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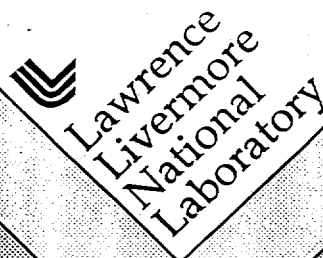
**A New Synthesis of TATB Using Inexpensive Starting
Materials and Mild Reaction Conditions**

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A NEW SYNTHESIS OF TATB USING INEXPENSIVE STARTING MATERIALS AND MILD REACTION CONDITIONS*

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Abstract

1,3,5-Triamino-2,4,6-trinitrobenzene (TATB) is currently manufactured in the USA by nitration of the relatively expensive and domestically unavailable 1,3,5-trichlorobenzene (TCB) to give 2,4,6-trichloro-1,3,5-trinitrobenzene (TCTNB) which is then aminated to yield TATB. Elevated temperatures (150 °C) are required for both reactions. There is a need for a more economical synthesis of TATB that also addresses current environmental issues. We have recently discovered that 1,1,1-trimethylhydrazinium iodide (TMHI) allows the amination of nitroarenes at ambient temperature via Vicarious Nucleophilic Substitution (VNS) of hydrogen. TMHI reacts with 2,4,6-trinitroaniline (picramide) in the presence of strong base (NaOMe or *t*-BuOK) to give TATB in over 95% yield. TMHI and picramide can be obtained from either inexpensive starting materials or surplus energetic materials available from demilitarization activities. TMHI, for example, is obtained from the reaction of methyl iodide with 1,1-dimethylhydrazine (unsymmetrical dimethylhydrazine, UDMH). UDMH (30,000 metric tons) is currently available in the former Soviet Union as a surplus rocket propellant. Transition of the new TATB synthesis from the laboratory to the pilot plant is projected.

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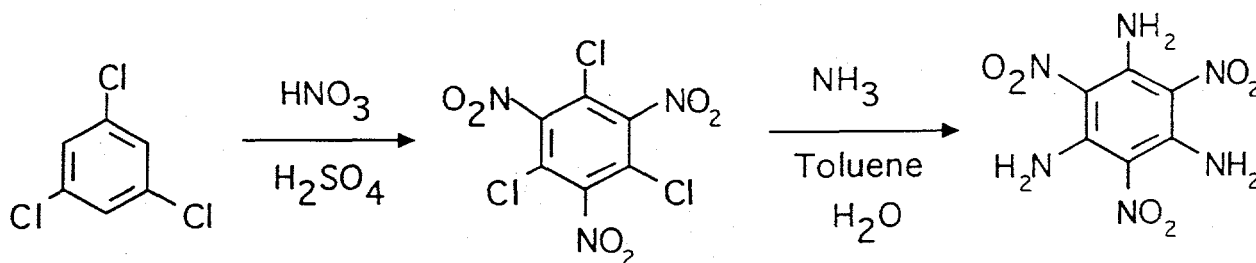
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Introduction

The objective of this program is to provide a convenient and relatively inexpensive preparation of 1,3,5-triamino-2,4,6-trinitrobenzene (TATB). TATB is a reasonably powerful high explosive whose thermal and shock stability is considerably greater than that of any other known material of comparable energy.¹ It is used in military applications because of its significant insensitivity to thermal and shock environments. There is also interest in employing TATB in the civilian sector for deep oil well explorations where heat-insensitive explosives are required. TATB had been priced at \$20 to \$50 per pound when it was being produced on an industrial scale in the U.S. TATB is currently available to customers outside of the Department of Energy (DOE) for \$100 per pound.² This paper describes a new synthesis of TATB developed at Lawrence Livermore National Laboratory (LLNL) that utilizes inexpensive starting materials and mild reaction conditions.

