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# Tuning the Surface Chemistry of Metal Organic Framework Nodes: Proton Topology of the Metal-oxide-like $Zr_6$ nodes of UiO-66 and NU-1000

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**KEYWORDS.** Metal-organic frameworks,  $Zr_6$  node, proton topology, UiO-66, NU-1000

**ABSTRACT:** Some metal organic frameworks (MOFs) incorporate nodes that are nanoscale metal oxides, and the hydroxy-containing functional groups on them provide opportunities for introducing catalytic sites with precisely defined structures. Investigations have been done to understand the structures of these groups on nodes and node vacancies, because, in prospect, atomic-scale modulation of the composition, areal density, and/or siting of the groups would open up possibilities for exquisite tuning of the siting and performance of subsequently anchored catalytic units (e.g., single metal ions, pairs of metal ions, or well defined metal-ion-containing clusters). We have combined infrared (IR) spectroscopy and density functional theory (DFT) to demonstrate tuning of these sites, namely, hydrogen-bonded OH/OH<sub>2</sub> groups on the  $Zr_6$  nodes of the MOFs UiO-66 and NU-1000 via the intermediacy of node methoxy (or ethoxy) groups formed from methanol (or ethanol). Methoxy (or ethoxy) groups on node vacancy sites are converted to a structure incorporating one vacant Zr site and one terminal OH group per face by reaction with water. Our results highlight how the combination of DFT and IR spectroscopy facilitate the determination of the identity and chemistry of the functional groups on MOF node vacancies and defect sites.

## INTRODUCTION

The nodes of some metal-organic frameworks (MOFs) are essentially small, uniform metal oxide clusters, exemplified by the  $Zr_6$  nodes of UiO-66,<sup>1</sup> UiO-67,<sup>1</sup> and NU-1000.<sup>2</sup> The strong coordination bonds between these nodes and carboxylate oxygen atoms in the linkers help to stabilize these MOFs, which can be used at high temperatures (in the range of 300–400 °C<sup>3</sup>) and are therefore attractive in prospect for application as catalysts and catalyst supports. The crystal structures of these MOFs have been determined by X-ray diffraction crystallography,<sup>1,2</sup> but this technique is not sufficient to determine the identities and placement of small *non*-structural ligands (*i.e.*, ligands other than linkers) such as hydroxyl and aqua groups on the nodes.

The nonstructural ligands on the nodes of UiO-66 and UiO-67 are associated with missing linkers<sup>4,5</sup> (and, in some cases, nodes<sup>6-8</sup>). The nodes of UiO-type MOFs are nominally (or ideally) twelve-linker-connected, but in real samples the num-

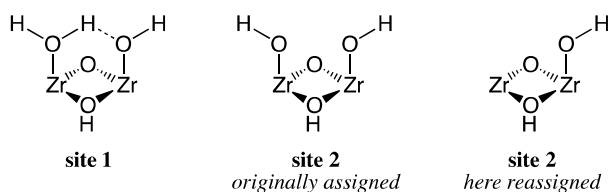
ber is less. The hexa-metal(IV) nodes of NU-1000, in contrast, are inherently eight- rather than twelve-connected, thus leaving numerous sites for binding of nonstructural ligands, even in the absence of structural defects.<sup>5</sup> The node-ligated aqua and hydroxo species are Brønsted acids<sup>9,10</sup> comparable to those on the surfaces of metal oxides and zeolites, and the corresponding MOFs are therefore comparable to these materials as catalysts and catalyst supports, albeit with potentially different properties owing to the nanoscale of the nodes and/or local confinement by the framework.

Because MOFs in prospect have structures that can be tuned more precisely than those of metal oxides and even zeolites, they offer excellent opportunities to create unique, uniform surface sites that exhibit new, controllable reactivities and catalytic properties—and, correspondingly, they offer additional opportunities for the anchoring of unique catalytic sites such as metal complexes.<sup>4,5,11-13</sup> Therefore, there is a strong incentive for investigating the precise molecular structures of MOF nodes and their vacancies, and for tailoring them when

possible for specific functions, a subject that is still at an early stage of investigation. An important recent contribution in this respect is the work of Ling and Slater,<sup>14</sup> who investigated the dynamic acidity and interconversion of alternative defect terminations of UiO-66 in aqueous solution, extending the understanding<sup>15</sup> with respect to the proton topology of such node faces in water.

Using a combination of infrared (IR) spectroscopy and density functional theory (DFT) calculations,<sup>16</sup> we previously inferred that the NU-1000 nodes and their proton topology are well represented as  $[\text{Zr}_6(\mu_3\text{-O})_4(\mu_3\text{-OH})_4(\text{OH})_4(\text{OH}_2)_4]^{8+}$ , with each structural vacancy occupied by a combination of terminal OH and OH<sub>2</sub> groups at sites referred to as **site 1** (Scheme 1). This same terminating functionality was also identified at the missing linker sites of UiO-66 and UiO-67 nodes.<sup>4</sup> Moreover, a second topology was observed for the nodes of UiO-66 and designated in the earlier work<sup>4,5</sup> as **site 2**. The topology for **site 2** was originally assigned as a node face having a pair of terminal OH groups, as illustrated in Scheme 1. The chemistries of these two sites were investigated in experiments with iridium carbonyl complexes as probes, and the chemisorbed iridium carbonyls were characterized by extended X-ray absorption fine structure (EXAFS) and IR spectra, with the latter demonstrating different electron-donor properties of the two distinct sites.<sup>4,5</sup>

An open challenge is to tune those groups terminating defects or vacancies on MOF node faces. Methods for interconverting such groups on individual faces are lacking and would be of particular value to the extent that they might modulate the properties of nodes and their reactivities for applications in catalysis and adsorption.



