

INTERIM FINAL TECHNICAL REPORT
September 1, 1992 through August 31, 1993Project Title: **PROTOCOLS FOR THE SELECTIVE CLEAVAGE OF
CARBON-SULFUR BONDS IN COAL**

DE-FC22-92PC92521

Principal Investigator: Mark Bausch, Southern Illinois
University at CarbondaleProject Manager: Ken Ho, Illinois Clean Coal
Institute**ABSTRACT**

Summarized in this INTERIM FINAL TECHNICAL REPORT for our project "Protocols for the Selective Cleavage of Carbon-Sulfur Bonds in Coal" are results of research pertaining to chemical reactions that aim to selectively cleave C-S bonds in model compounds as well as Illinois coal. Chemical reactions that result in carbon-sulfur bond cleavage are an essential aspect of any protocol designed to remove organic sulfur from coal. In the second year of our project "Protocols for the Selective Cleavage of Carbon-Sulfur Bonds in Coal" we have completed investigations of reactions in which organic sulfur-containing coal model compounds are subjected to different conditions of temperature, solvent mixtures, reagents, and radiation. We have also undertaken a series of reactions in which physically cleaned Illinois coal has been subjected to many of the same reaction conditions that were shown, via the use of model sulfides, to result in substantial C-S bond cleavage and/or sulfur oxidation. Therefore, summarized in this interim final report are results of our investigations of the photooxidation reactions of coal model sulfones and sulfides; the photolytic desulfurization of coal; and various other topics, including a summary of our endeavors aimed at initiating C-S bond cleavage reactions using oxidation/chlorination/desulfurization protocols, and various tellurium reagents. Important experiments remain to be completed on this project; therefore, our efforts in these areas will continue through the end of calendar year 1993.

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EXECUTIVE SUMMARY

The following sets of experiments were slated for completion during the two-year funding period for "Protocols for the Selective Cleavage of Carbon-Sulfur Bonds in Coal":

- 1. Allow various coal model compounds (species that contain aliphatic and aromatic C-S bonds) to be subjected to various conditions. The variables in these studies include different substrates, solvents, bases, added reagents (catalysts, oxidizing agents, electron acceptors), temperature and pressure. Photochemical reactions of these coal model compounds will also be studied by varying the wavelength of light, using various solvents with or without sensitizers, oxidizing the compounds and subsequently hydrolyzing the products of the reaction.
- 2. Subject coal and coal model compounds to ultrasonic irradiation for different periods of time. The effects of adding catalysts or oxidizing agents will also be determined. Different substrates, solvents and oxidizing agents, along with ultrasonic irradiation, altogether will form a matrix of results that will allow determination of optimum conditions that will accomplish the desired C-S bond cleavage on coal samples.
- 3. Heat the coal model compounds containing both aliphatic and aromatic sulfur, oxygen, carbon, and nitrogen at various temperatures, using a Differential Scanning Calorimeter(DSC). Observe the transitions in the sample brought about by C-S, C-O, C-C and C-N bond cleavages.
- 4. Investigate differences in the bond strengths in the radical cations derived from coal model compounds 10-phenylthio-9-methylanthracene and 9-phenylthiomethyl-anthracene. We will attempt to ascertain the differences in reactivities of different types (i.e. aromatic or aliphatic carbon) of carbon-sulfur bonds, under various reaction conditions, and correlate these differences with their respective bond strengths.

We have been granted a no-cost extension to continue work on this project through the end of calendar year 1993. Therefore this document is an interim final report and includes discussion of the following experiments, topics, and results:

- Mixtures of BPS and NaOH/KOH, when heated for three hours at a final temperature of 330°C and final pressure of 250 psi, remains largely unchanged, when the reaction is carried out in THF, morpholine, or quinoline.
- Benzyl phenyl sulfide (BPS) was converted to its sulfone

(benzyl phenyl sulfone). The sulfone was then chlorinated with N-chlorosuccinimide, forming α -chloro-benzyl phenyl sulfone. The resulting α -chloro-benzyl phenyl sulfone was subjected to room temperature hydrolysis by allowing it to react with NH_3 in water. Benzaldehyde was isolated in 40% yield. Also observed was phenyl disulfide.

- Benzyl phenyl sulfide, when treated with N-chloro-succinimide followed by ammonium hydroxide, forms benzaldehyde and dibenzyl disulfide, in near quantitative yields.

- NaTeH , when allowed to react with dibenzothiophene sulfone at temperatures less than 220°C , does not react. Dibenzothiophene sulfone, when allowed to react with tellurium powder alone, is recovered unchanged. When NaTeH was allowed to react with dibenzothiophene sulfone at temperatures greater than 220°C , C-S bond cleavage was observed. Products isolated include dibenzothiophene as well as the tellurium analogue of dibenzothiophene. In addition, NaTeH , when allowed to react with benzyl phenyl sulfone, forms the corresponding sulfide in high yields.