Current Preparation of TATB

TATB is produced in the USA by nitration of the relatively expensive and domestically unavailable 1,3,5-trichlorobenzene (TCB) to give 2,4,6-trichloro-1,3,5-trinitrobenzene (TCTNB) which is then aminated to yield TATB (Scheme 1).³ Elevated temperatures (150 °C) are required for both reactions. The major impurity encountered in this process is ammonium chloride. The inclusion of



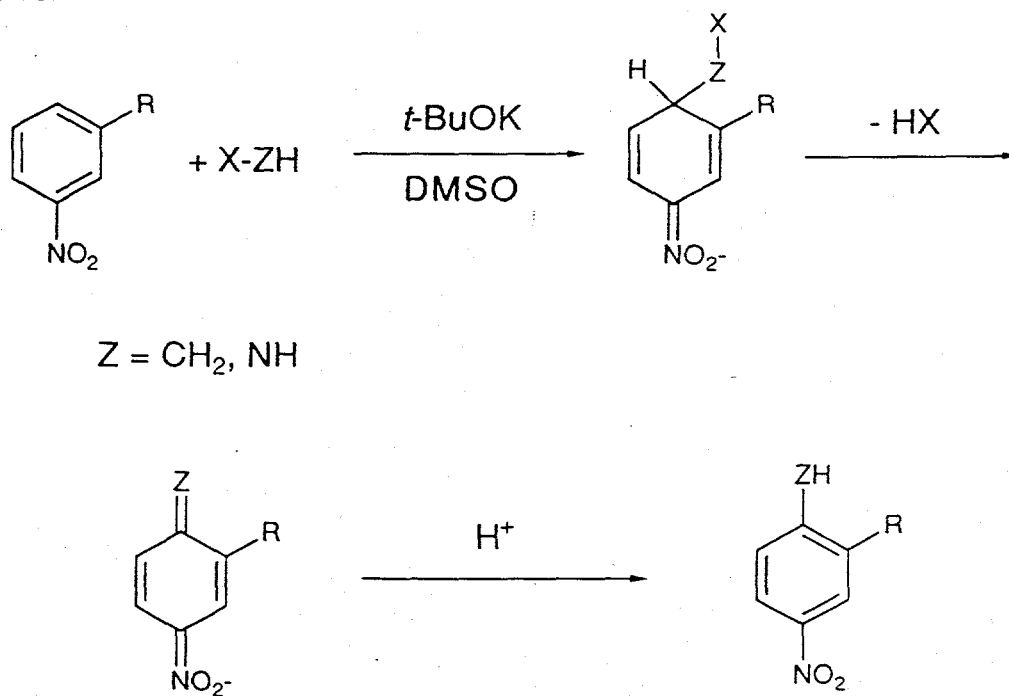
Scheme 1. Current synthesis of TATB.

2.5% water during the amination step significantly reduces the ammonium chloride content of the TATB.⁴ Low levels of chlorinated organic impurities have also been identified. These impurities include 2,4,6-trichloro-1,3,5-

trinitrobenzene (TCTNB), 1,3-dinitro-2,4,5,6-tetrachlorobenzene, 1,3-dinitro-2,4,6-trichlorobenzene and their partially aminated products.⁵ A synthesis of TATB that replaces TCTNB with a non-chlorinated starting material is clearly desirable.

Vicarious Nucleophilic Substitution of Hydrogen

Vicarious nucleophilic substitution (VNS) of hydrogen is a well-established procedure for the introduction of carbon nucleophiles into electrophilic aromatic rings.^{6,7} The reaction involves the addition of a carbanion bearing a leaving group (X) to an electrophilic aromatic ring and subsequent rearomatization by loss of the leaving group through elimination as HX (Scheme 2, Z=CH₂). This reaction has been applied to a wide variety of nitroarenes and nitro-substituted heterocycles.⁷



Scheme 2. Vicarious nucleophilic substitution of hydrogen.

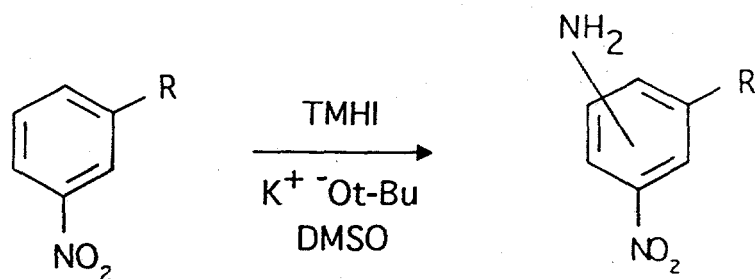
By analogy, VNS reactions can also take place with amine nucleophiles. Such reagents are of the common form X-NH₂, where X is an auxiliary group capable of stabilizing a negative charge, thus driving rearomatization of the σ -