**Scheme 1.** Simplified depictions of the Zr<sub>6</sub> node-face topologies, **site 1** and **site 2**, the latter both as originally assigned and as reassigned here in light of new results, for the MOFs UiO-66 and NU-1000. Note that the originally assigned **site 2** was for a Zr<sub>6</sub> node with the otherwise “missing” proton from **site 1** moved to another face of the core (*i.e.*, all of the sites in this scheme have zero net charge and spin).

Carboxylate-containing functional groups<sup>17,18</sup> or, alternatively, phosphonate groups<sup>19</sup> have been reported to reactively replace the hydroxyl groups on the Zr<sub>6</sub> nodes through a solvent-assisted ligand incorporation (SALI) method. We now report how to incorporate methoxy ligands on the node surfaces and through their intermediacy to convert MOF node **site 1** into the alternative MOF node **site 2**. Importantly, we here reassign the nature of **site 2** to involve only a single terminating OH group together with a vacant Zr coordination site. Our results demonstrate a facial ligand exchange on the vacancies of these metal-oxide-like MOF nodes and indicate a subtlety in the acid-base properties of the nodes akin to that noted by Ling and Slater.<sup>14</sup>

## EXPERIMENTAL AND COMPUTATIONAL METHODS

**Synthesis of UiO-66.** ZrCl<sub>4</sub> (99.99%, 0.125 g, 0.54 mmol) and 1 mL of concentrated HCl modulator (37 wt%, Sigma-Aldrich) were dissolved in 15 mL of dimethylformamide (DMF, 99.8%, Sigma-Aldrich) in an 8-dram vial by using ultrasound for 5 min. The linker precursor benzene-1,4-dicarboxylic acid (0.123 g, 0.75 mmol, ≥98.0%, Sigma-Aldrich) was dissolved in the solution with ultrasound applied for about 15 min. The vials were stored at 353 K for 24 h. The white precipitates that formed were isolated by centrifugation after cooling of the samples to room temperature. The solids were washed with DMF three times to remove unreacted precursors and with acetone six times to remove DMF, and the powder was then dried at room temperature.

**Post treatment of UiO-66 in Methanol.** UiO-66 (20 mg) was placed in 2 mL of methanol (≥99.9%, Sigma-Aldrich, for HPLC) at 298 K and held for 48 h. The solution was decanted and replaced three times over the course of the experiment. The powder was dried in air at 298 K for 30 min. The residual methanol was then removed by treatment under dynamic vacuum overnight at 423 K.

**Synthesis of NU-1000.** A sample consisting of 97 mg of ZrOCl<sub>2</sub>·8H<sub>2</sub>O (0.30 mmol, 98.0%, Sigma-Aldrich) and 2700 mg (22 mmol) of benzoic acid (≥99.5%, Sigma-Aldrich) was mixed with 8 mL of DMF in a 6-dram vial and ultrasonically dissolved. The clear solution was incubated at 353 K for 1 h. After cooling of the solution to room temperature, 40 mg (0.06 mmol) of 1,3,6,8-tetrakis(*p*-benzoic acid)pyrene (H<sub>4</sub>TBAPy<sup>20</sup>) was added, and the mixture was sonicated for 20 min. The yellow suspension was held at 353 K for 24 h. After cooling of the sample to room temperature, a yellow polycrystalline material was isolated by filtration (35 mg of activated material, 54% yield) and washed with DMF and then activated with HCl. The solid was subsequently washed twice with DMF and six times with acetone. The powder was dried at room temperature.

**Post treatment of NU-1000 in Methanol.** NU-1000 (20 mg) was placed in 2.0 mL of methanol (≥99.9%, Sigma-Aldrich, for HPLC) at 343 K and held for 48 h. The solution was decanted and replaced three times over the course of the experiment. The methanol was then removed by treatment under dynamic vacuum at 298 K, and the sample was activated under dynamic vacuum at 423 K for 12 h.

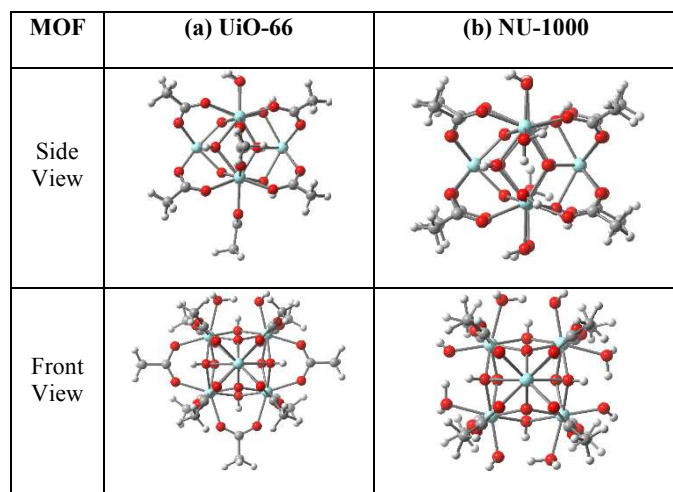
**IR Spectroscopy.** A Bruker IFS 66v/S spectrometer with a spectral resolution of 2 cm<sup>-1</sup> was used to collect transmission IR spectra of power samples. Approximately 30 mg of solid sample in an argon-filled glove box was pressed into a thin wafer and loaded into a cell that served as a flow reactor (In-situ Research Institute, South Bend, IN). The cell was sealed and connected to a flow system that allowed recording of spectra while reactant gases flowed through the cell at reaction temperature. Each spectrum is the average of 64 scans.