- Dibenzothiophene sulfone, when allowed to react with a mixture of NiCl_2 and NaBH_4 (at ca. 15°C for one hour), was transformed (in 30% yield) into biphenyl.

- Photooxidations of phenyl sulfide, benzyl phenyl sulfide, thiophene, benzothiophene, and dibenzothiophene have been carried out under five sets of reaction conditions, all of which include irradiation of the sulfide with a Hanovia lamp: (a) adding oxygen in the presence of Rose Bengal in methanol solution; (b) adding oxygen in the presence of 9,10-dicyanoanthracene in acetonitrile solution; (c) adding oxygen in the presence of anthracene in acetonitrile solution; (d) adding oxygen in the presence of no sensitizer in several solvents; and (e) "blank" studies in which the sulfides were irradiated for similar time periods in the same solvents as those mentioned previously in the absence of oxygen. Results from these experiments suggest that electron transfer-mediated photooxidations (i.e. those in which 9,10-dicyanoanthracene or anthracene was used as a sensitizer) were more likely to occur (using the present set of sulfides) than the Rose Bengal-mediated reactions. Results from experiments in which benzyl phenyl sulfide and dibenzothiophene were photooxidized in the absence of sensitizer indicates that these substrates are quite reactive in our photooxidative protocol. In essence, they appear to react with molecular oxygen, in the presence of light, without any added catalyst, forming both the analogous sulfoxides and sulfones. Finally, experiments carried out in which the sulfides were subjected to the Hanovia lamp in the absence of oxygen reveal that little or no reaction occurs. In other words, the oxygen is essential in these transformations.

- IBC-108 has been subjected to the same photooxidative protocol that we have developed in our work with the sulfide

model compounds. The photooxidation protocol, when carried out in acetonitrile, appears to result in a 10% reduction in the organic sulfur content of the coal. Surprisingly, when carried out in other solvents (such as pyridine, dimethyl sulfoxide, and methanol), the desulfurization appears to be much greater (ca. 70%). Experiments in this area are continuing as the latter results need to be verified and rechecked.

•DSC experiments have been completed using IBC-108 and photooxidized IBC-108 as substrate. Analyses of our initial results indicates that the photooxidation treatment does not substantially overwhelm or degrade the backbone of the coal macromolecule. Experiments in this area are continuing.

OBJECTIVES

Our long range objective in the research funded by ICCI has been to carry out experiments that will aid in solving the problem of coal desulfurization--from the perspective of inventing, understanding, and perfecting new carbon-sulfur bond breaking chemical reactions. Experiments carried out during the past three months included studies of reactions in which base mixtures were allowed to react with sulfur-containing coal model compounds under high temperature (Parr bomb) conditions. Additional experiments carried out during the present reporting period were aimed at cleaving carbon sulfur bonds in α -chlorosulfones, with the aid of various tellurium-based reagents. Other experiments carried out during this reporting period included studies of photochemical and photo-oxidation reactions of sulfur-containing coal model compounds in the presence/absence of sensitizers. It is our objective to develop a set of low-temperature reaction conditions that will enable substantial C-S bond cleavage to be affected in Illinois coal. The described experiments are the first steps in achieving some understanding at the molecular level of a newly-invented process. Details of photooxidative experiments involving IBC-108 are also found in this summary.

Examination of the four tasks listed below reveals details of the experiments planned to aid in the efforts to reach the stated objective:

TASK 1: Carry out experiments that aim to selectively cleave C-S bonds in coal model compounds benzyl phenyl sulfide, phenyl sulfide and dibenzothiophene, as well as physically cleaned Illinois coal. Main thrusts of these investigations include radical- and Lewis acid-initiated, as well as caustic- and photooxidatively-assisted C-S bond cleavages.

TASK 2: Evaluate the effect of ultrasonic radiation on the effectiveness of several C-S bond cleavage protocols.

TASK 3: Utilize differential scanning calorimetry (DSC) techniques in examinations of various methods of cleaving C-S bonds in coal model compounds, as well as gaining understanding of the degree of depolymerization in Illinois coals (coals that contain varying amounts of organic and inorganic sulfur) when these coals are subjected to depolymerization protocols. Also planned are DSC studies of several Illinois coals subjected to a variety of pretreatment conditions.

TASK 4: Carry out experiments that enable determinations of the bond energies for aliphatic and aromatic C-S bonds in the

radical ions derived from several coal model compounds. Communicate these results to the coal-desulfurization community through publications in scientific journals.