intermediate adduct (Scheme 2, Z=NH). One of the first examples of amination by VNS of hydrogen was provided by Meisenheimer and Patzig who reacted 1,3-dinitrobenzene with hydroxylamine in the presence of strong base to yield 2,4-dinitrophenylene-1,3-diamine.⁸ Recently, a number of more active aminating reagents such as 4-amino-1,2,4-triazole⁹ and substituted sulfenamides¹⁰ have been developed. These reagents were designed to be good nucleophiles that easily add to electrophilic aromatic rings and also possess good leaving groups (Scheme 2, X)

Use of 1,1,1-Trimethylhydrazinium Iodide as a VNS Reagent

Although quaternary hydrazinium compounds have been known for over one hundred years,^{11,12} they have not been employed as VNS reagents. We examined 1,1,1-trimethylhydrazinium iodide (TMHI)¹³ $[(CH_3)_3N^+-NH_2 I^-]$ for use as a VNS reagent. We reasoned that TMHI would be sufficiently nucleophilic to substitute into nitro-substituted aromatic rings but would be superior to the previous examples because the leaving group would be the neutral trimethylamine instead of a stabilized anionic species. In addition, there is a possibility that the hydrazinium halide would react with base to form the neutral ylide species, $[(CH_3)_3N^+-NH^-]$, which may be the reactive species in the amination process. Indeed, when TMHI was reacted with various nitro-substituted aromatics the amino functionality was introduced in good to excellent yields. We found that the number of amino-groups which may be added to the electrophilic aromatic ring is equal to the number of nitro groups present on the ring.

A study of product yields and distribution of various 3-substituted nitrobenzene derivatives was performed using TMHI (Scheme 3) (Table 1).¹⁴ The results of our study were compared with the findings reported for 4-amino-1,2,4-triazole (ATZ).⁹ ATZ was found to be regioselective, giving substitution exclusively para- to the nitro group. TMHI gave all possible product isomers thereby showing no selectivity but presumably greater reactivity. There was a general tendency for TMHI to yield products in which the amine substitution occurs ortho- to the nitro group as the major components although some exceptions were noted. We are



Scheme 3. Amination of 3-substituted nitroaromatics.

currently investigating the use of more sterically crowded 1,1,1-trialkylhydrazinium halides derived from UDMH in an attempt to influence the regioselectivity of the aminating reagent.

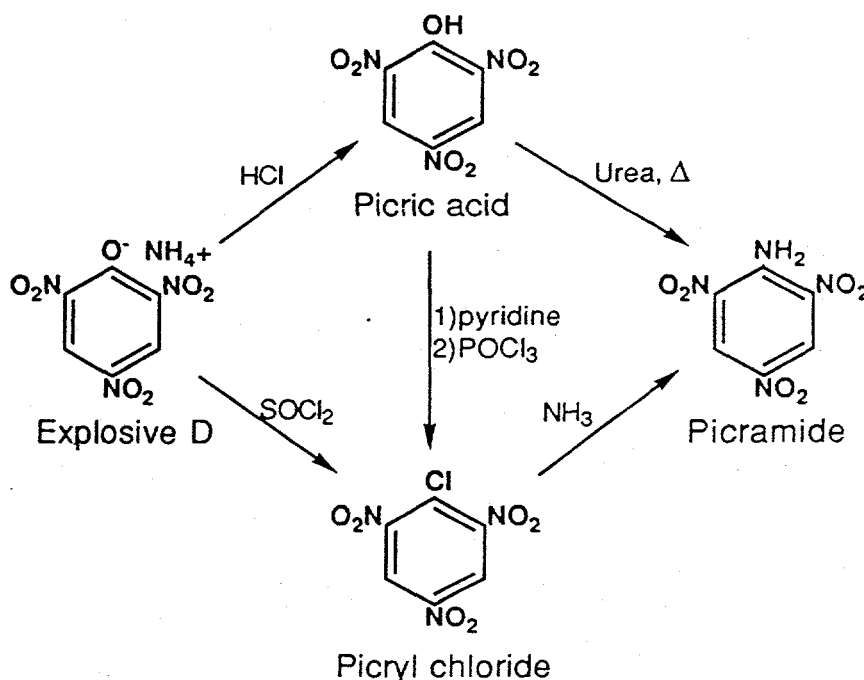
Table I. Amination of 3-Substituted Nitrobenzenes