**Isotopic Exchange Experiments.** UiO-66 (20 mg) was placed in 2.0 mL of methanol-D<sub>4</sub> (99.5%, Cambridge Isotope Laboratories) and held at 293 K for 48 h. The solution was decanted and replaced three times over the course of the experiment. The methanol-D<sub>4</sub> was removed by evacuation by treatment under dynamic vacuum overnight at 423 K.

**Models and Computational Details.** *Cluster Model:* NU-1000. Calculations were performed starting from the mix-S topology of the Zr<sub>6</sub>-NU-1000 node,  $[\text{Zr}_6(\mu_3\text{-O})_4(\mu_3\text{-$

OH)<sub>4</sub>(OH)<sub>4</sub>(OH<sub>2</sub>)<sub>4</sub>]<sup>8+</sup>, reported by Planas et al.<sup>16</sup> The Zr<sub>6</sub> node of NU-1000 was modeled as a finite cluster extracted from an optimized periodic unit cell to allow a detailed investigation of the plausible new topologies of NU-1000. The UiO-66/67 cluster model was obtained from Yang et al.,<sup>4</sup> with the organic linkers kept as acetate (CH<sub>3</sub>CO<sub>2</sub><sup>-</sup>) groups as in our previous studies (Figure 1).<sup>4,16</sup> Moreover, CM5 partial atomic charges<sup>21</sup> were computed for **site 1**, **site 2**, and other relevant structures (Table S2).

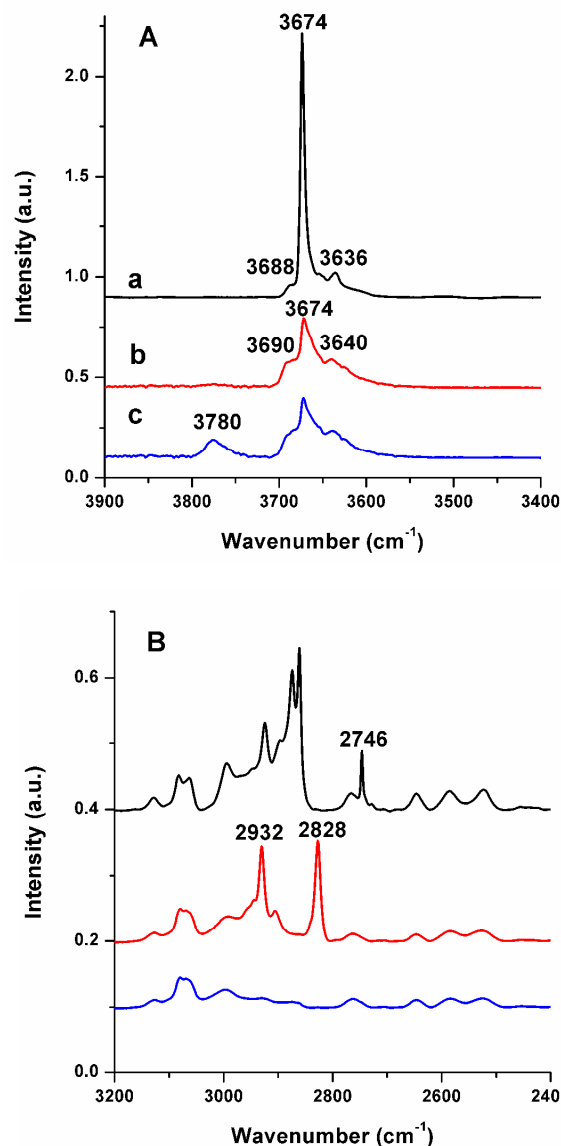
**Kohn-Sham Density Functional Calculations.** Clusters were optimized using the M06-L<sup>22</sup> density functional as implemented in the *Gaussian 09* software package,<sup>23</sup> together with the def2-TZVP<sup>24,25</sup> basis set for C, H, O, and Zr atoms. Associated effective core potentials were used on Zr atoms. In all cases, the positions of all atoms were optimized, except for the C and H atoms of the capping acetate groups. The nature of all stationary points was verified by analytical computation of vibrational frequencies, which were also used to compute 298.15 and 343 K thermochemical quantities with all frequencies below 50 cm<sup>-1</sup> replaced by 50 cm<sup>-1</sup> when computing free energies. The SMD continuum solvation model was employed to account for methanol solvation effects when comparing calculated results with those of experiments done in bulk solution,<sup>26</sup> except for water and methanol themselves as solutes, for which experimental values/estimates were employed. For details, see the Supporting Information, SI.



**Figure 1.** Schematic representation of acetate cluster model used to calculate various topologies for (a) UiO-66 and (b) NU-1000. Color code: Carbon, gray; hydrogen, white; zirconium, cyan; oxygen, red.

## RESULTS

The IR spectrum of a UiO-66 sample prepared in a synthesis modulated by HCl<sup>4,27</sup> and washed with acetone is shown in Figure 2. The band at 2746 cm<sup>-1</sup> has been assigned to the hydrogen-bonded OH/OH<sub>2</sub> pairs of **site 1** on the Zr<sub>6</sub> node; the sharp band at 3674 cm<sup>-1</sup>, with two weak shoulders, at 3688 and 3636 cm<sup>-1</sup>, derives from a combination of the non-hydrogen-bonded OH groups of **site 1** together with the (perforce, also non-hydrogen-bonded) μ<sub>3</sub>-OH bands,<sup>4,16</sup> with the majority of the intensity of the 3674 cm<sup>-1</sup> band coming from the terminal groups (*vide infra*).

