Photocatalytic Oxidation of Dissolved Mn^{2+} by TiO_2 and the Formation of Tunnel Structured Manganese Oxides

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ABSTRACT

The redox reaction of manganese (Mn) is of great environmental, geological, and public health significance, as Mn oxides control the distribution and electron flow of numerous nutrients and contaminants in natural and engineered environments. Current understanding on the oxidation pathways of Mn(II) to Mn(III/IV) mainly focuses on biotic processes due to their much higher oxidation rates than those associated with abiotic processes. This study demonstrates rapid photocatalytic oxidation of Mn^{2+}(aq) under circumneutral conditions catalyzed by naturally abundant semiconducting TiO_2 minerals. Notably, the photocatalytic oxidation rates are comparable to or even higher than those of reported biotic/abiotic processes. In addition, the rapid photocatalytic oxidation leads to the formation of large tunnel structured Mn oxides (todorokite and romanechite) on the surface of TiO_2. These findings suggest that photocatalytic oxidation of Mn^{2+}(aq) by natural semiconducting minerals is likely an important yet previously overlooked pathway for understanding the occurrence of natural Mn oxide coatings on rock surfaces. In addition, considering the increasing input of photo-reactive engineered nanoparticles into environmental systems, this study shows the potential impacts of nanoparticles on influencing natural redox cycles.

Keywords: Mn oxides, Ti oxide, nanoparticles, photochemistry, oxidation

1. INTRODUCTION

Mn(III/IV) (oxyhydr)oxides (hereafter Mn oxides) are a group of ubiquitous natural minerals in terrestrial and aquatic settings. They play important roles in numerous elemental cycles and affect the electron flow in nature.[1-9] They have high reduction potential ($E_{\mathrm{MnO_2}/Mn^{2+}}$ = ~500 mV at pH 7)[10] and high specific surface area (around 10 to 200 m^2/g)[11, 12] and are among the most reactive minerals and the most significant solid oxidants in nature. Thus, understanding the redox reactions of Mn in fresh water and high salinity aqueous systems (e.g., sea water) has been considered a key to elucidate the geochemical electron cycles in the history of Earth and Mars.[13-16] With their strong redox reactivity and high adsorption capacity, Mn oxides and related compounds are also widely used in energy and environmental engineering systems for energy storage, water treatment, and contaminant removal.[17-21] For example, recent studies demonstrated the application of Mn oxides for harvesting energy from wastewater and salinity gradient[22, 23] as well as for the degradation of antimicrobial agents, organic contaminants, and heavy metals[23-28].

Because of the important roles that Mn oxides play in natural and engineered systems, many studies have explored the kinetics and mechanisms for their formation and transformation, as well as redox reactions involving Mn oxides. Previous studies showed that the abiotic homogeneous oxidation of Mn^{2+}(aq) by molecular oxygen is kinetically sluggish and takes years even though it is thermodynamically favorable.[29, 30] Mineral surface catalyzed heterogeneous oxidation of Mn^{2+}(aq) by molecular oxygen takes 5–2,800 days (half-life) under circumneutral conditions.[29, 31] In contrast, microbially mediated processes via enzymes (e.g., multicopper oxidase) or reactive oxygen species (ROS, such as superoxide) show much faster oxidation (1–69 days half-life).[31-34] Because of the widespread presence of Mn-oxidizing microorganisms and the fast oxidation rate, biotic processes have been generally accepted as the most significant contributor to

the oxidation of Mn^{2+}(aq) and formation of Mn oxides in natural low temperature environments. Interestingly, recent studies showed that photochemically generated superoxide by dissolved organic matter or nitrate can also enable indirect photooxidation of Mn^{2+}(aq) (i.e., oxidation of Mn^{2+}(aq) by photochemically generated oxygen-related species) at rates comparable to biotic processes.[35-39] These exciting results shed lights on the potentially overlooked contribution of photochemistry, an abiotic process, to the oxidation of Mn^{2+}(aq) and formation of Mn oxides.

Notably, previously known biotic/abiotic processes typically lead to the formation of highly disordered and poorly crystalline layered Mn oxides (LMO) that are structurally similar to vernadite (δ-MnO_2) or hexagonal birnessite. Although diverse tunnel structured Mn oxides (TMO) occur ubiquitously in many environmental settings[40], the detailed mechanisms of their formation have not been clearly resolved in laboratory studies under circumneutral conditions. Most studies showed the transformation of LMO to TMO under pH or temperature conditions significantly deviating from low temperature circumneutral conditions.[41-44] In our recent study under circumneutral pH and ambient temperature condition, we demonstrated redox cycling driven transformation of LMO to TMO.[45] However, although this study provided a new angle for understanding the transformation of LMO to TMO under fluctuating natural conditions such as those at oxic/anoxic interfaces, direct formation of TMO from the oxidation of Mn^{2+}(aq) was rarely observed in laboratory conditions and its feasibility and mechanism remain a puzzle.

Natural semiconducting minerals (such as metal oxides and sulfides) are widespread in nature and capable of direct redox reactions at conduction or valence band via photo-excitation.[46] Although evidence for the direct heterogeneous photocatalysis of Mn^{2+}(aq) (i.e., oxidation through the electron transfer from Mn^{2+}(aq) to valence band of natural semiconducting mineral) is lacking, it is thermodynamically feasible as the valence band position of many natural minerals is higher

than the reduction potential of Mn(III/IV)/Mn^{2+}(aq). Thus, a photo-excited hole in a valence band is thermodynamically feasible to oxidize Mn^{2+}(aq). A recent observation on the photon-to-electron conversion by natural semiconducting minerals on rock surfaces supports the potential of mineral photocatalysis and the subsequent direct photocatalytic oxidation of Mn^{2+}(aq) by natural semiconducting minerals.[47] A recent study described the oxidation of Mn^{2+}(aq) on natural semiconducting minerals (Fe and Ti oxides), but the high concentration of Mn^{2+}(aq) (14 mM) used in the study was not representative of natural conditions.[48] Systematic studies are thus desired to further explore the reaction kinetics and mechanisms of mineral catalyzed photocatalytic oxidation of Mn^{2+}(aq) under environmentally relevant conditions, as well as the structure of the formed Mn oxides.

