

Conf 413-113 /

LA-UR- 93-2129

Title:

THERMODYNAMICS AND SUBLIMATION CHEMISTRY OF PLUTONIUM -
OXYGEN - CHLORINE

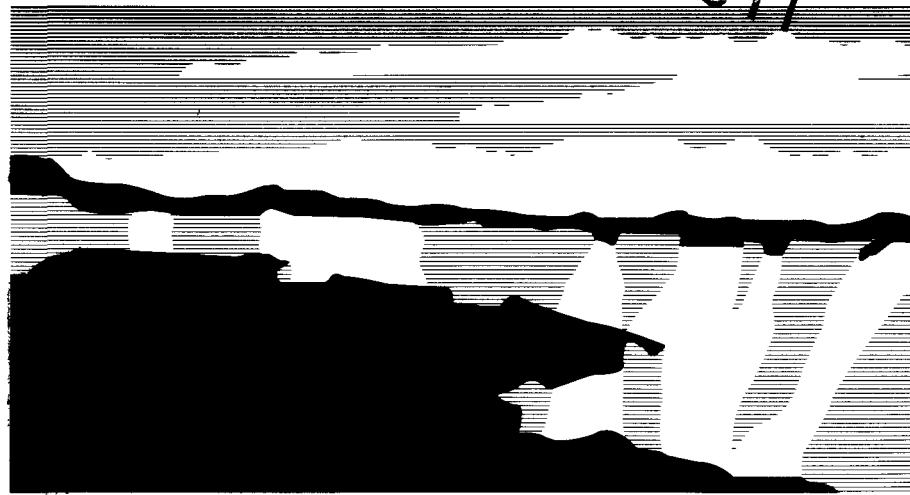
Author(s):

M. A. Williamson, and P. D. Kleinschmidt

Submitted to:

15th Midwest High Temperature & Solid State Chemistry Conf.
June, 14-16, 1993
Cornell University, Ithaca, NY

MASTER



Los Alamos
NATIONAL LABORATORY

Los Alamos National Laboratory, an affirmative action/equal opportunity employer, is operated by the University of California for the U.S. Department of Energy under contract W-7405-ENG-36. By acceptance of this article, the publisher recognizes that the U.S. Government retains a nonexclusive, royalty-free license to publish or reproduce the published form of this contribution, or to allow others to do so, for U.S. Government purposes. The Los Alamos National Laboratory requests that the publisher identify this article as work performed under the auspices of the U.S. Department of Energy.

1. DISTRIBUTION OF THIS DOCUMENT IS UNLIMITED

RP Form No 836 R5
ST 2629 10/91

DISCLAIMER

This report was prepared as an account of work sponsored by an agency of the United States Government. Neither the United States Government nor any agency Thereof, nor any of their employees, makes any warranty, express or implied, or assumes any legal liability or responsibility for the accuracy, completeness, or usefulness of any information, apparatus, product, or process disclosed, or represents that its use would not infringe privately owned rights. Reference herein to any specific commercial product, process, or service by trade name, trademark, manufacturer, or otherwise does not necessarily constitute or imply its endorsement, recommendation, or favoring by the United States Government or any agency thereof. The views and opinions of authors expressed herein do not necessarily state or reflect those of the United States Government or any agency thereof.

DISCLAIMER

Portions of this document may be illegible in electronic image products. Images are produced from the best available original document.

Thermodynamics and Sublimation Chemistry of PuOCl

**M.A. Williamson and P.D. Kleinschmidt
Nuclear Materials Technology Division
Los Alamos National Laboratory
Los Alamos, New Mexico**

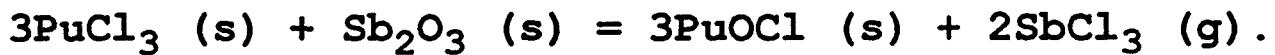
Introduction

Pyrochemical processing of plutonium produces spent salts that contain plutonium bearing species. One of the species is PuOCl. The behavior of pure PuOCl at high temperatures and PuOCl in the liquid salt matrix is not well characterized. The purpose of the study is to establish the vapor pressure and vapor composition for the sublimation of PuOCl. The results of the sublimation study and the thermodynamic properties associated with the sublimation process will be presented.