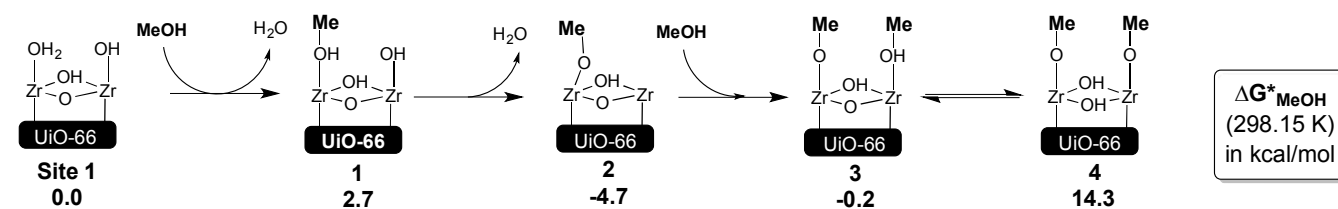


**Figure 2.** IR spectra in the (A) 3400–3900 cm<sup>-1</sup> region and (B) 2400–3200 cm<sup>-1</sup> region characterizing (a) UiO-66 and (b) UiO-66 after treatment in liquid methanol at 298 K for 48 h and (c) UiO-66 after treatment in water vapor/helium at 393 K for 30 min. The UiO-66 samples before and after treatment with liquid methanol or water vapor were activated at 393 K or 423 K, respectively, by treatment under high vacuum for 12 h.

In the first step to generate **site 2** in this work, UiO-66 (modulated by HCl and washed with acetone) was treated in liquid methanol for 48 h at 298 K, followed by decanting of the methanol and drying of the sample at 298 K for 30 min, followed by activation under high vacuum at 423 K for 12 h to remove residual methanol. Analysis indicated that **site 1** was completely removed from the nodes after the treatment, as evidenced by the disappearance of the 2746 cm<sup>-1</sup> band and a marked decrease in intensity of the 3674 cm<sup>-1</sup> band (that component associated with the terminating groups, *vide supra*), with a concomitant growth of new bands at 2932 and 2828 cm<sup>-1</sup> (Figure 2). These two new bands are consistent with the C–H stretches of methoxy groups.

It has correspondingly been reported that methanol reacts with the surface hydroxyl groups of bulk zirconia powder to form a methoxy-terminated surface accompanied by the release of water,<sup>28,29</sup> with the Zr–OCH<sub>3</sub> species exhibiting IR bands at 2923 and 2817 cm<sup>-1</sup>. Similarly, methoxy groups on other metal oxides have been characterized by IR bands, at 2960 and 2849 cm<sup>-1</sup> for Al<sub>2</sub>O<sub>3</sub>,<sup>30</sup> 2920 and 2815 cm<sup>-1</sup> for Fe<sub>2</sub>O<sub>3</sub>,<sup>31</sup> and 2958 and 2858 cm<sup>-1</sup> for SiO<sub>2</sub>.<sup>32</sup>

We thus infer that methanol reacts similarly with the hydroxyl groups of MOF **site 1** of UiO-66 to form Zr–OCH<sub>3</sub> at the missing linker sites. To check the assignment, the sample after methanol treatment and activation was immersed in excess methanol-D<sub>4</sub> at 298 K for 48 h. After the ensuing isotopic exchange, the bands at 2932 and 2828 cm<sup>-1</sup> had disappeared, with the simultaneous growth of a sharp new band at 2060 cm<sup>-1</sup> (SI, Figure S1). The frequency of this band agrees well with that of the IR band at 2062 cm<sup>-1</sup> observed for Zr–OCD<sub>3</sub> on zirconia powder.<sup>33</sup>



**Figure 3.** Relative free energies in kcal/mol (298.15 K, liquid methanol containing 1% H<sub>2</sub>O) of possible products from methanol substitution.

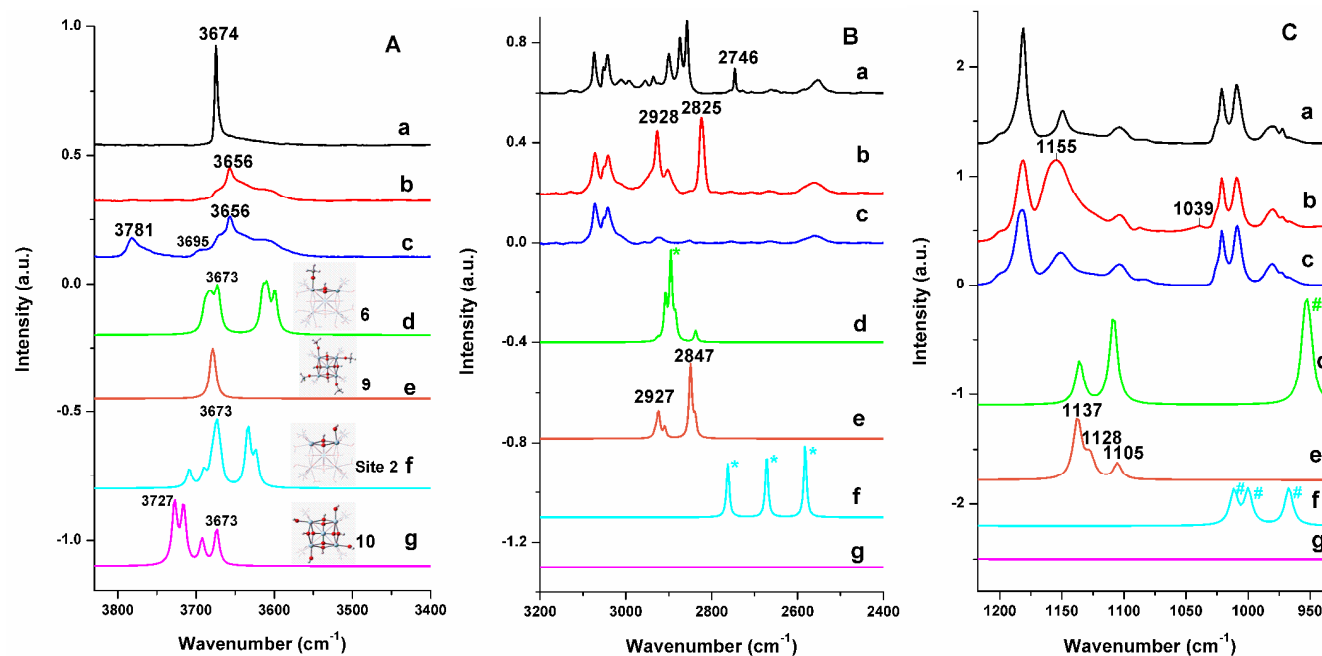
To gain further insight into the nature of the methanol substitution process, potential structures and their relative free energies were computed from density functional theory (M06-L functional, see the Methods section for computational details). As illustrated in Figure 3, the thermodynamically most favorable structure is **2**, which is formed through the substitution of one water molecule by one methanol molecule followed by the loss of the terminating hydroxide as a second water molecule after abstraction of the coordinated methanol proton. (Note that the loss of two equivalents of water becomes still more favorable with methanol containing less than the 1% water that was assumed simply as a convenient standard-state condition.) Structure **2** is predicted to be more stable relative to the adsorption of a second methanol molecule (tautomeric structures **3** and **4**), even in liquid methanol.