In natural systems, one of the most photoreactive semiconducting minerals is TiO_2. Ti is the 9th most abundant element in Earth's crust, and rutile is the most commonly occurring structure of natural TiO_2.[49] Anatase is a polymorph of TiO_2.[50, 51] Due to their strong photoreactivity, both rutile and anatase nanoparticles are used extensively in photocatalytic energy and environmental applications.[46, 51-53] Because of the widespread anthropogenic use of TiO_2 in the last two decades, many studies have reported rapidly increasing concentrations of engineered TiO_2 nanoparticles in surface waters and wastewater treatment systems.[53-57] While extensive studies have been conducted to evaluate the toxicity of TiO_2 nanoparticles in nature, their photocatalytic effects on natural redox cycles remains elusive. Thus, investigating the roles of TiO_2 minerals in the photocatalytic oxidation of Mn^{2+} and formation of Mn oxides have important implications for both natural and engineered systems.

In this study, we show rapid photocatalytic oxidation of Mn^{2+}(aq) in the presence of rutile and anatase (representing natural minerals and emerging engineered nanoparticles of TiO_2,

respectively). The observed oxidation rates are comparable to or even faster than currently known biotic processes. In addition, we show that the direct formation of TMO with large tunnel size, such as todorokite (3 × 3) and romanechite (2 × 3), occurs via heterogeneous nucleation on the surface of TiO_2. Considering the abundance of natural semiconducting minerals, our findings confirm the importance of direct photocatalytic oxidation for explaining the oxidation of Mn(II) and the formation and diversity of Mn oxides in environmental systems. These results also point to the potential influence of environmental nanoparticles on natural element and electron cycles.

2. EXPERIMENTAL SECTION

Materials and reagents

Rutile and anatase TiO_2 phases were obtained from Alfa Aesar and were confirmed phase pure by X-ray diffraction (XRD). Their specific surface areas were measured by using Brunauer–Emmett–Teller (BET) method with N_2 gas adsorption (Autosorb-1-MP surface pore analyzer, Quantachrome Corp.) and are 5.5 m^2/g for rutile and 51.5 m^2/g for anatase. $MnCl_2$ stock solution was used to achieve 100 μM Mn^{2+}(aq) in artificial sea water (ASW) and deionized water (DI). ASW was prepared with 0.42 M NaCl, 0.025 M $MgSO_4$, 0.0091 M $CaCl_2$, 0.0089 M KCl, and 0.0024 M $NaHCO_3$.[58] The pH value of the ASW experiments remained at 7.9 ± 0.1 during the 8 h reaction without any adjustment. The pH value in the DI experiments was initially adjusted to 7.5 ± 0.1 and later adjusted every hour using dilute NaOH or HCl. Throughout the reaction, pH stayed within the range of 7.5–7.8. TiO_2 particles (0.1 g/L) were dispersed in 180 mL ASW or DI solutions and sonicated for 10 min before starting the photocatalytic reactions. To understand the effect of cations on the rapid photocatalytic oxidation and formation of tunnel structured Mn oxides, parallel experiments were conducted by reacting anatase particles in DI water containing

only one cation (0.42 M NaCl, 0.025 M $MgSO_4$, or 0.0091 M $CaCl_2$; consistent with their corresponding concentrations in ASW). The suspension pH was initially adjusted to 7.5 ± 0.1, was adjusted every hour, and stayed within the range of 7.5–7.8 throughout the experiments. These experiments are referred to as single cation experiments.

Photocatalytic oxidation of Mn^{2+}(aq)

Photochemical experiments were initiated by illuminating the prepared TiO_2 suspension by a 450 W Xe-arc lamp light source (Newport). The light passed through a 10 cm IR water filter to avoid the increase of temperature. We used a borosilicate reactor with a quartz window (1 inch diameter) facing the light. The reaction suspension was magnetically stirred over the entire reaction process. Aliquots (0.3 mL) of the suspension was taken every hour to analyze the concentration of oxidized Mn(III,IV) using the leucoberbelin blue (LBB, Sigma Aldrich) colorimetric method at 625 nm on an UV-vis spectrophotometer (Cary 60, Agilent). Calibration of the LBB method used $KMn^{VII}O_4$.[59] Because 1 mole of Mn(VII) oxidizes 5 moles of LBB, using the obtained calibration with Mn(VII), the amount of oxidized Mn^{2+} can be converted to Mn(III) or Mn(IV) equivalent by considering the oxidation state of Mn oxides. For example, 10 µM Mn(VII) is equivalent to 25 µM Mn(IV) or 50 µM Mn(III). Because it is difficult to accurately determine the oxidation state of the photochemically formed Mn oxides, we used Mn(III) equivalent calculation in order to compare the electron flow under different reaction conditions. All experiments were conducted in replicates.

