respectively, compared with the other fractions. Moreover, the amount of radiogenic Ar in the heavy-liquid treated fraction is depleted. The rest of the clinoptilolite fractions exchanged for three and five days show similar amounts of radiogenic Ar (Fig. 3a).

Table 2. Results of analysis (ppm) of solutions from 24-hr ion-exchange experiments using atomic absorption spectrometry. The K-rich clinoptilolite was exchanged for 3 and 5 days with each of the cations listed. Standard WP591#1 was analyzed for comparison with accepted data. Solutions were changed every 24 hours.

Chloride Solutions	Sample ID (27073/CPT#1)	Ca	Cs	K	Na
CaCl ₂ -3 days	3-1	>1500	<2	505	690
	3-2	>1500	<2	330	204
	3-3	>1500	<2	240	78
	3-3 (duplicate)	>1500	<2	245	80
CaCl ₂ -5 days	5-1	>1500	<2	512	672
	5-2	>1500	<2	354	232
	5-3	>1500	<2	270	90
	5-4	>1500	<2	222	42
	5-5	>1500	<2	181	23
CsCl-3 days	3-1	680	>>1000	1750	910
	3-2	140	>>1000	700	112
	3-3	45	>>1000	313	24
CsCl-5 days	5-1	680	>>1000	1740	910
	5-2	135	>>1000	700	111
	5-3	45	>>1000	313	24
	5-4	27	>>1000	194	10
	5-5	15	>>1000	126	8
	5-5 (duplicate)	15	>>1000	130	8.5
KCl-3 days	3-1	710	<1	>3500	915
	3-1	710	<1	>3500	940
	3-2	150	<1	>3500	130
	3-3	40	<1	>3500	20
KCl-5 days	5-1	677	<1	>3500	915
	5-2	160	<1	>3500	149
	5-3	37	<1	>3500	26
	5-4	17	<1	>3500	3
	5-5	8	<1	>3500	<1
NaCl-3 days	3-1	670	<1	700	>>1000
	3-2	140	<1	433	>>1000
	3-3	57	<1	295	>>1000
NaCl-5 days	5-1	652	<1	707	>>1000
	5-2	170	<1	477	>>1000
	5-3	57	<1	307	>>1000
	5-4	30	<1	235	>>1000
	5-4 (duplicate)	30	<1	235	>>1000
	5-5	14	<1	192	>>1000
WP591#1*	10-1	28		16	85
WP591#1†		27		16	88.4

* EPA reference standard solution; † Published chemical data for the reference solution

