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4f-electron behavior in the cerium oxides Ce_2O_3 and CeO_2 under pressure

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4f-electrons play a key role in the high-pressure behavior of the 4f rare-earth metals. The element cerium presents the prototypical case with nominally just one 4f electron. The participation of that f-electron in the bonding of compounds such as the cerium oxides, Ce_2O_3 and CeO_2 , governs their material properties. To investigate this 4f electron behavior we have studied the magnetic susceptibility for both oxides at atmospheric pressure and angle dispersive X-ray diffraction of Ce_2O_3 with and without a pressure-transmitting medium (PTM) in a diamond anvil cell up to nearly 70 GPa. In contrast to previous data we find that the Ce atom in Ce_2O_3 behaves like a trivalent Ce^{3+} ion. Up to the highest pressure Ce_2O_3 remains in the hexagonal phase (space group 164, P-32/m1) typical for the lanthanide sesquioxides. The isothermal bulk modulus and its pressure derivative for the quasi-hydrostatic case are $B_T = 107 \pm 2$ GPa, $B' = 5.2 \pm 0.2$ and for the case without PTM $B_T = 99 \pm 2$ GPa, $B' = 7.0 \pm 0.2$. Since X-ray emission spectroscopy of the Ly ($4d_{3/2} \rightarrow 2p_{1/2}$) transition in cerium metal is sensitive to the 4f-electron occupancy, we have followed the high-pressure dependence of this line for both Ce_2O_3 and CeO_2 and compared it with spectra taken from Ce metal under pressure and the zero f-electron element lanthanum. In both cases, no change of the respective lineshape was observed, indicating that the 4f-electron configuration in Ce_2O_3 is stable up to at least 50 GPa and that the volume collapse from the high-symmetry, low-pressure fluorite structure to the lower-symmetry orthorhombic phase of CeO_2 at 31 GPa is not driven by the f electrons.

Introduction

Ce_2O_3 and CeO_2 are technologically important materials that are often encountered together in applications. Both can be considered as the two extreme brackets of CeO_x with x ranging from 1.5 to 2 where x can vary almost seamlessly over a wide part in between. They often occur together in reaction pathways that e.g. eliminate pollutants as in the use of ceria in catalytic converters or allow thermochemical splitting of water into O_2 and H_2 possibly useful in the

pursuit of energy storage [1-4]. The variability in x allows the material to either release or store oxygen and explains its great usefulness. Changing x from $x = 1.5$ to 2 results in a crystallographic structure change from a hexagonal lattice (SG 164, P-32m1, the so-called A-type in the context of rare-earth sesquioxides with lattice constants $a = 3.891$ and $c = 6.059$ Å, 1 formula unit per cell [5]) to face centered cubic lattice (fluorite structure, SG 225, Fm-3m, $a = 5.411$ Å, 4 formula units per cell).

Since CeO_2 is thermodynamically stable there has been more research performed on it while relatively little is known about the less stable Ce_2O_3 . Strangely enough, there appear to be more controversies surrounding CeO_2 . One of them is caused by conflicting reports whether CeO_2 is paramagnetic [6,7] or diamagnetic [8]. Ce_2O_3 , on the other side, is definitely paramagnetic. While CeO_2 is thermodynamically stable, Ce_2O_3 converts into CeO_2 at ambient conditions, which allows experimental results of Ce_2O_3 to be affected by contamination with CeO_2 unless great care is exercised.

Part of the theoretical challenge of describing these materials lies in the coherent description of f-electronic orbitals, bands, and their occupation which has produced many publications on the subject of the valence and f-electron configuration [6-7, 9-21]. Most research agrees that the Ce in Ce_2O_3 ought to have (about) one localized 4f-electron while it is highly debated whether the 4f electron is completely delocalized in CeO_2 ($4f^0$ configuration, tetravalent) or whether a partial charge remains localized resulting in a mixed-valent configuration [6-7, 9-17]. Early theoretical efforts found that CeO_2 would have 0.5 f-electrons localized [10] and intermediate valence [14]. This was challenged by the results of an optical absorption experiment [6, 7] which appeared to impose a limit of at most 0.05 localized f-electrons in CeO_2 favoring a purely f^0 electronic configuration. The picture of a tetravalent Ce^{4+} ion, in CeO_2 , however, with a purely f^0 electronic configuration in the ground state is too simplistic [14, 22] since the hole L in the oxygen 2p level interacts with the 4f resulting in a $4f^1L$ charge transfer configuration. More sophisticated approaches to treat both cerium oxides consistently become increasingly difficult to implement. A first principles calculation of the solid-state properties of both CeO_2 and Ce_2O_3 had to distinguish between a core state model (CSM) for Ce_2O_3 localizing the f-electron explicitly and a valence band model (VBM) delocalizing the f-electron for CeO_2 [18]. Approaches employing screened hybrid density functionals appeared more promising in reproducing experimental results [21]. A later attempt using LDA allowed the description of both oxides at zero temperature and pressure by varying the so-called U-parameter in the LDA+U formalism, which had to be set to 5 - 6 eV or higher for both Ce_2O_3 and CeO_2 [19, 20]. The accuracy of the LDA + U approach was thought to provide a superior description of the cerium oxides [23].

Qi et al. expanded on the use of the LDA+U method and calculated the pressure dependence of structural parameters for Ce_2O_3 [24] as well as its elastic constants. Their work culminated in a

prediction that the hexagonal structure of Ce_2O_3 would become unstable at 30 GPa [24], close to the pressure at which CeO_2 undergoes its volume collapse from a high symmetry fcc to a lower symmetry orthorhombic phase. This transition starts at 31 GPa and is completed at 38 GPa under non-hydrostatic conditions (no pressure transmitting medium)[25], under more quasi-hydrostatic conditions (4:1 methanol ethanol mixture) the transition continues for another 15 GPa [26]. The prediction is somewhat surprising since the hexagonal structure of the sesquioxides is considered rather stable [27]. For CeO_2 the theoretical data can be compared to results for lattice constants and bulk moduli of several experimental studies [25, 26, 28, 29] whereas apparently there exists no equation of state study for Ce_2O_3 up to now.

Because the 4f electrons already have a large impact on the cerium oxides in their ambient state, our study focuses on the behavior of the 4f electrons under changing length scales by tuning the interatomic distances via pressurization in diamond anvil cells (DAC). For reference, in pure metallic cerium the f-electrons drive the famous volume collapse from the γ - to the α -phase, which occurs at a pressure lower than 1 GPa [30], and for which the mechanism was highly debated [31-33]. The same questions regarding localization and itinerancy, magnetic moments, their screening and how this affects the volume evolution / collapse of metallic cerium also pertain to the behavior of CeO_2 and Ce_2O_3 . In our present work we therefore investigate the pressure-volume room temperature equation of state for Ce_2O_3 up to 70GPa, the pressure dependence of its lattice constants as well as the bulk modulus B and its pressure derivative B' . We also examine the possibility that the thermodynamic instability of Ce_2O_3 leads to a disproportionation back into Ce metal and CeO_2 under pressure ($2 \text{ Ce}_2\text{O}_3 \rightarrow \text{Ce} + 3 \text{ CeO}_2$). Additionally, we characterize (at ambient pressure) our CeO_2 and Ce_2O_3 samples with regard to their magnetic susceptibility, which is intimately tied to their 4f occupancy. To assess the occupancy under pressures unattainable for magnetic susceptibility studies, we turn to X-ray emission spectroscopy of the Ly line ($4d_{3/2} - 2p_{1/2}$) that exhibits a satellite due to an exchange interaction with a 4f-electron [34, 35] similar to other rare-earths materials [36, 37]. We compare our experimental observations of the Ly lineshape for Ce_2O_3 with those of metallic γ - and α -cerium, elemental lanthanum, and theoretical lineshapes of Ce_2O_3 calculated previously by others in the framework of an Anderson impurity model with full multiplet couplings [34]. Finally, we estimate the 4f-electron occupancy based on the integrated intensity of the Ly satellite.