INTRODUCTION AND BACKGROUND

The organic sulfur found in coal consists primarily of sulfur atoms covalently bonded to aliphatic or aromatic carbon atoms contained in the backbone of the coal macromolecule (there may be a small percentage of S-S and S-H bonds in coal as well). The sulfur present in Illinois coal is problematic because the combustion of high-sulfur coal results in the formation of species implicated in environmental problems related to acid rain. The relevance and significance of the experiments planned for the present funding period is perhaps best described by acknowledgment of the following two statements:

- (a) In order to remove the organic sulfur from coal, carbon-sulfur bonds must be broken.
- (b) Chemists' knowledge and awareness of reactions that lead to the selective cleavage of carbon-sulfur bonds, and in particular, sulfur atoms bonded to sp^2 (i.e. aryl) carbon atoms, is limited.

The experiments outlined in the four TASKS listed earlier in this interim final report are aimed at developing an increased understanding of methods that enable cleavage of C-S bonds. The longer term goal of these experiments is the invention of new chemical reactions, or perfection of existing chemical reactions. The attainment of reaction conditions that allow the selective cleavage of carbon-sulfur bonds in coal and coal model compounds is therefore the main goal of "PROTOCOLS FOR THE SELECTIVE CLEAVAGE OF CARBON-SULFUR BONDS IN COAL". An important outcome of these studies is the resulting increase in chemists' knowledge of carbon-sulfur bond cleavage reactions. Successful completion (successful in terms of affecting carbon-sulfur bond cleavages) of the studies described in this proposal thus increases the likelihood of the deployment of a simple and cost-effective protocol designed to remove organic sulfur from Illinois coal.

Much of the coal mined in Illinois suffers from its high sulfur content. Methods have been developed that are quite effective at removing the inorganic sulfur from coal. Unfortunately, the problem of organic sulfur present in Illinois coal remains to be solved. A major reason for the lack of progress in the area of organic sulfur removal is the lack of knowledge concerning carbon-sulfur bond cleavage reactions. Completion of the experiments outlined in TASKS 1-4 will at least partly remedy this situation. It is our desire to convert what we learn from these studies of carbon-

sulfur bond breaking reactions into a simple protocol that will offer a reasonable and economically viable solution to the problems of using Illinois coal.

The experiments summarized in this interim final report are best described as fundamental studies of carbon-sulfur bond cleavage reactions, using both coal and model compounds. At present, none of these experiments describe a process that is cost-effective in terms of removing organic sulfur from coal. It is our desire to develop a practical carbon-sulfur bond-breaking technology from the research described in this proposal, and transition this new technology to the Illinois coal community.

EXPERIMENTAL PROCEDURES

TASKS 1-4 describe experiments that have synthetic, colorimetric, and electrochemical components. The synthetic component of our work is typical of the procedures that practicing organic chemists follow. A uv-visible spectrophotometer is utilized in the collection of appropriate acid-base and kinetic data.

Other experimental techniques and apparatus utilized in our work during the past three months include a Parr high pressure reactor vessel, GC techniques, FT-NMR techniques, DSC techniques, as well as other chromatographic techniques.

RESULTS AND DISCUSSION

TASK 1: Carry out experiments that aim to selectively cleave C-S bonds in coal model compounds benzyl phenyl sulfide, phenyl sulfide and dibenzothiophene, as well as physically cleaned Illinois coal. Main thrusts of these investigations include photo-oxidatively assisted C-S bond cleavage reactions.

1.1. CHEMICAL DESULFURIZATION OF COAL MODEL SULFIDES

• Introduction

As far as coal is concerned, it is likely that sulfides are the primary functional group that contain organic sulfur.. During the past 24 months of ICCI funding for this project, we have carried out several different reactions that aimed to thermally cleave C-S bonds in model sulfides such as benzyl phenyl sulfide (BPS). In many of these reactions, substantial degradation of the BPS was observed. Unfortunately, the product that results from these reactions often includes substantial (i.e. greater than 50%) amounts of intractable tars. When carried out in typical organic solvents, the products of reactions between BPS and caustics

were particularly difficult to evaluate and identify. We therefore felt it was important to determine the maximum time and temperature at which little or no reaction occurred when BPS was allowed to react with caustic.

In other experiments, we have developed a new strategy for breaking aliphatic C-S bonds which involves chlorination followed by hydrolysis. Results from these experiments are also found in this section.