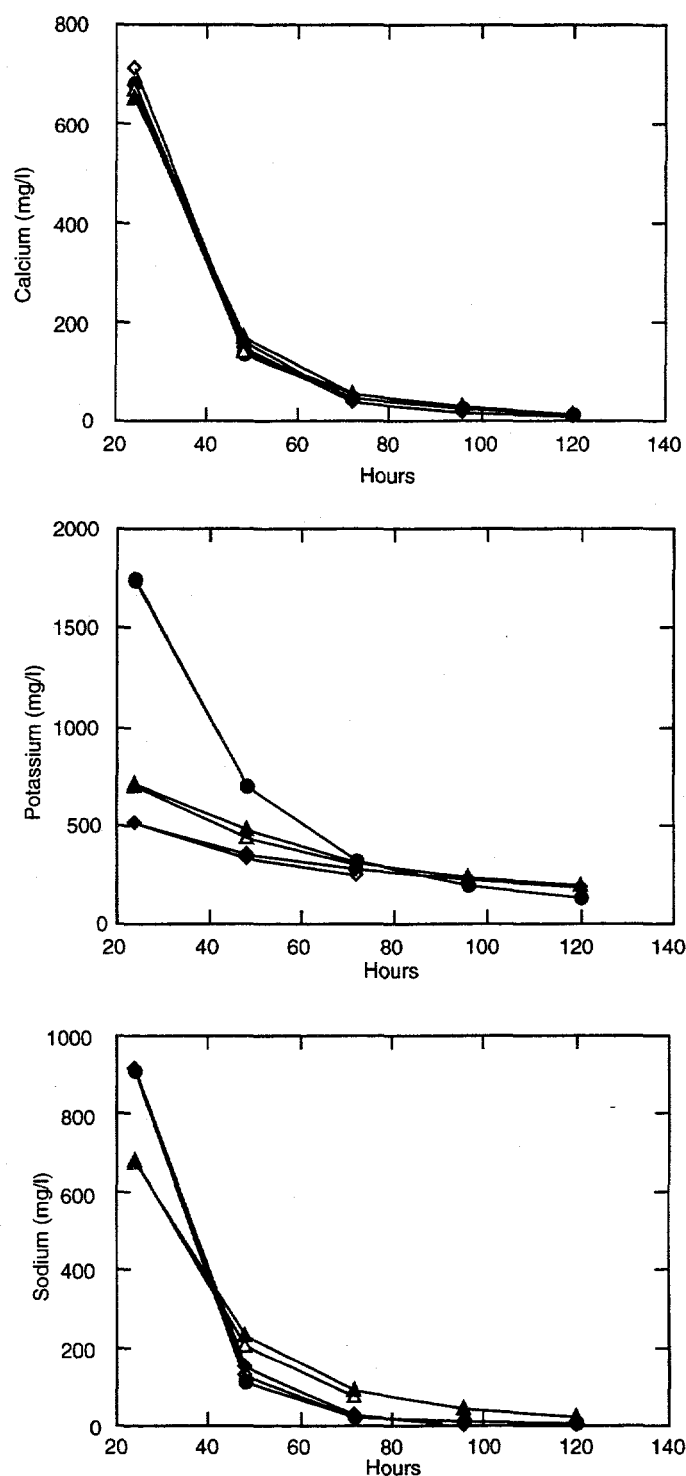


Figure 2. Cation concentrations of Ca^{2+} , K^+ , and Na^+ in solutions from ion-exchange experiments. Ca -, Cs -, K -, Na -chloride solutions were used for three and five days indicated by open and solid circles, respectively. Most of the cations were replaced within 70 hours. Similar amounts of Ca^{2+} ions were replaced by Cs^+ , K^+ , and Na^+ in A, whereas Cs^+ replaced more K^+ compared with Ca^{2+} and Na^+ in B. More Na^+ ions were replaced by Cs^+ and K^+ as shown in C. Circles, triangles, and diamonds represent cations of Cs , K , and Na in A; Cs , Na , and Ca in B; and Cs , K , and Ca in C.

Table 3. Radiogenic argon and potassium contents of clinoptilolite fractions from dehydration and cation-exchange experiments.

Sample ID	Time	Temperature (°C)	Solution	K ₂ O (wt %)	⁴⁰ Ar * † (10 ⁻¹¹ moles/g)
27073/CPT#1‡				5.28	0.8127 ± 23.78
27073/CPT#1§				5.34	0.6248 ± 22.83
27073/CPT#1	16 hours	50	unsaturated	5.23	1.0777 ± 23.92
27073/CPT#1	16 hours	100	unsaturated	5.45	1.2683 ± 20.37
27073/CPT#1	16 hours	150	unsaturated	5.49	1.0280 ± 23.63
27073/CPT#1	16 hours	200	unsaturated	5.55	0.4620 ± 76.42
27073/CPT#1	1 week	100	saturated	5.48	0.7520 ± 21.83
27073/CPT#1	1 month	100	saturated	5.67	1.5193 ± 11.68
27073/CPT#1	3 months	100	saturated	5.17	1.0927 ± 16.69
27073/CPT#1	5.4 months	100	saturated	5.15	1.1363 ± 15.01
27073/CPT#1	3 days	50	1 M KCl	9.22	1.0415 ± 16.30
27073/CPT#1	5 days	50	1 M KCl	8.45	1.2087 ± 12.54
27073/CPT#1	3 days	50	1 M CaCl ₂	3.09	0.9307 ± 18.71
27073/CPT#1	5 days	50	1 M CaCl ₂	2.67	1.0435 ± 15.93
27073/CPT#1	3 days	50	1 M NaCl	3.36	0.9141 ± 16.16
27073/CPT#1	5 days	50	1 M NaCl	2.93	0.4980 ± 29.43
27073/CPT#1	3 days	50	1 M CsCl	1.23	0.8580 ± 17.05
27073/CPT#1	5 days	50	1 M CsCl	1.04	1.8218 ± 6.24