Experimental

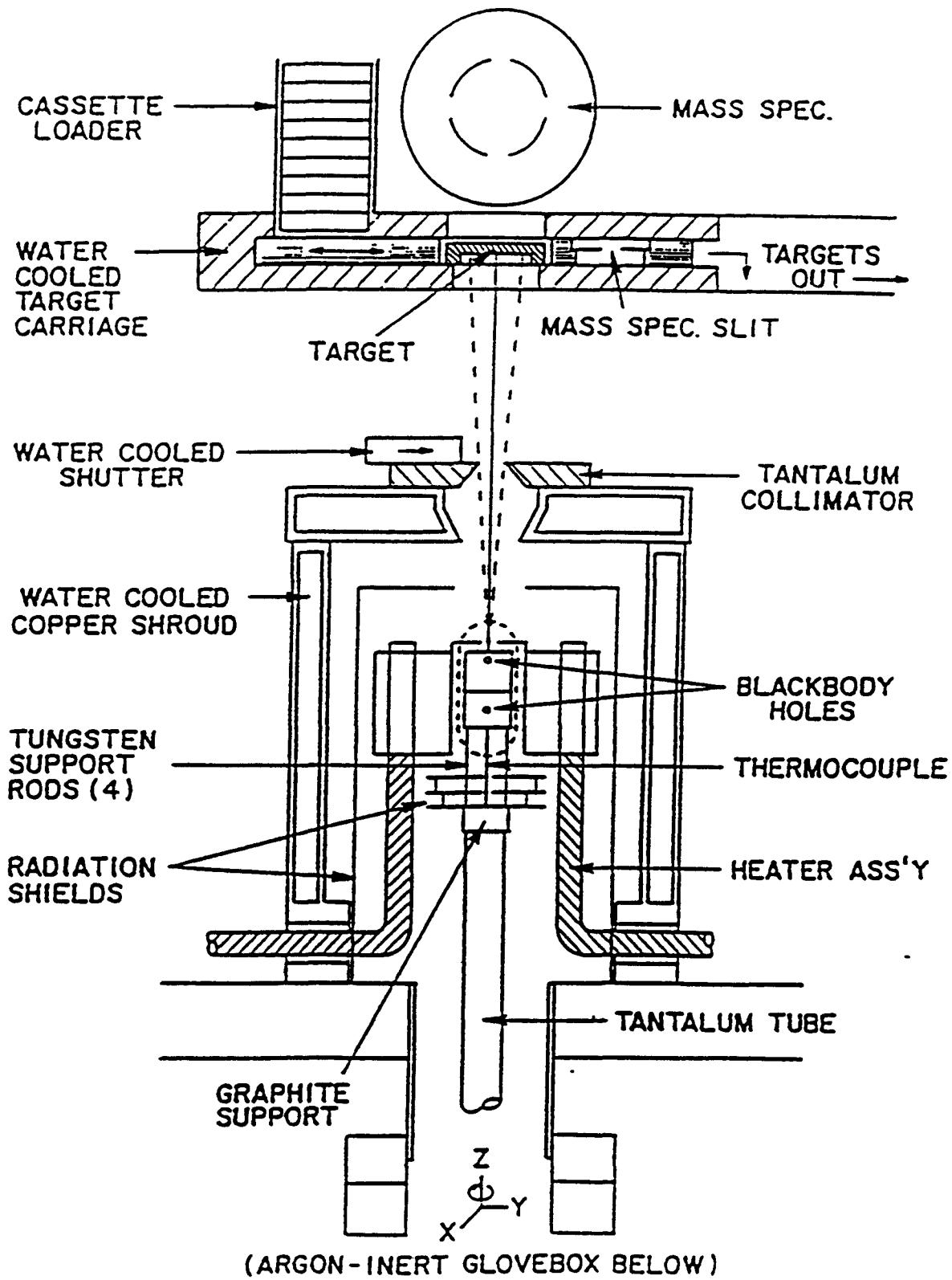
The sublimation of PuOCl was studied by Knudsen effusion mass spectrometry. A diagram of the apparatus is given in Figure 1. The PuOCl sample was contained in a platinum cup inside a platinum effusion cell which was contained inside a molybdenum effusion cell. The temperature of the sample was measured with a chromal-alumel thermocouple. The vapor species that were monitored included; PuCl₄, PuCl₃, PuOCl, PuO, and Pu. The vapor pressure of PuOCl was measured in the temperature range 1095 - 1270 K. The mass spectrometer was calibrated by studying the sublimation of UF₄.