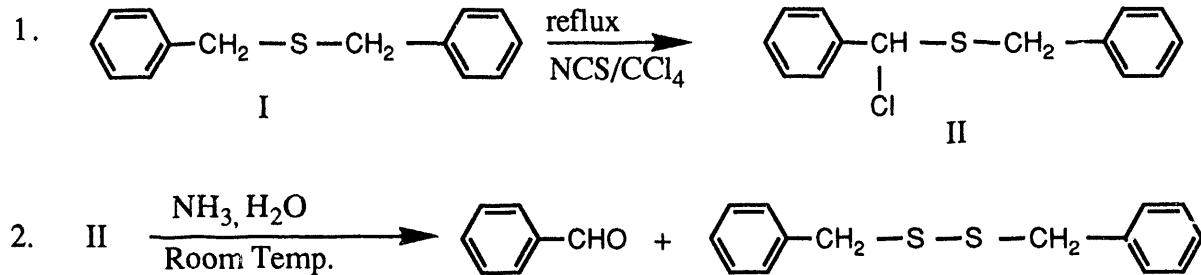
•Results

CAUSTIC/THERMOLYSIS PROTOCOL

In separate experiments, benzyl phenyl sulfide was allowed to react with KOH/NaOH mixtures, in THF, morpholine, and quinoline solvents, for 3 hrs. The final temperature for these reactions was 330°C; the final pressure was 250 psi. The main conclusion to be drawn from these experiments is that benzyl phenyl sulfide was recovered essentially unchanged in all of these reactions. It may therefore be possible to carry out similar reactions at slightly higher temperatures in an effort to study the transformation of BPS into other products and identify these products in reactions in which the mass balances can be accurately accounted.

CHLORINATION/HYDROLYSIS PROTOCOL

The reaction between dibenzyl sulfide (I) and N-chlorosuccinimide (NCS) followed by treatment of ammonium hydroxide showed that the carbon and sulfur bond was broken with one product being benzaldehyde and another product being dibenzyl disulfide.



In a related reaction, benzyl phenyl sulfide, when treated with N-chlorosuccinimide followed by ammonium hydroxide, forms benzaldehyde and dibenzyl disulfide, in near quantitative yields.

That the chlorination of the aliphatic carbon is an essential part of this reaction was verified by monitoring similar reactions in which phenyl sulfide was used as a substrate. In this reaction, we now believe, with the aid of TLC analyses, that approximately 20% of the starting material has been transformed into as yet unidentified products.

1.2. CHEMICAL DESULFURIZATION OF COAL MODEL SULFONES

• Introduction

It is also likely that sulfides and their oxides exist in coal. We have therefore added sulfones to our list of coal desulfurization model compounds. As described in previous quarterly reports, we have investigated tellurium-based reagents in reactions that aim to cleave C-S bonds as well as chlorination/hydrolysis protocols. In this report, additional results are described for these and related reactions. Also found in this section are initial results from reactions in which NiCl_2 was used as a reagent to cleave C-S bonds in organic sulfones.

• Results

CHLORINATION/HYDROLYSIS PROTOCOL

The substrate α -chloro-benzyl phenyl sulfone was subjected to room temperature hydrolysis by allowing it to react with NH_3 in water. Benzaldehyde was isolated in 40% yield. Also observed was phenyl disulfide.

TELLURIUM PROTOCOL

The utility of NaTeH as a reagent in C-S bond cleavage reactions has also been investigated. We first allowed NaBH_4 to react with dibenzothiophene sulfone. There was no reaction observed. We then allowed the same sulfone to react with a mixture of NaBH_4 and tellurium. Below 220°C , we observed no reaction. However, upon heating up to 500°C , we observed that biphenyl formed in 20-30% yields, while dibenzothiophene was formed in yields of about 30%. There are several more (as yet) unidentified products in this reaction.

When allowed to react with NaTeH at 300°C , dibenzothiophene sulfone formed substantial amounts of biphenyl, dibenzothiophene, and the tellurium analogue of dibenzothiophene. The products of these reactions have been isolated and confirmed by NMR, GC, and MS.

NICKEL CHLORIDE/SODIUM BOROHYDRIDE PROTOCOL

Benzothiophene sulfone was allowed to react, at ca. 15°C and 1 hour, with a mixture of NiCl_2 and NaBH_4 . Isolated as a main product (ca. 30% yield) in this reaction was biphenyl, indicative of substantial C-S bond cleavage and extrusion of SO_2 . IBC 108 has also been subjected to analogous reaction conditions. We await the analyses of the coal treated in this fashion.

1.3. PHOTOOXIDATION OF ORGANIC SULFIDES

As described in prior quarterly reports, experiments have been carried out for the series of the coal model sulfides (phenyl sulfide, benzyl phenyl sulfide, dibenzothiophene, thiophene, and benzothiophene).

