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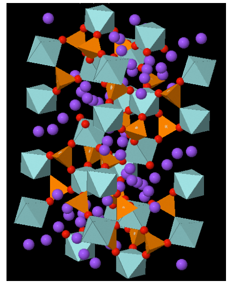

The Influences of Excess Sodium on NaSICON Materials Chemistry

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Nelson Bell, Cynthia Edney, Jill S. Wheeler,

PI: David Ingersoll

Thanks for generous support from Dr. Imre Gyuk and the DOE Office of Electricity Delivery and Energy Reliability

Electrical Energy Storage Applications and Technologies
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San Diego, CA USA

Exceptional service in the national interest

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Sodium-based Battery Development

<u>Program Focus</u>: Develop sodium-based battery chemistries for large scale energy storage

- Sodium-iodine, sodium-bromine, sodium-air, sodium-insertion, sodium-metal, *etc*.

Professor Eric Wachsman

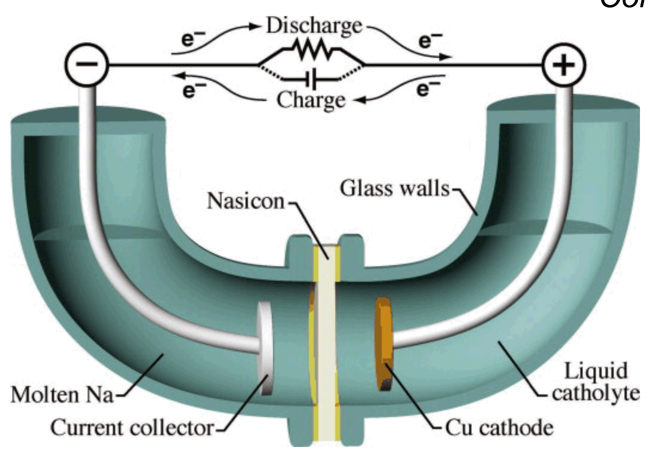
“Higher Conductivity NASICON Electrolyte for Room Temperature Solid-State Sodium Ion Batteries”

University of Maryland Energy Research Center

Dr. Sai Bhavaraju

“Development of Sodium-Iodine Battery for Large-Scale Energy Storage”

CERAMATEC TOMORROW'S CERAMIC SYSTEMS



Professor Robert Kee

“Computational modeling of sodium-copper-iodide secondary batteries”

Colorado School of Mines 1874

Dr. David Ingersoll (PI)

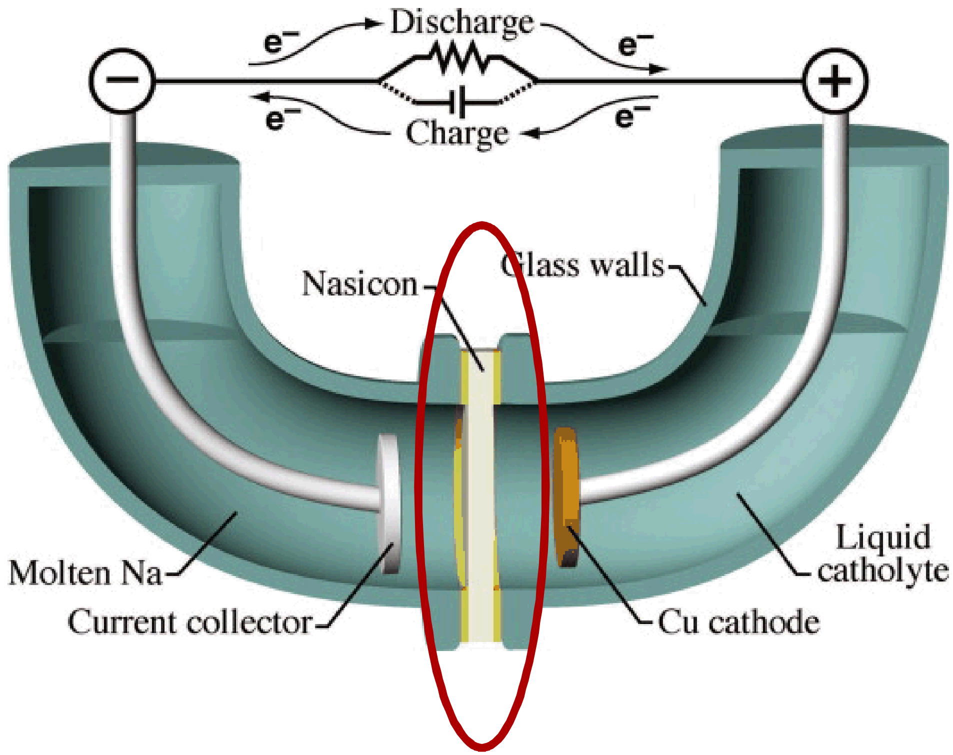
“Sodium Based Battery Development”

Sandia National Laboratories

H. Zhu, S. Bhavaraju, and R. Kee. “Computational model of a sodium–copper-iodide rechargeable battery,” Electrochimica Acta (2013), http://dx.doi.org/10.1016/j.electacta.2013.09.010

Ceramic Solid State Electrolyte Separators

The ceramic separator is central to Na-battery performance!



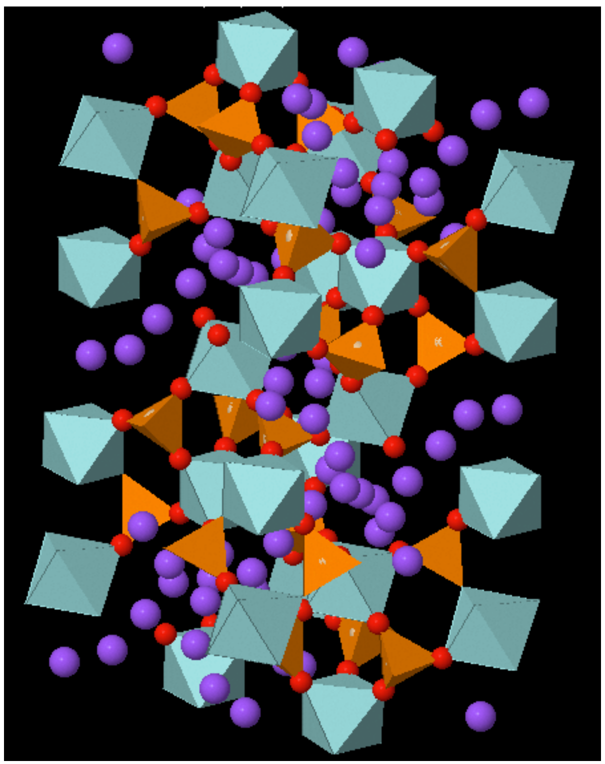
<u>Ceramic requirements:</u>

- High ionic conductivity
- High electrical resistivity
- Robust stability in extreme chemical environments
- Facile, low cost synthesis

NaSICON Ceramic Electrolytes

What is NaSICON? (Sodium (**Na**) **S**uper **I**onic **Con**ductor)

$$Na_{1+x}Zr_2P_{3-x}Si_xO_{12} \xrightarrow{x = 2} Na_3Zr_2PSi_2O_{12}$$



Key NaSICON attributes:

- High ionic conductivity (up to 10^{-2} S/cm at RT)
- High electrical resistivity
- Robust stability in extreme chemical environments **?**
- Facile, low cost synthesis **?**

These qualities all depend on the materials chemistry of the ceramic!