The PuOCl was prepared by the following reaction;



The PuCl₃ was prepared by the chlorination of liquid plutonium metal.

COSS SECTIONAL VIEW OF UHV TARGET/MASS-SPECTROMETER KNUDSEN EFFUSION APPARATUS



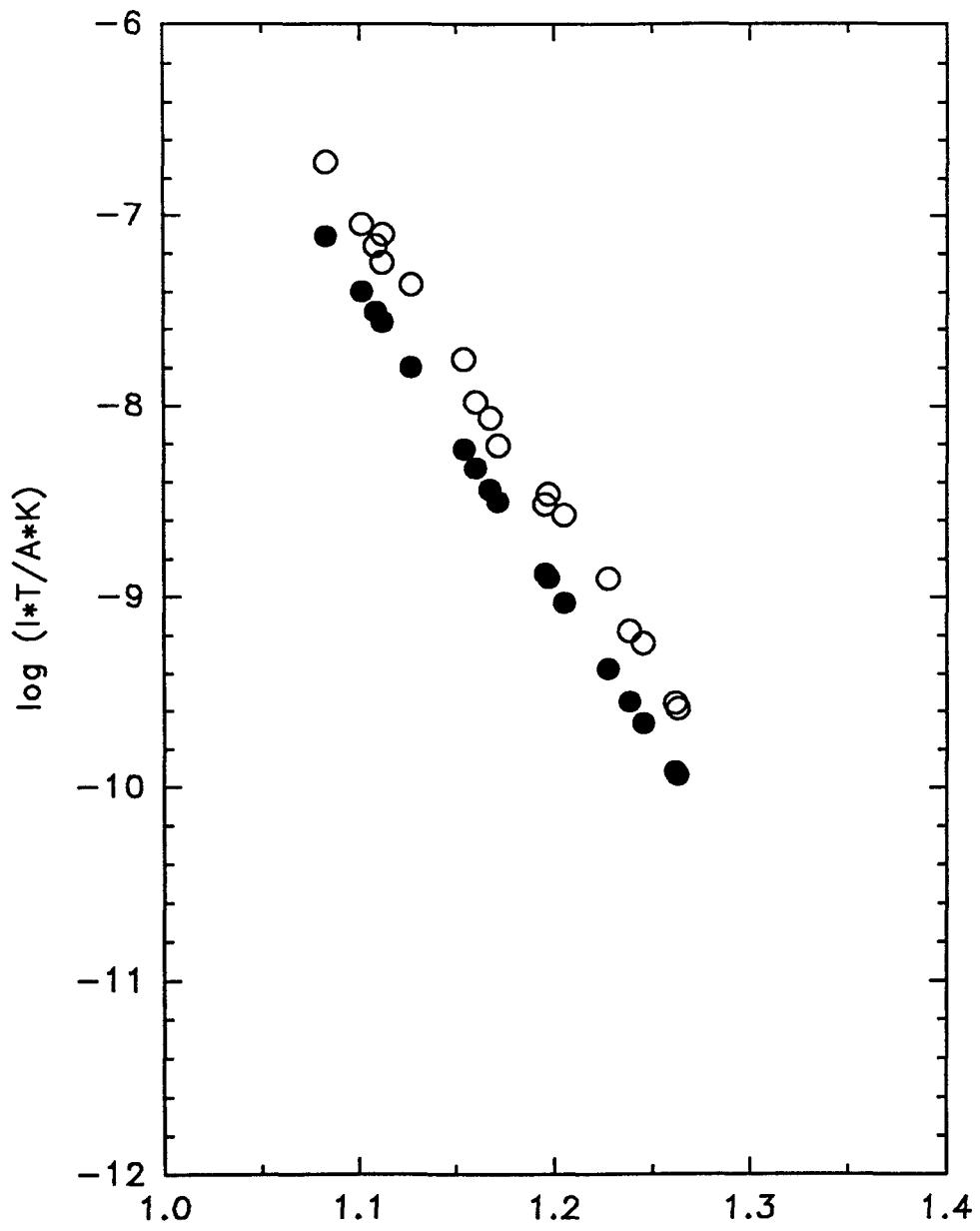
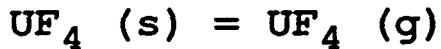


Figure 2. $\log(I*T)$ vs $1000 K/T$ for the sublimation of $UF_4(s)$; open circles— $\log(I*T)$ data, closed circles—Hildenbrand's $\log(P/atm)$ data.