Experimental Details

Measurements on CeO_2 were performed on a commercially acquired 99.9% CeO_2 powder (Alfa Aesar). Our Ce_2O_3 was synthesized starting from 99.9% cerium metal (Alfa Aesar) and CeO_2 that were weighed to give the correct stoichiometry, pressed into a pellet and then heated to 1350 °C for 10 hours under UHP argon flow. The resulting powder was golden-yellow in color,

consistent with other reports of Ce_2O_3 synthesis [38]. Ce_2O_3 is not thermodynamically stable: exposed to air at slightly elevated temperature it rapidly converts to CeO_2 . This conversion apparently even proceeds when the material is stored in a bottle at room temperature. Indeed, our first experiments with a commercially acquired Ce_2O_3 sample had to be repeated with material produced in house, because the commercially acquired material was found, on subsequent analysis, to have changed into 80 % CeO_2 (balance Ce_2O_3) over the years. Therefore, to prevent degradation and contamination, our Ce_2O_3 was synthesized and stored in an argon glove box with less than 0.1 ppm of oxygen and less than 0.5 ppm of water present. A portion of the as-produced Ce_2O_3 powder was loaded into a hermetically sealed x-ray diffraction holder, and x-ray diffraction (see Fig. 1) indicated a purity of at least 99%, the lattice constants were 3.8908 and 6.0619 Å in excellent agreement with the literature [5]. In particular, no peaks indicative of the presence of CeO_2 were found in the freshly synthesized material.

Ambient-pressure magnetic characterization of the cerium oxides used a Quantum Design SQUID magnetometer with an applied field $H=0.1$ T. Gel caps were used to contain the powders of both CeO_2 and Ce_2O_3 . Great care was taken to limit any air exposure for the Ce_2O_3 sample: the powder was loaded into a gel cap and straw assembly inside of the glove box, and that assembly was transferred within a sealed bag of argon to the magnetometer, where it was rapidly inserted in the helium atmosphere of the venting air-lock.

Handling of the material and placing small crystallites into diamond anvil cells were performed in another glove box. For the quasi-hydrostatic samples the DAC were additionally charged with neon as pressure-transmitting medium. We chose rhenium as gasket material for the angle dispersive x-ray (ADX) diffraction experiments and used copper for pressure calibration [39]. For the x-ray emission spectroscopy (XES) experiments we employed beryllium as the gasket material with ruby as pressure calibrant. Culet sizes of the diamonds ranged from 300 to 500 μm . Two DAC were prepared for the quasi-hydrostatic EOS experiments, one for the non-hydrostatic one, and several for the x-ray emission experiments investigating either CeO_2 or Ce_2O_3 .

ADX diffraction and XES experiments were carried out at the Advanced Photon Source at the High Pressure Collaborative Access Team (HPCAT) Sector 16 IDB and 16 IDD beamlines, respectively. The typical incident wavelength for ADX diffraction was ~ 0.4 Å. For the XES, we used high energy x-rays (11 or 18 keV) entering axially through the diamonds, and collected the L_γ x-rays (from 6007 to 6090 eV, peak at 6052 eV) leaving through the beryllium gasket under 90 degrees angle to the incoming beam. The energy was scanned using a standard four-inch, spherically bent (333) Si analyzer crystal on a 1m Rowland geometry situated 90 degrees to the incoming X-ray beam with a resolution of 1 eV.

Results and Discussion

Ambient-pressure Magnetic Characterization of the Ce Oxides

The temperature-dependent magnetic susceptibilities χ of the cerium oxides are shown in Fig. 2. Immediately evident is the magnitude of χ for Ce_2O_3 relative to that of CeO_2 : the former exhibits a magnetic susceptibility that is at least 20 times larger than the latter over the entire temperature range. The solid lines through the data points for both oxides are fits to a modified Curie-Weiss expression:

$$\chi(T) = \chi_0 + f \frac{C}{T - \Theta_{CW}} \quad (\text{Eq. 1})$$

where χ_0 is a constant susceptibility, f is the molar fraction of Ce, C is the Curie constant, and Θ_{CW} is the Curie-Weiss temperature.

For Ce_2O_3 , the results of this fit yield a full Ce^{3+} moment (measured $\mu_{\text{eff}} = 2.57 \mu_B$ versus expected Hund's rule value of $\mu_{\text{eff}} = 2.54 \mu_B$), $\chi_0 = 4.5 \times 10^{-3} \mu_B/\text{f.u.}$, and a $\Theta_{CW} = -12.5 \text{ K}$, suggesting weak antiferromagnetic correlations between the Ce ions. The observed full, trivalent Ce moment is perhaps not surprising in Ce_2O_3 , as the oxidation state would be expected to be Ce^{3+} ($[\text{Xe}]4f^1$). However, the measured effective moment is in contrast to the report of Pinto, et al., which suggested an effective Ce moment half that of the trivalent value [40]. Pinto, et al. note that this low effective moment could be due to crystalline field effects, but another explanation could be the presence of CeO_2 impurities, as they are clearly evident from their published diffraction pattern. No evidence for magnetic ordering is seen in our magnetization data at $H = 0.1 \text{ T}$.

While Ce_2O_3 may be expected to exhibit the classic trivalent rare-earth element moment, CeO_2 should be tetravalent (Ce^{4+}), with a filled outer-shell electron configuration ($[\text{Xe}]$) and thus a diamagnetic susceptibility. Indeed, Laachir, et al. has reported a diamagnetic susceptibility for an ultra-pure sample of CeO_2 (<6 ppm ferromagnetic impurities) [8], although no temperature dependence was included in their report. Our temperature-dependent magnetic susceptibility data for CeO_2 , on the other hand, are best described by equation 1 with $\chi_0 = 3.0 \times 10^{-4} \mu_B/\text{f.u.}$ (paramagnetic not diamagnetic) and a $\Theta_{CW} = 0 \text{ K}$. The lack of obvious diamagnetism suggests that the picture of a $4f^0$ tetravalent Ce ion may be too simple of a description of CeO_2 . The Curie tail seen in the susceptibility of CeO_2 is suggestive of the presence of impurities. If these impurities are assumed to be trivalent Ce ions, then applying equation 1 to our CeO_2 specimen implies that only about 0.5% of the Ce ions are Ce^{3+} rather than Ce^{4+} .