* Radiogenic

† Determined from decay constants and isotopic abundance of ⁴⁰K according to Steiger and Jager, 1977

‡ Untreated bulk sample

§ Organic heavy-liquid treated bulk sample

¶ Estimated total percentage error (Cox and Dalrymple, 1967)

A depletion pattern in the radiogenic Ar contents is noted in the clinoptilolite fractions heated at different temperatures for 16 hours (Fig. 3b). About 70% of the radiogenic Ar was lost during the heating at 200 °C compared with the first sample heated at 50 °C, with most of the Ar diffusing out during the higher temperature (200 °C) heating (Table 3). The clinoptilolite aliquots heated in saturated conditions at lower temperature (100 °C) yielded variable radiogenic Ar contents, especially those fractions heated for a week and one month (Fig. 3c). The other two fractions heated for longer times contain similar amounts of Ar.

IV. DISCUSSION

The time of alteration of the silicic tuff to clinoptilolite in the Sucker Creek Formation is unknown. The radiogenic Ar contents for the clinoptilolite fractions reported here are minimum

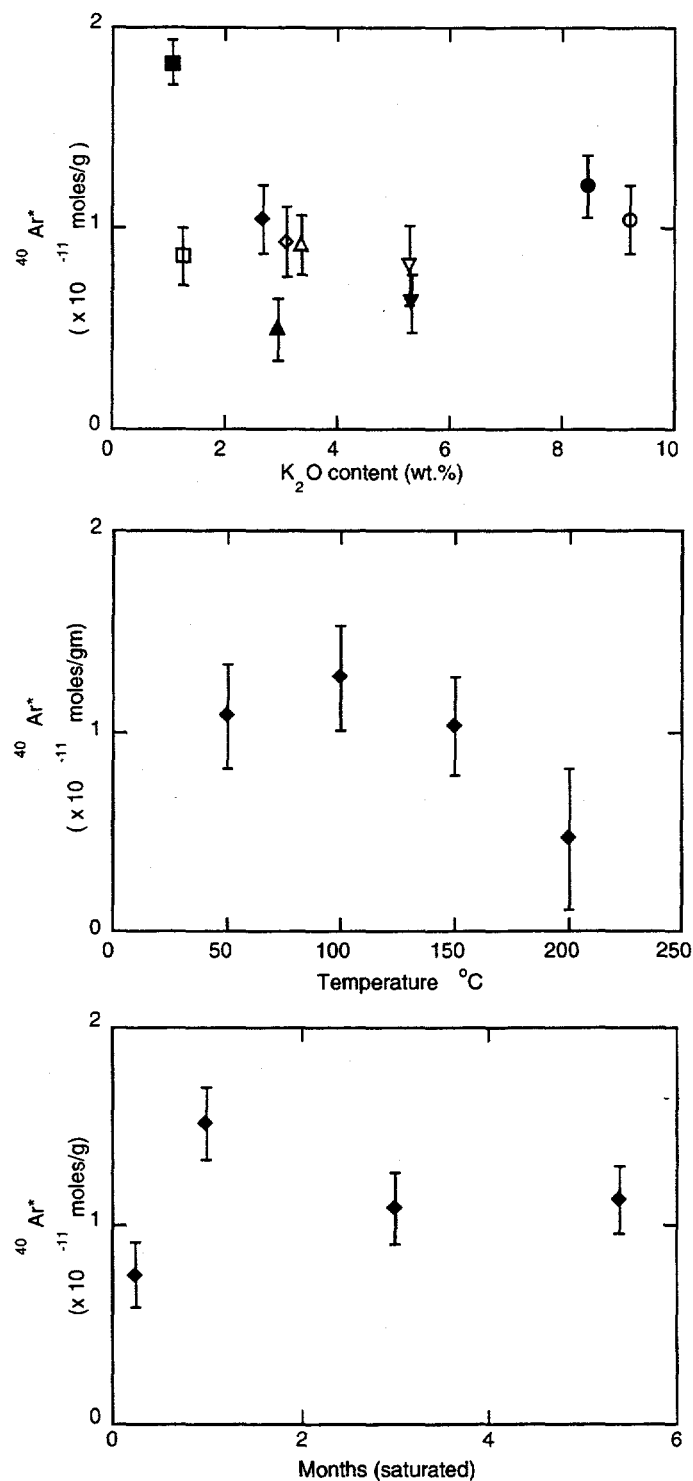


Figure 3. Radiogenic Ar content plots of ion-exchanged clinoptilolite fractions. Except for the inverted triangles, open and solid symbols in (A) represent clinoptilolite fractions ion exchanged for five and three days, respectively. Symbols represent Ca^{2+} (diamonds), Cs^+ (square), K^+ (circle), Na^+ (triangle); also shown are untreated (inverted triangle) and heavy-liquid treated (solid inverted triangle) sample results. Other panels show results from unsaturated heating at various temperatures for 16 hours (B) and saturated heating at 100°C for several time periods (C). Uncertainties are cumulative percentage errors.

values and do not represent the amount accumulated since the crystallization of the clinoptilolite from the alteration of the silicic shards. Isotopic ages were not calculated from the measured radiogenic Ar contents of the clinoptilolite fractions because the samples used in ion-exchange and dehydration experiments have modified K_2O and Ar contents. Calculated apparent isotopic ages from such experiments are thus meaningless. It has been reported that the clinoptilolite and other coexisting minerals in the Sucker Creek Formation were modified by post-diagenetic cation exchange during low-temperature (75-100 °C) hydrothermal alteration (Altaner and Grim, 1990). The smaller amounts of radiogenic Ar in the untreated bulk clinoptilolite fraction may thus be related to loss of Ar during the long-term hydrothermal heating of the zeolites. Moreover, loss of water during dehydration in an unsaturated environment appears to enhance loss of radiogenic Ar (WoldeGabriel, et al., 1993).