Calibration



The intensity of $\text{UF}_3^+/\text{UF}_4$ at 45 eV was monitored as a function of temperature.

<u>T (K)</u>	<u>I*T (A*K)</u>	<u>P (atm)^a</u>	<u>C (atm/A*K)^b</u>
791.8	2.601e-10	1.156e-10	0.445
792.6	2.778e-10	1.211e-10	0.436
866.7	1.765e-08	5.966e-09	0.338
835.6	3.475e-09	1.264e-09	0.364
902.5	6.920e-08	3.118e-08	0.451
853.7	6.235e-09	3.162e-09	0.508
803.0	5.762e-10	2.185e-10	0.379
887.7	4.362e-08	1.600e-08	0.367
836.6	3.060e-09	1.331e-09	0.435
807.6	6.689e-10	2.824e-10	0.422
829.9	2.698e-09	9.395e-10	0.348
923.5	1.936e-07	7.748e-08	0.400
908.1	8.955e-08	3.991e-08	0.446
899.6	7.943e-08	2.740e-08	0.345
814.9	1.250e-09	4.216e-10	0.337
862.1	1.048e-08	4.776e-09	0.456
856.7	8.698e-09	3.668e-09	0.422
899.8	5.655e-08	2.765e-08	<u>0.489</u>
		average	0.410