Of particular interest to us was the photooxidation of

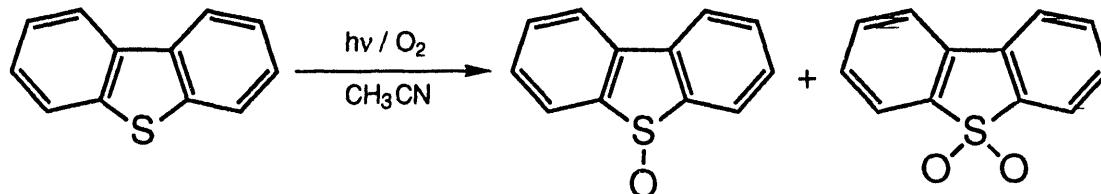
dibenzothiophene (DBT). As reported previously, this compound when irradiated for 22 hours, in the presence of 9,10-dicyanoanthracene (DCA) and oxygen, yielded oxygenated products (dibenzothiophene sulfoxide (DBTO) and dibenzothiophene sulfone (DBTO₂)) with about 60% yield. The yields were only slightly lower when anthracene was used as a sensitizer. This result is particularly interesting from the point of view of coal chemistry since polyaromatic hydrocarbons (potential electron transfer sensitizers) are abundant in coal. Our previously reported results indicate that singlet oxygen is probably not involved in this reaction since formation of the oxygenated products is not retarded by the addition of the ¹O₂ quencher - β -carotene.

Interestingly, when no sensitizer was used, and dibenzothiophene was exposed to oxygen and light only, dibenzothiophene was isolated with 15% yield.

The acetonitrile used in our experiments was 99% pure. The one percent of impurities could be a source of the sensitizer. To evaluate this possibility 99.9% (HPLC grade) solvent was used in the experiments described in these report. The results of the photooxygenations of DBT are summarized in Table 1 (next page).

Furthermore, the role of molecular oxygen was evaluated for all the compounds studied by carrying their oxidations in the absence of light, with all other conditions unchanged.

TABLE 1



<u>CONDITIONS</u>	<u>YIELDS [%]</u>	
DCA sens./99% solvent	33	28
DCA sens./99.9% solvent	29	41
AN sens./99% solvent	37	11
no sens./99% solvent	15	0
no sens./99.9% solvent	28	13
β -carotene/99% solvent	11	0
no light/99% solvent	0	0
no oxygen/99.9% solvent	0	0

- Results

The same light source (200 W, medium pressure, Conrad-Hanovia 654A36 lamp, filtered through the Pyrex 7740 absorption sleeve) and standard photochemical reactor were used as described in previous reports. Oxygen was bubbled

through 150 ml of the 0.01 M solution of the sulfide in acetonitrile. Sensitizer concentration was 0.0001 M. Reactions were monitored by GC and TLC. Products were isolated by column chromatography or by centrifugal, preparative, thin layer chromatography using a Harrison Research Chromatotron (Model 7924T). The separated compounds were then characterized by GC, TLC, and ¹H and ¹³C NMR (DMSO-D₆ solution).

Three experiments have been carried out for dibenzothiophene. First, 22 hour unsensitized photooxidation was carried out in HPLC grade acetonitrile (99.9%) pure. The reaction gave unexpected results. The yields of oxygenated products instead of being lower were significantly higher than those obtained in 99% pure solvent. The yield for DBTO was 28% and for DBTO₂ 13%. About 50% of the substrate was recovered.

DCA sensitized photooxidation of DBT, in HPLC grade solvent, was carried out to evaluate the role of the solvent quality on the sensitized reaction. In the 22 hour reaction the yield of DBTO was 29% and DBTO₂ 41%. 9% of DBT was recovered. Thus, only the sulfone formation seemed to be enhanced when acetonitrile of higher purity was used. However, the increase of the total yield of the oxygenated products was not as dramatic as in the unsensitized reaction. The purity of acetonitrile probably plays important role in stabilization of the intermediate [DBT⁺ O₂⁻•], formed in the unsensitized reaction.

The third reaction involved 22 hours of bubbling of oxygen through the acetonitrile (99.9% pure) solution of DBT but without irradiation. 95% of DBT remained unreacted (GC) and no other compounds were detected by GC or NMR in the product mixture.

Similar results were obtained for other model sulfides exposed to the same conditions¹. Their conversions were low (GC) and only traces of unidentified products were detected.

This work shows that dibenzothiophene alone, in the present of light and oxygen, when dissolved in acetonitrile, is converted to the corresponding sulfoxide (in 28% isolated yield) and sulfone (in 13% isolated yield).

The results are important from the point of view that coal undergoes changes upon weathering, and that much of the chemistry associated with the weathering of coal may in fact be catalyzed by light.