Zeolite separations from coexisting authigenic quartz, K-feldspar, and clays using sedimentation methods in deionized water and organic heavy liquids have minimal effects on the K_2O contents of alkali zeolites. This is indicated by the similarity in K_2O contents of untreated and organic heavy-liquid treated clinoptilolite fractions (Table 3). However, the radiogenic Ar content of the untreated and heavy-liquid treated fractions are lower compared with most of the untreated aliquots from the heating experiments of the same sample (Fig. 3a, Table 3). Although the depletion in the two fractions may be explained by partial loss of Ar, it is not clear why the untreated bulk sample lost more radiogenic Ar than most of the the bulk untreated fractions used for the heating experiments.

The K and Ar contents of the clinoptilolite fractions were variably affected during the ion-exchange experiments with the 1 M Ca-, Cs-, K-, and Na-chloride solutions (Fig. 3a). Most of the Ca-, Cs-, K-, and Na-clinoptilolite samples from the three- and five-day exchanges exhibited minor variations in their radiogenic Ar contents except for the Cs- and Na-clinoptilolite fractions exchanged for five days. Although the smaller amount of radiogenic Ar in the Na-clinoptilolite may be explained by loss of Ar, the source of the larger amount in the Cs-clinoptilolite is unclear. However, two of the untreated clinoptilolite fractions from the saturated (100 °C /1 month) and unsaturated (50 °C/16 hours) heating experiments contain similar amounts of radiogenic Ar (Table 3). Thus, the larger amounts in the three fractions may be real, and most of the other samples lost

some of their radiogenic Ar during the ion-exchange experiments. On the other hand, those samples with larger amounts of radiogenic Ar may be contaminated by few primary feldspar grains although the X-ray patterns do not indicate this to be the case. The larger amount of radiogenic Ar in the Cs-clinoptilolite fraction is not likely due to gain in radiogenic Ar because the similar ionic radius of Cs^+ (1.91 Å) compared with Ar (1.9 Å) probably minimized loss and/or gain of Ar from the clinoptilolite structure. According to Barrer (1978), ions and molecules traverse the same channel and their access to the zeolite structure is dependent on the size and shape of the diffusing ion or molecule.