Ce₂O₃ Crystal Structure under Pressure

Fig. 3 shows selected angle dispersive X-ray pattern from a sample without a pressure transmitting medium containing copper as pressure calibrant [39]. Up to the highest pressure, both quasi-hydrostatic and non-hydrostatic samples remained in the hexagonal symmetry and the structure was consistent with that of P3-2/m1 (SG 164). After pressure release, the sample returned back to the literature values for lattice constants and volume within experimental uncertainty.

Fig. 4 shows our measured lattice constants c and a for Ce₂O₃ up to 70 GPa. Fig. 5 displays the quasi-hydrostatic and non-hydrostatic equation of state for Ce₂O₃. The Vinet equation of state

$$P(V) = 3B_0 \frac{(1 - \sqrt[3]{x})}{\sqrt[3]{x}^2} \exp[1.5(B'_0 - 1)(1 - \sqrt[3]{x})] \quad \text{with } x = V/V_0 \quad (\text{Eq. 2})$$

was fitted to the EOS data. We find $B_0 = 107 \pm 2$ GPa and $B'_0 = 5.2 \pm 0.2$ for the quasi-hydrostatic data and $B_0 = 99 \pm 2$ GPa and $B'_0 = 7.0 \pm 0.2$ for the non-hydrostatic case. Both EOS are basically identical up to about 25 GPa where the non-hydrostatic EOS diverges and becomes somewhat stiffer.

Qi et al. [24] calculated the lattice constants and equation of state using the LDA + U approach in the framework of density functional theory. Their predicted bulk modulus is $B_0 = 135$ GPa and its pressure derivative $B'_0 = 3.96$ – quite different from our values – but their $p(V)$ data fit well onto the present isotherms (see Figs. 4 and 5). While their calculated B_0 is quite a bit higher than our measured one, they also propose a bulk modulus based on the Voigt-Reuss-Hill averaging scheme of their calculated elastic constants C_{ij} for polycrystalline material and that scheme yields a bulk modulus $B(0\text{GPa}) = 109$ GPa, very close to our experimental value for quasi-hydrostatic conditions. Their calculated anisotropic behavior of the lattice constants at 25 GPa is $\Delta a/a \sim 1.9\%$ and $\Delta c/c \sim 10.5\%$. Experimentally, at 25 GPa, we find $\Delta a/a \sim 2.1\%$ and $\Delta c/c \sim 10.4\%$, in very good agreement. The anisotropy continues on: At 70 GPa we find that $\Delta a/a \sim 4.6\%$ whereas $\Delta c/c \sim 18.5\%$. In other words, most of the more than 25% volume reduction at 70 GPa originates from compression along the c -axis under pressure. Table 1 contains a compilation of the present data and a comparison with theoretical calculations mostly based on the LDA+U formalism. The predictions for Ce₂O₃ show substantial variance while also presenting a larger bulk modulus than our measurements.

At 30 GPa, Qi et al. find an anomalously high bulk modulus (289 GPa) and a negative C_{44} elastic constant, implying a predicted structural change away from the hexagonal unit cell for pressures in excess of approximately 30 GPa. However, up to the highest pressure measured, the hexagonal structure persists experimentally. Determining whether this predicted phase

transition is hindered by kinetic effects or finite-temperature entropic contributions will require additional theoretical work.

Persistent f-electron Occupancy in Ce₂O₃ under Pressure

Owing to Hund's rules, the 4f occupancy of the lanthanides manifests clear temperature-dependent behavior in the magnetic susceptibility, making these measurements sensitive probes of the f-electron configuration. However, the experimental capabilities to perform magnetic susceptibility under pressure typically do not exceed 2 GPa. We therefore use the satellite of the Ly emission ($4d_{3/2} \rightarrow 2p_{1/2}$) following the excitation of a 2p electron into the continuum in Ce as a proxy for the presence of a localized 4f electron as has been done in numerous previous studies of other lanthanides [34-37, 42]. The satellite is caused by the exchange interaction of the f-electron with the 2p hole and therefore provides a simple answer to whether a localized 4f-electron exists [43]. The volume collapse in cerium metal from the γ - to the α -phase is accompanied by a significant reduction of the Ly satellite tracking the f-occupancy closely [35, 44], thus highlighting the utility of XES to track changes in the nature of the bonding of the f-electrons.

Tanaka et al. [34] calculated the $4d \rightarrow 2p$ lineshape of the Ly radiation of Ce₂O₃ and found that it looked very similar to that of the Ce³⁺ ion where the lower energy satellite can be assigned to the low spin final state ($4d^9 4f^1$) ¹P. They found that effects of the hybridization of the 2p valence band with the 4f orbitals are almost cancelled out by “phase matching” of the involved wavefunctions, meaning that the hybridization affects both intermediate and final state in nearly the same manner [34, 45] and the signature of an electronic charge redistribution between intermediate and final states does not appear [45]. Our experimental findings reveal that the actual lineshape differs somewhat from the calculations. The experimentally observed lineshapes are broader, similar to the way the observations of the lineshapes of Ce metal are broader than the modified extended atomic calculations [35] since only a small number of orbitals are allowed to hybridize in the calculations.

In order to visualize changes in the f-electron occupancy, we show in Fig. 6 the Ly emission of lanthanum, taken from a sample in a DAC at 1 GPa pressure. The La Ly emission spectrum serves as our true zero f-electron system baseline, showing a lack of a satellite and providing a hallmark lineshape for the main peak [45]. Additionally and for comparison, the spectra of Ce₂O₃ (at the low pressure of 1 GPa inside a DAC) and CeO₂ outside a DAC are included in Fig. 6. The long exposure times of several hours for these measurements required that Ce₂O₃ and La were sealed in a DAC to guard against their chemical reactivity. Fig. 6 also shows the lineshapes of cerium metal below and above the 15% volume collapse at 0.75 GPa, which coincides with a 30% drop in f-occupancy. Fig. 6 shows that the line shape of cerium metal at 0.3 GPa is basically

identical to the one of Ce_2O_3 , the largest of the satellites. Other nominally trivalent cerium compounds (such as Ce_2S_3 and CeSi_2 , not shown) also exhibit the same Ly lineshape.

The XES spectra of Ce_2O_3 under pressure are plotted in Fig. 7. Up to the highest pressure of 50 GPa, the satellite for Ce_2O_3 does not appear to change. The lack of changes under pressure shows that any changes in the bonding and thus in the EOS are not due to f-electron involvement, at least up to 50 GPa. As the XES lineshapes are very similar for Ce_2O_3 and cerium metal, so are their respective f-occupancies. Without being able to address the possibility of the magnetic moment generated by the 4f electron being screened by the 5d band electrons, as in the case for cerium metal, and thus affecting the value of the magnetic susceptibility, we can say that localized f-electrons in Ce_2O_3 do persist under pressure. Ce_2O_3 thus represents a material with a stable 4f-electron occupancy (trivalent) and crystal structure up to at least 50 GPa.