Task Focus: NaSICON Ceramic Solid State Electrolytes

Sandia National Laboratories

- Understanding the materials chemistry of the solid-state ion-conductor NaSICON
- Correlating material chemistry to materials properties (e.g., chemical stability, ionic conductivity, ceramic integrity)
- Designing improvements to NaSICON through processing and composition to optimize performance for Na-based batteries

Our approach takes advantage of materials chemistry and characterization capabilities at SNL to enable innovative improvements in energy storage systems.

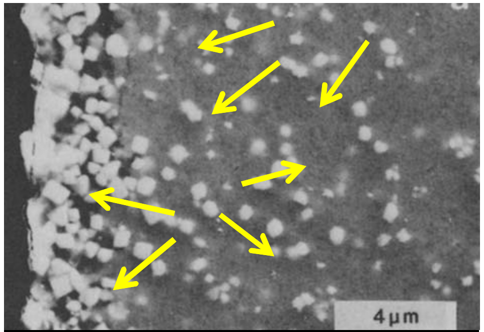
NaSICON Materials Chemistry

NaSICON performance depends on phase chemistry!

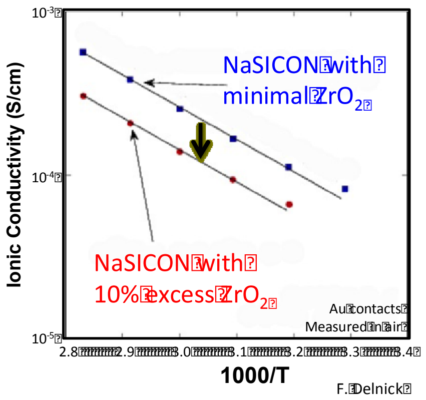
Secondary phase formation can have a significant impact on:

- *ionic conductivity*
- *structural integrity*
- *chemical stability*

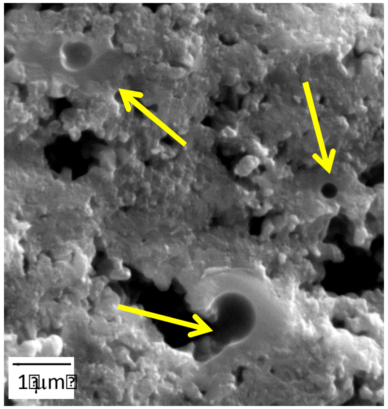
ZrO_2
(monoclinic and tetragonal)



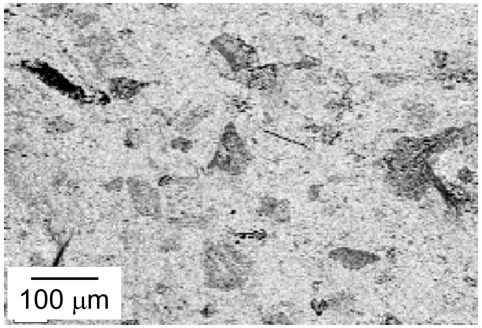
Ahmad, et al. *Sol. St. Ionics*. (1987) 24: 89-



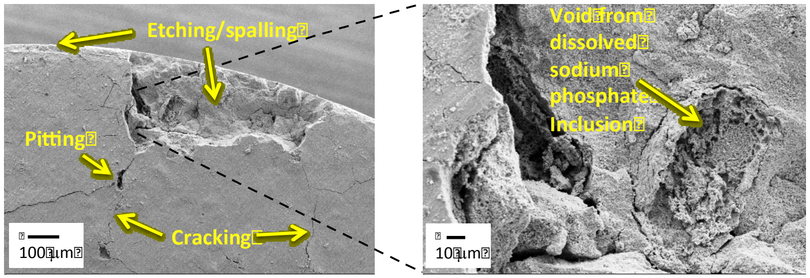
Glassy inclusions



Sodium silicates



Sodium phosphates



Phase Dependence on Processing

Phase composition of NaSICON depends on processing

- Solid state processing of NaSICON ceramics typically involves an extended high temperature firing stage (>1200°C, >12 hours)

“Decomposition” of NaSICON

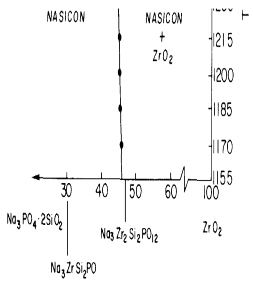


Fig. 1. Tentative phase relations in the Nasicon–ZrO_2 phase field (from X-ray analyses of sintered specimens).

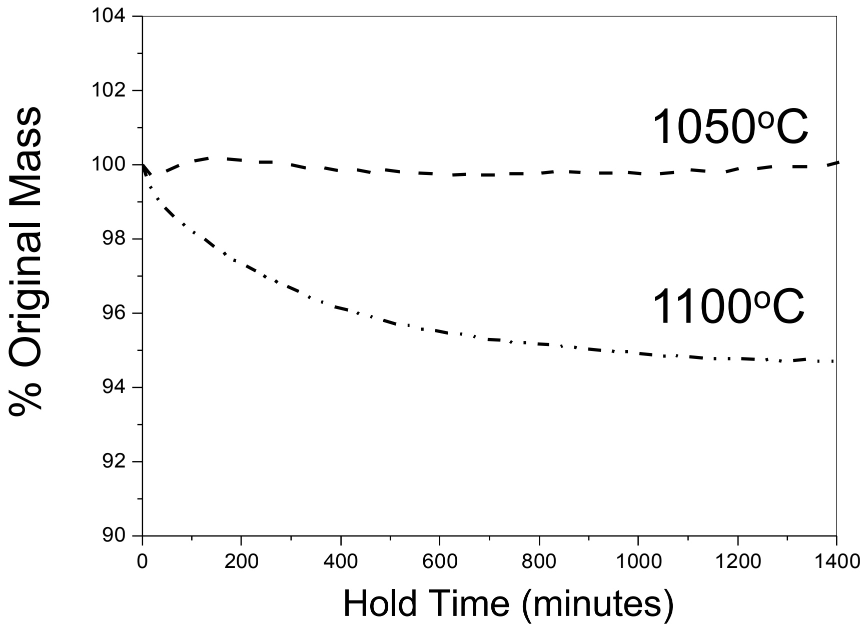
Another potentially economical approach [4] for the formation of Nasicon is the mechanical mixing, calcination (1150–1160°C), milling and subsequent sintering of $ZrSiO_4$ and Na_3PO_4 powder mixtures. Bar samples of milled

ponents. These mixtures were heated between 90 and 120°C where a glassy transparent matrix formed due to the presence of the polyfunctional acid. The mixtures were then pyrolized to their component oxides by heating to 400°C for 4 h. Residual carbon was removed by calcining the material at 900°C. The resultant “soft cakes” were then milled using dense α-alumina media leading to a mean particle size of 2 μm. Bar samples were uniaxially pressed and isostatically pressed. Sintering in air was performed at temperatures between 1175 and 1300°C. X-ray examination of the calcined powder indicated that the material was basically non-crystalline. Two detectable phases were present: Nasicon and trace amounts of ZrO_2. The $x = 2.3$ composition had only a trace of ZrO_2 while the $x = 2.0$ composition had slightly more.