Despite less K^+ depletion, the Na-clinoptilolites indicate an opposite pattern of Ar depletion compared with the Cs-clinoptilolites. The Na-clinoptilolite exchanged for 5 days lost half of its Ar compared with a similar fraction exchanged for three days (Table 3). The significant loss of Ar may be related to the smaller ionic radius of Na (1.1 Å or 1.24 Å, depending on the coordination number) and to loss of water. It is suggested that Na^+ and Ca^{2+} occupy water sites in a clinoptilolite structure during ion exchange, if there are no other locations, and that the high content of Na^+ and K^+ and their lower hydration energies result in less total water in clinoptilolite than a Ca-rich end member (Berger et al., 1970; Tsitsishvili et al., 1992).

Compared with the Cs- and Na-clinoptilolites, the Ar contents in the Ca- and K-clinoptilolites exhibit less variability (Fig. 3a). However, the Ca-clinoptilolites contain lesser amounts of radiogenic Ar than the K-clinoptilolite fractions. The radiogenic Ar contents in the K-clinoptilolites exchanged for three and five days are almost identical and the samples contain amounts of radiogenic Ar similar to most of the untreated clinoptilolite fractions from the saturated and unsaturated heating experiments.

The results of the ion-exchange experiment suggest the possible effects of post-diagenetic chemical interactions between zeolites and major and trace element cations in ground water. The K_2O contents of the ion-exchanged samples were affected significantly (Table 3). In the diagenetic environment, the effects of ground water chemistry on alkali zeolites can be evaluated by comparing the radiogenic Ar contents or isotopic ages of coexisting authigenic mineral phases (e.g., illite/smectite and K-feldspar) in the altered rock. Such studies can indicate whether the alkali zeolites gained or lost K and Ar during post-crystallization alteration. However, this type of

information does not exist for zeolites of the Sucker Creek Formation and little can be said about the geological history of the clinoptilolite-rich sample used in this experiment.

Post-crystallization modification of the K_2O and Ar contents of zeolites in the diagenetic environment may depend on 1) temperature, 2) ground and/or pore water composition, 3) permeable fluid pathways, and 4) fluid-rock interaction. Clauer (1982) reported a mid-Miocene diagenetic age for a phillipsite and implied that induration of the host rock was responsible for the cessation of diagenetic reactions between pore fluids and phillipsites in mid-Eocene ocean floor sediments. Similar interpretations were made by Ijima (1987) concerning the incomplete transformation of alkali zeolites to analcime in marine sediments, which he attributed to impermeability in the host rock and/or a slow rate of reaction that arrested the diagenetic reactions.

The radiogenic Ar contents of alkali zeolites may also be affected by post-diagenetic heating and/or dehydration as indicated by the preliminary results of the experiments. The zeolite-rich tuffs at Yucca Mountain are characteristic of a low-temperature diagenetic environment consisting of smectite, clinoptilolite, and opal-CT with an upper stability limit of about 100 °C (Iijima, 1978; Bish, 1989). Because water appears to minimize Ar diffusion from the clinoptilolite structure, loss of Ar during dehydration may be a major factor contributing to the depletion of radiogenic Ar contents from K-bearing clinoptilolite fractions in unsaturated environment. Alternatively, clinoptilolite existing in a partially saturated environment may contain less water, facilitating the more rapid diffusion of Ar.

In one experiment, loss of Ar from a clinoptilolite fraction heated for 16 hours at 200 °C is probably related to the loss of water from the mineral structure. This loss of Ar from clinoptilolite occurred at a temperature (200 °C) that is higher than the upper stability limit of alkali zeolites and other coexisting authigenic minerals (e.g., smectite and opal-CT). Dehydration is known to cause increased atomic displacement and disorder in the tetrahedral framework of alkali zeolites (Bish, 1988; Armbruster, 1993), that is likely to result in Ar loss from the framework structure. Variable isotopic ages (2-13 Ma) were measured on clinoptilolite fractions from the saturated and unsaturated zones at Yucca Mountain, Nevada, that have been attributed to dehydration of clinoptilolite in the unsaturated zone resulting in Ar loss and younger isotopic ages (WoldeGabriel et al., 1993). This conclusion is consistent with the preliminary experimental results reported here.