Several control experiments were also conducted by using anatase in ASW, as this system showed the fastest oxidation among all tested conditions. (1) Dark experiments were conducted to differentiate photocatalytic oxidation (i.e., Mn(II) oxidation by photochemically generated holes)

153 from mineral surface catalyzed oxidation (i.e., heterogeneous Mn(II) oxidation by molecular
154 oxygen on mineral surface[60, 61]). (2) To determine the photoactive wavelength region of anatase
155 for Mn^{2+}(aq) oxidation, 400 or 550 nm optical cut-off filters (Newport) was used to cut-off lights
156 lower than 400 or 550 nm wavelength, respectively. The optical filter was placed between the light
157 source and reactor. Because anatase showed much higher oxidation rate as compared to rutile, the
158 control experiment was conducted using anatase to clearly distinguish the effective wavelength
159 region for the photocatalytic oxidation of Mn^{2+}(aq) by TiO_2. (3) Because TiO_2 can generate
160 superoxide upon photoexcitation[51], to confirm the contribution of superoxide for the oxidation of
161 Mn^{2+}(aq), scavenge experiments were conducted by adding 0.1 μM superoxide dismutase (SOD,
162 Sigma Aldirch) in the suspension to scavenge superoxide during the reaction. (4) A control
163 experiment was conducted in the absence of TiO_2, which showed no oxidation of Mn^{2+}(aq) in both
164 ASW and DI water conditions.
165
166 **Calculation of quantum yield**
167 Using the photocatalytic oxidation of Mn^{2+}(aq) by TiO_2 with 450 W Xe-arc lamp, we
168 calculated the quantum yield of rutile and anatase in ASW and DI. Multiplying irradiance of 450
169 W Xe-arc lamp (Figure S1), wavelength, unit conversion factor, absorbance at each wavelength
170 provides photons generated by a certain wavelength of 450 W Xe-arc lamp light. The exposed area
171 and penetration depth were multiplied in the reactor. To obtain total photons generated by anatase
172 or rutile by 450 W Xe-arc lamp, the calculated number of photons per wavelength using Equation
173 1 below were integrated from 250 to 800 nm (Table S4).
174
175 $$I\left(\frac{photons}{nms}\right) = I(W/cm^2nm) \times Wavelength(nm) \times 5.035 \times 10^{15}(photons/s \cdot nm) \times$$

$$Abs.(1/cm) \times penetration\ depth\ (cm) \times area\ (cm^2) \quad \text{Eq. (1)}$$

We used the concentration of Mn(III) equivalent, which represents successfully transferred holes from TiO_2. Because the Mn(III) equivalent concentration reached a plateau after 6 h for reaction in rutile suspension, we used the concentration at 6 h for the calculation of quantum yield.

$$QY\ (\%) = \frac{total\ oxidized\ Mn^{2+}(aq)\ for\ 6\ hrs}{total\ generated\ photons\ from\ TiO2\ for\ 6\ hrs} \times 100 \quad \text{Eq. (2)}$$

Using the obtained quantum yields of TiO_2 in ASW and DI, we obtained the estimated oxidation rates of Mn^{2+}(aq) by TiO_2 under natural sunlight exposure. The spectrum of natural sunlight (reference E-490-00; Figure S1) is obtained from the National Renewable Energy Laboratory. Using Equation 1 and integration, the total number of photons were obtained under sunlight irradiation. By multiplying the total number of photons and quantum yield, we obtained the estimated concentration of Mn(III) equivalent under sunlight irradiation in 6 h reaction. The oxidation rates under sunlight condition were estimated by linearizing the Mn(III) equivalent concentration between 0 and 6 h.

Characterization of heterogeneously nucleated Mn oxides on TiO_2

At each time point, the reacted suspension was syringe filtered (0.2 μm) and analyzed by LBB colorimetric method to quantify the amount of Mn oxides formed on the surface of TiO_2 (heterogeneous nucleation) and in the solution (homogeneous nucleation). To identify the phase and oxidation state of Mn oxides, reacted solids were collected at the end of photocatalysis by repeated centrifugation and rinsing by DI water, followed by freeze-drying. The dried samples

were analyzed using X-ray photoelectron spectroscopy (XPS), synchrotron XRD (SXRD), and synchrotron X-ray absorption spectroscopy (XAS), as detailed below.

Mn 3p XPS spectra, which show better sensitivity of Mn oxidation state than Mn 2p spectra[62, 63], was obtained on a K-alpha XPS system with monochromatic Al Kα radiation (1486.6 eV) (ThermoFisher Scientific). Although Mn 3s spectra are also useful for analyzing Mn oxidation state, the use of Mn 3s in this study was inappropriate due to its overlap with Mg 1s spectra for ASW samples. Energy calibration used C 1s spectra (284.6 eV). Fitting of Mn 3p spectra was conducted with allocations of Mn(II), Mn(III), and Mn(IV) at 47.8, 48.6, and 49.8 eV, respectively.

SXRD was measured at Beamline 17-BM-B (51.324 keV, $\lambda = 0.24157$ Å) at the Advanced Photon Source (APS, Lemont, IL, US). Mn K-edge X-ray absorption spectroscopy (XAS) analysis was conducted at Beamline 4-1 at the Stanford Synchrotron Radiation Lightsource (SSRL, Menlo Park, CA, US), Beamline 12-BM-B at APS, and Beamline 6-BM at National Synchrotron Light Source-II (NSLS-II, Upton, NY, US). Both XANES (X-ray absorption near edge structure) and EXAFS (extended X-ray absorption fine structure) data were collected. The monochromators were detuned by 40% to avoid higher order harmonics. Energy calibration used Mn foil. Multiple scans (2–6) were collected for each sample, averaged, and normalized for further analysis. Analysis of the Mn XANES spectra for each sample showed no evidence of photo-reduction under the X-ray beam.

XAS data analysis was performed using the programs SIXPACK [64] and Ifeffit [65]. Linear combination fitting (LCF) of the Mn XANES region was conducted to determine the relative percentage of Mn(II), Mn(III), and Mn(IV) species and the average oxidation state (AOS) following previous procedures.[66] To obtain statistically meaningful fitting results, all possible combinations were fitted with at most four references among 12 references in Table S1. Among

the obtained 781 different fitting results, 30 fitting results with the lowest R factors were selected and averaged to determine AOS and percentages of Mn valence states. Considering the typical 10% error for LCF, values lower than 5% were considered negligible. Details on the reference compounds and results for LCF analysis are in Table S1 and Table S2, respectively.