An evaluation of selected sintering conditions for the two compositions studied is given in table 1. The existence of single-phase Nasicon is

R.S. Gordon, *et al. Solid State Ionics.* ***3/4*** *(1981) 243-248.*

Loss of volatile species (e.g., Na and P)

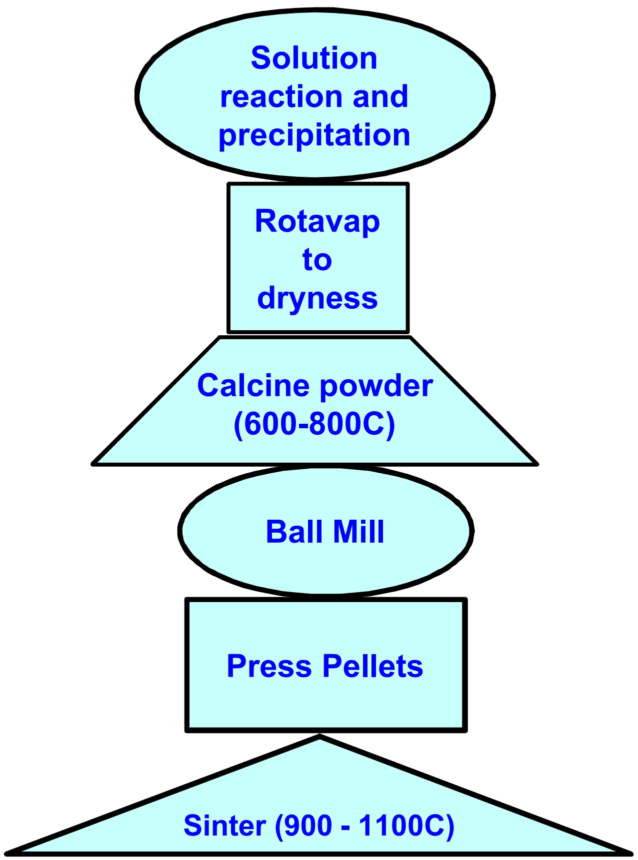


High temperature processing leads to deleterious secondary phases!

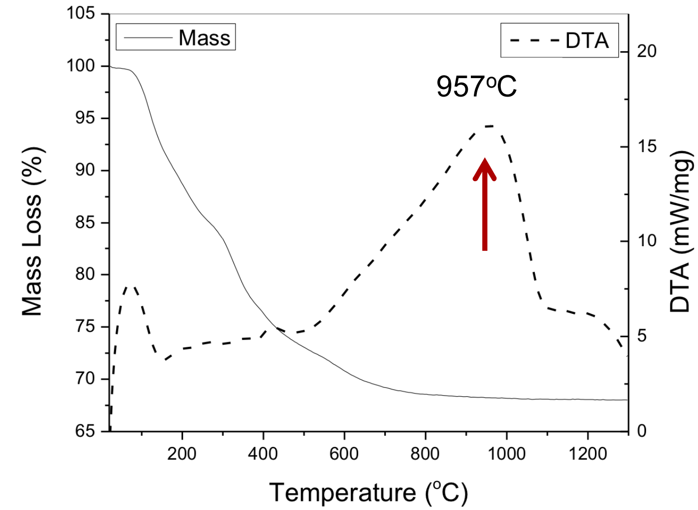
Will a lower temperature process resolve phase impurity?

“Low” Temperature Sol-Gel NaSICON

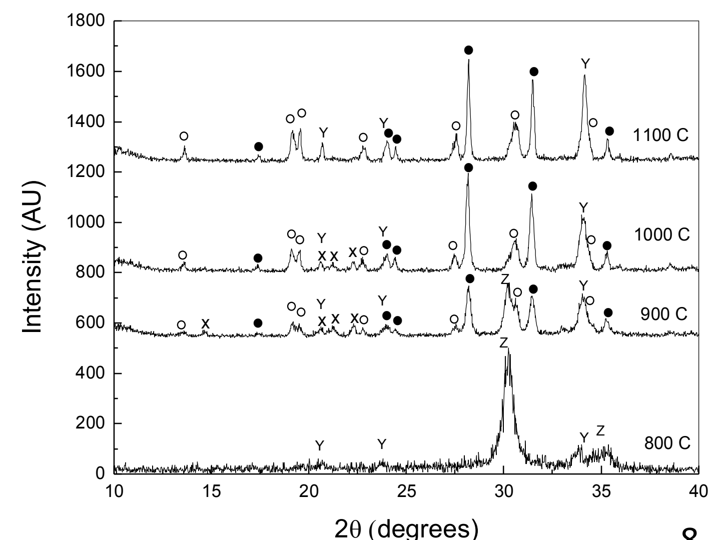
Sol-gel processing



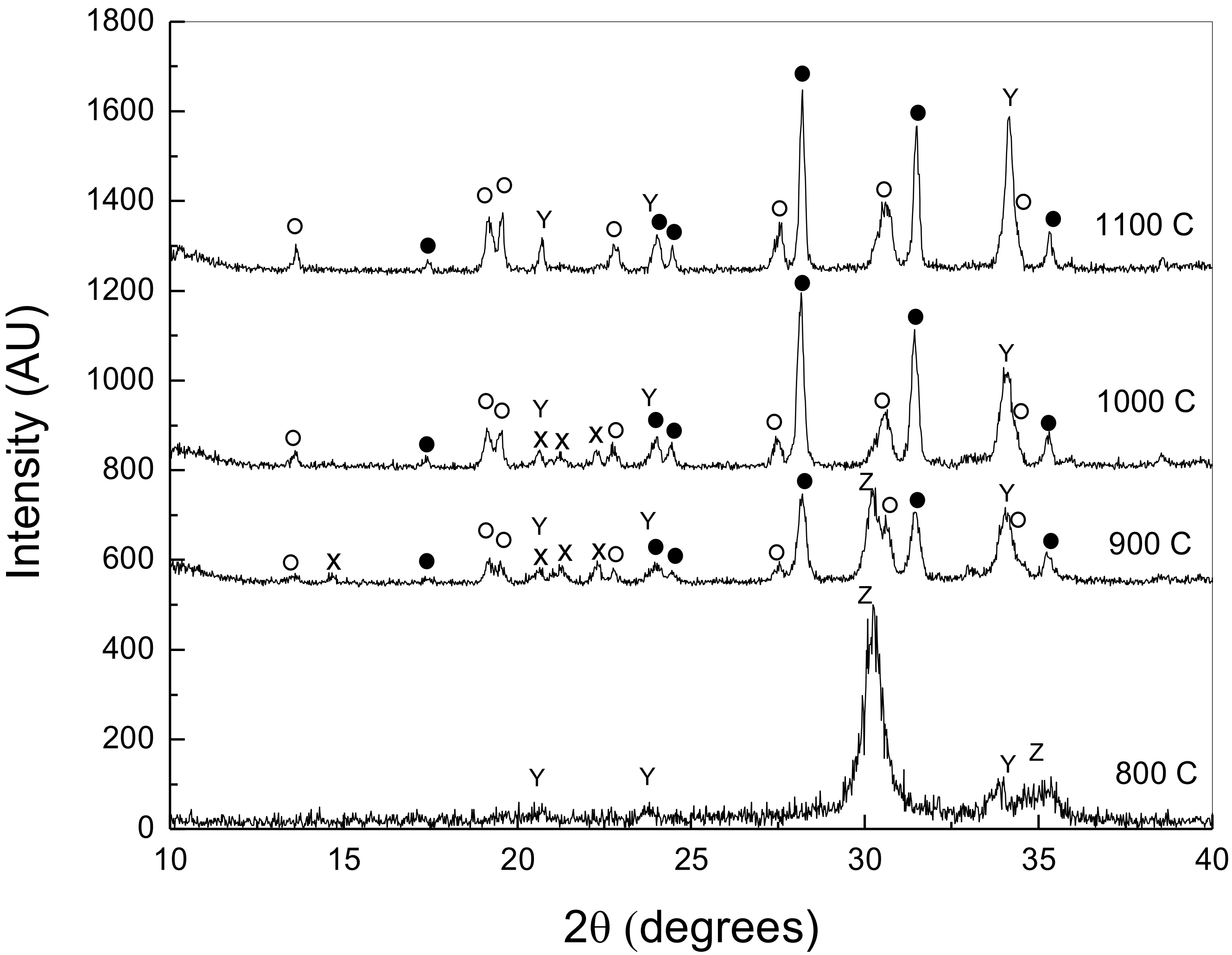
Thermal analysis identifies NaSICON formation temperature.