According to Barrer and Vaughan (1969), Ar diffusion can be minimized by rehydration of clinoptilolite as long as it is not heated above 110 °C. This is consistent with the results of the older isotopic ages measured on clinoptilolite fractions from the saturated zone at Yucca Mountain (WoldeGabriel et al., 1993). Moreover, the importance of saturation is demonstrated by minimal Ar loss from clinoptilolite fractions that were heated at 100 °C in a saturated environment for more than five months. Radiogenic Ar contents of clinoptilolite fractions from the saturated and unsaturated heating experiments (Figs. 3b and 3c) clearly reflect the variation pattern indicated by the clinoptilolite K/Ar data from the saturated and unsaturated zones at Yucca Mountain. This similarity confirms the importance of water in maintaining the retention of Ar within K-bearing zeolites. Clinoptilolite isotopic ages increase toward the lower part of the unsaturated zone at Yucca Mountain (WoldeGabriel et al., 1993), and this may be related to the proximity of the dated samples to the static water level. The preliminary results from this study will be used in experiments that will be conducted to document Ar mobility in hydrated and dehydrated clinoptilolite fractions at different conditions that possibly reflect the zeolite diagenetic environment.

V. SUMMARY AND CONCLUSIONS

The radiogenic Ar contents from the clinoptilolite fractions used in the experiment represent minimum amounts and are consistent with loss of Ar. Results from the ion exchange indicate that K⁺ contents are significantly modified compared with the Ar composition of the same samples. K-bearing zeolites in a permeable diagenetic environment can be modified during water-rock interaction especially if the fluids are saline and/or alkaline in composition. Ar depletion was apparent in a clinoptilolite fraction dehydrated by heating at different temperatures. A similar result was obtained from a Na-clinoptilolite exchanged for five days.

Samples heated in a saturated environment at 100 °C for various time periods indicate irregular variations in their Ar contents unlike those samples from the unsaturated heating. Water plays an important role in maintaining Ar retention in the clinoptilolite framework. This is consistent with the clinoptilolite K/Ar data from the saturated zone at Yucca Mountain that yield older isotopic ages compared with those samples from the same section in the unsaturated zone.

Thus, if K-bearing zeolites are used in conjunction with other dateable coexisting authigenic minerals, information from such minerals may be useful in understanding the diagenetic history of a zeolite-rich environment.

ACKNOWLEDGMENTS

This work was supported by the Yucca Mountain Site Characterization Project Office as part of the Civilian Radioactive Waste management Programs of the U.S. Department of Energy. This project is managed by the U.S. Department of Energy Yucca Mountain Site Characterization Project. Comments by David Bish, David Broxton, David Vaniman, and Schon Levy improved the manuscript. In compliance with LANL-YMP-QP-03.23, R2, the DTN # is LA000000000121.001.

REFERENCES

- Altaner, S. P. and Grim R. E., "Mineralogy, chemistry, and diagenesis of tuffs in the Sucker Creek Formation (Miocene), Eastern Oregon," *Clays and Clay Minerals* **38**, No. 6, 561-572 (1990), YMP accession no. MOL.19941220-0115.
- Ames, L.L., Jr., "Cation sieve properties of the open zeolites chabazite, mordenite, erionite, and clinoptilolite," *Am. Mineral.* **45**, 689-700 (1961), YMP accession no. NNA.921026.0010.
- Armbruster, T., "Dehydration mechanism of clinoptilolite and heulandite: single-crystal X-ray study of Na-poor, Ca-, K-, Mg-rich clinoptilolite at 100K," *Am. Mineral.* **78**, 260-264 (1993), YMP accession no. NNA.930625.0035.
- Barrer, R.M., and Vaughan, J.H., "Sorption and diffusion of rare gases in heulandite and stilbite," *Surface Science* **14**, 77-92 (1969), YMP accession no. NNA.930629.0002.
- Barrer, R. M., "Cation-exchange equilibria in zeolites and feldspathoids," In *Natural Zeolites: Occurrences, Properties, Use* (eds. Sand, L. B. and Mumpton, F.A.), Pergamon Press, Elmsford, New York, 385-395 (1978). Readily Available.
- Bell, T. E., "Deposition and diagenesis in the Brush Basin Member and upper part of the Westwater Canyon Member of the Morrison Formation, San Juan basin, New Mexico," In *A Basin Analysis Study-The Morrison Formation, Grants Uranium Region, New Mexico* (eds. C. E. Turner-Peterson, E.S. Santos, and N. S. Fishman); American Association of Petroleum Geologists Studies in Geology **22**, 77-91 (1986), YMP accession no. MOL.19941220-0116.