3. RESULTS AND DISCUSSION

Rapid photocatalytic oxidation of Mn^{2+}(aq)

In the presence of TiO_2 minerals, Mn^{2+}(aq) in both ASW and DI was quickly oxidized (Figure 1). The oxidation rates with different mineral phases and aqueous conditions (1.4 to 8.0 μM hr^{-1}) (Figure S2) are all comparable to or faster than those of biotic processes (~1.2 ± 0.2 μM hr^{-1}) [37, 67-69]. This finding suggests that photocatalytic oxidation of Mn^{2+}(aq) by TiO_2 is a feasible scenario in nature due to the abundance of natural semiconducting TiO_2 minerals[49] and increasing concentrations of anthropogenic nanoparticles in environmental systems.[53-57, 70] Indiscernible oxidation under dark condition indicates that both surface catalyzed oxidation of Mn^{2+}(aq) by TiO_2 (i.e., heterogeneous oxidation) and homogeneous oxidation of Mn^{2+}(aq) by dissolved oxygen in solutions are negligible compared to photocatalytic oxidation (Figure S3). Anatase shows about two times faster oxidation rates than rutile in both ASW and DI (Figure 1 and Figure S2). This difference might be from the higher specific surface area of anatase (51.5 m^2 g^{-1}) than rutile (5.5 m^2 g^{-1}). However, surface area normalized concentrations of Mn(III) equivalent showed that the oxidized Mn^{2+}(aq) is about an order of magnitude higher on the surface of rutile than that of anatase (Figure S4), indicating that the difference of surface area might not be the dominant controlling factor for the higher oxidation rates in anatase suspensions. This suggests that the higher oxidation rates in anatase suspensions is likely due to the higher photo-reactivity of anatase for redox

245 reaction.[51]

246 Except for the photocatalytic oxidation of Mn^{2+}(aq) in ASW-anatase suspension, all other
247 conditions showed slowed oxidation after 6 h (Figure 1). This is possibly due to the aggregation
248 of TiO_2 particles over time, attachment of TiO_2 particles to the photochemical reactor over time,
249 or the blockage of surface active sites by the nucleated Mn oxide.

250 Interestingly, both anatase and rutile show faster oxidation in ASW than that in DI. The
251 pH difference in ASW (7.9 ± 0.1) vs DI (7.5−7.8) might partially contribute to the difference in
252 oxidation rate. Our thermodynamic calculations using Visual MINTEQ showed no significant
253 changes in Mn(II) aqueous speciation in ASW and DI conditions (data not shown). Thus we
254 hypothesize that the difference in oxidation rate is likely due to the presence of large amount of
255 cations (e.g., Na^+, K^+, Ca^{2+}, and Mg^{2+}) in ASW that can facilitate the oxidation of Mn^{2+}(aq) and
256 formation of Mn oxides.[44, 71, 72] This is confirmed by the single cation experiments using solutions
257 containing only one cation at the corresponding ASW concentration (i.e. 0.091 M Ca^{2+}, 0.025 M
258 Mg^{2+}, or 0.420 M Na^+). The photocatalytic oxidation of Mn^{2+}(aq) by anatase in single cation
259 experiments with Mg^{2+} (~8.2 μM hr^{-1}) or Ca^{2+} (~6.6 μM hr^{-1}) (Figure S5A) shows oxidation rates
260 similar to ASW systems (8.0 ± 0.6 μM hr^{-1}). Interestingly, in single cation experiment with Na^+,
261 we observed the fastest oxidation rate (~11.3 μM hr^{-1}) among all the tested experimental conditions.
262 This indicates the profound role of cations in the photocatalytic oxidation and formation of Mn
263 oxides on TiO_2 particles. The different phase of Mn oxides formed in the single cation experiments
264 (Figure 2) also supports the influence of cations on the Mn oxide structure, as discussed below.
265 Since this study focuses on the photocatalytic oxidation of Mn^{2+}(aq) by TiO_2 nanoparticles and
266 consequent formation of TMO, detailed mechanistic understanding of the roles of cations and
267 anions in the oxidation rate is beyond the scope of this work. Further studies are warranted to

explain the photocatalytic reactions of TiO_2 under varied aqueous conditions.

Formation of large tunnel structured Mn oxides

The rapid photocatalytic oxidation of Mn^{2+}(aq) led to the formation of Mn oxides on the surface of TiO_2. Mn(III,IV) species were not detected by the LBB method in the filtrate of the reaction suspension at all times (Figure S3). Considering that homogeneously nucleated Mn oxides are 6–30 nm in diameter,[73] which can readily pass through 0.22 μm filters, the absence of Mn oxides at all times in the filtrate likely indicates the heterogeneous nucleation of Mn oxides on the surface of TiO_2 instead of homogeneous nucleation of Mn oxide nanoparticles in solution.

Interestingly, we also observed the formation of TMO with large tunnels under all reaction conditions (Figure 2A and 2B). Considering that LMO (such as δ-MnO_2 and birnessite) are the typical phase of homogeneously nucleated Mn oxides through biotic/abiotic processes,[36, 37, 74] the formation of TMO in our system also strongly suggest the heterogeneous nucleation of Mn oxides on TiO_2 through photocatalytic oxidation of Mn^{2+}(aq). SXRD analyses of the 8-h samples in anatase-ASW and rutile-ASW suspensions (Figure 2A and 2B) indicate the formation of todorokite (3 × 3 tunnel size) with diffraction peaks at 9.6 and 4.8 Å d-spacing. Although the large d-spacing can also occur from buserite, which has similar structure to birnessite but with a 10 Å interlayer spacing, the interlayer of buserite (10 Å interlayer spacing) collapses to birnessite (~7 Å interlayer spacing) upon dehydration due to weakly bound interlayer H_2O in the interlayer positions [40, 75, 76]. Thus, the large d-spacing of our fully dehydrated samples (obtained under vacuum freeze-drying for 1 day) indicates that the nucleated Mn oxide on TiO_2 in ASW is todorokite. In anatase-DI and rutile-DI suspensions, SXRD showed only weak diffraction at ~4.9 Å, which occurs from all large tunnel structured Mn oxides (3 × 3, 3 × 2, and 2 × 2 tunnel size)