$$\log (I*T) = -(15606 \pm 238)/T + (10.17 \pm 0.28)$$

$$\Delta H^\circ (298 \text{ K}) = 312.5 \pm 4.6 \text{ kJ/mol}$$

^a Hildenbrand's Data¹

$$\log (P/\text{atm}) = -(15691 \pm 109)/T + (9.90 \pm 0.11)$$

$$\Delta H^\circ (298 \text{ K}) = 313.8 \pm 2.1 \text{ kJ/mol}$$

^b $C = P/I*T$

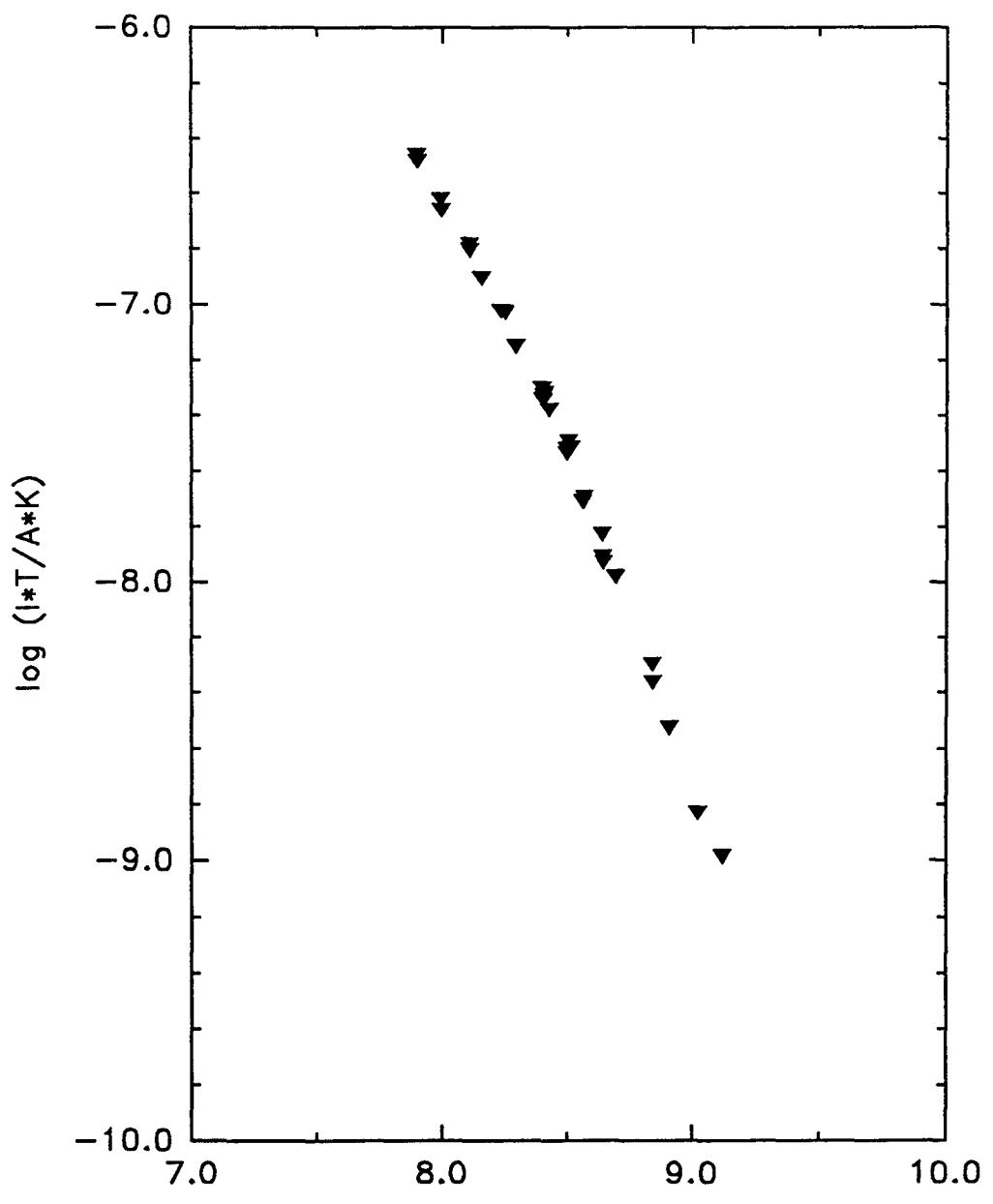


Figure 3. $\log(I*T)$ vs $10000 K/T$ for the sublimation of PuOCl.

PuOCl Sublimation



The intensity of $\text{PuCl}_2^+/\text{PuCl}_3$ at 45 eV was monitored as a function of temperature.

<u>Point</u>	<u>T (K)</u>	<u>I*T (A*K)</u>	<u>log (I*T/A*K)</u>
3	1150.6	1.065e-08	-7.973
8	1189.7	4.911e-08	-7.309
11	1190.9	5.103e-08	-7.292
12	1096.8	1.049e-09	-8.979
15	1176.9	3.063e-08	-7.514
18	1177.0	2.951e-08	-7.530
19	1131.1	4.394e-09	-8.357
21	1131.4	5.083e-09	-8.294
29	1157.7	1.251e-08	-7.903
32	1157.2	1.197e-08	-7.922
33	1190.1	4.786e-08	-7.320
36	1157.9	1.507e-08	-7.822
37	1176.1	3.257e-08	-7.487
38	1174.9	3.098e-08	-7.509
56	1190.8	4.650e-08	-7.332
62	1190.3	4.629e-08	-7.334
63	1186.8	4.256e-08	-7.371
64	1167.4	2.064e-08	-7.685
68	1168.2	1.983e-08	-7.703
73	1108.5	1.499e-09	-8.824

$$\log (I*T/A*K) = -(23668 \pm 371)/T + (12.58 \pm 0.32)$$

$$\log (P/\text{atm}) = -(23668 \pm 371)/T + (12.13 \pm 0.32)$$

$$\Delta H^\circ (298 \text{ K}) = 459.13 \pm .71 \text{ kJ/mol}$$

$$\Delta S^\circ (298 \text{ K}) = 244.0 \pm 6.1 \text{ J/mol K}$$

PuOCl Sublimation



$$L_1 \equiv [1-x \text{ PuCl}_3 + x \text{ Pu}_2\text{O}_3] \text{ (l)}$$

The intensity of PuCl_2^+ /PuCl₃ at 45 eV was monitored as a function of temperature.