1.4. PHOTOOXIDATION OF IBC-108 COAL

That light and oxygen (in the presence of sensitizer) can cleave or transform chemical bonds points to the power of these reagents in certain reactions. Published reports claim that atmospheric weathering leads to desulfurization of coal.² This observation was followed by laboratory

¹Reaction times were equal to those of the photolytic runs i.e.: 75 minutes for PS, 1 hour for BPS and 5 hours for BT.

experiments in which samples of coal were desulfurized in a two stage process (photooxidation and hydrolysis of products)³. Up to 35% reduction in sulfur was reported when coal samples (suspensions in ethanol) were irradiated at 254 nm and then hydrolyzed.

In our attempts to achieve coal desulfurization via a photooxidative process, we have exposed IBC-108 to the standard unsensitized photooxidative conditions followed by hydrolysis in NaOH solution.

Results

The experiments were carried out as described previously. A sample of IBC 108 (250 mg) was suspended in 150 ml of a given solvent, and irradiated for 72 hours. The coal-acetonitrile suspension was stirred with a magnetic stirrer and agitated by oxygen that was delivered through teflon tubing. A 200 W medium pressure Conrad-Hanovia lamp, modified by a Pyrex 7740 light filter, was used as a light source. After the irradiation was completed the coal was filtered, weighed, and vacuum dried at 130°C.

The most recent quarterly report indicated that IBC-108, when subjected to photooxidative conditions in acetonitrile solution, was desulfurized to the extent of about 10%. Hydrolysis of the photooxidized coal resulted in an additional 10% desulfurization.

We can now report results of the quantitative sulfur analysis for similar experiments in which pyridine, dimethyl sulfoxide, and methanol were used as solvents in the photooxidative protocols. In the Table below, Rob Bartelsmeyer of Carterville IL carried out the sulfur analyses.

Table I. Sulfur Analyses of IBC-108 Coal. Determination of the Effect of Photooxidation.

sample and description	% sulfur
IBC-108 photooxidized coal (pyridine)	0.32
IBC-108 photooxidized coal (DMSO)	0.75
IBC-108 photooxidized coal (MeOH)	0.57
IBC-108 (untreated)	2.53

Inspection of the data in Table I reveals that these experiments result in organic sulfur decreases by ca. 70% as a result of photooxidation alone. We are currently repeating these experiments to ensure that these results are valid and reproducible. Of note is that the photooxidative procedure

²Chandra, D.; Chakrabarti, J. N.; Swamy, Y. V. *Fuel* **1982**, *61*, 204.

³Mathur, V. K.; Murphy, S. *Am. Chem. Soc. Div. Fuel Chem.* **1988**, *33* (1), 232.

renders about 30-50% of the coal soluble in the respective solvents.

TASK 2: Evaluate the effect of ultrasonic radiation on the effectiveness of several C-S bond cleavage protocols.

RESULTS: No experiments were undertaken for this task for this period.

TASK 3: Utilize differential scanning calorimetry (DSC) techniques in examinations of various methods of cleaving C-S bonds in coal model compounds, as well as gaining understanding of the degree of depolymerization in Illinois coals (coals that contain varying amounts of organic and inorganic sulfur) when these coals are subjected to depolymerization protocols. Also planned are DSC studies of several Illinois coals subjected to a variety of pretreatment conditions.

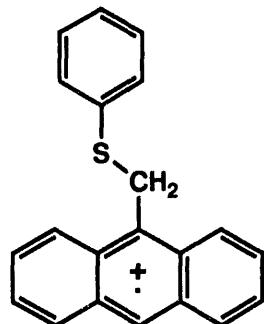
RESULTS: We have decided to carry out our initial experiments on IBC-108 itself, as well as photooxidized IBC-108. Our intent in these experiments is to determine what it is that the DSC-based results can tell us about coal, as well as how our photooxidative protocol changes the thermal properties of IBC-108.

As described previously, our initial results have proven to be difficult to analyze and well as somewhat irreproducible. Inspection of a plot of heat vs temperature for UC-1 (i.e. IBC-108 that has been photooxidized) reveals a maximum at ca. 100°C (probably water, an inflection point at about 350°C, and a minimum at about 500°C. After 500°C, the curve continues to rise upward. An identical experiment completed on UC-4 (i.e. untreated IBC-108) is remarkably similar, except for the absence of the inflection point and a somewhat earlier minimum (about 450°C). At this time we interpret the similarity in these data to mean that the photooxidative process does not overwhelmingly change the nature of the coal macromolecule, at least for the samples that we have examined to date. It is likely that the aforementioned inflection point is the result of some moderate depolymerization of the coal macromolecule that results from the photooxidative process. Work in this area is continuing, with particular emphasis on the photooxidized coal samples described in TASK 1 of this report.