(Figure 2A and 2B). To further identify the structure of the Mn oxides formed in DI experiments, we conducted Mn K-edge EXAFS analysis and identified the formation of romanechite-like structure (2 × 3 tunnel size) (Figure 2C). Specifically, the region between 7 and 10 $Å^{-1}$ in k space is a well-known "indicator region" for Mn oxide structure.[77, 78] In this region, Mn oxides formed in DI experiments have two peaks at ~8.7 and 9.2 $Å^{-1}$, consistent with the spectra of romanechite and is distinctively different from the spectra of todorokite formed in ASW (Figure 2C) and hollandite (2 × 2 tunnel size, Figure S6). These results indicate an important observation that heterogeneous nucleation on natural mineral surfaces can facilitate the direct formation of TMO with large tunnels. Specifically, todorokite is one of the most ubiquitous natural Mn oxides, yet its formation in low temperature environments is poorly constrained due to the difficulties in synthesizing this phase in laboratory settings through the transformation of LMO to TMO. Most previous studies employed reaction conditions that do not represent circumneutral low temperature environments, such as the transformation of Mg^{2+}-intercalated LMO using reflux or hydrothermal reactions at temperatures above 100 °C [41, 42, 44]. Our findings suggest that photocatalytically promoted direct heterogeneous nucleation of Mn oxides on natural minerals might be an alternative pathway for the formation of todorokite. In our system, the underlying TiO_2 minerals serve as templates to facilitate the heterogeneous nucleation of Mn oxides, which might have lowered the structural strain and formation energy for TMO. In addition, this observation also suggest a new phase selection pathway (i.e., formation of TMO via heterogeneous nucleation of Mn oxides) as compared to previously observed common formation of LMO induced by homogeneous nucleation of Mn oxides upon Mn(II) oxidation [36, 37, 74]. Future microscopic studies are warranted to further investigate the mechanism, structure, and morphology of the heterogeneously nucleated TMO, such as the use of high resolution transmission electron microscopy (HRTEM) to analyze the

micromorphology, potential presence of poorly crystalline LMO, and relationships with the mineral substrate.

Prior studies showed the importance of Mn(III) and large cations (e.g., Mg^{2+}, Ca^{2+}) on the transformation of LMO to TMO with large tunnels.[41, 42, 44, 71, 79-81] In our system, Mn oxides produced from all tested conditions contain considerable amount of structural Mn(III) as probed by XPS and XANES (Figures 3 and S7 and Table S2), which is necessary to form TMO. The results show that Mn oxides formed in the presence of anatase have more Mn(III) in the structure than those formed in the rutile system. Specifically, Mn oxides formed in the anatase-ASW suspension show a higher content of Mn(III) than those formed in rutile-ASW suspension based on the fitting of Mn 3p XPS spectra. This feature can be readily observed from the lower peak position of Mn 3p spectra of anatase-ASW sample than that of rutile-ASW (Figure 3). It is worth noting that obtaining accurate ratios of Mn(III/IV) using XPS and XANES is challenging, due to errors resulting from data fitting and the intrinsic differences in these techniques (i.e., XANES obtains bulk information whereas XPS detects near-surface structure).[82] Nonetheless, linear combination fitting (LCF) of Mn XANES spectra also shows that the nucleated Mn oxides in anatase-ASW suspension have higher content of Mn(III) than those formed in rutile-ASW suspension (Figure S7 and Table S2). Despite the difference in Mn(III) content, the Mn oxides formed in both anatase and rutile suspensions have the same structure (i.e., todorokite in ASW and romanechite in DI).

Interestingly, although the Mn oxides formed in rutile-DI and rutile-ASW suspensions show similar Mn(III) contents, they have different structures. This suggests that the observed different tunnel sizes between ASW (todorokite, 3 × 3) and DI (romanechite, 2 × 3) systems mainly resulted from the presence of cations in ASW, not from the difference in Mn(III) content. Indeed,

results from the single cation experiments (DI water containing only Mg^{2+}, Ca^{2+}, or Na^{+} at their corresponding ASW concentrations) support the relative importance of different cations in ASW in mediating the TMO formation. Among all anatase-single cation experiments (Mg^{2+}, Ca^{2+}, or Na^{+}), the Mn oxides produced in Mg^{2+} experiment shows the most relevant structure to Mn oxide formed in anatase-ASW suspension, as revealed by linear combination fitting of Mn K-edge EXAFS spectra (Figure S5B and Table S3). This observation is consistent with prior studies showing the necessity of Mg^{2+} intercalation into LMO to enable the hydrothermal transformation of LMO to todorokite.[41, 42, 44, 71, 79-81] Ca^{2+} also likely contributes to the formation of todorokite, based on the similarity in EXAFS indicator region for Mn oxides produced in the Ca^{2+} single cation experiment as compared to todorokite spectra. We can rule out the contribution of Na^{+} to the formation of todorokite in ASW based on the formation of romanechite-like structure (Figure S5B and Table S3). Overall, results from the single cation experiments indicate that, regardless of the mineral catalyst, large cations (e.g., Mg^{2+} and Ca^{2+}) in ASW are main factors responsible for the structural difference during heterogeneous nucleation of Mn oxide on TiO_2 in ASW vs DI water. Detailed mechanistic investigation on the effects of cations on the oxidation of Mn^{2+}(aq) and the structure of nucleated Mn oxides is beyond the scope of this study and warrants further work.

Reaction mechanisms

Our mechanistic study further reveals that photo-excitation of TiO_2 enables the rapid oxidation of Mn^{2+}(aq) by the generation of (1) holes at valence band (direct oxidation) as well as (2) superoxide (indirect oxidation), as explained below.