<u>Point</u>	<u>T (K)</u>	<u>I*T (A*K)</u>	<u>log (I*T/A*K)</u>
2	1205.9	7.202e-08	-7.142
4	1233.2	1.589e-07	-6.799
5	1233.6	1.663e-07	-6.779
22	1214.7	9.623e-08	-7.017
24	1226.1	1.263e-07	-6.899
26	1266.2	3.534e-07	-6.452
28	1266.0	3.351e-07	-6.475
42	1212.3	9.515e-08	-7.022
49	1251.1	2.414e-07	-6.617
55	1250.7	2.208e-07	-6.656

$$\log (I*T/A*K) = -(16503 \pm 488)/T + (6.57 \pm 0.39)$$

$$\log (P/\text{atm}) = -(16503 \pm 488)/T + (6.12 \pm 0.39)$$

PuOCl Sublimation

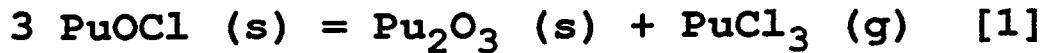
Vapor Species

The vapor species that were observed were PuCl_3^+ , PuOCl^+ and their fragments.

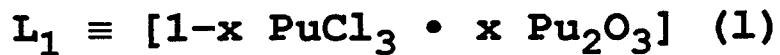
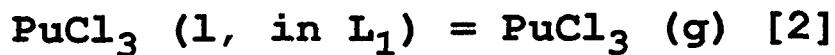
Intensity-temperature products were converted to pressures with the UF_4 calibration constant, $C=0.410 \text{ atm}/\text{A}^*\text{K}$, and the fragmentation corrected cross section, 0.870.

Vaporization Reactions

$T < 1190 \text{ K}$

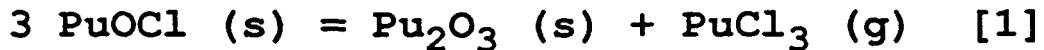


$T > 1190 \text{ K}$



PuOCl Sublimation

Thermodynamics



$$\Delta H^\circ \text{ (298 K)} = 459.13 \pm .71 \text{ kJ/mol}$$

$$\Delta S^\circ \text{ (298 K)} = 244.0 \pm 6.1 \text{ J/mol K}$$

What is the heat of formation of $\text{Pu}_2\text{O}_3 \text{ (s)}$?

$$\Delta H^\circ \text{ (298 K)} = -1682.2 \text{ kJ/mol}$$

$$\Delta H^\circ \text{ (298 K)} = -1656 \text{ kJ/mol estimated}^2 \text{ by Bessman et al.}^2$$

What is the entropy of $\text{PuCl}_3 \text{ (g)}$?

$$S^\circ \text{ (298 K)} = 382.8 \text{ J/mol K}$$

$$S^\circ \text{ (298 K)} = 367.4 \text{ J/mol K Fuger et al.}^3$$

The entropy of $\text{PuCl}_3 \text{ (g)}$ derived in this study is more in line with the value calculated by statistical methods, 380.2 J/mol K.

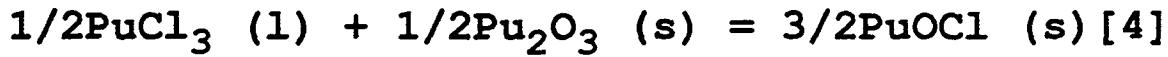
PuOCl Sublimation

Thermodynamics

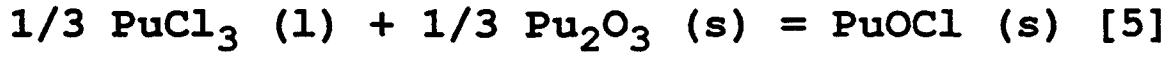
What is the free energy of formation of PuOCl calculated from the vaporization data assuming the peritectic decomposition of PuOCl occurs at 1190 K?

$$\Delta\bar{G}_{\text{PuCl}_3} = (-8.231 \cdot T) - 74920 \quad T=1190-1275 \text{ K}$$

$$\Delta\bar{G}_{\text{PuCl}_3} = -84.72 \text{ kJ} \quad T=1190 \text{ K}$$



$$\Delta G^\circ [4] = -42.36 \text{ kJ/g-atom}$$



$$\Delta G^\circ [5] = -28.24 \text{ kJ/mol}$$

$$\Delta G^\circ_{f,\text{PuOCl}} = -28.24 + \frac{1}{3}(-708) + \frac{1}{3}(-1396.3)$$

$$\Delta G^\circ_{f,\text{PuOCl}} = -729.6 \text{ kJ/mol}$$

$$\Delta G^\circ_{f,\text{PuOCl}} = -742.8 \text{ kJ/mol} \quad \text{Weigel et al.}^4$$

A phase diagram has been proposed for the PuCl_3 - Pu_2O_3 system, Figure 4, based on the PuCl_3 -PuOCl phase diagram, Figure 5, and the sublimation data obtained in this study.