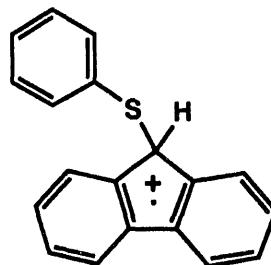
TASK 4: Carry out experiments that enable determinations of the bond energies for aliphatic and aromatic C-S bonds in the radical ions derived from several coal model compounds. Communicate these results to the coal-desulfurization community through publications in scientific journals.

RESULTS: Rather involved experiments involving photomodulated

voltammetry have been completed in attempts to determine redox potentials for electron transfer reactions involving various organic radicals and cations. The essence of these experiments is that we are trying to perfect a device which will enable us to evaluate redox potentials for transient organic species. Experiments are now underway in which we are determining the oxidation potentials for variously substituted fluorenyl radicals and methylanthracenyl cations. These data, combined with other redox data, will enable us to determine the C-S bond strengths in molecules such as **4** and **5**



4



5

These results will be summarized in our final report.

RESEARCH OUTLOOK

In the area of photooxidations of Illinois coal and organic sulfides, my research group and I are making important and real progress in areas that logically lead to new experiments in coal chemistry. We have yet to complete all of the necessary experiments. We have been granted a no-cost extension that will enable completion of these experiments. Therefore, planned in the next 4 months are additional experiments aimed at photo-oxidizing Illinois coal, as well as carrying out additional wet chemical reactions on raw coal samples and DSC analyses of the resulting coal samples. All of the TASKS will be reported in the FINAL report due at the end of December, 1993.

PROJECT MANAGEMENT REPORT
June 1 through August 31, 1993

Project Title: **PROTOCOLS FOR THE SELECTIVE CLEAVAGE OF CARBON-SULFUR BONDS IN COAL**

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COMMENTS

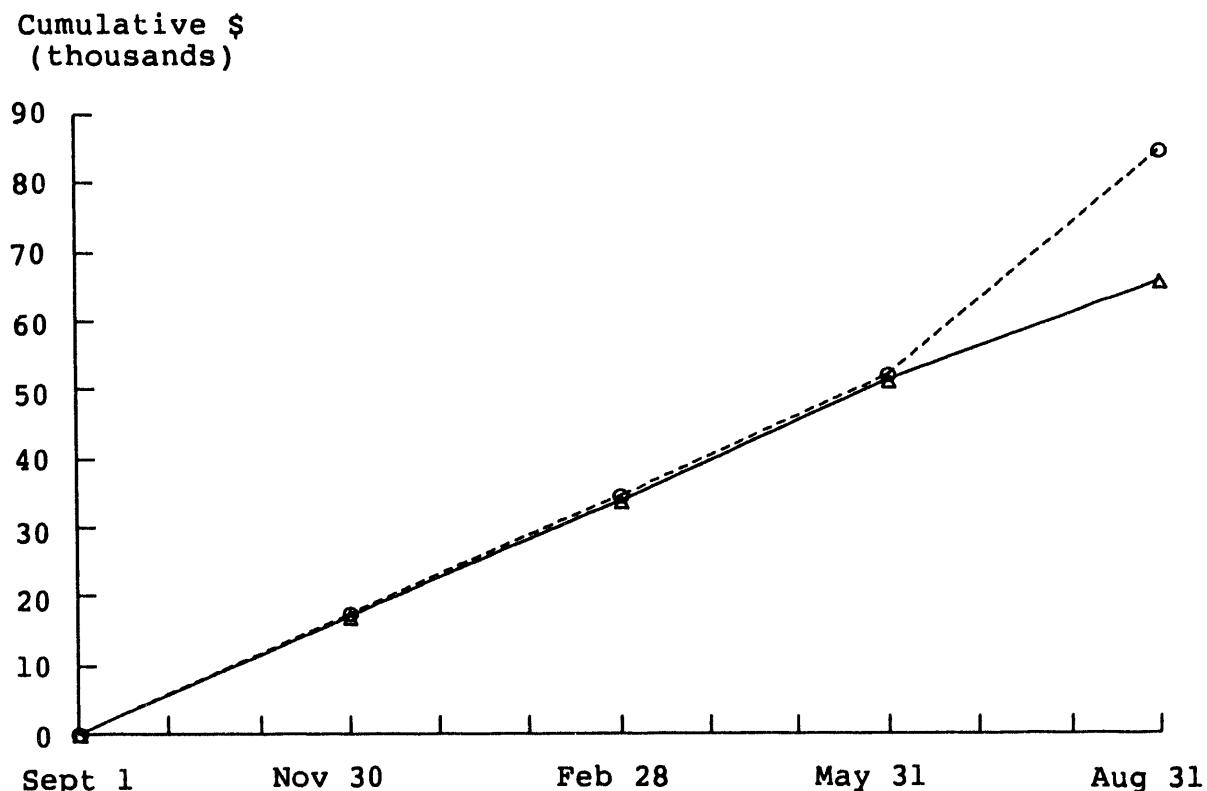
This project has received permission from ICCI to be extended through December 31, 1993.