CeO_2 under pressure

Unlike Ce_2O_3 , CeO_2 does not retain its ambient-pressure crystal structure up to very high pressures. Instead and near 30 GPa, CeO_2 undergoes an fcc-orthorhombic phase transition with an accompanying volume collapse of nearly 10%, comparable in size to the isostructural collapse of Ce metal [30]. Fig. 8 shows a comparison of the normalized equations of state of Ce_2O_3 and CeO_2 through its volume collapse together with previous results by other authors. At low pressures Ce_2O_3 is more compressible than CeO_2 , but with increasing pressure CeO_2 makes up for the lack in volume change when it undergoes its volume collapse from the fcc to the orthorhombic phase [25, 26]. Whether the f-electrons of CeO_2 play a role in the volume collapse, as is the case for Ce metal, has been an open question.

While Ce_2O_3 showed a small satellite similar in magnitude to that of Ce metal (Fig. 6), the lineshape of CeO_2 reveals an even smaller satellite intensity. No calculations of the Ly lineshape exist for CeO_2 but it can be compared to the experimentally observed lineshape of La. La possesses no 4f-electron and its lineshape shows no satellite (see Fig. 6). Comparing the La lineshape with the one of CeO_2 , a small, remnant satellite in CeO_2 is apparent (see Fig. 6). This remnant can be modelled with the same lineshape parameters of Ce_2O_3 or Ce metal but with reduced amplitude.

XES spectra for CeO_2 under pressure are shown in Fig. 9. Up to the highest pressure of 48 GPa, and identical to the behavior of Ce_2O_3 , the satellite intensity does not appear to change. This pressure range encompasses the structural transition in CeO_2 from the high-symmetry fluorite fcc to the low-symmetry orthorhombic phase which starts at ~ 31 GPa and is completed at 38 GPa in non-hydrostatic conditions [25]. The persistence of the small satellite in the Ly emission of CeO_2 suggests that the volume collapse from the high symmetry fcc to the low-symmetry

orthorhombic phase is not driven by changes in the f-electron configuration, either through delocalization or hybridization.

Assuming the relationship with the f-electron occupancy also holds for the satellite area of CeO_2 , one can estimate the number of f-electrons for both oxides. Correlating the satellite area of cerium metal at pressures below the volume collapse (γ -Ce) with an f-electron occupancy of 0.97 [44], one finds $n_f \sim 1.09 \pm 0.1$ for Ce_2O_3 , and $n_f \sim 0.35 \pm 0.05$ for CeO_2 (Note: this value is larger than the upper-limit, 0.005, of impurity Ce^{3+} ions as determined from magnetic susceptibility). The non-zero value of n_f for CeO_2 indicates that the Ce ions cannot be modeled by a simple closed-shell picture lacking f-electrons. The extracted values of n_f are lower than recent hybrid density functional calculations, which find $n_f = 1.31$ for Ce_2O_3 and $n_f = 0.80$ for CeO_2 [21], but compare well with bond valence methods finding $n_f = 0.27$ for CeO_2 and $n_f = 1$ for Ce_2O_3 [46]. Less recent calculations in the framework of the Anderson impurity model with a filled valence band give a range from $0.38 < n_f < 0.52$ for CeO_2 and $n_f \sim 1$ for Ce_2O_3 [47]. Ref. [11] finds an f-occupancy of about 0.6 for CeO_2 via 3d core level photoemission and an older calculation [10] finds 0.5.

Conclusion

While the experimental properties depend heavily on whether the 4f electron is approximately localized (Ce_2O_3) or more delocalized (CeO_2) the volume reduction under pressure cannot be explained by the changing involvement of the 4f-electrons. In fact, based on our observation of the satellite feature in the Ly lineshape, the properties of the 4f electrons do not change with pressure, at least up to 50 GPa. In particular, the volume collapse in CeO_2 from the high symmetry fcc structure to the low symmetry orthorhombic playing out over a range of ~ 10 GPa starting at 31 GPa is not caused by a change in the 4f electron behavior. Previous XES experiments and calculations on cerium metal have established that the area of the satellite structure follows the ground state moment and occupancy rather closely [35, 44]. It is therefore argued that Ce_2O_3 possesses ~ 1.09 localized f-electrons and $\text{CeO}_2 \sim 0.35$.

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Table 1

Compound	B (GPa)	B' (GPa)
Ce₂O₃		
Calc. a, b, c, d) LDA + U	130 ; 144.7; 135; 150	- ; - ; 4.96 ; -
Calc. c)	109 (VRH average)	
Calc. d) LDA + DMFT	160	
Calc. e, f) LDA (CSM)	165.8 ; 125.9	
Present work (quasi –hydrostatic)	107 ± 2	5.2 ± 0.2
Present work (non –hydrostatic)	99 ± 2	7.0 ± 0.2
CeO₂		
Calc. a, b) LDA + U	214 ; 213.7	
Calc. f) LDA (VBM)	214.7	
Calc. g) screened hybrid density functional	206.1	
Expt. fcc / ort h)	230 ± 10 / 304 ± 25	4.0 / 4.0
fcc i)	236 ± 4	4.4 ± 0.4
	220 ± 4	4.4 ± 0.4
fcc j)	204	

- a) [19]
- b) [20]
- c) [24]
- d) [41]
- e) [18]
- f) [27]
- g) [21]
- h) [25]
- i) [26, 29]
- j) [28]