X-ray diffraction shows evolution of crystalline phases.



Sol-Gel NaSICON Phase Evolution



X-ray diffraction shows the presence of $Na_3Zr_2PSi_2O_{12}$ (o) *and*

- Tetragonal ZrO_2 (Z)
- Monoclinic ZrO_2 (•)
- Na_3PO_4 (Y)
- $Na_2Si_2O_5$ (X)

secondary phases.

Monoclinic ZrO_2 appears to form from conversion of metastable tetragonal ZrO_2.

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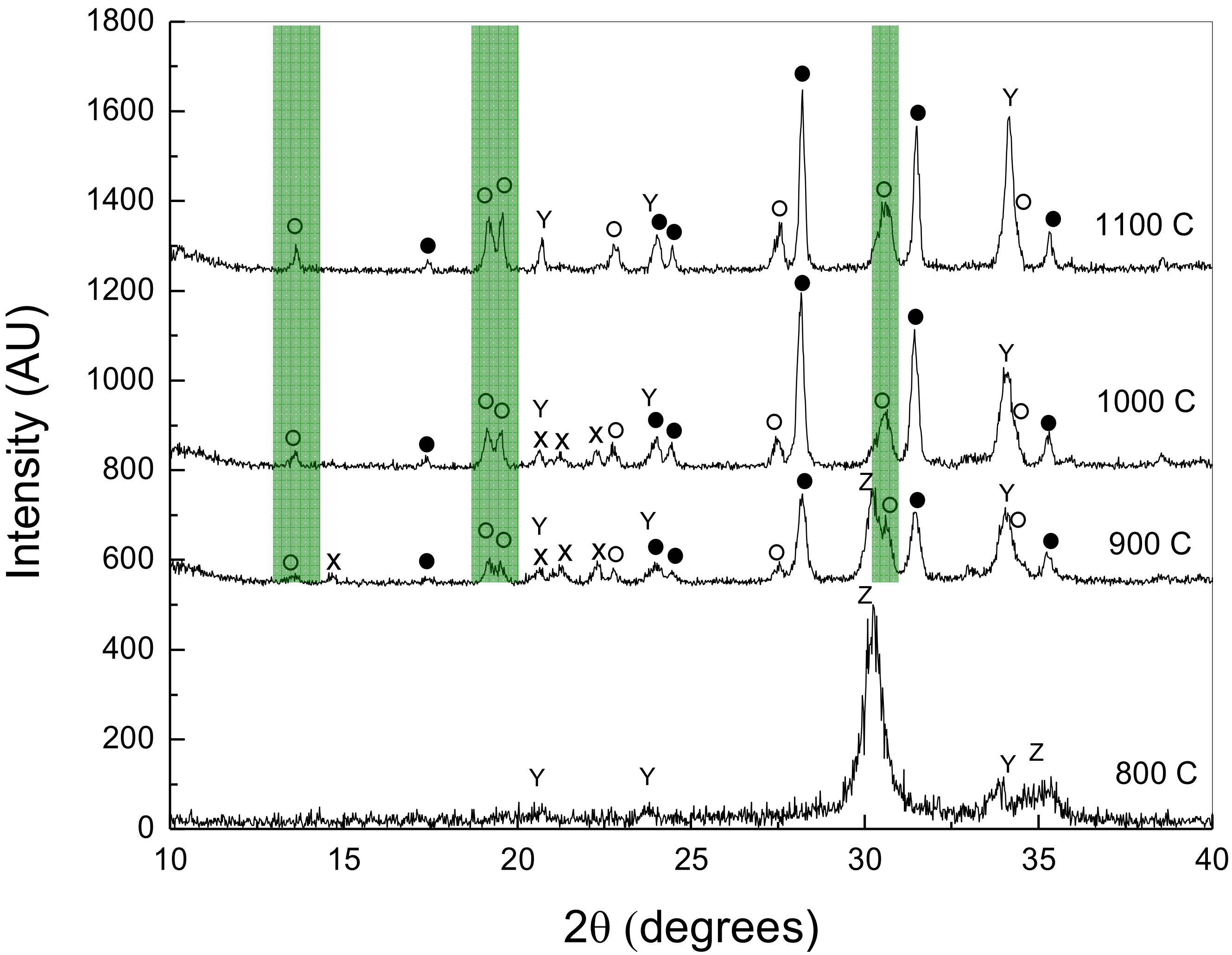

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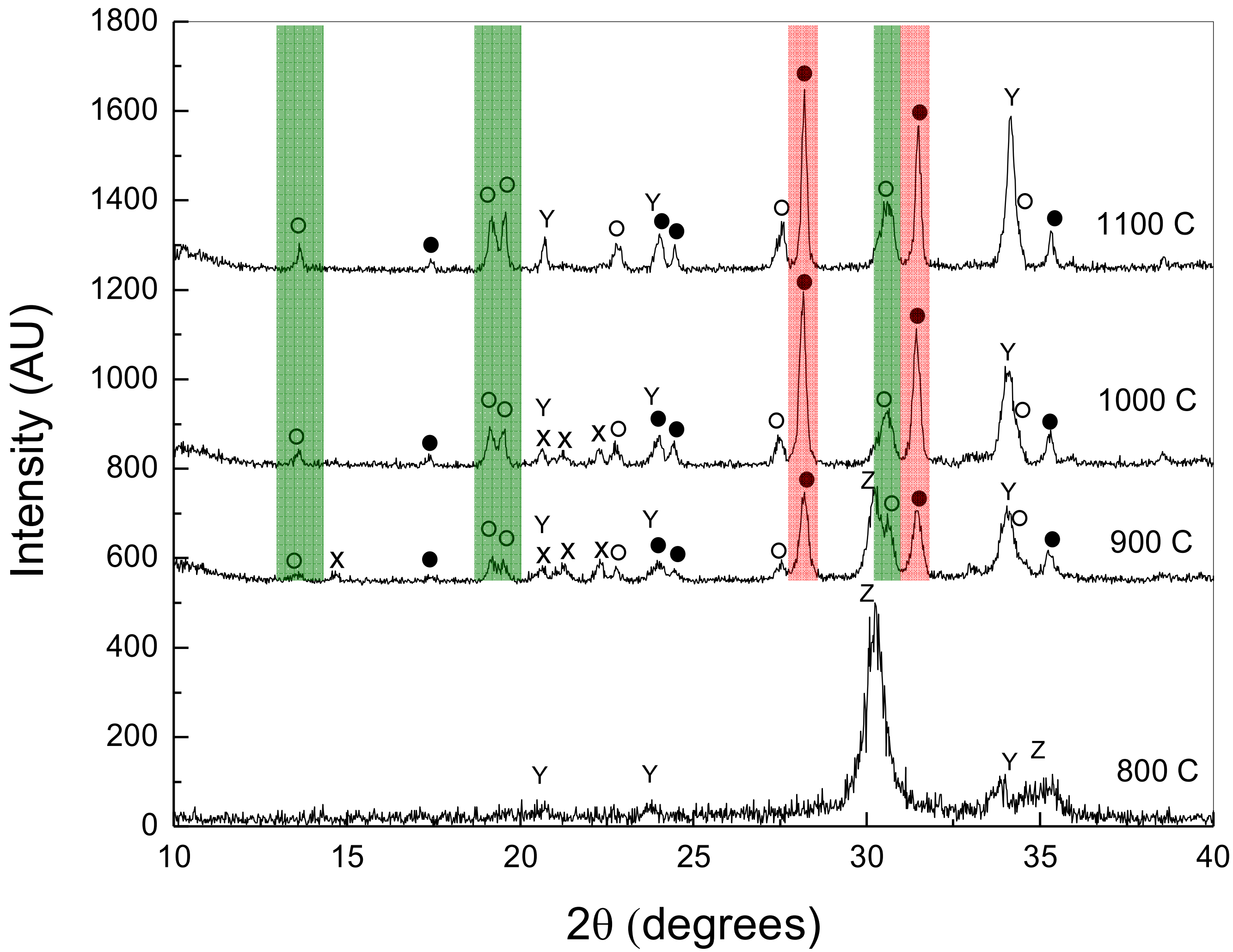