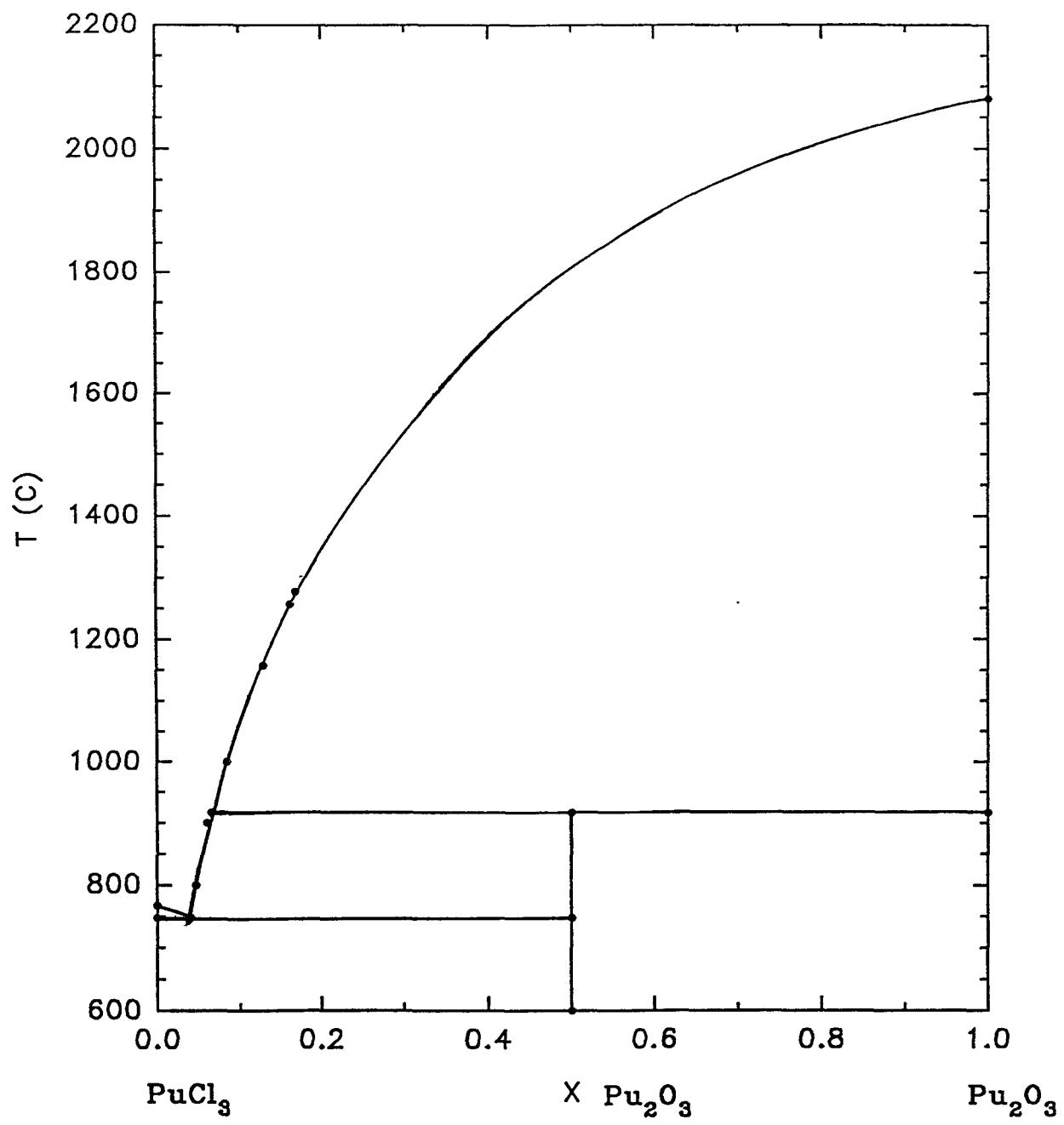


Figure 4. Proposed PuCl_3 - Pu_2O_3 phase diagram.