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Figures

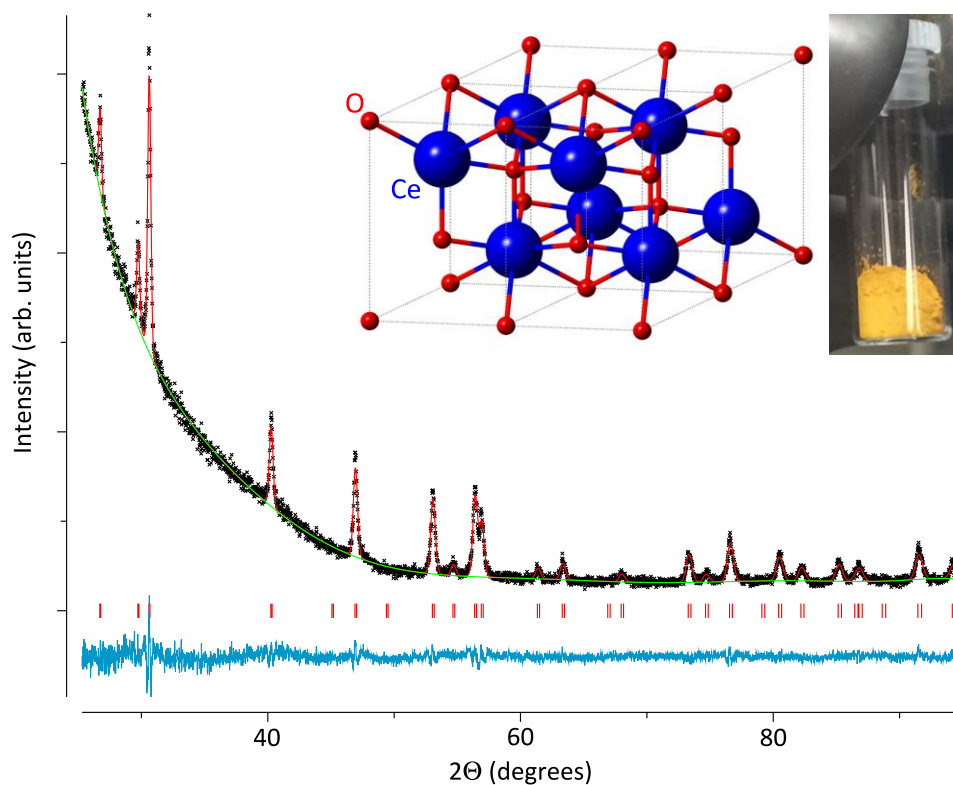


Fig. 1: X-ray diffraction ($\text{Cu K}\alpha$, 1.54059 \AA) results at ambient conditions for our synthesis of Ce_2O_3 . The black crosses are the data points, the green line is the background (with low-angle intensity dominated by the hermetically sealed holder), the tick-marks are the Bragg peaks for the Ce_2O_3 structure, and the blue line at the bottom represents the difference between the data and the refinement. The refinement yielded an $R_{\text{wp}}=5\%$. The crystal structure (left inset) is hexagonal, the so-called A-type (SG 164) of the lanthanide sesquioxides. The right inset shows the golden-yellow color of Ce_2O_3 powder.

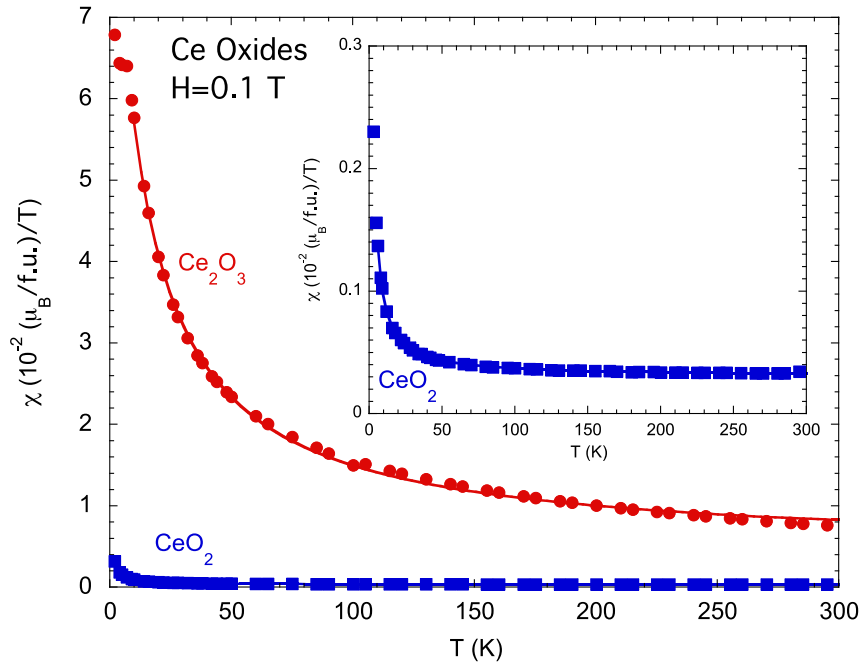


Fig. 2): Magnetic susceptibility in a 0.1 T magnetic field for Ce_2O_3 (red circles) and as-received CeO_2 (blue squares). The lines are fits to the sum of a Curie-Weiss law and a constant susceptibility (see text).

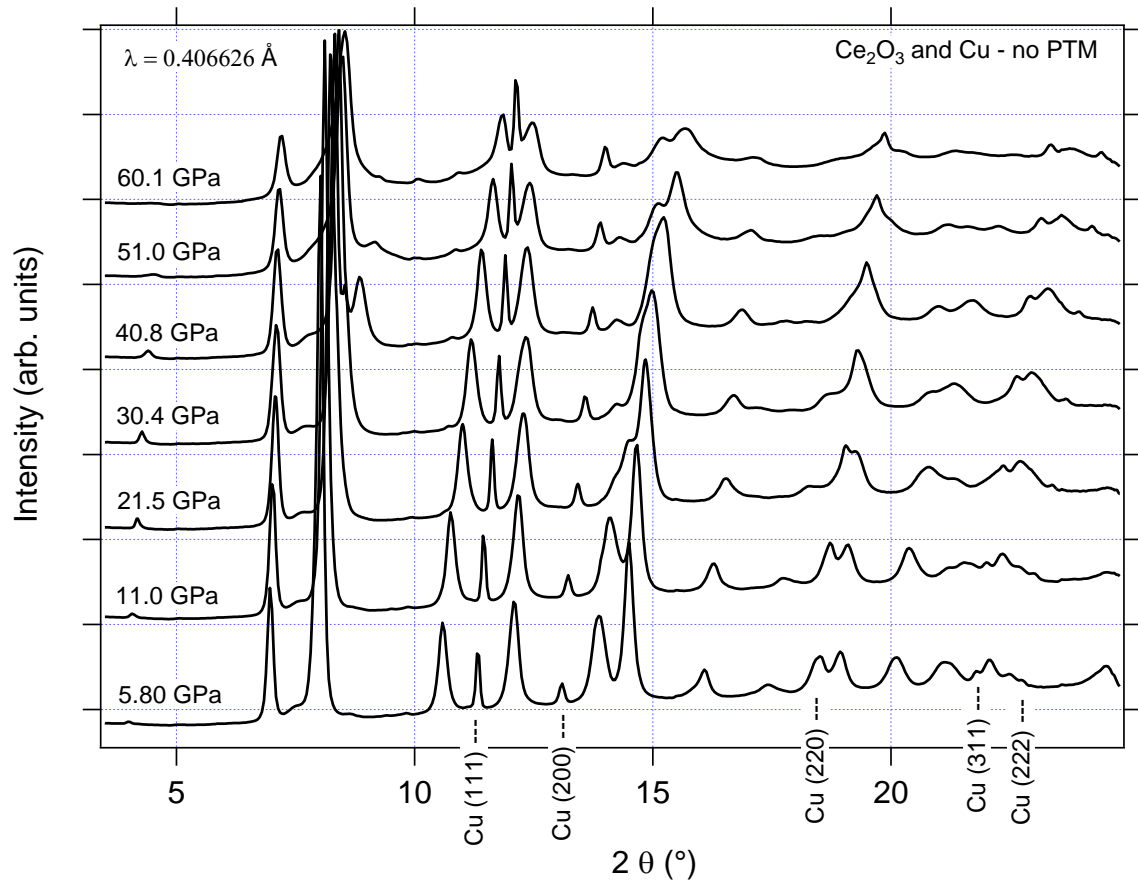


Fig. 3: Angle dispersive X-ray diffraction patterns of Ce_2O_3 under compression in the DAC at different pressures. Pressure calibrant was copper. No pressure-transmitting medium was used.