- Bernat, M., Bieri, R. H., Koide, M., Griffin, J. J., and Goldberg, E. D., "Uranium, thorium, K, and Ar in marine phillipsites," *Geochem. Cosmochim. Acta* **34**, 1063-1071 (1970), YMP accession no. NNA.930629.0003.
- Berger, I. A., Chandler, J. C., and Zubovic, P., "An infrared study of water in heulandite and clinoptilolite. *Am. Mineral.* **55**, 825-840 (1970), YMP accession no. MOL.19941220.0117.
- Bish, D.L., "Effects of composition on the dehydration behavior of clinoptilolite and heulandite," In *Occurrence, Properties, and Utilization of Natural Zeolites* (eds., Kallo, D., and Sherry, H.S.), Akadémiai Kiadó, Budapest, 565-576 (1988), YMP accession no. NNA.890525.0001.
- Bish, D.L., "Evaluation of past and future alterations in tuff at Yucca Mountain, Nevada, based on clay mineralogy of drill cores USW G-1, G-2, and G-3," *Los Alamos National Laboratory Report*, LA-10667, 40 p (1989), YMP accession no. NNA.890126.0207.
- Bish, D. L. and Aronson, J. L., "Paleogeothermal and paleohydrologic conditions in silicic tuff from Yucca Mountain, Nevada," *Clays and Clay Minerals* **41**, 148-161 (1993), YMP accession no. NNA.940323.0292.
- Boles, J. R., "Composition, optical properties, cell dimension, and thermal stability of some heulandite-group zeolites," *Am. Mineral.* **57**, 1463-1493 (1972), YMP accession no. MOL.19941220.0118.
- Boles, J. R. "Zeolites in low-grade metamorphic rocks," In *Mineralogy and Geology of Natural Zeolites* (eds., Mumpton, F. A.), *Mineralogical Society of America* **4**, 103-134 (1977), YMP accession no. MOL.1994122.0119.
- Clauer, N., "Strontium isotopes of Tertiary phillipsites from the Southern Pacific: Timing of the geochemical evolution," *J. Sed. Petrol.* **88**, 1003-1009 (1982), YMP accession no. NNA.930720.0001.
- Cox, A. and Dalrymple, G. B., "Statistical analysis of geomagnetic reversal data and the precision of potassium-argon dating," *J. Geophys. Res.* **72**, No. 10, 2603-2614 (1967), YMP accession no. NNA.930629.0006.
- Elliot, C. W., Aronson, J. L., Matisoff, G., and Gautier, D. L., "Kinetics of the smectite to illite transformation in the Denver basin: clay mineral, K-Ar data, and mathematical model results," *Amer. Associ. of Petr. Geol. Bull.* **75**, 436-462 (1991), YMP accession no. NNA.930625.0037.
- Feng, X. and Savin, S. M., "Oxygen isotope studies of zeolite, stilbite, analcime, heulandite, and clinoptilolite: II. Kinetics and mechanisms of isotopic exchange between zeolites and water vapor," *Geoch. et Cosmochim. Acta* **57**, 4219-4238 (1993), YMP accession no. MOL.19941220.0120.