Computed C–H stretching vibrations for the OCH<sub>3</sub> group in **2** are predicted, after the application of a scaling factor of 0.958 (a standard value for M06-L<sup>35</sup>), to occur at 2913 and 2812 cm<sup>-1</sup>; these compare favorably with the experimental values of 2932 and 2828 cm<sup>-1</sup> (three peaks are computed for the relevant C–H stretches, the symmetric stretch with strong intensity at lower frequency, and two asymmetric stretches, one with strong and one with weak intensity, respectively—experiment does not resolve these two peaks and we take the lowest of the computed values, which provides a peak separation most consistent with the two peaks observed experimentally). With respect to structure **3**, which the calculations show to be energetically less favorable, a vibration is predicted at 2618 cm<sup>-1</sup> for the stretch of the methanol OH group hydrogen bonded to the adjacent methoxide; no such stretch was observed experimentally.

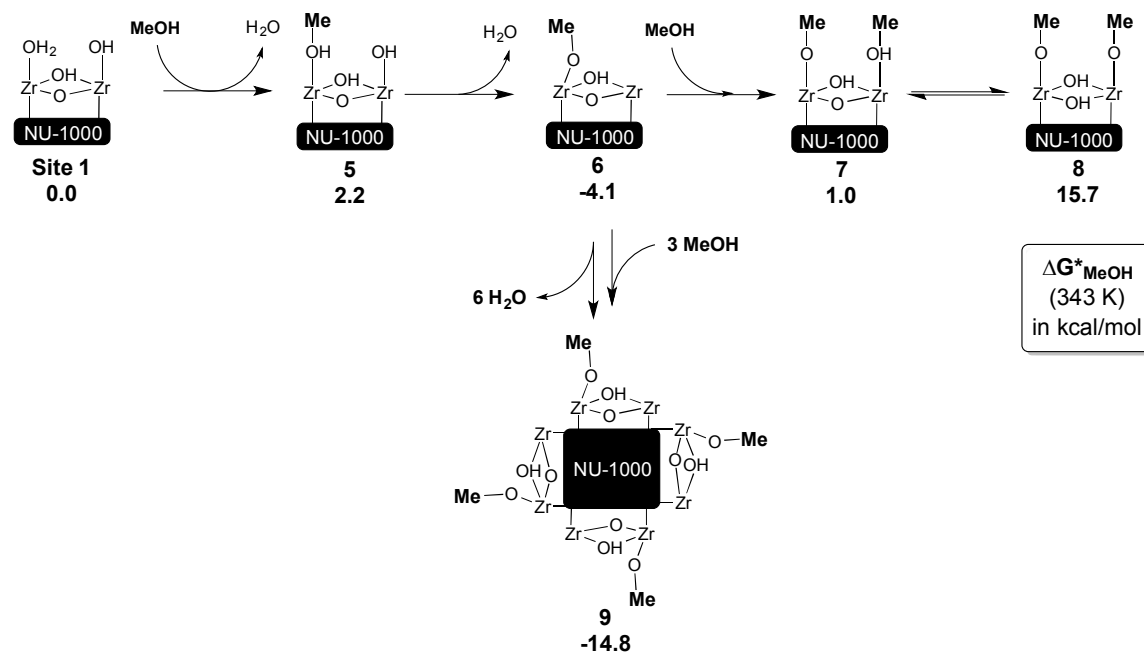
We attempted to carry out the same conversion with UiO-67 but found that this MOF is not stable under the methanol treatment conditions. After 2 h of soaking in methanol, the UiO-67 completely lost its crystallinity.<sup>4</sup>

To check our ability to identify any physisorbed (as opposed to chemisorbed) methanol on the UiO-66 (and to exclude the possibility of its presence in the aforementioned samples), the methanol-reacted UiO-66 was brought into contact with a pulse of methanol in helium at 373 K (~100 molecules per MOF node; methanol vapor was introduced into the IR cell containing the MOF by bubbling helium carrier gas through liquid methanol). The IR spectrum (SI, Figure S2) indicated the presence of the expected physisorbed methanol on UiO-66, as judged by new IR bands at 2945 and 2841 cm<sup>-1</sup>, which match closely those observed for methanol physisorbed on zirconia powder and those observed for methanol vapor.<sup>34</sup>