For Mn^{2+} oxidation by the generated holes at valence band (direct oxidation): we conducted parallel experiments using different optical cut-off filters in the anatase-ASW suspension (which

shows the fastest oxidation) to elucidate the mechanisms of photocatalytic oxidation of Mn^{2+}(aq). We found that the rapid oxidation occurs via the photo-excitation of electrons and holes from anatase (Figure 4A). Specifically, in the presence of 400 nm cutoff optical filter (which cuts off wavelengths at < 400 nm), the oxidation of Mn^{2+}(aq) is significantly suppressed. This indicates that the rapid oxidation of Mn^{2+}(aq) occurs mainly from the photo-excitation in UVA region (320–400 nm), which matches the band-gap of anatase (3.2 eV). With 550 nm cutoff optical filter (cutoff < 550 nm), little oxidation of Mn^{2+}(aq) occurred (Figure 4A) due to completely suppressed photo-excitation.

For Mn^{2+} oxidation by superoxide (indirect oxidation): TiO_2 is capable of generating superoxide because its conduction band located at lower reduction potential than that of oxygen (O_2) / superoxide (O_2^-) (Figure 4C). Prior studies show that superoxide can readily oxidize Mn^{2+}(aq) while hydroxyl radical and hydrogen peroxide cannot.[36, 68] In our control experiment with superoxide scavenger SOD, both the extent and rate of Mn^{2+} oxidation decreased slightly, indicating the contribution of superoxide (produced from photocatalysis) in the indirect oxidation of Mn^{2+} (Figure 4D). However, considering the small decrease in Mn^{2+} oxidation caused by SOD as compared to those caused by the cut-off filters (Figure 4D), direct electron transfer from Mn^{2+}(aq) to photo-generated holes is still considered as the dominant oxidation mechanism. In addition, the occurrence of superoxide and the observed direct electron transfer from Mn^{2+}(aq) to a hole in the valence band of TiO_2 indicate that oxygen is the electron acceptor which suppresses the recombination between the photo-generated electrons and holes.

4. ENVIRONMENTAL IMPLICATIONS

To assess the relative importance of this photocatalytic oxidation process in nature, we

compare the oxidation rates obtained from this study to previously observed representative oxidation rates. Using the experimentally obtained photocatalytic oxidation rate, spectra of 450 W Xe-arc lamp and natural sunlight, and by calculating the quantum yields (Table S4 and Figure S1), we extrapolated the laboratory oxidation rates (450 W Xe lamp) to natural sunlight condition. We note that oxidation processes are controlled by many factors (e.g., pH, initial Mn^{2+} concentration, dissolved oxygen concentration, temperature, solution chemistry, etc.). Nevertheless, the comparison of oxidation rates in Figure 5 provides an initial visual comparison of the relative importance of different oxidation pathways. As shown in the experimental and calculated results with TiO_2 (Figure 5), the calculated oxidation rates under natural sunlight condition are comparable to or faster than that of microbial oxidation processes. Natural semiconducting minerals occur in a wide variety of natural systems. Besides TiO_2 and Fe oxides, other natural semiconducting minerals such as ZnS, FeS, etc. may also serve to photocatalytically oxidize Mn^{2+}(aq) (or other species). However, different mineral catalysts might insert different photocatalytic oxidation mechanisms.

Our findings highlight the importance of a previously overlooked pathway: the photocatalytic oxidation of Mn^{2+}(aq), which might significantly contribute to Mn redox balance in nature. Meanwhile, the observed fast photocatalytic oxidation of Mn^{2+}(aq) raises the potential concern of semiconducting nanoparticles in influencing environmental redox systems. Through anthropogenic activities, TiO_2 and other semiconducting nanoparticles are released to a wide variety of environmental systems. Based on the observed rapid photocatalytic oxidation by TiO_2 in this study, accumulation of the photo-active nanoparticles in nature might influence the balance of Mn redox. As an example of such anthropogenic impact, a recent study reported the occurrence of elevated concentration of dissolved Mn from the reduction of natural Mn oxides by emerging

406 contaminants and fertilizers in US groundwater, impacting more than 2.6 million people
407 consuming the Mn contaminated groundwater.[83] In contrary, the increase of semiconducting
408 nanoparticles in environmental systems might promote the oxidation of Mn^{2+}(aq) and formation
409 of Mn oxides, thus affecting the fate and transport of many other elements and contaminants. For
410 example, Mn oxides are known to oxidize the more mobile As(III) into the less mobile As(V)
411 species, whereas they can oxidize the less mobile Cr(III) species into more mobile and toxic Cr(VI)
412 species [3, 84-87].

413 This study also demonstrates the direct formation of todorokite through heterogeneous
414 nucleation on the surface of natural minerals under circumneutral conditions in the presence of
415 large cations. To our knowledge, this is the first observation of direct formation of todorokite in
416 laboratory settings under environmentally relevant conditions (Table S5). The results suggest that
417 environmental conditions, such as solution chemistry, mineral surface, etc., may govern the
418 polymorphism of natural Mn oxides, and the latter in turn might potentially serve as a fingerprint
419 to trace environmental conditions at the time of Mn oxide formation.

ASSOCIATED CONTENT

Supporting Information.

Reference compounds used for XANES LCF (Table S1), results of XANES LCF (Table S2), results of EXAFS LCF (Table S3), results on the calculation of quantum yields (Table S4), experimental conditions of the representative biotic/abiotic oxidation of Mn^{2+}(aq) to Mn(IV) (Table S5), irradiance spectra of Xe-arc lamp and sunlight (Figure S1), linear initial oxidation rates (Figure S2), results of dark and filtrate control test (Figure S3), surface area normalized concentration of Mn(III) equivalent (Figure S4), effect of cations on the photocatalytic oxidation rates (Figure S5), EXAFS spectra of the nucleated Mn oxides and references (Figure S6), and XANES data (Figure S7).