Projected and Estimated Expenditures by Quarter

Quarter*	Types of Cost	Direct Labor	Materials and Supplies	Travel	Major Equipment	Other Direct Costs	Indirect Costs	Total
Sept. 1, 1992	Projected	12,250	1,636	0	500	1,500	1,589	17,475
to Nov. 30, 1992	Estimated Actual	12,250	1,636	0	0	1,500	1,539	16,925
Sept. 1, 1992	Projected	24,500	3,272	0	500	3,000	3,127	34,399
to Feb. 28, 1993	Estimated Actual	24,500	3,272	0	0	3,000	3,077	33,849
Sept. 1, 1992	Projected	36,750	4,908	500	500	4,500	4,716	51,874
to May 31, 1993	Estimated Actual	36,750	4,908	500	0	4,500	4,666	51,324
Sept. 1, 1992	Projected	62,955	6,545	1,000	0	6,000	7,700	84,700
to Aug. 31, 1993	Estimated Actual	48,955	4,908	850	0	5,000	5,971	65,684

*Cumulative by quarter.

COSTS BY QUARTER

Protocols for the Selective Cleavage of
Carbon-Sulfur Bonds in Coal

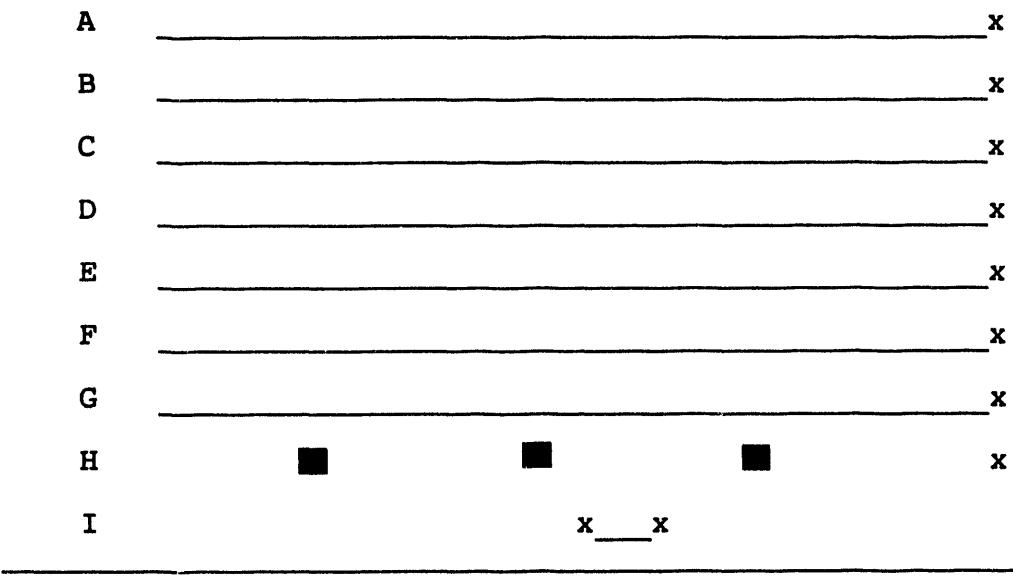
Months and Quarters

O = Projected Expenditures -----

Δ = Estimated Actual Expenditures _____

Total Illinois Clean Coal Institute Award \$84,700

SCHEDULE OF PROJECT MILESTONES



Begin S O N D J F M A M J J A
 Sept. 1
 1992

Milestones:

- A. Post-doctoral research associate employed
- B. Graduate research assistant employed
- C. Graduate research assistant employed
- D. Task 1 undertaken; second Bausch group Task 1 review
2/1/93
- E. Task 2 undertaken; second Bausch group Task 2 review
3/1/93
- F. Task 3 undertaken; second Bausch group Task 3 review
4/1/93
- G. Task 4 undertaken; second Bausch group Task 4 review
5/1/93
- H. Technical reports and project management reports
prepared and submitted
- I. Publication prepared and submitted: "Lewis Acid-
Promoted C-S Bond Cleavage Chemistry"

A high-contrast, black and white image showing a series of geometric shapes. At the top, there are four vertical rectangles of varying widths, with the second and fourth being white and the first and third being black. Below this is a large, solid black rectangle. A diagonal line, consisting of a thick black segment and a thin white segment, cuts across the middle. At the bottom, there is a large, solid black U-shaped or inverted C-shaped area containing a white, rounded rectangular shape.

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