R	Total Yield (%)	position of NH ₂ ^a	% isomer
H	85	2	61
		4	39
CH ₃	84	2	38
		4	35
		6	27
Cl	82	2	32
		4	49
		6	19
COOH	95	4	71
		6	29
OCH ₃	66	2	90
		4	10
F	84	2	45
		4	47
		6	8
I	76	2	45
		4	38
		6	17
CN	41	2	20
		4	44
		6	36

^a Relative to NO₂

Starting Materials for TATB Synthesis

This project originally began as part of demilitarization activities related to the chemical conversion of energetic materials into higher value products.^{15,16} Consequently, our synthesis of TATB has been designed so that either surplus energetic materials or inexpensive, commercially available chemicals can be used to make the necessary starting materials. Several million pounds of ammonium picrate (Explosive D) are available for disposal in the USA. Ammonium picrate can be converted to 2,4,6-trinitroaniline (picramide) by the pathways shown in Scheme 4. The reaction of picric acid with urea at elevated temperature (173 °C) has been reported to provide picramide in 88% yield.¹⁷ The analogous reaction of Explosive D with urea to directly provide picramide has not been reported. Picramide is also readily accessible from nitration of the commercially available and relatively inexpensive 4-nitroaniline¹⁸.



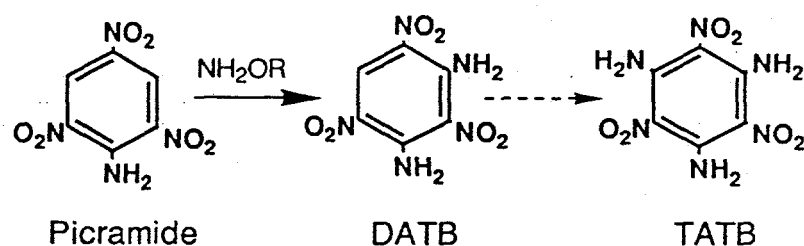
Scheme 4. Synthesis of picramide from Explosive D.

1,1-Dimethylhydrazine (unsymmetrical dimethylhydrazine, UDMH) has been used as a liquid rocket propellant. Thirty thousand metric tons of UDMH in Russia are available for disposal in a safe and environmentally responsible manner.¹⁹ 1,1,1-trimethylhydrazinium iodide (TMHI), our VNS reagent of choice,

is readily prepared through the alkylation of UDMH by methyl iodide.²⁰ Alternatively, TMHI can be prepared directly from hydrazine and methyl iodide.¹³

Amination of Picramide by Hydroxylamine Derivatives

We initially explored the conversion of picramide to DATB and TATB using hydroxylamine (Scheme 5, R = H) and aqueous base as reported for the conversion of 4,6-dinitrobenzofuroxan (DNBF) to give 5,7-diamino-4,6-dinitrobenzofuroxan (CL-14).²¹ DATB was obtained in a low yield (16%) after removal of unreacted picramide.

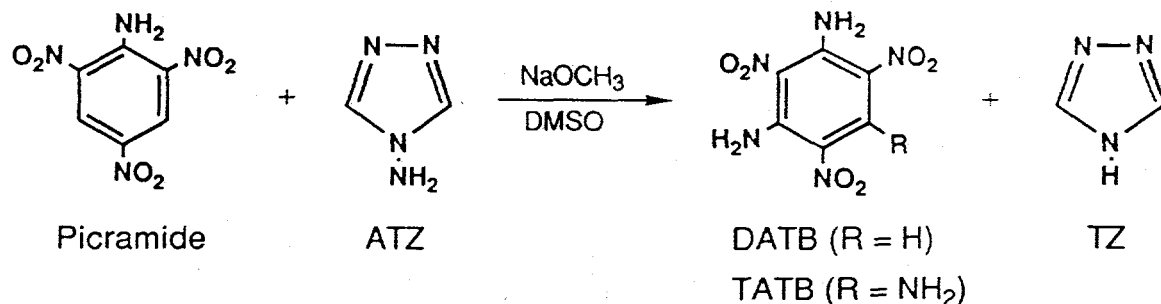


Scheme 5. Reaction of picramide with NH₂OR (R = H, CH₃, C₆H₅CH₂)