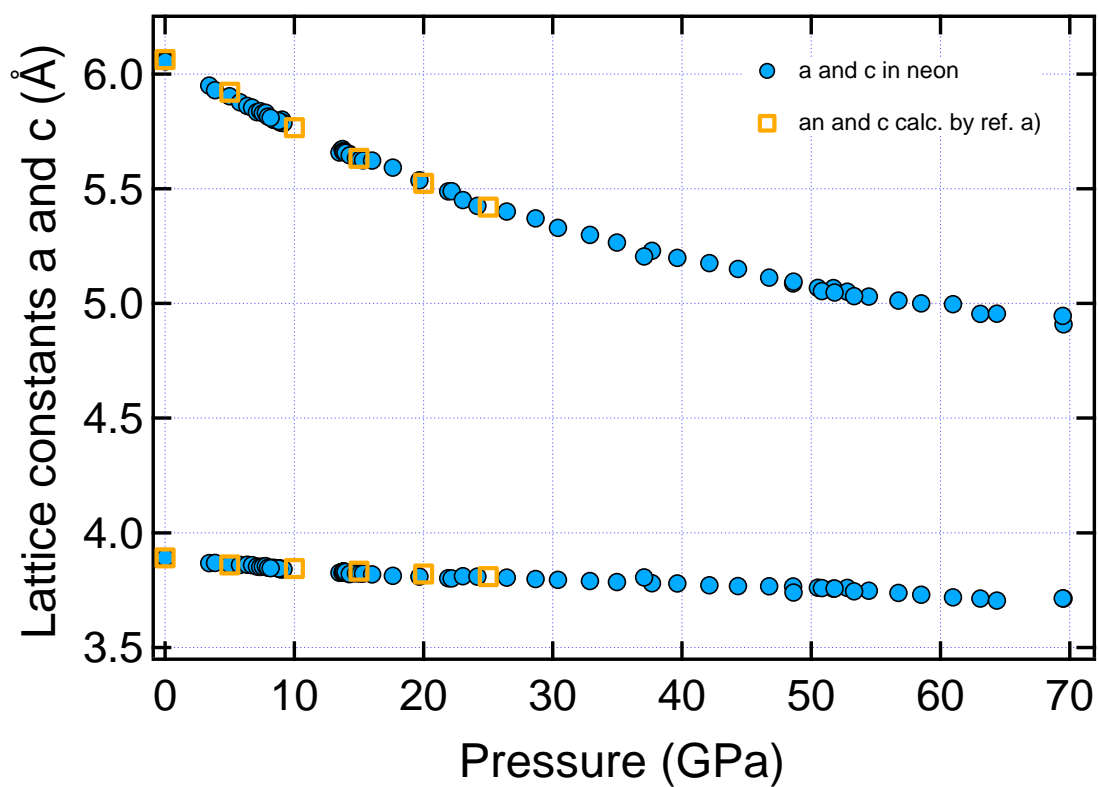


Fig. 4: Pressure dependence of the lattice constants a and c for Ce_2O_3 compressed quasi-hydrostatically with a neon pressure-transmitting medium. Ref a): [24].

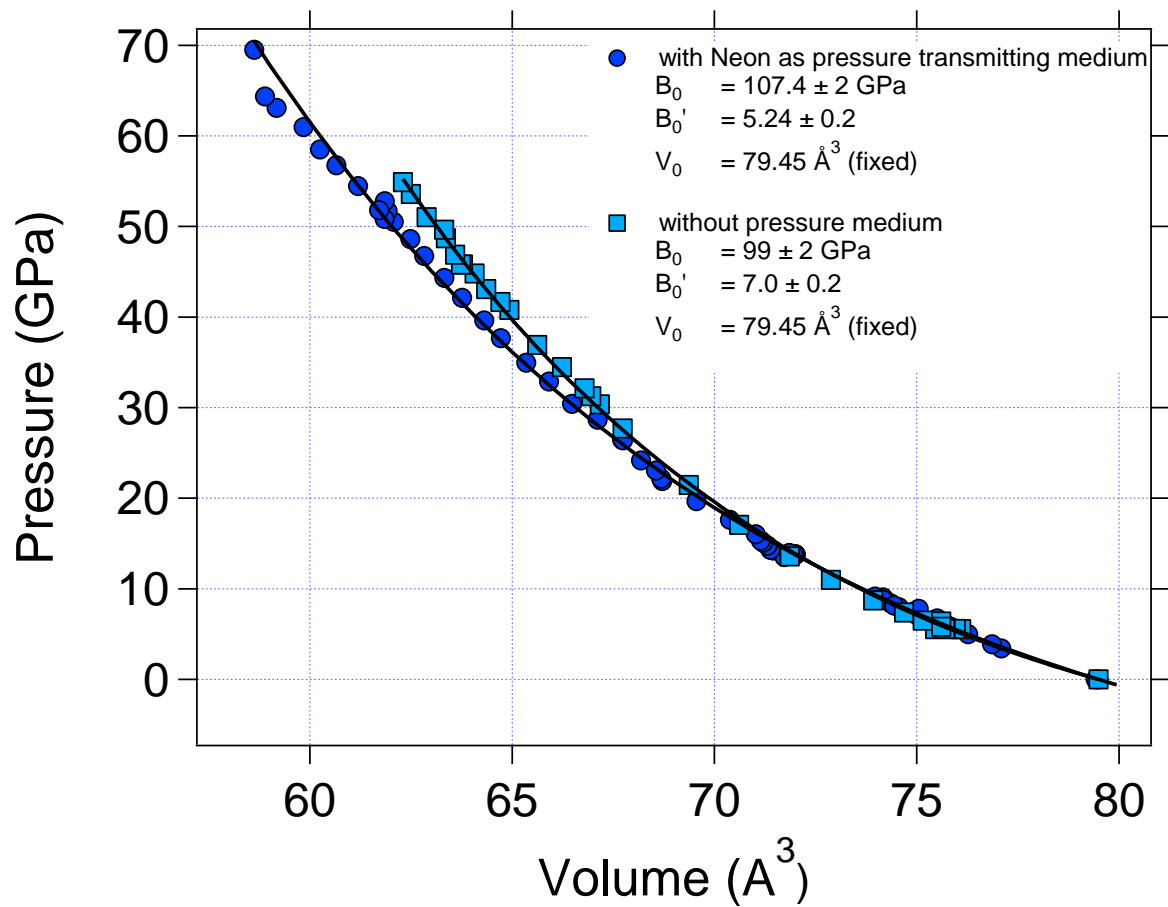


Fig. 5: Pressure-volume room temperature isotherms (equation of state – EOS) for Ce_2O_3 under quasi-hydrostatic (with neon as PTM) and non-hydrostatic conditions. The zero pressure volume was fixed at 79.45 \AA^3 for both.