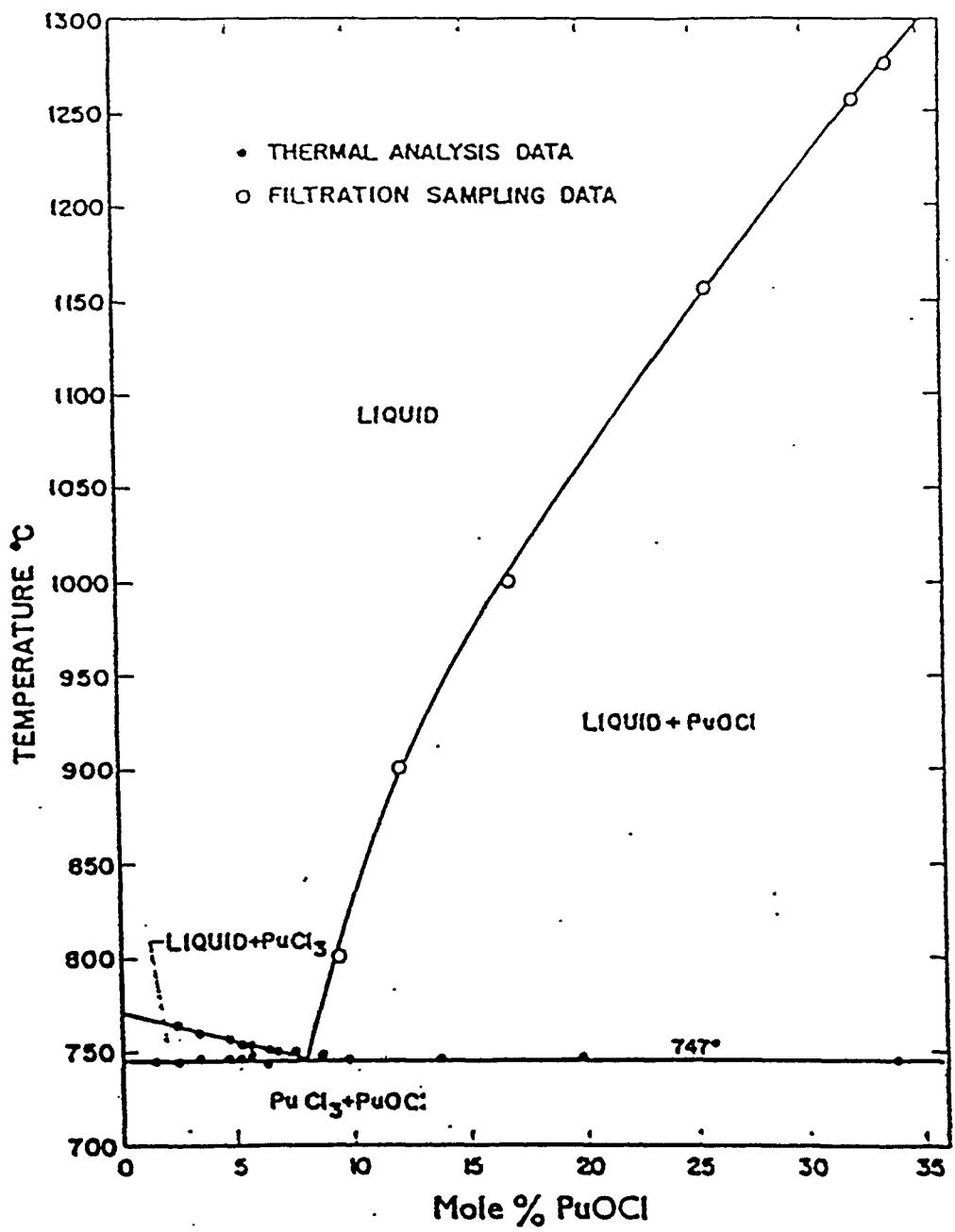


FIG. 5 —Partial phase diagram of the PuCl_3 - PuOCl system.⁵