The replacement of aqueous base with sodium methoxide in anhydrous methanol or DMSO raised the yield of DATB from 16% to over 50%. The use of alkoxy derivatives (R = CH₃, C₆H₅CH₂) gave exclusively DATB in 87-91% crude yields. The reaction of picramide with hydroxylamine or its O-alkyl derivatives never yielded more than trace quantities of TATB.

Amination of Picramide by ATZ

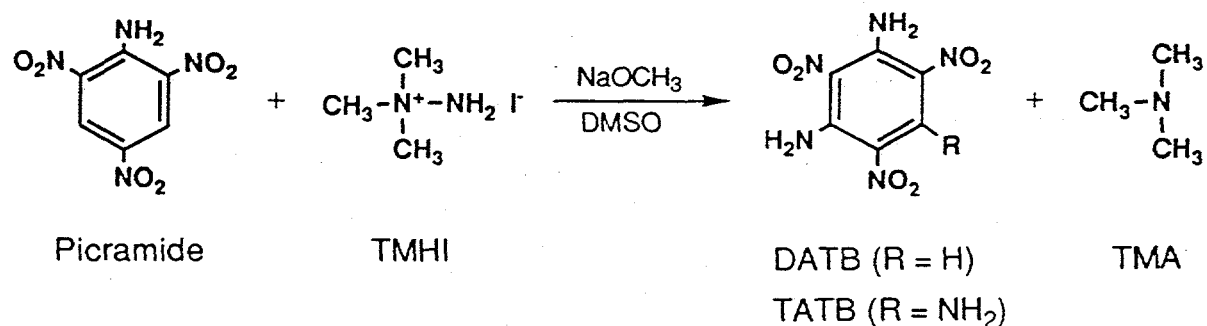
Picramide reacts with 4-amino-1,2,4-triazole (ATZ) in dimethylsulfoxide in the presence of strong base (NaOMe or KO^t-Bu) at room temperature (24 h) to yield, depending on the quantity of ATZ used, either DATB or TATB and 1,2,4-triazole (TZ) (Scheme 6).



Scheme 6. Reaction of picramide with ATZ.

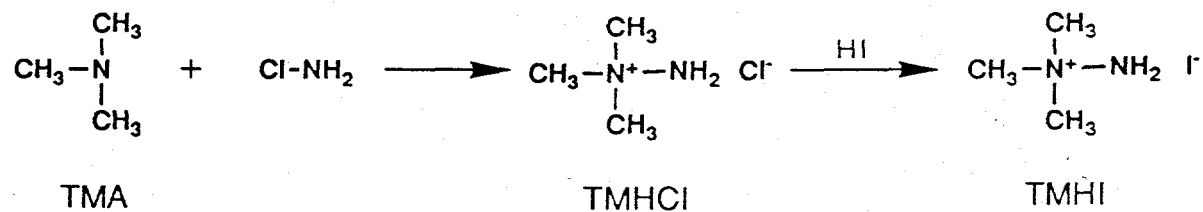
Amination of Picramide by TMHI

When TMHI replaces ATZ in the reaction with picramide, the reaction is complete in 3 h to yield DATB (70-80%) or TATB (95-98%) and trimethylamine (TMA) (Scheme 7).



Scheme 7. Reaction of picramide with TMHI.