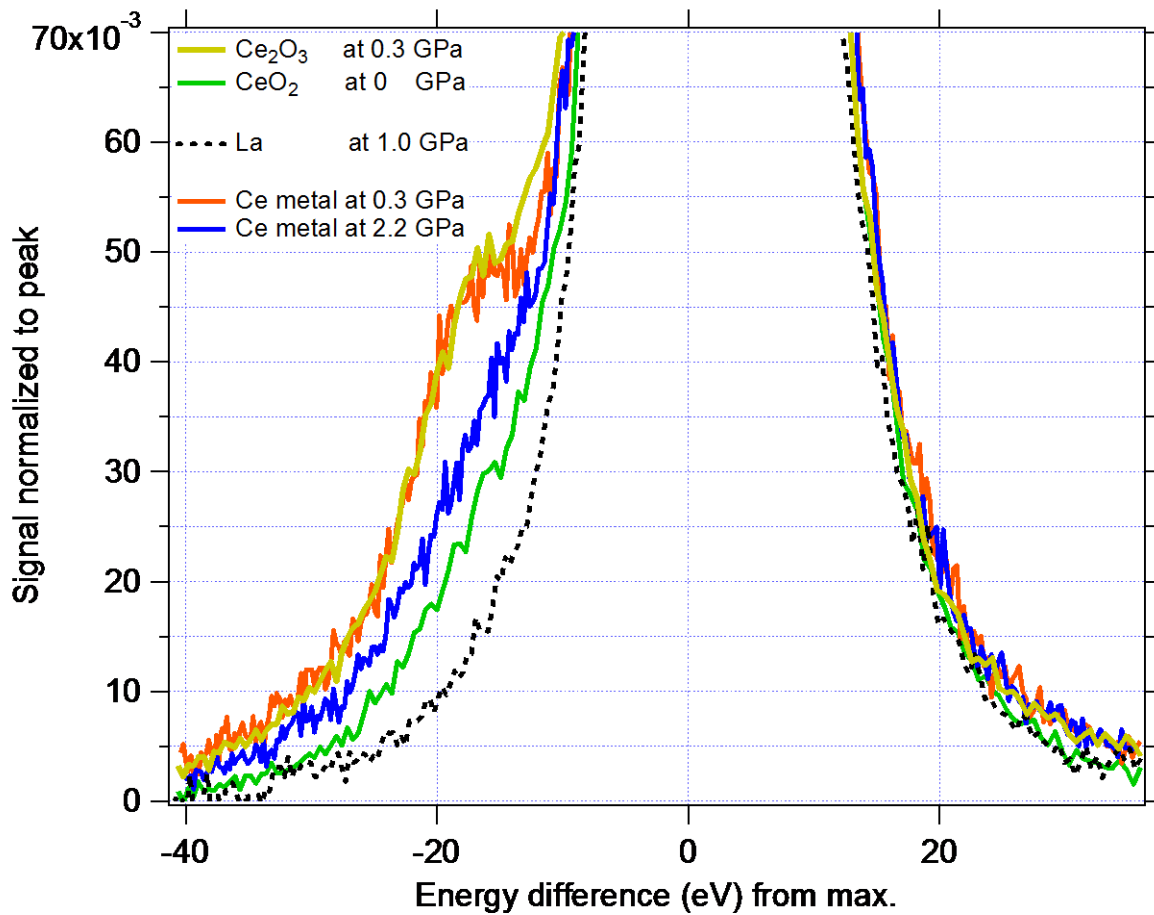


Fig. 6: X-ray emission spectra of the Ly emission for Ce_2O_3 , CeO_2 , La and metallic cerium at two different pressures – at 0.3 GPa below the 15% volume collapse and at 2.2 GPa above the volume collapse. With the exception of CeO_2 all samples were contained inside a DAC confined by a beryllium gasket.

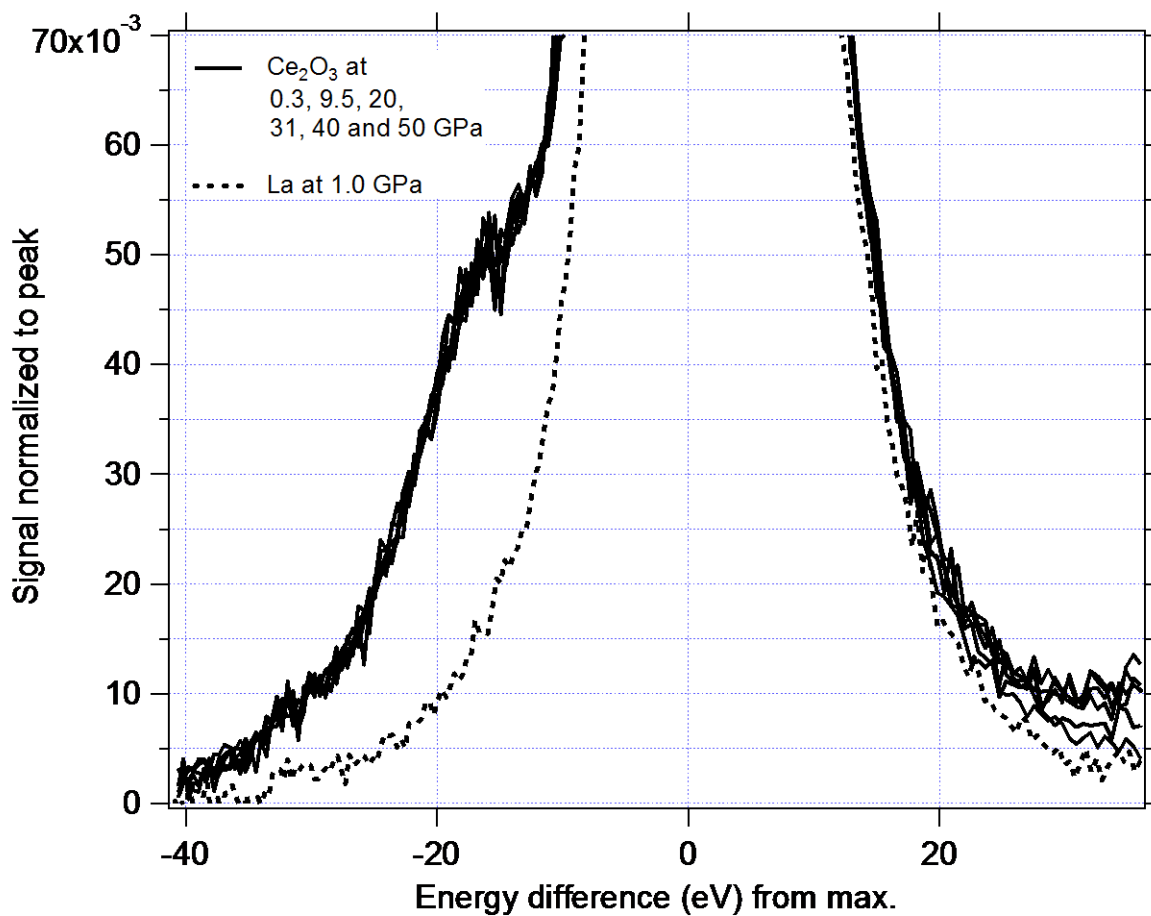


Fig. 7: Ly emission spectra for Ce_2O_3 at pressures of 0.3, 9.5, 20, 31, 40 and 50 GPa (solid lines). Also shown is the line-shape of the zero-f-electron element lanthanum. Within the experimental uncertainty the curves are indistinguishable from each other and do not change with pressure. The dip at -15 eV is an artefact.