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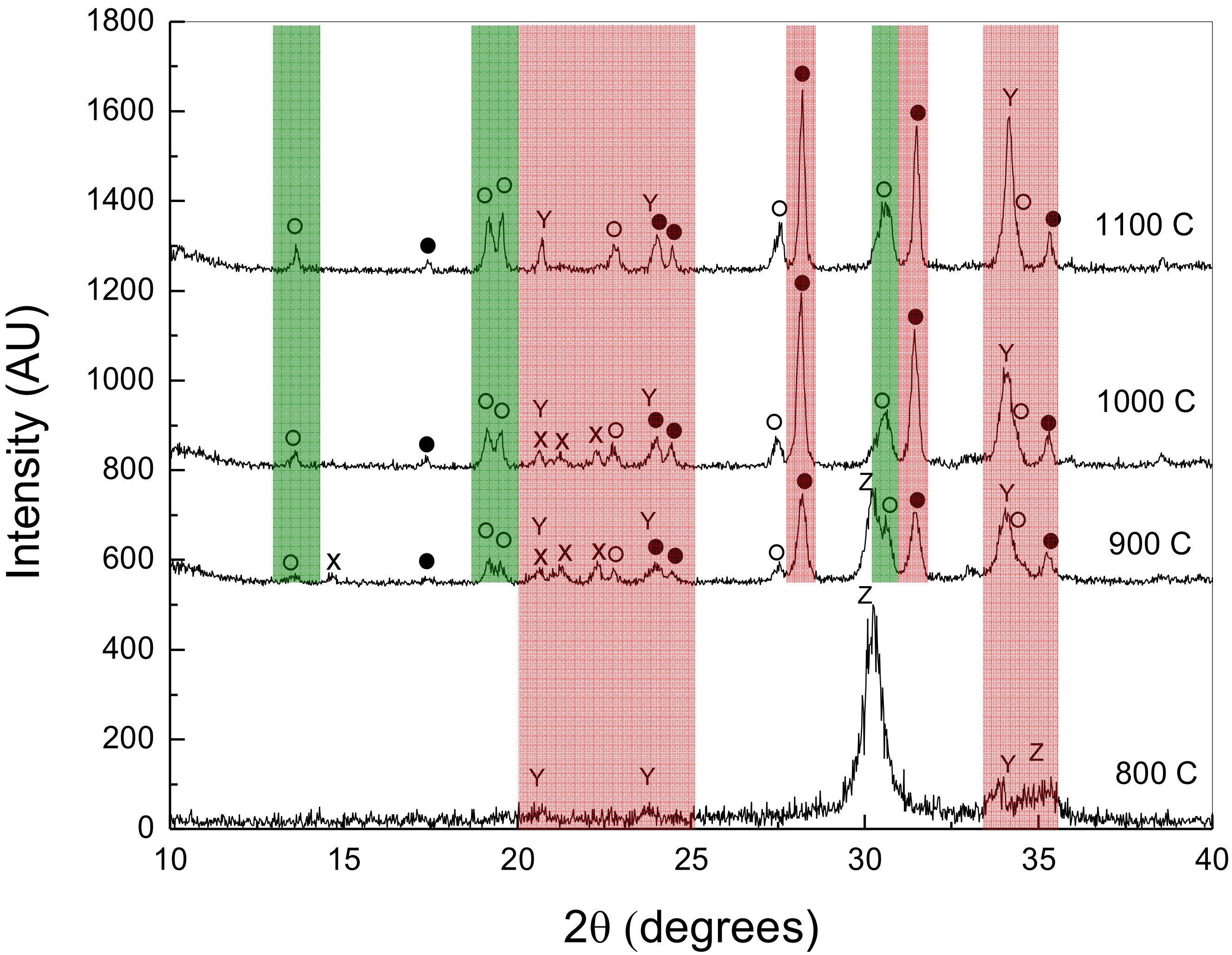

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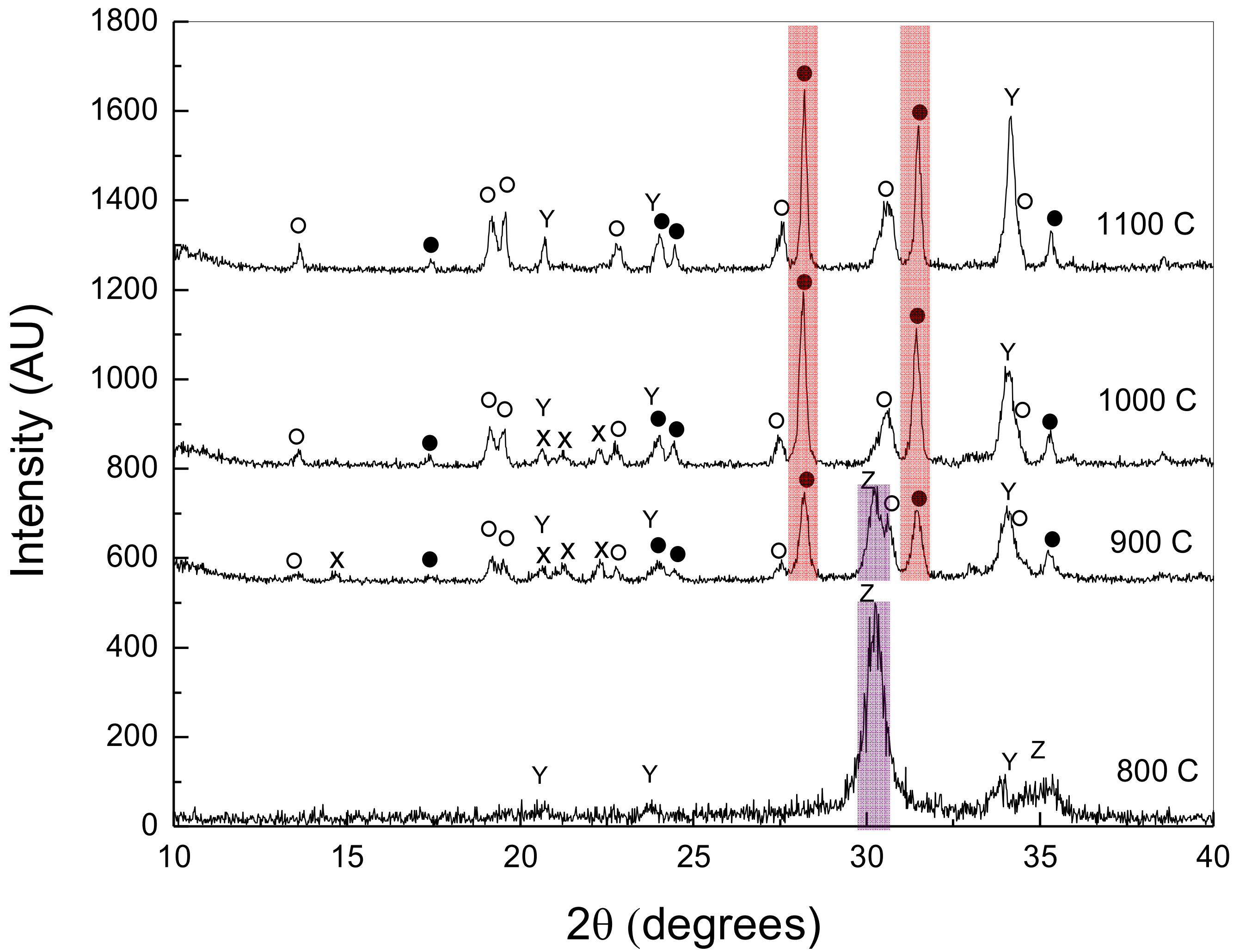