Extending the work from UiO-66 to NU-1000, we conducted similar methanol washing experiments and obtained the IR spectra shown in Figure 4. The sharp band at 2746 cm<sup>-1</sup> observed for pre-methanol-washed NU-1000 (spectrum **a**) is similar to that found for pre-methanol-washed UiO-66 modulated with HCl and washed with acetone, and it is similarly assigned<sup>4,16</sup> to the hydrogen-bonded OH/OH<sub>2</sub> stretch associated with **site 1**. We soaked this NU-1000 in neat methanol at 343 K for 48 h and then removed the methanol under high vacuum at 298 K followed by further high-vacuum treatment at 423 K, avoiding contact with moisture at all steps. Following these treatments, the band characterizing the hydrogen-bonded OH/OH<sub>2</sub> species, at 2746 cm<sup>-1</sup>, disappeared, and two new bands appeared, at 2928 and 2825 cm<sup>-1</sup> (Figure 4, spectrum **b**), again attributed to methoxy group C–H stretches as observed for UiO-66. Simultaneously, the band at 3674 cm<sup>-1</sup>, assigned to non-hydrogen-bonded μ<sub>3</sub>-OH and OH and OH<sub>2</sub> stretches,<sup>4,16</sup> disappeared, with a new, broader band appearing, centered at 3657 cm<sup>-1</sup>, which we assign to the μ<sub>3</sub>-OH stretches in the newly formed surface species. Furthermore, a band was observed at 1155 cm<sup>-1</sup>, which we associate with the ν<sub>CO</sub> stretch of the methoxy group. (We note that up to 30–50% of the methoxy groups adsorbed on bulk zirconia are observed to be coordinated in a μ<sub>2</sub> fashion to two zirconium atoms,<sup>28,29</sup> as judged by a lower-frequency ν<sub>CO</sub> stretch observed in the spectrum of the bulk. In spectrum **b**, we observe a very low-intensity peak at 1039 cm<sup>-1</sup> that might possibly be attributed to such a species, but DFT does not find such a structure to be stationary, and so we infer that if such species were present, they must be at very low concentration, and that the different local environment of the MOF node compared with a bulk zirconia surface mitigates their formation.)



**Figure 4.** A, Experimental IR spectra in the 3400–3900  $\text{cm}^{-1}$  region, B, in the 2400–3200  $\text{cm}^{-1}$  region, and C, in the 940–1250  $\text{cm}^{-1}$  region characterizing NU-1000 (a); NU-1000 after treatment with liquid methanol at 343 K for 48 h (b); and sample depicted in b after contact with water vapor/helium at 393 K for 30 min (c). NU-1000 and NU-1000 after treatment with liquid methanol were activated at 393 or 423 K, respectively, under high vacuum for 12 h. Theoretical IR spectra of model **6** (d), **9** (e), **site 2** (f), and **10** (g) in the C–H and O–H stretching regions. The bands were scaled with a factor of 0.958. Bands labeled with an asterisk represent the hydrogen bond O–H stretching of the aqua ligand. Bands labelled with a pound symbol represent bending of aqua ligands. Color code: carbon, gray; hydrogen, white; zirconium, cyan; oxygen, red.



**Figure 5.** Relative free energies (343 K, liquid methanol, MeOH, containing 1%  $\text{H}_2\text{O}$ ) of possible products from methanol substitution. Only one face of NU-1000 was modified in the calculations shown on the upper line. Consistent with prior work indicating little communication between faces with respect to thermochemical changes associated with node decoration, the energy predicted for structure **9**, having the topology of **6** on all four NU-1000 faces ( $-14.8 \text{ kcal mol}^{-1}$ ), is just slightly less than four times the value for a single face.

The same computational procedure described above for UiO-66 was employed to characterize the effect of methanol washing on NU-1000, and analogous structure **6** was found again to be the most thermodynamically favored species in contact with methanol solution (Figure 5). For the structure having the topology of **6** on all four NU-1000 faces (Figure 5, structure **9**), the peaks associated with methoxy group C–H bond stretching are computed to occur at 2927 and 2847  $\text{cm}^{-1}$ , again in satisfactory agreement with the experimental values. The computed  $\nu_{\text{CO}}$  stretch **9** is found at 1129  $\text{cm}^{-1}$  (averaged over peaks with strong C–O stretching character separated by 26  $\text{cm}^{-1}$ ), whereas the corresponding peak for the less stable **7** (when considering substitution on a single face) is predicted to be on average at 1049  $\text{cm}^{-1}$ ; the former value is in considerably better agreement with experiment than the latter, supporting the inference of the predominant presence of the topology of **6** on each face (*i.e.*, **9**) after washing and activation of the MOF.

In additional experiments, we observed that the reaction between methanol and hydroxyl groups on NU-1000 nodes also takes place at 298 K, as shown by the results of Figure S3 in SI, but the conversion under these conditions was incomplete.

We tested the reactivity of node-surface-bound methoxy groups with water, finding that the treatment generated **site 2**. After treatment of the UiO-66 sample that had been exposed to methanol at 298 K for 48 h and the NU-1000 sample that had been exposed to methanol at 343 K for 48 h with water vapor at 393 K for 30 min (water vapor was introduced into the IR cell containing the MOF by bubbling helium carrier gas through liquid water), the bands in the IR spectrum characterizing the C–H vibrations at 2932 and 2828  $\text{cm}^{-1}$  for UiO-66 and the C–H vibrations at 2928 and 2825  $\text{cm}^{-1}$  and the C–O vibration at 1155  $\text{cm}^{-1}$  for NU-1000 decreased substantially (Figure 2c and Figure 4c).