The TMA released can be used to regenerate the TMHI that is consumed during the reaction. The reaction of TMA with chloramine (ClNH₂) yields trimethylhydrazinium chloride (TMHCl) which is treated with hydroiodic acid to yield TMHI (Scheme 8).²²

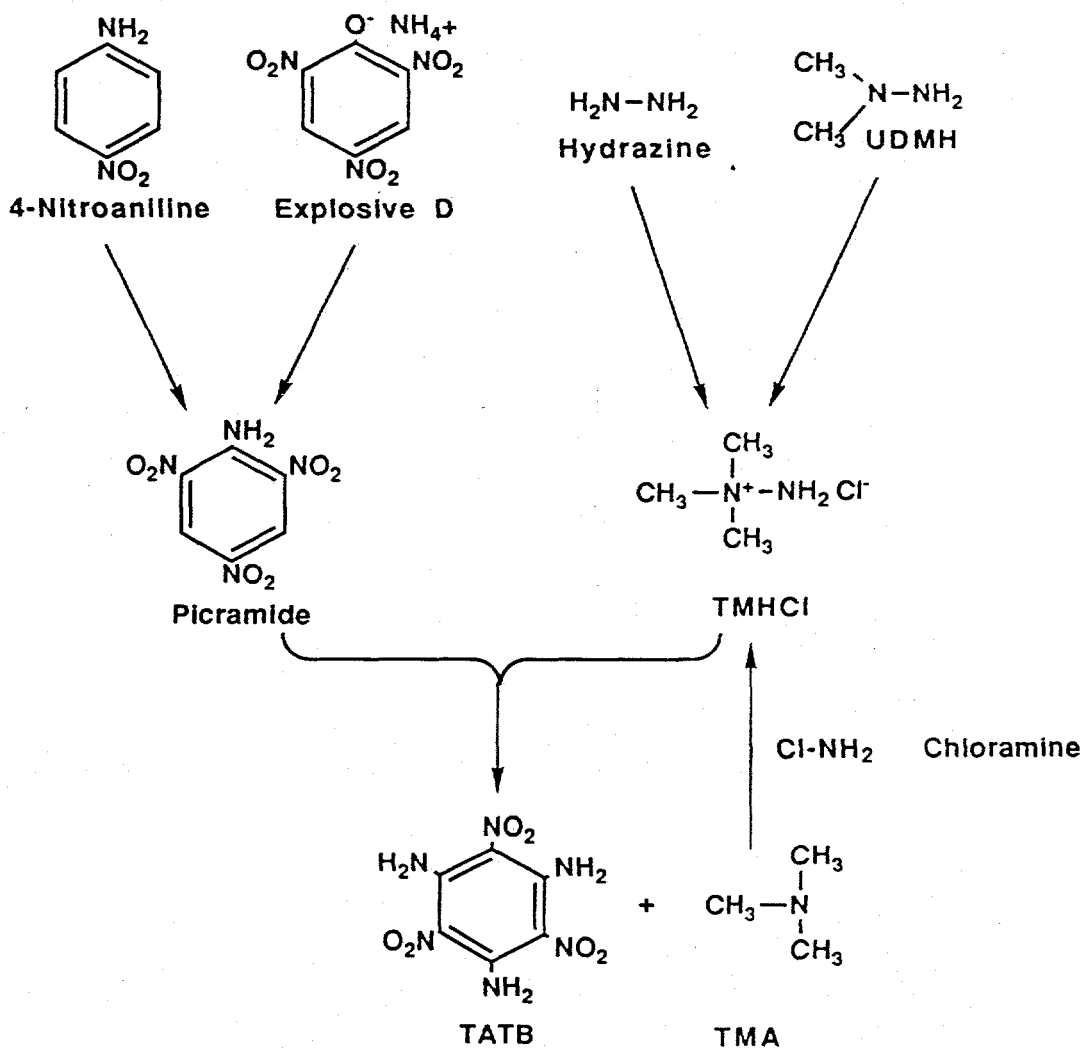


Scheme 8. Conversion of TMA to TMHI..

The use of TMHCl for nucleophilic aminations of nitroarenes is under investigation.

Summary

The new synthesis of TATB that we intend to scale up from the laboratory bench to pilot plant is outlined in Scheme 9. The starting materials are relatively inexpensive and can be obtained from demilitarization programs (Explosive D, UDMH) or commercial sources (4-nitroaniline, hydrazine). We have eliminated



Scheme 9. Synthesis of TATB from inexpensive starting materials.

the use of chlorinated arenes in response to environmental concerns and will recycle solvents and materials whenever possible.²³ Trimethylamine, a noxious and moderately toxic gas (bp 3 °C) released during the VNS reaction, will be reacted with chloramine to regenerate the TMHCl consumed in the synthesis of TATB.

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