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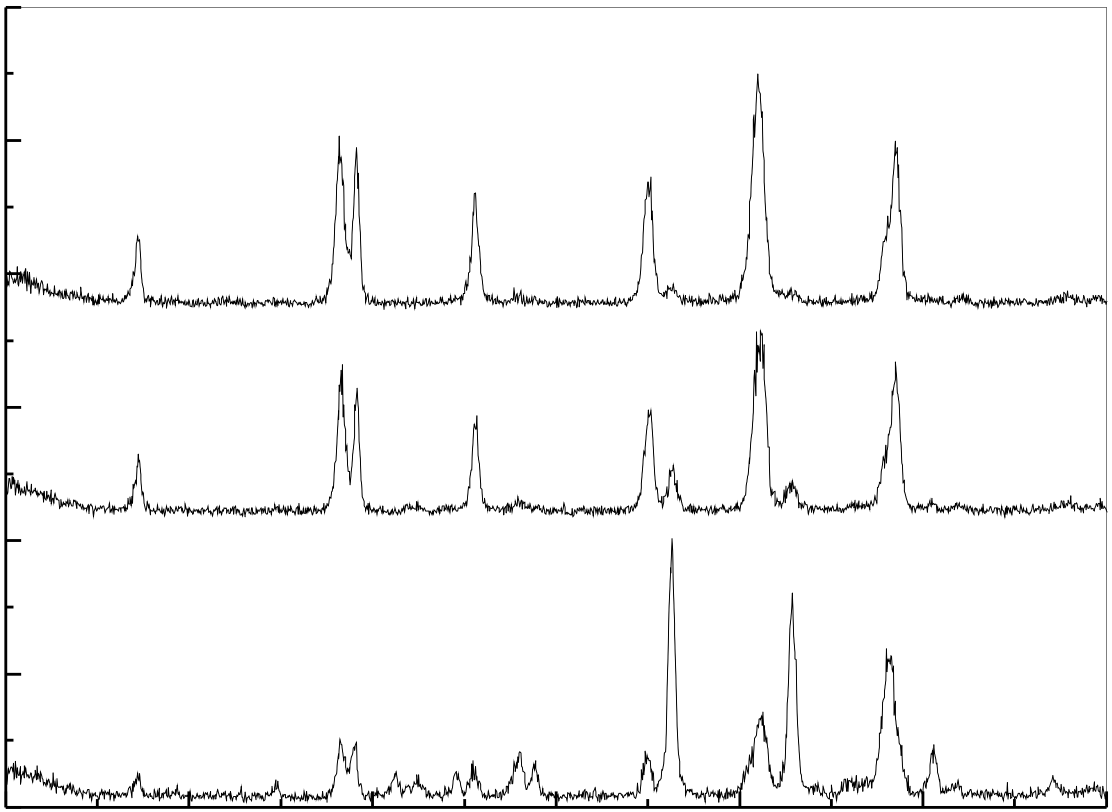
Lessons from Low Temperature Processing

- Phase evolution during heating is complex!
- Lower processing temperatures result in significant secondary phase formation.
- Secondary phase are not formed just from high temperature processes, but can be residual from incomplete low temperature conversions.
- Higher temperatures appear to be needed for complete phase conversion, but high T°C is expected to lead to secondary phases.

What Next?

Excess Sodium Addition

NaSICON with excess sodium fired at 1000°C shows dramatically cleaner phase chemistry!



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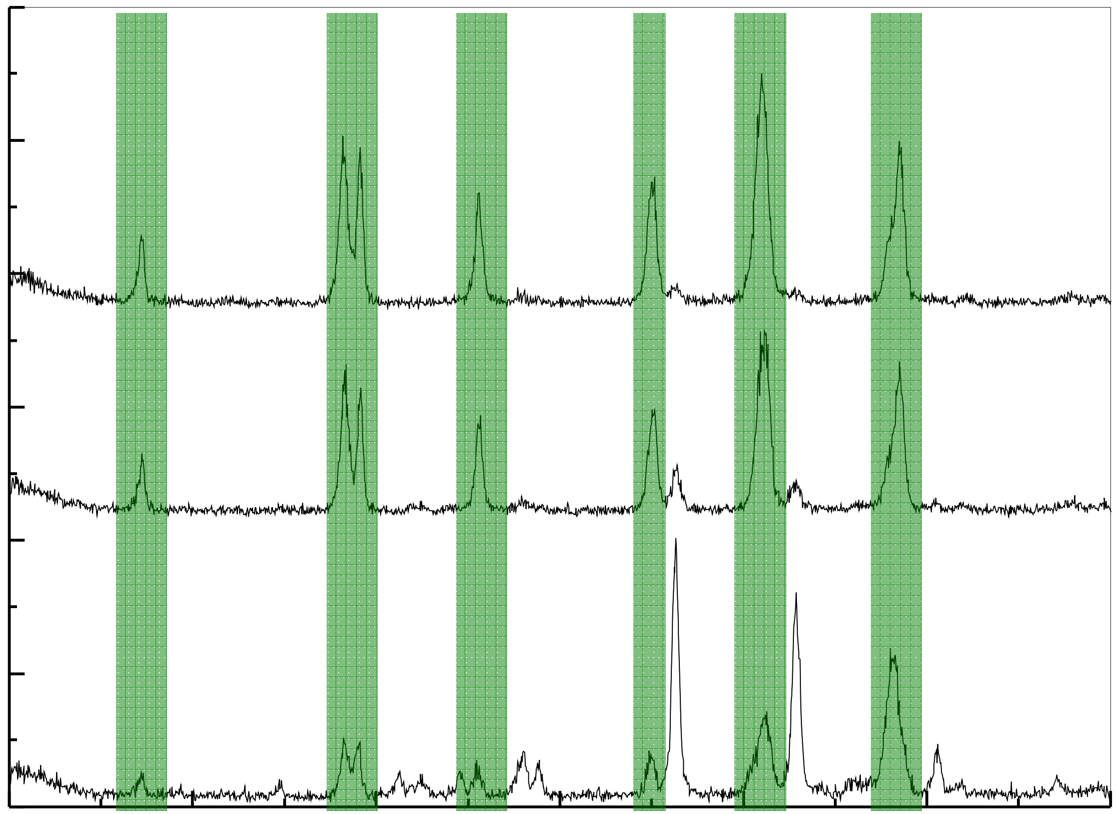