- Govindaraju, K., Special issue of *Geostandards Newsletter*, XIII, **21** (1989), YMP accession no. MOL.19941220.0121.
- Gundogdu, M. N., Bonnet-Courtois, C., and Clauer, N., "Isotope and chemical signatures of sedimentary smectite and diagenetic clinoptilolite of lacustrine Neogene basin near Bigadic, Western Turkey," *Appl. Geochem.* **4**, 635-644 (1989), YMP accession no. NNA.930629.0009.
- Hay, R. L., "Stratigraphy and zeolites diagenesis of the John Day Formation of Oregon," *Univ. of Calif. Publ. Geol. Sci.* **42**, 199-262 (1963), YMP accession no. NNA.930629.0009.
- Hay R.L., "Geologic occurrences of zeolites," In *Natural Zeolites: Occurrences, Properties, Use* (eds., Sand L. B. and Mumpton F. A.); Pergamon Press, Elmsford, New York, 135-143 (1978), YMP accession no. NNA.870406.0096.
- Hoffman, J., Hower, J., and Aronson, J. L., "Radiometric dating of time of thrusting in the disturbed belt of Montana," *Geology* **4**, 16-20 (1976), YMP accession no. MOL.19941220.0122.
- Iijima, A., "Geological occurrences of zeolites in marine environments," In *Natural Zeolites: Occurrences, Properties, Use* (eds., Sand, L.B., and Mumpton, F.A.); Pergamon Press, Elmsford, New York, 175-198 (1978). Readily Available.
- Koul, S. L., Chdderton, L. T., and Brooks, C. K., "Fission track dating of zeolites" *Nature* **294**, 347-350 (1981), YMP accession no. NNA.930625.0038.
- Mumpton, F. A., "Natural zeolites: a new industrial mineral commodity," In *Natural Zeolites: Occurrences, Properties, Use* (eds. Sand, L. B. and Mumpton, F.A.); Pergamon Press, Elmsford, New York, 3-27 (1978). Readily Available.
- Odin, G. S., and 35 collaborators, "Interlaboratory standards for dating purposes," In *Numerical Dating in Stratigraphy* (ed., Odin, G. S.), Wiley, New York, 124-150 (1982). Readily Available.
- Sheppard, R. A., Gude, A. J. 3rd, and Mumpton, F. A., "Sheaville zeolite deposit, Sheaville, Oregon," In *Zeo-Trip '83* (ed. Mumpton, F. A.), State University of New York, Brockport, 25-31 (1983), YMP accession no. MOL.19941220.0123.
- Steiger, R.H., and Jager, E., "Subcommission on geochronology: Convention on the use of decay constants in geo- and cosmochemistry," *Earth and Planet. Sci. Lett.* **36**, 359-362 (1977), YMP accession no. NNA.930629.0016.
- Tsitsishvili, G. V., Andronikashvili, T. G., Kirov, G. N., and Filizova, L. D., "Natural Zeolites," Ellis Horwood, New York, 18-34 (1992). Readily Available.
- Turner, C.E. and Fishman, N. S., "Jurassic Lake T'oo'dichi': A large alkaline, saline lake, Morrison Formation, eastern Colorado Plateau," *Geological Society of America Bulletin*, **103**, 538-558 (1991), YMP accession no. MOL.19941220.0124.

- Vaughan, D. E. W., "Properties of natural zeolites," In Natural Zeolites: Occurrences, Properties, Use (eds. Sand, L. B. and Mumpton, F.A.); Pergamon Press, Elmsford, New York, 353-371 (1978), YMP accession no. NNA.921026.0011.
- WoldeGabriel, G. and Goff, F., "Temporal relations of volcanism and hydrothermal systems in two areas of the Jemez volcanic field, New Mexico," *Geology* **17**, 986-989 (1989), YMP accession no. NNA.930629.0019.
- WoldeGabriel, G., Broxton, D. E., Bish D. L., and Chipera, S. J., "Mineralogy and clinoptilolite K/Ar results from Yucca Mountain, Nevada, USA: A potential high-level radioactive waste repository site," Los Alamos National Laboratory Report, LA-12652, 44 p. (1993), YMP accession no. NNA.930921.0037.