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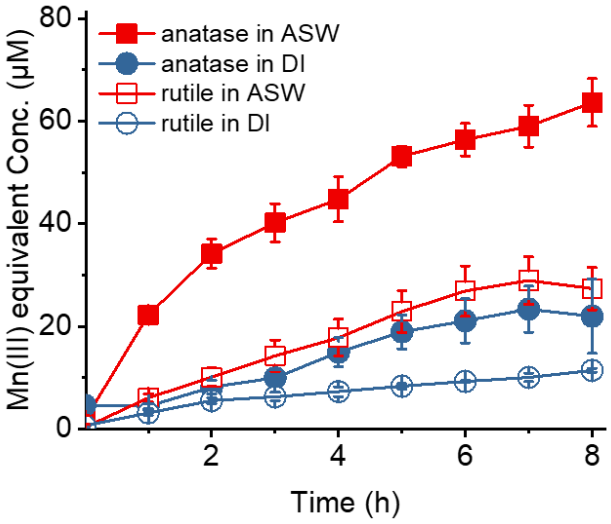

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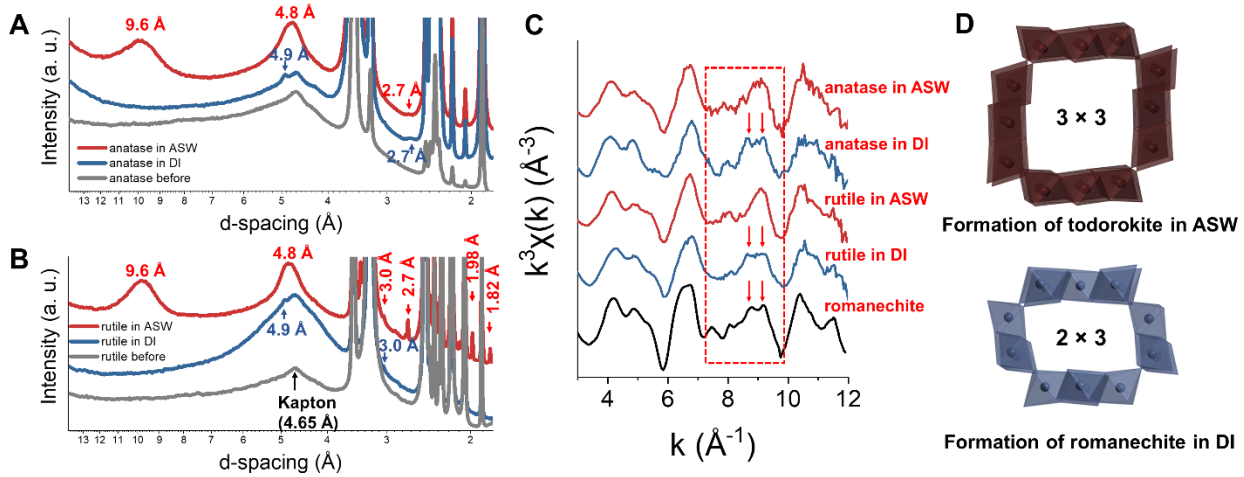
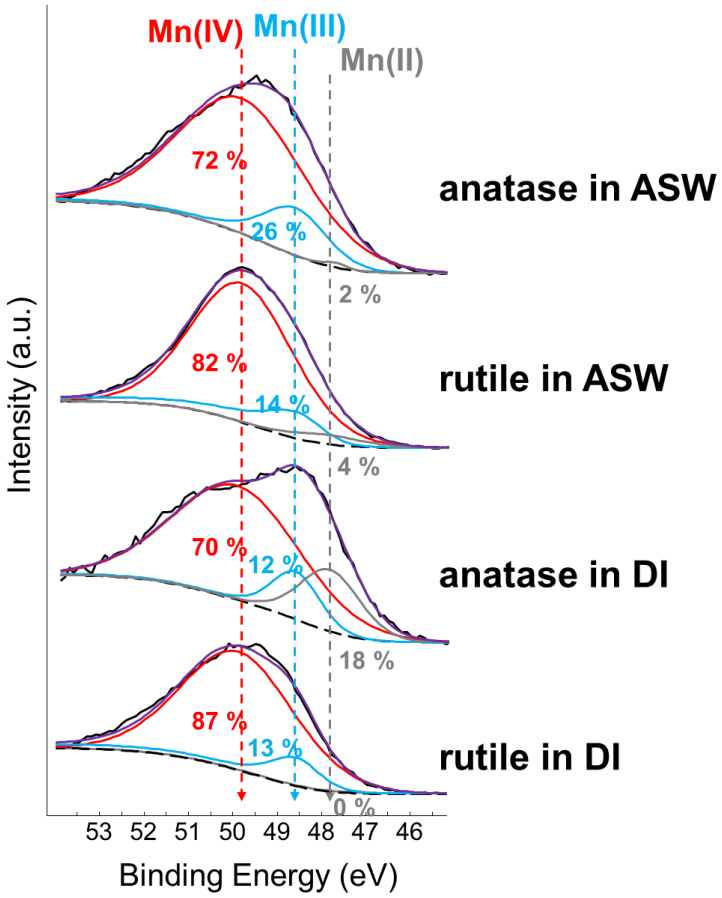
689 **Figure 1.** Rapid photocatalytic oxidation of Mn^{2+}(aq) by anatase and rutile in artificial seawater

690 (ASW) and deionized (DI) water.