Moreover, a new band was observed at 3781  $\text{cm}^{-1}$ , indicating the presence of new terminal OH groups. Methanol produced in the reaction of the methoxy groups with water was observed in the effluent gas by mass spectrometry (Figure S5, SI). Having calibrated our mass spectrometer with pulses of methanol, we determined how much methanol desorbed as a result of the reaction with water, finding the number of methoxy groups per NU-1000 node to be 3.9. Thus, the number of missing linker sites calculated from the quantified methanol signal (assuming only one methoxy group is bonded to each missing linker site per UiO-66 node), 1.84/12, is in good agreement with the value of 1.75/12 determined by thermogravimetric analysis of the MOF.<sup>36</sup> Consequently, we conclude that structure **2** (and structure **9**) is an accurate representation of the structure for UiO-66 (and for NU-1000) after the methanol treatment, respectively.

The methoxy groups are evidently needed as intermediate species to facilitate the conversion from **site 1** to **site 2** on the node vacancies. When we treated UiO-66 or NU-1000 samples incorporating **site 1** with just water in a closed container at 393 K for 12 h, there was no observed conversion of **site 1**. Nor was **site 2** converted to **site 1** under such conditions.

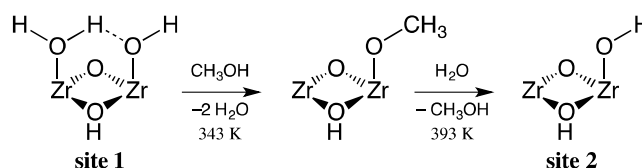
The calculated IR spectra of NU-1000 structures **6**, **9**, one monohydroxy-substituted NU-1000 **site 2**, and a NU-1000 fully substituted with four **sites 2** (designated as **10**) are presented in Figure 4. A peak at about 3727  $\text{cm}^{-1}$  was observed for **10**, representing the O–H stretching, which is 54  $\text{cm}^{-1}$  blue shifted compared with the experimental value of 3781  $\text{cm}^{-1}$ .

The  $\mu_3$ -O–H stretching is observed at about 3673  $\text{cm}^{-1}$  for **6**, NU-1000 **site 2**, and **10**, whereby the structure **6** band is in good agreement with Figure 4.A.a, which represents the MOF after methanol treatment at 298.15 K. A difference of 22  $\text{cm}^{-1}$  was observed for NU-1000 **site 2** and **10**. For details, see Table S5 in the SI.

CM5 partial charge analysis was performed for the  $\text{Zr}_6$  nodes having **site 1** and **site 2** topologies, as well as for structures **2**, **6**, and **10** (Table S2). The charge increase for the low-coordinate Zr in UiO-66 is +0.05 a.u. on going from **site 1** to structure **2**, and +0.07 a.u. on going from **site 1** to **site 2**. For NU-1000, the same change of +0.07 a.u. was observed, both on going from **site 1** to structure **5**, and also at every undercoordinated Zr on going from **site 1** to **10**. As might be expected, eliminating an aquo ligand increases the electrophilicity of the undercoordinated metal atom in every instance.

To generalize the chemistry, we used ethanol or 1-propanol instead of methanol in the reaction with **site 1** on UiO-66 nodes, as shown by the results in Figures S6 and S7 in the SI. The C–H bands for ethoxy groups appear in the respective IR spectra at 2972, 2930, and 2871  $\text{cm}^{-1}$ , and these disappeared after contact of the ethoxylated sample with water vapor at 423 K for 1 h, with a simultaneous increase in the 3780  $\text{cm}^{-1}$  band characterizing **site 2**. The C–H bands for *n*-propoxy groups appeared in the IR spectrum at 2962, 2932, 2877, and 2859  $\text{cm}^{-1}$ . In contrast to methoxy and ethoxy groups, the propoxy groups were not removed by the water vapor treatment at 423 K for 3 h, and we speculate that the lack of reactivity in this instance may be an indication that the alkyl chain of 1-propanol is sufficiently long for favorable van der Waals interactions with the surrounding organic linkers to reduce its effective volatility compared with those of methanol and ethanol.

## DISCUSSION



**Scheme 2.** Chemistry inferred for face topology changes on the MOF  $\text{Zr}_6$  nodes.

The chemistry of these various face topology changes on the  $\text{Zr}_6$  nodes is represented in Scheme 2. We emphasize that the methanol (or ethanol) treatment is crucial for the conversion of **site 1** to **site 2** at moderate temperature, as treatment of either UiO-66 or NU-1000 with water or helium or exposure to high vacuum at 393 K does not lead to loss of the aquo ligand, although treatment at temperatures above 473 K with the sample under high vacuum does effect conversion (Figure S8, SI). Considering the reverse reaction, we observed that **site 2** is not converted to **site 1** by treatment with water at various temperatures, at least over shorter time scales.