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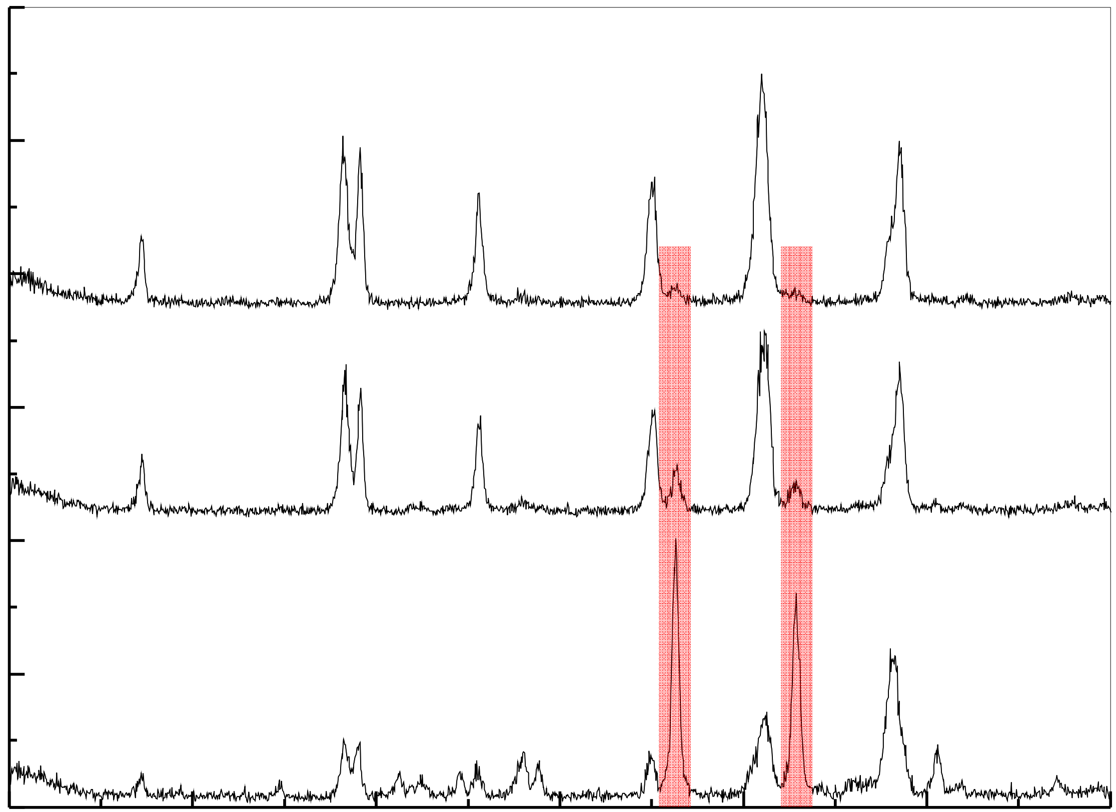
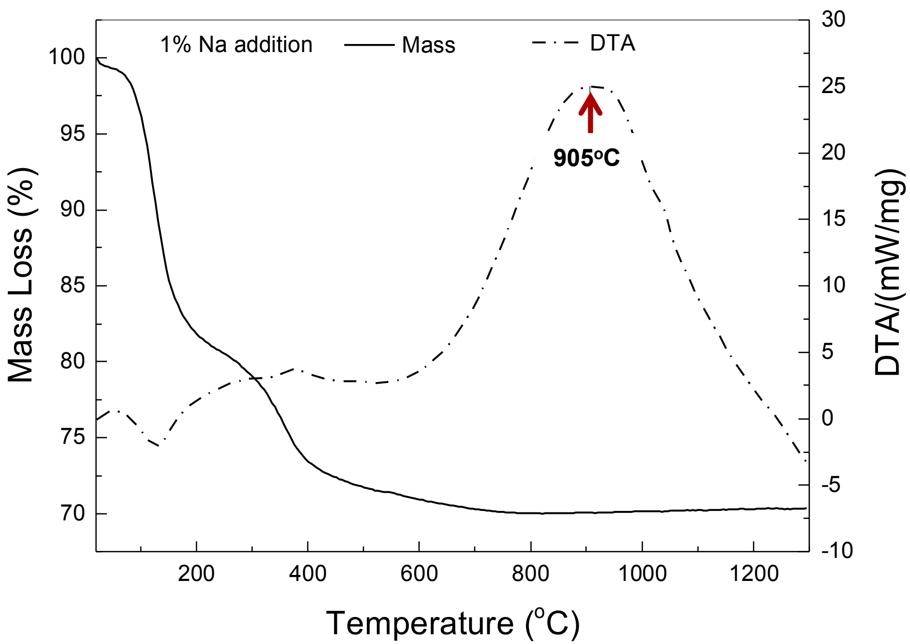
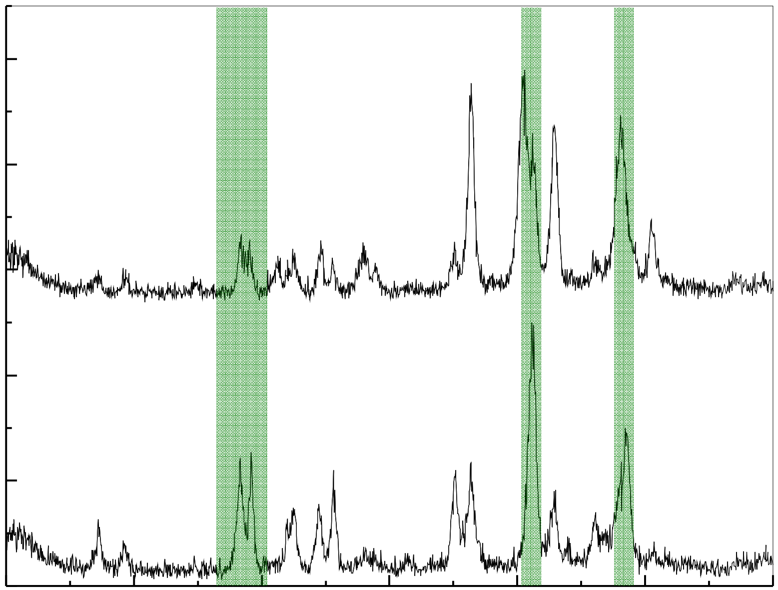


Excess Sodium Reduces Effective Processing Temperature

Thermal Analysis and XRD show NaSICON formation at lower temperatures with excess Na!





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Excess sodium addition appears to change the energetics of NaSICON conversion, likely by affecting mass transport in liquid phase elements of sintering.

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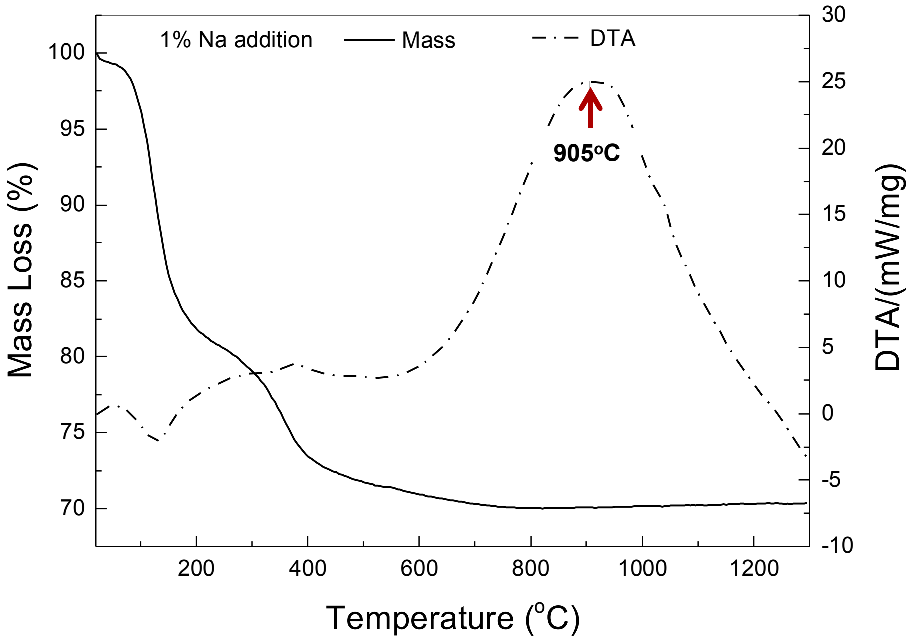