691

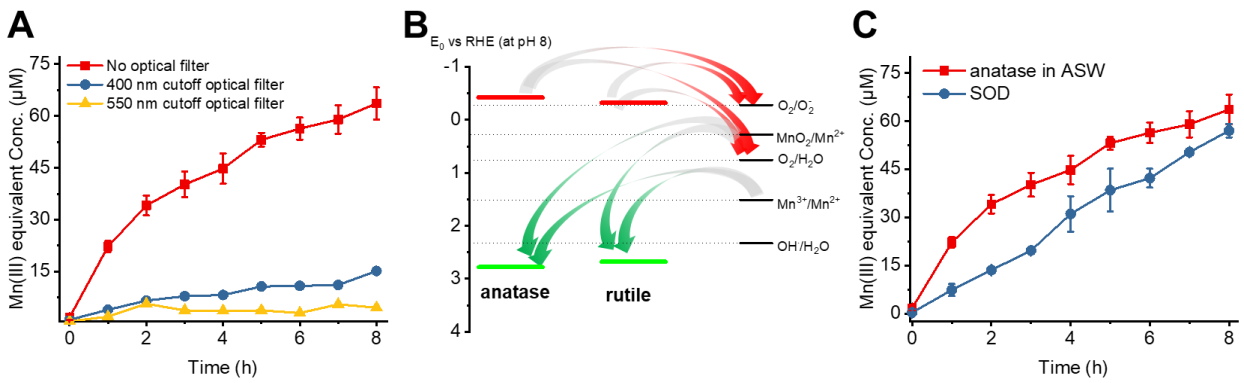
692 **Figure 2.** Formation of large tunnel structured Mn oxides on the surface of TiO_2. (A and B)
693 Synchrotron XRD analyses show that the nucleated Mn oxide on the surface of both anatase and
694 rutile in ASW is todorokite (3 × 3 tunnel structure). In DI system, weak diffraction occurred at 4.9
695 Å. Because several tunnel structured Mn oxides (romanechite (2 × 3), hollandite (2 × 2), and
696 todorokite (3 × 3)) have diffraction peak at this position, it is challenging to identify the phase of
697 Mn oxide using SXRD alone. (C) EXAFS analysis confirms that Mn oxides formed in DI system
698 are structurally similar to romanechite. (D) Schematic structures of todorokite (3 × 3 tunnel size)
699 and romanechite (2 × 3 tunnel size).



700

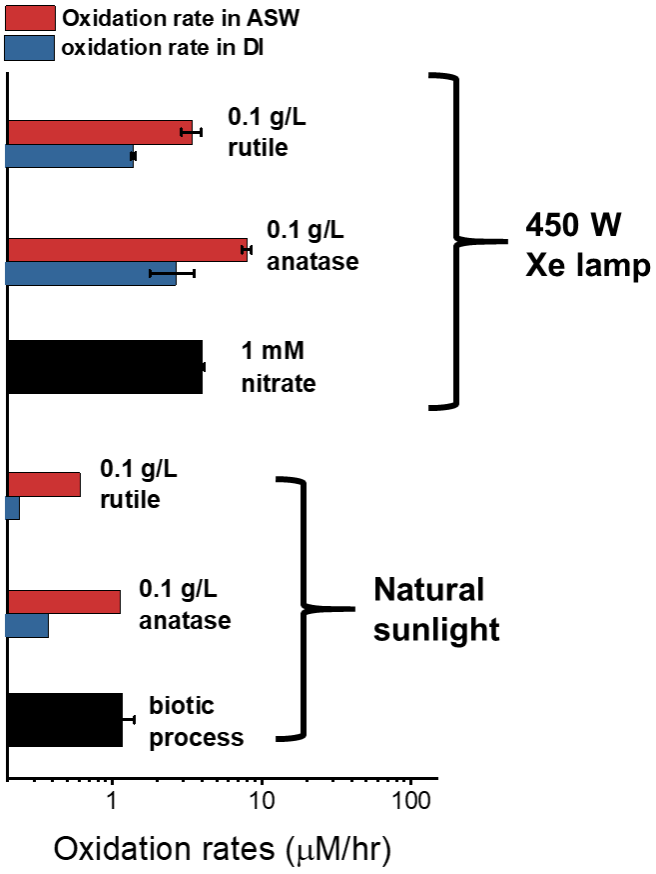
701

702 **Figure 3.** Mn 3p XPS spectra of Mn oxides formed on the surface of TiO_2. Fitting results show
703 that all the Mn oxides formed under varied conditions have Mn(III). Mn oxides formed in the
704 presence of anatase have more Mn(III) than those formed in the rutile system.



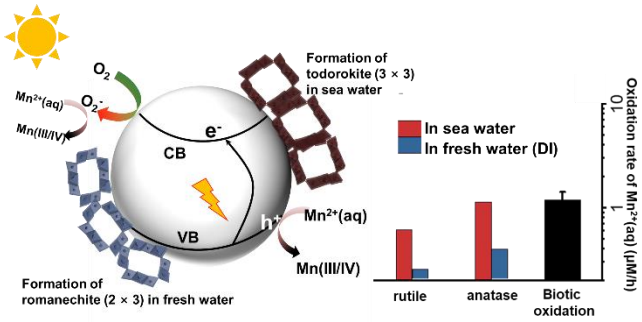
705

706 **Figure 4.** Mechanistic understanding of the photocatalytic oxidation of Mn^{2+}(aq) in the presence
707 of TiO_2. (A) Control tests with 400 or 550 nm optical cut off filters suggest that the rapid
708 photocatalytic oxidation of Mn^{2+}(aq) with TiO_2 occurs from the photo-excited electron-hole pair
709 mainly at wavelength < 400 nm. (B) The positions of valence and conduction bands show that the
710 photocatalytic oxidation of Mn^{2+}(aq) with TiO_2 is thermodynamically feasible via both direct
711 electron transfer and indirect reaction of Mn^{2+}(aq) with photocatalytically generated superoxide.
712 (C) The decrease of oxidation rate in the presence of superoxide dismutase (SOD; a superoxide
713 scavenger) indicates that both superoxide and photo-excited holes contribute to the rapid oxidation
714 of Mn^{2+}(aq) in the presence of TiO_2 and light. In addition, the occurrence of superoxide indicates
715 that oxygen is the electron acceptor, which enables the rapid direct electron transfer from Mn^{2+}(aq)
716 to a hole in the valence band of TiO_2 by blocking the recombination of electron-hole pair.



717

718 **Figure 5.** Comparison of the photocatalytic oxidation rates of Mn^{2+}(aq) in this study (using Xe
719 lamp or calculated to natural sunlight condition) with previously reported oxidation rates in the
720 presence of nitrate or via biotic processes.



721

722 **TOC graphic**