That the aquo ligand of **site 1** is not lost upon treatment with hot water is perhaps not surprising, particularly given that the gas-phase binding enthalpy of the aquo ligand to Zr in **site 1** is

1 predicted to be 34 and 38 kcal/mol in UiO-66 and NU-1000,  
 2 respectively. Soaking of the sample in alcohol, however, per-  
 3 mits an equilibrium exchange driven by concentration differ-  
 4 ences (and favorable solvation of lost water) that permits ac-  
 5 cess to the monoalkoxy substituted facial topology (structures  
 6 **2** and **6** in Figures 3 and 5, respectively), and in contrast to the  
 7 situation in water, binding of an additional neutral alcohol is  
 8 *not* predicted to be favorable. This reasoning sets up a scenario  
 9 whereby subsequent water treatment reverses the concentra-  
 10 tion imbalance, and swapping of hydroxide for methoxide is  
 11 computed to be almost thermoneutral on conversion to a local  
 12 **site 2**. What is puzzling, of course, is that treatment of **site 2**  
 13 with high-temperature water does not lead to speedy regenera-  
 14 tion of **site 1**. It may be that the hydrophobic nature of the  
 15 local environment in contact with the MOF face, which de-  
 16 rives from the proximity of the aromatic linkers, leads to a  
 17 local aqueous structure that is clathrate-like, and that such  
 18 structuring makes it energetically difficult for a water mole-  
 19 cule near the face to re-coordinate to Zr. Assessing such a sit-  
 20 uation would require simulations with explicit solvent (and a  
 21 suitable model for computing Zr–O interactions) that go be-  
 22 yond the scope of this work, but this is a topic meriting further  
 23 study. It is known that fully dehydrated NU-1000 reverts to  
 24 fully hydrated NU-1000 (*i.e.*, all faces having **site 1** topology)  
 25 when exposed to benchtop humidity for a period of several  
 26 weeks,<sup>37</sup> and so we are confident that the barrier is indeed  
 27 kinetic in nature.

28 The chemistry reported here is unique to MOFs because the  
 29 conversion involves unique terminal OH and/or OH<sub>2</sub> groups  
 30 on MOF nodes, and comparable groups are evidently lacking  
 31 in zeolites and metal oxides. The uniform structures of the  
 32 MOFs allow unprecedentedly precise control of surfaces with  
 33 these groups, which are essential for some separations by ad-  
 34 sorption,<sup>9</sup> for catalysis,<sup>38,39</sup> and for anchoring of metal com-  
 35 plexes.<sup>4,5,40</sup>

36 We note, however, that the  $\nu_{\text{OH}}$  bands of terminal OH  
 37 groups of **site 2** on UiO-66 and NU-1000 have full width at  
 38 half-maximum (fwhm) values of about 26 and 23 cm<sup>-1</sup>, respec-  
 39 tively. In contrast, the value characterizing the hydroxyl band  
 40 at 3630 cm<sup>-1</sup> in the IR spectrum of SiOHAl groups of HY zeo-  
 41 lite (Figure S9 in the SI) is only 9 cm<sup>-1</sup>. A comparison of these  
 42 values indicates that the degree of uniformity of hydroxyl **sites**  
 43 **2** on the Zr<sub>6</sub> MOF nodes is markedly less than the degree of  
 44 uniformity of comparable hydroxyl sites on the zeolite, con-  
 45 sistent with our previously reported results derived from the  
 46 fwhm values characterizing  $\nu_{\text{CO}}$  bands of Ir(CO)<sub>2</sub> species an-  
 47 chored to the MOF nodes and to HY zeolite.<sup>4</sup> There is more  
 48 yet to be learned about the subtle chemistry of these MOF  
 49 nodes and their surface sites.

## 50 CONCLUSIONS

51 In summary, methoxy groups on the Zr<sub>6</sub> nodes of UiO-66 and  
 52 NU-1000 (Scheme 1), characterized by IR bands at 2932 and  
 53 2828 cm<sup>-1</sup> and identified by IR spectroscopy and DFT calcula-  
 54 tions, arise from the reaction of the MOF node hydrogen-  
 55 bonded OH/OH<sub>2</sub> groups (**site 1**) with methanol, and the result-  
 56 ant node surface structure is converted into one vacant Zr site  
 57 and one terminal OH group upon reaction with water. To the  
 58 best of our knowledge, this is the first example of the tuning  
 59 of the chemistry of reactive groups (hydroxyl groups) on MOF  
 60 nodes—and the intermediacy of methoxy or ethoxy groups is  
 crucial for control of this chemistry. Because the DFT calcula-

tions demonstrate that the various types of hydroxyl group on  
 the Zr<sub>6</sub> nodes have markedly different free energies, and be-  
 cause these groups perform differently from each other in  
 bonding to metal complexes,<sup>4,5</sup> we may expect the control of  
 these sites to open precise new chemistry, including that of  
 anchored metal complexes, on MOF nodes.

## ASSOCIATED CONTENT

**Supporting Information.** Experimental and computational de-  
 tails are reported, primarily IR spectra, Cartesian coordinates with  
 its associated energies. This material is available free of charge  
 via the Internet at <http://pubs.acs.org>.

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### Author Contributions

The manuscript was written through contributions of all authors.  
 All authors have given approval to the final version of the manu-  
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### Notes

The authors declare no competing financial interests.

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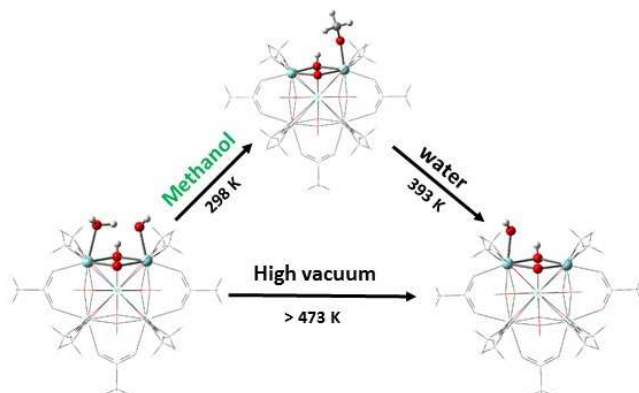
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## SYNOPSIS TOC

## Topology change on missing linker site of UiO-66 nodes



## Topology change on missing linker site of UiO-66 nodes