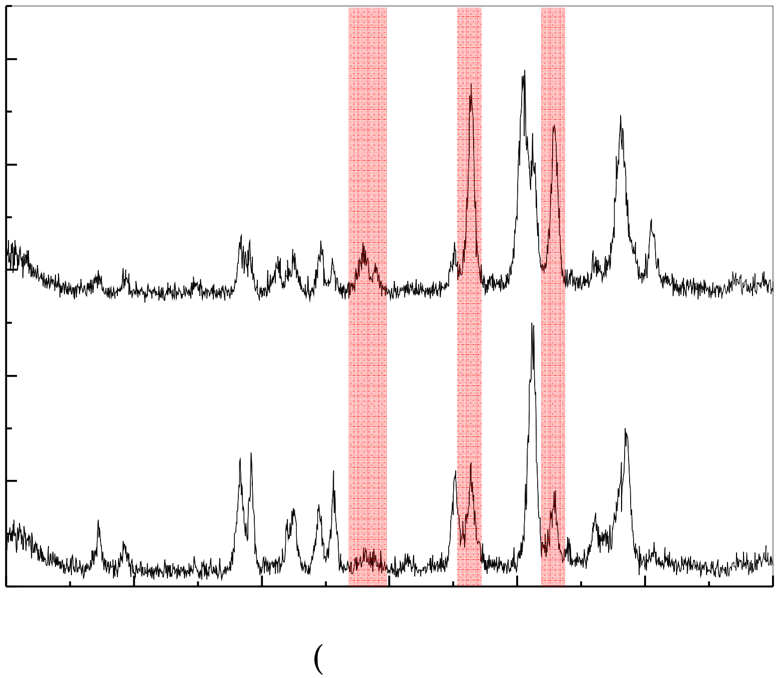

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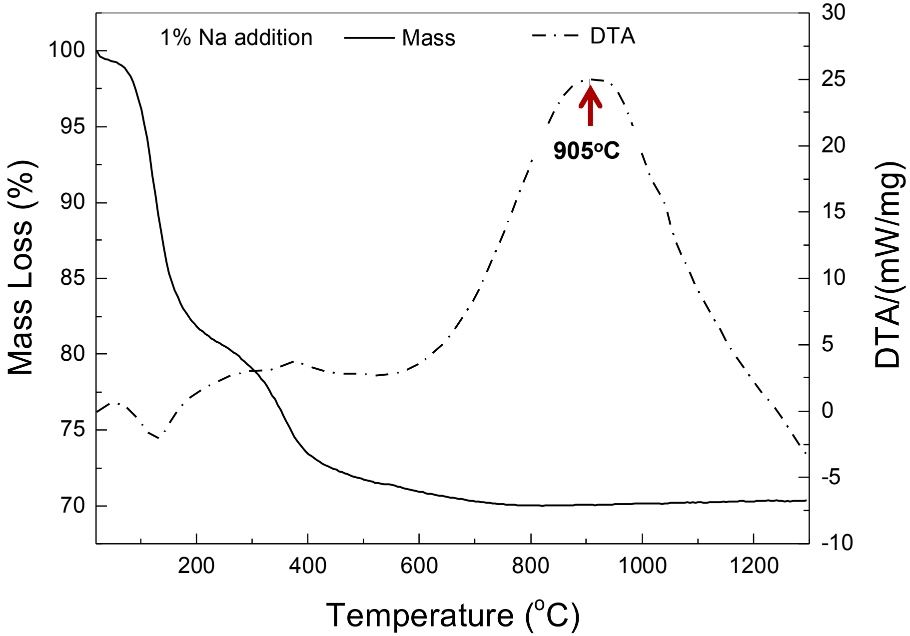

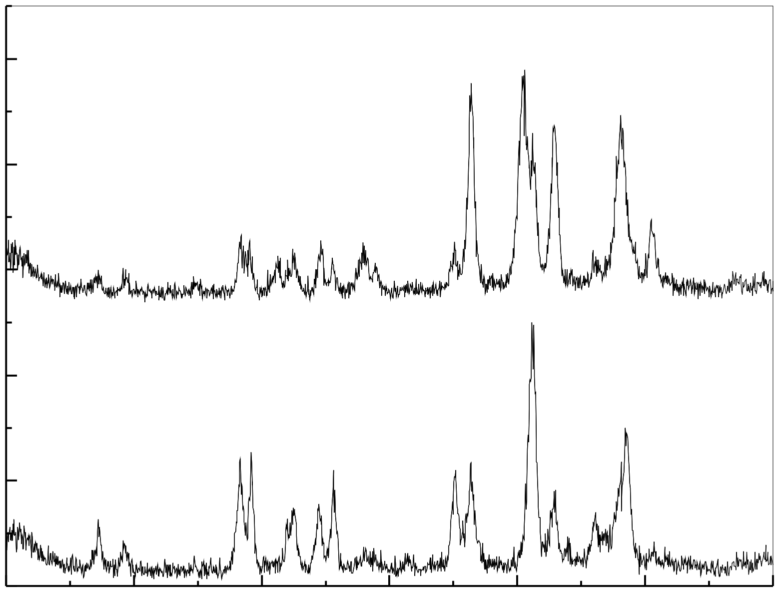

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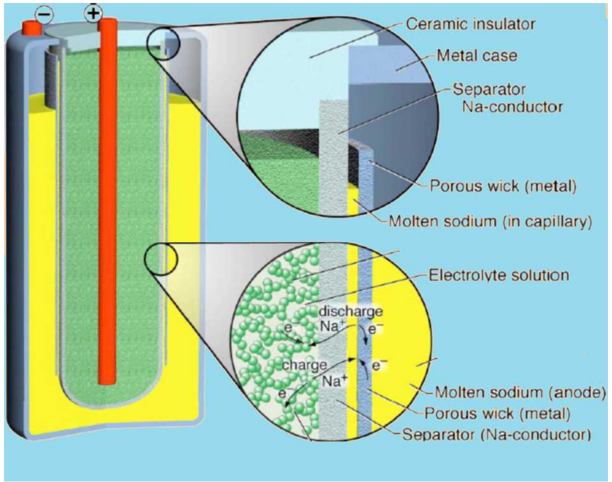
Summary/Conclusions

- NaSICON ceramics are *enabling* solid state electrolytes for *grid-scale and vehicle* Na-based batteries.
- Controlling secondary phase chemistry is critical to optimizing NaSICON performance.
- Reducing processing temperatures does not improve NaSICON phase purity.
- **Addition of small amounts of excess sodium dramatically reduces secondary phase formation at lower temperatures!**

Future Tasks

Targeting synthesis of improved NaSICON stability to enable integration into next generation Na-based batteries:

- Explore alternative mechanisms to reduce processing temperatures with high phase purity.
- Investigate alternative precursor pathways to control phase chemistry.
- Evaluate effects of phase chemistry on sodium ion transport/conductivity.
- Examine chemical stability of NaSICON as affected by additives (such as sodium).



Contact and Acknowledgements

- Dr. Erik D. Spoerke: edspoer@sandia.gov
- Principal Investigator: Dr. David Ingersoll: dingers@sandia.gov

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N.S. Bell, C. Edney, D. Ingersoll, and E.D. Spoerke, "The Influences of Excess Sodium on Low Temperature NaSICON Synthesis," *J. Mater. Chem. (*2013, in review).

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