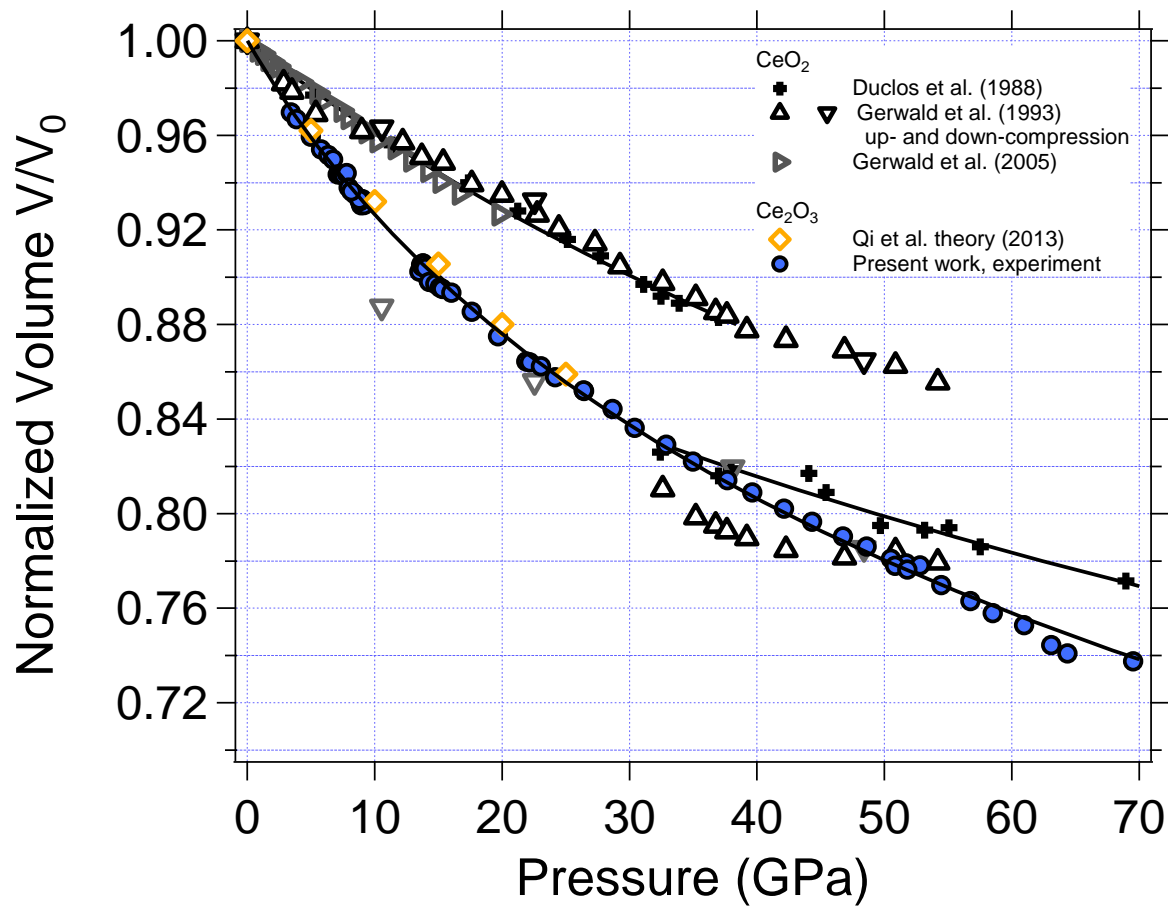


Fig. 8: Comparison of the normalized EOS for Ce_2O_3 and CeO_2 . The data by Duclos et al. [25] were taken without PTM, the data by Gerwald et al. with a PTM (4:1 methanol/ethanol mixture [26], 1993, and 16:3:1 methanol/ethanol/water mixture [29], 2005).

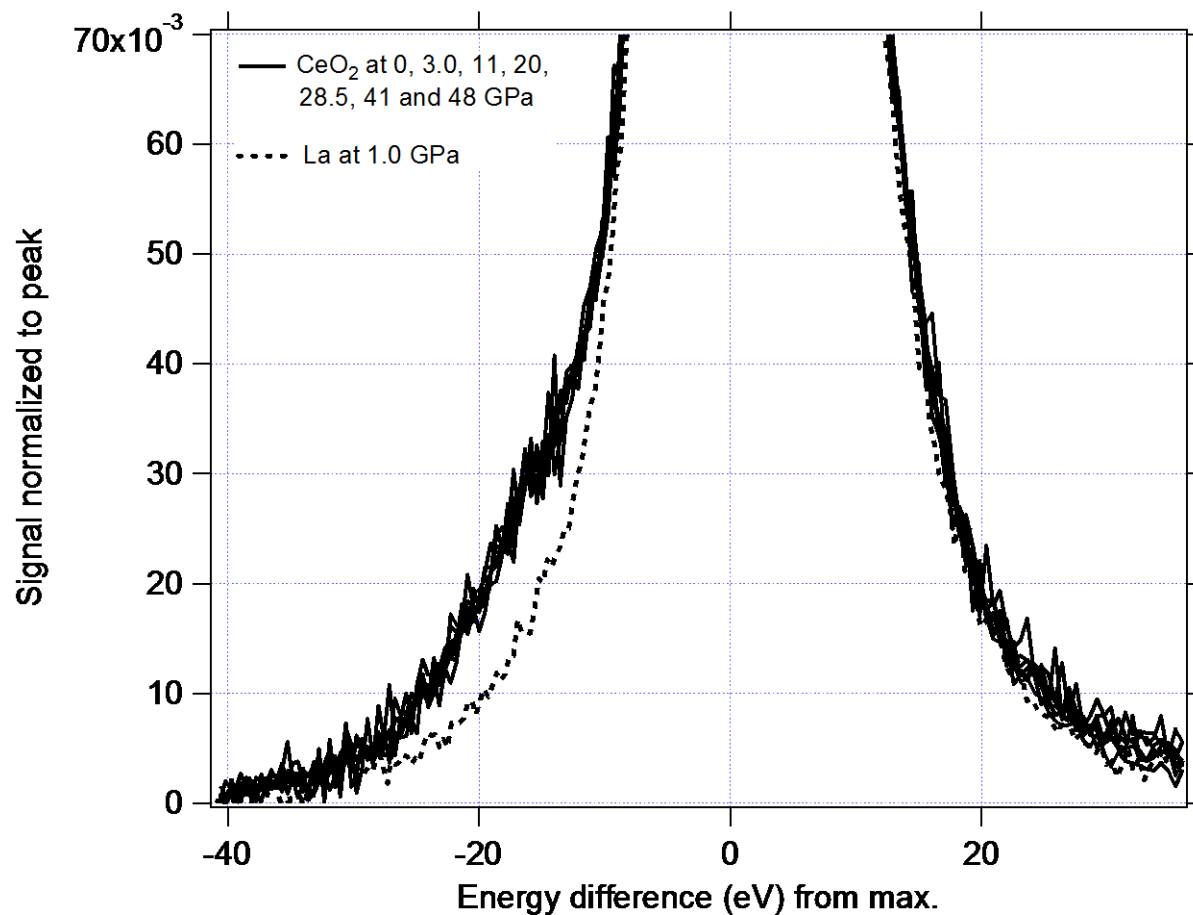


Fig. 9: Ly emission spectra for CeO_2 at pressures of 0, 3.0, 11, 20, 28.5, 41 and 48 GPa (solid lines). The line-shape of the zero-f-electron element lanthanum is again displayed for comparison. Within the experimental uncertainty the curves are indistinguishable from each other and as for Ce_2O_3 there is no discernible change with pressure.